

SYNTHESIS AND STRUCTURAL STUDIES OF TRANSITION  
METAL ALKYNE AND DIAZENE COMPOUNDS

by

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## ABSTRACT

A series of new heterotrimetallic alkyne clusters of general formula  $\text{CpNiCoM}(\text{CO})_6(\text{RC}_2\text{R}')$  ( $\text{M} = \text{Fe}, \text{Ru}, \text{Os}$ ;  $\text{R} = \text{R}' = \text{Ph}, \text{Et}$ ;  $\text{R} = \text{Ph}, \text{R}' = \text{Me}$ ) has been synthesized and the compounds characterized by their infrared, NMR and mass spectral data. In addition, four of the clusters had their solid state structures examined by X-ray diffraction. In all cases examined, the alkyne is oriented parallel to one edge of the trimetallic framework but the particular location varies. In the iron complex  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$ , it is parallel to the Ni-Co edge and all carbonyls are terminal. In the ruthenium and osmium complexes  $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  and  $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ , the orientation is parallel to the Ni-Ru or Ni-Os edge and one of the CO groups adopts a semi-bridging position across the Co-Ru or Co-Os bond. A similar arrangement is observed in an iron complex when one of the carbonyl groups bound to cobalt is replaced by triphenylphosphine to give  $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$ . The alkyne ligand was shown to be fluxional in solution and to involve rotation on only one face of the heterotrimetallic core.

In Chapter 2, the X-ray structure of  $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$ , a heterobimetallic 'ferrole' compound, is reported. The structure consists of a substituted cobaltacyclopentadienyl ring which is bound to the  $\text{Mn}(\text{CO})_3$  fragment by a metal-metal bond and by 4 electron  $\pi$ -bonding. Other significant structural differences are described.

Chapter 3 reports a study of the  $\eta^2$ -coordination of some two-electron donors in mononuclear cyclopentadienyl rhenium compounds. The X-ray structures of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  and  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  reveal an overall 'four legged piano stool' geometry with  $\eta^2$  coordination of both the  $\text{C}_2\text{Ph}_2$  and

$\text{N}_2\text{Ph}_2$  ligand. The  $\text{C}_2\text{Ph}_2$  group is bound symmetrically to the rhenium but the  $\text{N}_2\text{Ph}_2$  ligand is not. In solution,  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  retains its geometry, but for  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  IR and NMR data revealed the presence of isomers which are in dynamic equilibrium and involve both  $\eta^2$ - and  $\eta^1$ -bound  $\text{N}_2\text{Ph}_2$  with the  $\eta^1$  form being favoured in solution.

In Chapter 4, the X-ray structures of cis- $\text{Cp}^*\text{Re}(\text{CO})_2\text{I}_2$  and trans- $\text{Cp}^*\text{Re}(\text{CO})_2\text{Br}_2$  were determined. The rhenium is seven coordinate and the overall geometry is that of a typical  $\text{Cp}^*\text{ML}_4$  complex. The structure of cis- $\text{Cp}^*\text{Re}(\text{CO})_2\text{I}_2$  is slightly disordered limiting the precision of the observed bond lengths and angles.

## QUOTATION

William James used to preach the

"will to believe."

For my part, I should wish

to preach the "will to doubt."...

What is not wanted is not the will to believe

but the wish to find out, which is

the exact opposite.

BERTRAND RUSSELL

Skeptical Essays, 1928

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## LIST OF ABBREVIATIONS

br = broad

m = medium

q = quartet

s = strong

sh = sharp

t = triplet

Cp =  $\eta^5\text{-C}_5\text{H}_5$

Cp\* =  $\eta^5\text{-C}_5\text{Me}_5$

Et =  $\text{C}_2\text{H}_5$

Me =  $\text{CH}_3$

IR = infrared

NMR = nuclear magnetic resonance

GOF = goodness of fit

NOE = nuclear overhauser enhancement

THF = tetrahydrofuran

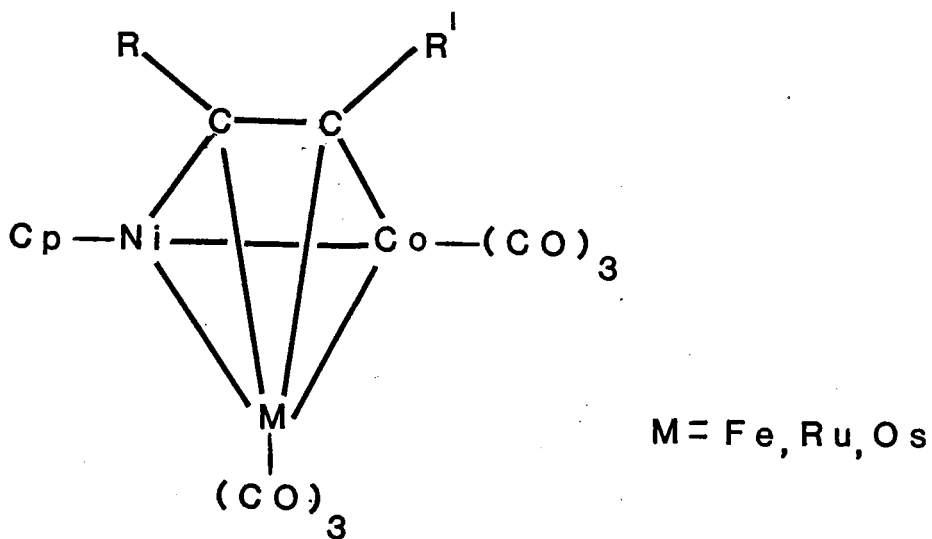
## LIST OF NUMBERED COMPOUNDS

- |  |  |
|--|--|
| $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ (1)                | $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$ (10)                       |
| $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$ (1a) | $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ (11)                         |
| $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ (2)                | $\text{CpRe}(\text{CO})_2(\text{MeC}_2\text{Ph})$ (12)                         |
| $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ (3)                | $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ (13)                         |
| $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$ (4)                | $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$ (14) |
| $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Et}_2)$ (5)                | $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$ (15)  |
| $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Et}_2)$ (6)                | <u>cis</u> - $\text{Cp}^*\text{Re}(\text{CO})_2\text{I}_2$ (16)                |
| $\text{CpNiCoFe}(\text{CO})_6(\text{MeC}_2\text{Ph})$ (7)                | <u>trans</u> - $\text{Cp}^*\text{Re}(\text{CO})_2\text{Br}_2$ (17)             |
| $\text{CpNiCoRu}(\text{CO})_6(\text{MeC}_2\text{Ph})$ (8)                | $\text{Os}_3(\text{CO})_{10}\text{Cl}_2$ (18)                                  |
| $\text{CpNiCoOs}(\text{CO})_6(\text{MeC}_2\text{Ph})$ (9)                |  |

## CHAPTER 1

HETEROTRIMETALLIC ALKYNE CLUSTERS1.1 Introduction

In this chapter the synthesis, structural characterization and dynamic behaviour in solutions of a series of heterotrimetallic (MM'M'') alkyne clusters<sup>a</sup> having the general composition shown in I-1 are detailed, where R and R' are substituents on the alkyne. The structure illustrated is representative but modifications of it are displayed by many of the examples as will be discussed.



I-1

Increasing interest in cluster compounds and their derivatives has led to many recent reviews on various aspects of their bonding<sup>1</sup>, ligand

<sup>a</sup> The term heterotrimetallic is ambiguous since it may be used to describe clusters with three different metals, such as MM'M'', as well as those with only two different metals, designated here as M<sub>2</sub>M'.

fluxionality<sup>2</sup> and preparation<sup>3,4</sup>. Only 15 years ago the number of well defined metal clusters was small, but today many hundreds of clusters are known and a large fraction of these have been structurally characterized by X-ray crystallographic studies. One of the main reasons for the interest in this field is based on the premise that metal clusters may be viewed as models for the chemisorption of small molecules on metal surfaces<sup>5,6</sup>, possibly during catalysis. The nature of the bonding of molecules on metal surfaces is not well understood at the present time. Analogies that might be drawn with homogeneous cluster species would be extremely beneficial since the details of structure and dynamic behaviour are more easily established for cluster chemistry than for surface chemistry. Recent assessments of the metal cluster to metal surface analogy have suggested a range of validity which improves as the cluster nuclearity increases<sup>7-9</sup>. A separate and distinct view is that cluster compounds may serve as incipient metal particles that may provide active aggregates of uniform size and composition when spread over a suitable support and the sheath of ligands stripped off. This is referred to as supported homogeneous catalysis<sup>10</sup> and is starting to receive an increasing amount of attention<sup>11</sup>.

Cluster compounds may serve as catalyst precursors in their own right rather than simply mimicking metal surfaces. Metal cluster chemistry provides a connection between the traditionally separate areas of homogeneous and heterogeneous catalysis and may offer advantages over both areas. To this end, a great deal of effort is being directed towards the use of clusters as catalysts for such varied reactions as hydroformylation<sup>12,13</sup>, isomerization of alkenes and alkynes<sup>13,14</sup>, hydrogen reduction of carbon monoxide<sup>15</sup>, oxidation of CO to CO<sub>2</sub><sup>16</sup> and many more.

In this context it is of interest to establish how simple organic

molecules interact with a multimetal system and this is partly the thrust of the present study which specifically probes the bonding of alkynes. An increasingly popular and important subset of cluster chemistry comprises those clusters that contain a mixture of different transition metals in their framework and this is a second aspect of this study in that it specifically utilizes triangular clusters comprising three different metals. In particular, because of the varying reactivities of the different metals present in mixed-metal clusters, multimetal homogeneous catalysts may show reactivity patterns significantly altered from those of homometallic clusters. Of course, the bonding of various organic molecules to these heterometallic clusters may be different than to the analogous homometallic clusters also, but a very useful consequence of using heterometallic clusters is their low symmetry. This allows the dynamics (H, CO or other ligand migrations) of the cluster or cluster derivatives to be probed by IR,  $^1\text{H}$  or  $^{13}\text{C}$  NMR, in greater detail than is often possible for homometallic clusters. This is a third aspect that we shall address in this study.

In order to set the research in context a brief survey of clusters and cluster alkynes in particular is provided in the next few sections.

### 1.1.1 Clusters

Metal clusters are molecules or molecular ions that contain three or more metals bonded to each other in such a way that they form a triangular or polyhedral array. We shall only be concerned with transition metal-based clusters and will merely reference the reader to the subgroup of post-transition metal cluster naked ions<sup>18</sup> like  $\text{Bi}_5^{3-}$  and  $\text{Pb}_9^{4-}$ , and to the cubane-like  $\text{M}_4\text{S}_4$  and  $\text{M}_4\text{X}_4$  clusters used as models for ferredoxins<sup>19,20</sup>. The framework of the transition metal cluster is surrounded and stabilized by



various ligands, commonly carbonyls, hydrides, phosphines and halides but may also include alkenes, alkynes, isonitriles, sulphur, arenes, nitrosyls and so on.

Clusters were originally synthesized by serendipitous methods and seldom by rational synthetic procedures. Gradually, as more was learned about clusters, designed synthetic strategies began to appear, but fortune still plays its role in the development of clusters and their chemistry. The particular methods used so far to synthesize homo- and heterometallic clusters have been well documented<sup>3,4,21,22</sup> and will not be reiterated here.

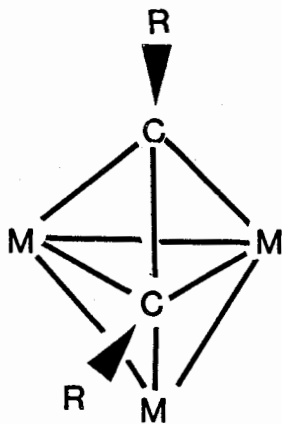
## 1.2 Cluster Alkynes

At the time work on this thesis began, a significant literature existed on the binding of alkynes to trimetallic and larger clusters. However, only one report of a heterotrimetallic (MM'M'') alkyne cluster existed, which was the synthesis of  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ <sup>23</sup>, although it was incorrectly formulated at that time. Most of the workers previous<sup>24</sup> to our study have clearly focussed on the synthesis and characterization of  $\text{M}_2\text{M}'$  heterotrimetallic alkyne clusters although during this investigation reports were published of MM'M'' heterotrimetallic alkyne clusters closely related to those discussed<sup>25</sup>. Specifically these two clusters are  $\text{Cp}_2\text{NiMoFe}(\text{CO})_5(\text{PhC}_2\text{CO}_2\text{i-Pr})$  and  $\text{CpNiCoFe}(\text{CO})_6(\text{PhC}_2\text{CO}_2\text{i-Pr})$ .

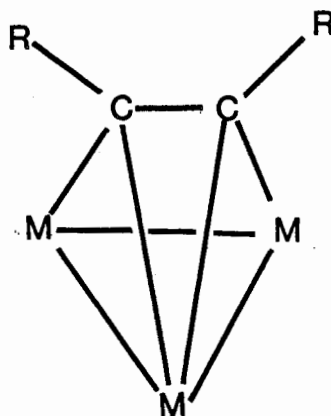
### 1.2.1 Bonding and Structure

Alkynes interact with one metal atom in the classic  $\sigma$ - $\pi$  bonding scheme described by Dewar<sup>26</sup> and Chatt and Duncanson<sup>27</sup>. This is briefly described

in the introduction to Chapter 4. The bonding of an alkyne to a single metal atom may arise in the initial stages in the reaction of an alkyne with a metal cluster but typically the isolated products have the alkyne interacting with more than one metal atom. With trimetallic alkyne clusters the two basic structures that are adopted are shown below as I-2 and I-3.



I-2

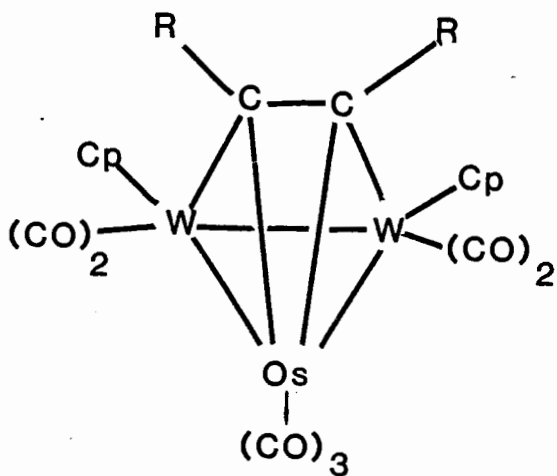


I-3

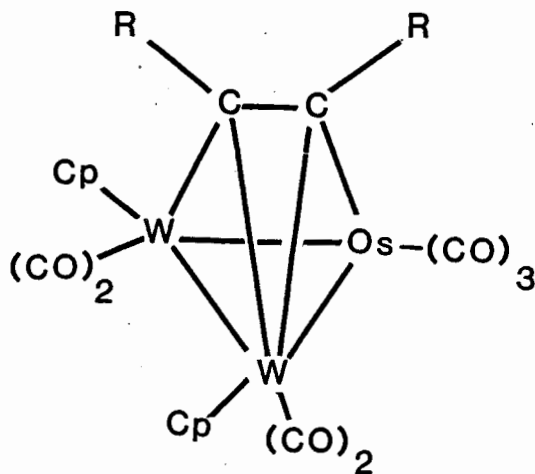
In structure I-2 the alkyne symmetrically bridges two metal atoms and the carbon of the alkyne which is above the  $M_3$  triangle is within bonding distance of all three metal atoms. This combines features of a  $\sigma$ -bonded and a side-on or ' $\pi$ ' bonded ligand and has been designated  $(\sigma + 2\pi)$  in reference to classical bonding schemes<sup>24</sup>. In current nomenclature it is often (and properly) designated  $\mu_3-(\eta^2-\perp)$ , where the  $\perp$  symbol refers to the perpendicular orientation of the alkyne  $C_2$  vector with respect to the symmetrically bridged metal-metal bond. The alternative form I-3 has the alkyne  $C_2$  vector above the metal plane and nearly parallel to the  $M_2$  vector but displaced towards the third metal atom to such an extent that both alkyne carbons are within bonding distance. While structure I-3 can be fully described in classical bonding terms and the alkyne as a  $(2\sigma + \pi)$

donor, it is properly designated as  $\mu_3-(\eta^2-||)$ .

The  $\mu_3-(\eta^2-||)$  alkyne interaction is more widely encountered than the  $\mu_3-(\eta^2-\perp)$  mode among homo- and heterometallic alkyne clusters, with 21 X-ray structures of the former listed in a recent review<sup>24</sup>, compared with only 3 structures for the latter geometrical type. Specifically only  $\text{Fe}_3(\text{CO})_9(\text{C}_2\text{Ph}_2)$ <sup>28</sup>,  $\text{Cp}_2\text{W}_2\text{Fe}(\text{CO})_6(\text{C}_2(\text{C}_6\text{H}_4\text{Me})_2)$ <sup>29</sup> and  $\text{Fe}_2\text{Ru}(\text{CO})_9(\text{C}_2\text{Ph}_2)$ <sup>30</sup> have been characterized as having a  $\mu_3-(\eta^2-\perp)$  alkyne ligand. The  $\mu_3-(\eta^2-||)$  alkyne types include  $\text{Cp}_3\text{Rh}_3(\text{CO})(\text{C}_2\text{Ph}_2)$ <sup>31</sup> and  $\text{Cp}_2\text{W}_2\text{Os}(\text{CO})_7[\text{C}_2(\text{C}_6\text{H}_4\text{Me})_2]$ <sup>32</sup>. The latter alkyne cluster is of interest since it has two isomers in the asymmetric unit which differ principally by rotation of the alkyne ligand above the trimetallic face. These are shown in I-4 and I-5. Molecule I-4



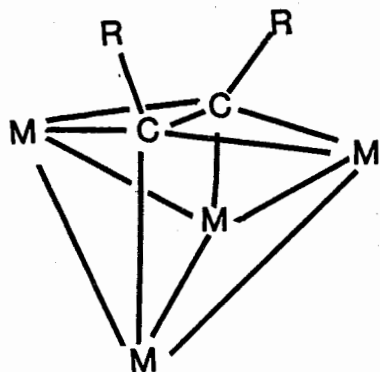
I-4



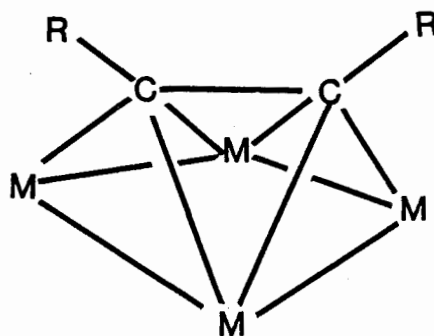
I-5

is essentially the symmetric isomer where the alkyne ligand bridges two tungsten atoms and has a  $\eta^2$ -linkage to the osmium atom. I-5 has the alkyne ligand spanning the W-Os bond while having a  $\eta^2$ -linkage to the second W atom. These two isomers, observed in the same crystal, describe the two possible rotamers of alkyne bonding to a  $\text{M}_2\text{M}'$  system.

There are two types of alkyne- $M_n$  structures known, the butterfly and the square-planar structure. These are represented by I-6 and I-7. The arrangement of metal atoms and alkyne carbons in the butterfly structure



I-6



I-7

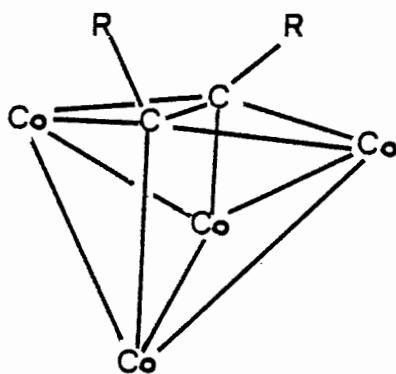
essentially completes an irregular octahedral geometry while in the square planar form the alkyne lies above the almost planar array of metal atoms. Heterotetrametallic alkyne clusters are fewer in number than the homometallic analogues but  $FeRu_3(CO)_{12}(C_2Ph_2)$  and  $Cp_2Ni_2Fe_2(CO)_6(C_2Ph_2)$  serve as examples of butterfly and square planar arrangements respectively<sup>33,34</sup>.

### 1.2.2 Synthesis of Cluster Alkynes

In reviewing this subject the synthetic strategies for the assembly of heterometallic alkyne clusters will be emphasized although the reactions are often general and could equally be used for the homometallic analogues. Indeed, many of the synthetic routes described were adopted or modified from methods used in the homometallic cluster alkyne field.

### Substitution of CO

This is one of the more common methods used for the synthesis of both homo- or heterometallic alkyne clusters and certainly is the most obvious. The appropriate cluster precursor is first assembled and then the alkyne is added to the cluster with consequent loss of the correct number of CO groups. The heterotrimetallic (MMM') cluster  $\text{Co}_2\text{Ru}(\text{CO})_9(\text{C}_2\text{R}_2)$ , can be synthesized in this fashion from  $\text{Co}_2\text{Ru}(\text{CO})_{11}$  and the desired alkyne ligand<sup>35</sup>. Rearrangement of the metal skeleton or breaking of metal-metal bonds is a possible consequence of this reaction since the alkyne fragment prefers to interact with more than one metal atom of a cluster. A well known example<sup>36</sup> is the reaction of  $\text{Co}_4(\text{CO})_{12}$  with an alkyne yielding  $\text{Co}_4(\text{CO})_{10}(\text{C}_2\text{R}_2)$  (I-8). Two CO ligands have been lost, a cobalt-cobalt bond has been severed and the alkyne inserted.

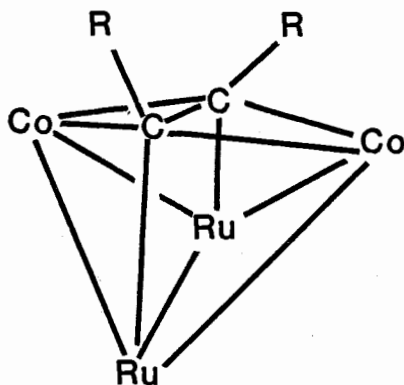


**I-8**

has been severed and the alkyne inserted. A  $\text{C}_{2v}$  arrangement of the metal atoms or a butterfly structure is the end product.

An interesting heterometallic example is the reaction of  $\text{Co}_2\text{Ru}_2(\text{CO})_{13}$  with alkynes<sup>37</sup>. As in  $\text{Co}_4(\text{CO})_{12}$ , the tetrahedral cluster loses two CO ligands and a butterfly cluster is obtained but exclusive formation of only

one isomer (I-9) out of three possibilities is observed; specifically, that in which the Co-Co bond is cleaved and not the Ru-Co or the Ru-Ru bond.



**I-9**

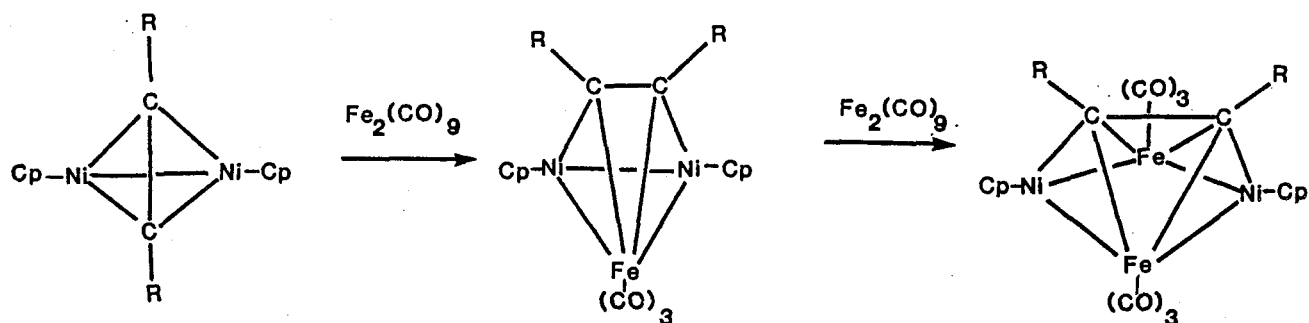
It should be noted here that, while this reaction may appear to be general, there are tetrahedral clusters which will not add alkynes in a similar fashion.  $\text{Cp}_2\text{Ni}_2\text{Fe}_2(\text{CO})_7$  has not been found to add an alkyne with loss of a carbonyl to form  $\text{Cp}_2\text{Ni}_2\text{Fe}_2(\text{CO})_6(\text{C}_2\text{R}_2)$ , even though this known compound has been synthesized by a different method and structurally characterized<sup>38</sup>.

#### Metal Fragment Condensation

This synthetic method is growing more popular as it often allows systematic addition of an unsaturated mononuclear fragment to an existing cluster. The unsaturated fragment (such as  $\text{Cr}(\text{CO})_5$ ) is commonly formed in situ or can be obtained by the loss of a labile ligand from a modestly stable organometallic fragment such as  $(\text{PPh}_3)_2\text{Pt}(\text{C}_2\text{H}_4)$ . This allows cluster growth to often appear more rational and not rely on empirical approaches with little predictability. However, generalizations are still difficult to

make and approaches only experimental.

The two most common fragments that are employed in this reaction are CpNi and  $\text{Fe}(\text{CO})_3$ . CpNi is usually obtained in situ<sup>38,39</sup> from  $\text{Cp}_2\text{Ni}_2(\text{CO})_2$  or  $\text{Cp}_2\text{Ni}$ , and  $\text{Fe}_2(\text{CO})_9$  is the most common precursor for  $\text{Fe}(\text{CO})_3$  fragments although  $\text{Fe}(\text{CO})_5$  and  $\text{Fe}_3(\text{CO})_{12}$  have also been employed<sup>2,24</sup>. A good example of this reaction is illustrated in Scheme I-1.  $\text{Cp}_2\text{Ni}_2(\text{C}_2\text{R}_2)$  (A), a



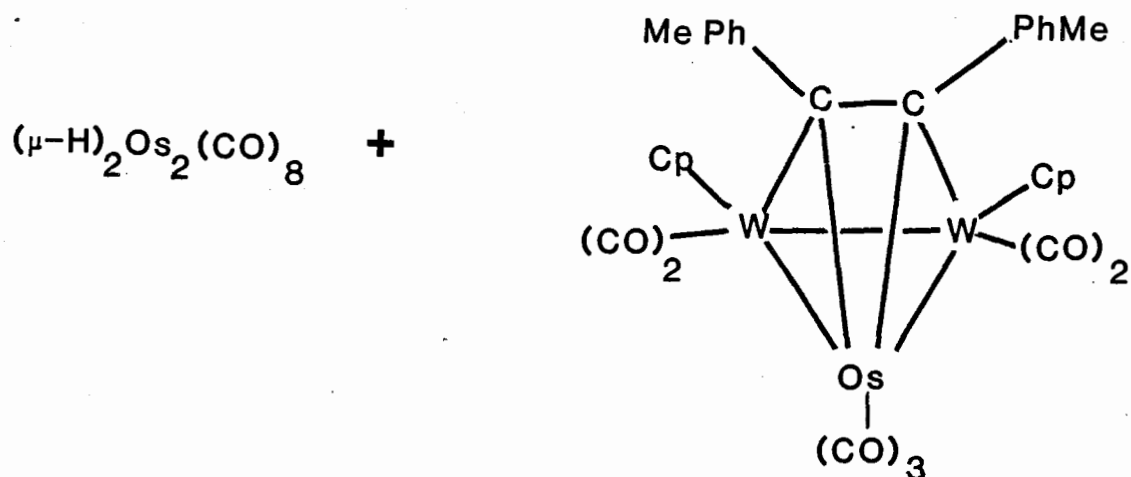
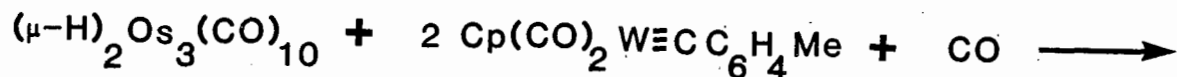
**SCHEME I-1**

dinuclear alkyne compound which is formally saturated (i.e., all metals have an 18 electron count) adds an  $\text{Fe}(\text{CO})_3$  group to form (B), a heterotrimetallic (MMM') cluster that can be isolated or add a further  $\text{Fe}(\text{CO})_3$  unit to form a heterotetrametallic ( $\text{M}_2\text{M}'_2$ ) alkyne cluster (C)<sup>38</sup>.

### Carbyne Coupling

Alkyne ligands have often been observed to cleave their carbon-carbon triple bond upon reaction with some transition metal complexes to form  $\mu_3$ -carbyne species<sup>41</sup>. However Stone<sup>42</sup> and Churchill<sup>32</sup> have recently utilized the reverse reaction, carbyne coupling, to form a number of cluster alkyne compounds. Both workers used mononuclear carbynes ( $\text{M}\equiv\text{C-R}$ ) as precursors in

this reaction. An example is given in Scheme I-2. The mononuclear carbyne compound couples with itself and adds an Os(CO)<sub>2</sub> fragment from the



**SCHEME I-2**

$(\mu\text{-H})_2\text{Os}_3(\text{CO})_{10}$  cluster in addition to scavenging a CC to form the MMM' alkyne cluster,  $\text{Cp}_2\text{W}_2\text{Os}(\text{CO})_7(\text{C}_2(\text{C}_6\text{H}_4\text{Me})_2)$ .

### Reactions of Carbonyl Anions

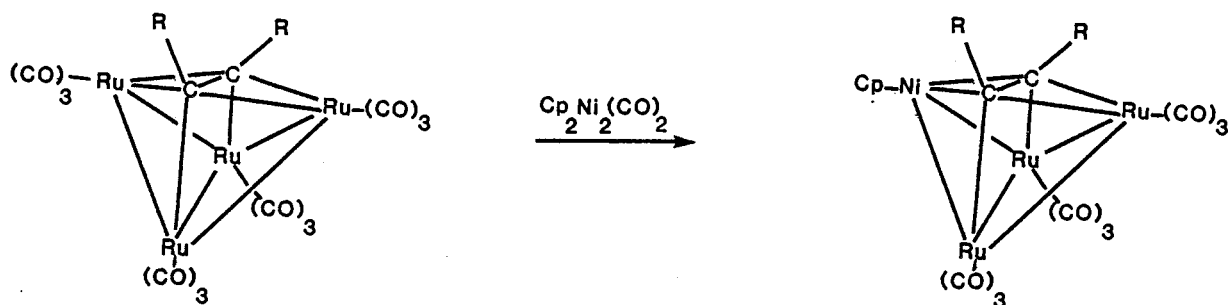
The synthesis of mixed-metal clusters via the addition of a carbonylmethylate to a closed metal carbonyl cluster has been pioneered by Geoffroy<sup>43</sup> and shown to be a versatile method. The compounds are usually isolated as the hydride clusters after acidification rather than as the anionic cluster. Only a few homo- or heterometallic alkyne clusters have been made in this fashion.  $\text{H}_2\text{Ru}_3(\text{CO})_9(\text{C}_2\text{Ph}_2)$  can be synthesized from



$\text{Ru}_3(\text{CO})_{12}$  and  $\text{C}_2\text{Ph}_2$  in basic methanol followed by acidification<sup>44</sup>.

### Metal Substitution or Metal Exchange Reactions

The specific exchange of one metal unit for another in a metal cluster was first discovered by Vahrenkamp<sup>45</sup>. Since that time the method has been extended to the synthesis of heterometallic alkyne clusters. The example shown in Scheme I-3 has a CpNi fragment substituting for a  $\text{Ru}(\text{CO})_3$  moiety. Only exchange with a wing ruthenium is observed and no other isomers were



SCHEME I-3 .

reported<sup>46</sup>. This method has been growing in popularity as it allows the synthesis of mixed-metal alkyne clusters which could not be made in any other fashion. Certainly dinuclear metal-metal bonded complexes<sup>47</sup> and dinuclear alkyne compounds<sup>38</sup> equilibrate easily with formation of the mixed metal complexes. However, this reaction has not yet reached the stage in cluster alkyne chemistry where products are totally predictable but it clearly demonstrates potential.

### Metal Extrusion

Only a few examples exist of the extrusion of a metal from a cluster

alkyne species to give a cluster of lower nuclearity. These include the proton induced fragmentation<sup>48</sup> of  $[\text{Co}_3\text{Ru}(\text{CO})_{12}(\text{C}_2\text{Ph}_2)]^-$  to give  $\text{Co}_2\text{Ru}(\text{CO})_{10}(\text{C}_2\text{Ph}_2)$  and the phosphine induced cleavage<sup>40</sup> of  $\text{Cp}_2\text{Ni}_2\text{Fe}_2(\text{CO})_6(\text{C}_2\text{Ph}_2)$  to yield  $[\text{CpNi}(\text{PR}_3)_2]^+[\text{CpNiFe}_2(\text{CO})_6(\text{C}_2\text{Ph}_2)]^-$ . This method is unlikely to become a general preparative reaction and will not be treated further.

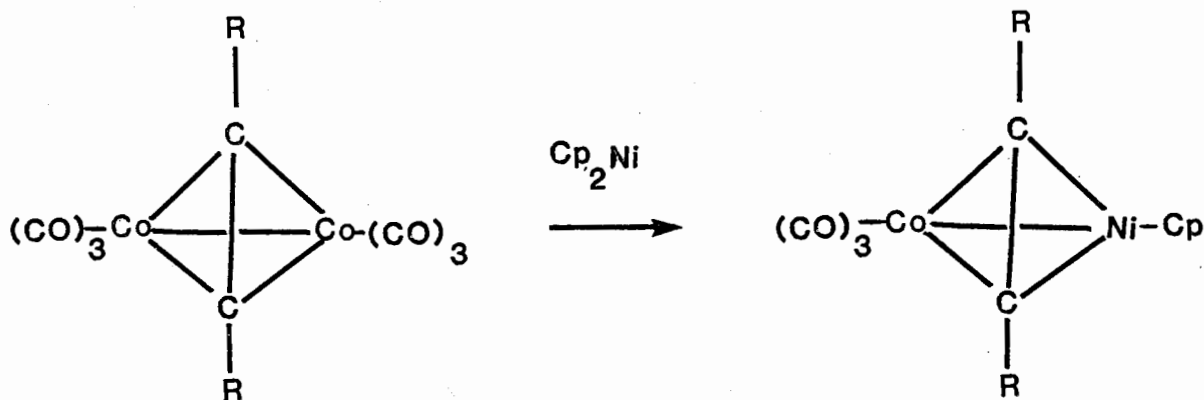
### 1.3 Synthesis and Characterization of the Heterotrimetallic Alkyne Clusters, $\text{CpNiCoM}(\text{CO})_6(\text{RC}_2\text{R}')$ (M = Fe, Ru, Os)

The subject of the present investigation is now described. The study entails developing a method for the synthesis of a family of heterotrimetallic alkyne clusters in which all three metal atoms are different (MM'M") and one of them is varied among the congeners Fe, Ru and Os. The solid state structures have been obtained by IR and X-ray methods of as closely representative members as possible comprising each metal of the iron group in turn. Because of inevitable difficulties in obtaining X-ray quality crystals, minor changes in the alkyne and some other modifications have been necessitated. The observed structures are not identical down the series, and the significant changes which occur will be discussed. The solid state structures are complemented by solution studies by IR and particularly dynamic  $^1\text{H}$  NMR methods. The clusters are found to be stereochemically non-rigid in solution and the nature of this non-rigidity involving reorganization of the alkyne on the trimetallic face will be described.

#### 1.3.1 Synthesis

It was decided to pursue the method of metal fragment condensation in order to synthesize the desired heterotrimetallic alkyne clusters.  $\text{CpNiCo}(\text{CO})_3(\text{RC}_2\text{R}')$  was chosen as the starting heterobimetallic alkyne compound because a variety of alkynes could be bound;  $\text{C}_2\text{Ph}_2$ ,  $\text{C}_2\text{Et}_2$  and  $\text{MeC}_2\text{Ph}$  were chosen. Also the compounds could be easily synthesized<sup>38</sup> in fair yield from  $\text{Co}_2(\text{CO})_8(\text{RC}_2\text{R}')$  and  $\text{Cp}_2\text{Ni}$  (or  $[\text{CpNi}(\text{CO})]_2$ ) by a metal

exchange reaction as shown below. In addition, the compounds were air stable



both in the solid state and solution and could be separated easily from other materials by column chromatography.

The reaction of excess  $\text{Fe}_2(\text{CO})_9$  with  $\text{CpNiCo}(\text{CO})_3(\text{RC}_2\text{R}')$  in hexanes produces the desired NiCoFe alkyne clusters,  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (1),  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (4) and  $\text{CpNiCoFe}(\text{CO})_6(\text{MeC}_2\text{Ph})$  (7) in reasonable yield (25-30%). These clusters can be separated from starting materials and side-products by repeated column chromatography.  $\text{Fe}_2(\text{CO})_9$  fragments to give  $\text{Fe}(\text{CO})_4$  and  $\text{Fe}(\text{CO})_5$  groups, which can recombine with  $\text{Fe}_2(\text{CO})_9$  to form  $\text{Fe}_3(\text{CO})_{12}$  which was always present and isolated in these reactions. The formation of  $\text{Fe}_3(\text{CO})_{12}$  is enhanced by an increase in the reaction temperature or by changing solvents from hexanes to diethyl ether or THF. Also, when THF or diethyl ether was used a substantial amount of the black  $\text{Fe}_3(\text{CO})_9(\text{C}_2\text{Ph}_2)$  was isolated<sup>28</sup>. Replacement of  $\text{Fe}_2(\text{CO})_9$  by  $\text{Fe}(\text{CO})_5$  led to an increased number of side-products, especially  $\text{Cp}_2\text{Ni}_2\text{Fe}_2(\text{CO})_6(\text{C}_2\text{Ph}_2)$ <sup>34</sup>, at the expense of the desired NiCoFe cluster. The use of  $\text{Fe}_3(\text{CO})_{12}$  rather than  $\text{Fe}(\text{CO})_5$  or  $\text{Fe}_2(\text{CO})_9$  led to very poor yields of 1 that were not improved upon changing solvent or temperature.

The formation of CpNiCoRu(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (2), CpNiCoRu(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>) (5) and CpNiCoRu(CO)<sub>6</sub>(MeC<sub>2</sub>Ph) (8) was effective when Ru(CO)<sub>5</sub>, but not Ru<sub>3</sub>(CO)<sub>12</sub> was employed. The yields were low (10-15%), but the unreacted Ru(CO)<sub>5</sub> could be recovered from the reaction as Ru<sub>3</sub>(CO)<sub>12</sub> and reconverted to Ru(CO)<sub>5</sub> and reused. The use of Ru<sub>3</sub>(CO)<sub>12</sub> led only to trace amounts of 2, 5 or 8 being formed.

The synthesis of the osmium analogues, CpNiCoOs(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (3), CpNiCoOs(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>) (6) and CpNiCoOs(CO)<sub>6</sub>(MeC<sub>2</sub>Ph) (9) was similar to that of the ruthenium clusters. Only Os(CO)<sub>5</sub> and not Os<sub>3</sub>(CO)<sub>12</sub> was successful at room temperature. When the reaction temperature was increased a small amount of NiCoOs cluster was obtained using Os<sub>3</sub>(CO)<sub>12</sub>, but this offered no advantage to using Os(CO)<sub>5</sub> at lower temperatures.

### 1.3.2 Characterization

The heterotrimetallic alkyne clusters have been characterized by solution and solid state (KBr) IR (Table 1), X-ray, <sup>1</sup>H NMR (Table 2), mass spectral analysis (Table 2) and by elemental analysis where possible (see Experimental). In addition, for each cluster the calculated isotopic distribution of the parent ion in the mass spectrum was compared to the experimental distribution.

The numbering scheme used for the clusters is shown below.

CpNiCoFe(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> ) - <u>1</u>	CpNiCoRu(CO) <sub>6</sub> (C <sub>2</sub> Et <sub>2</sub> ) - <u>5</u>
CpNiCoFe(CO) <sub>5</sub> (PPh <sub>3</sub> )(C <sub>2</sub> Ph <sub>2</sub> ) - <u>1a</u>	CpNiCoOs(CO) <sub>6</sub> (C <sub>2</sub> Et <sub>2</sub> ) - <u>6</u>
CpNiCoRu(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> ) - <u>2</u>	CpNiCoFe(CO) <sub>6</sub> (MeC <sub>2</sub> Ph) - <u>7</u>
CpNiCoOs(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> ) - <u>3</u>	CpNiCoRu(CO) <sub>6</sub> (MeC <sub>2</sub> Ph) - <u>8</u>
CpNiCoFe(CO) <sub>6</sub> (C <sub>2</sub> Et <sub>2</sub> ) - <u>4</u>	CpNiCoOs(CO) <sub>6</sub> (MeC <sub>2</sub> Ph) - <u>9</u>

#### 1.3.2a Infrared Spectroscopy

The solution (hexanes) IR spectra (listed in Table 1) did not vary

greatly between compounds of the same trimetallic framework, i.e., NiCoFe, with different alkynes. The IR spectra of 1, 4 and 7, (all with NiCoFe frameworks) are shown in Figure 1 for comparison. There are certainly some small shifts in wavenumber between the different species and variations in intensity, but all have the highest wavenumber absorption at about  $2068\text{ cm}^{-1}$  and the most intense band around  $2022\text{ cm}^{-1}$ . The NiCoRu and NiCoOs clusters differ from the NiCoFe analogues but are fairly similar among themselves. Again, the highest wavenumber absorption is around  $2068\text{ cm}^{-1}$  and the largest peak is now found at  $2040\text{ cm}^{-1}$ . In addition to sharp terminal  $\nu(\text{CO})$  absorptions, all the NiCoRu and NiCoOs alkyne clusters show a weak, broadened carbonyl peak in the  $1885\text{--}1895\text{ cm}^{-1}$  region. This is assignable to a  $\mu_2\text{-CO}$  carbonyl species. While broadened in solution, an IR of the cluster as a KBr pellet revealed the band more clearly, and an example is shown in Figure 2. While 2 has a very intense and sharp bridging CO band, both the intensity and broadness of this absorption vary from cluster to cluster. The assignment of the lowest wavenumber band to a doubly bridging CO is in agreement with the X-ray structures of 2 and 3 which both contain a semibridging CO.

1a, which has only five carbonyl ligands, displays a different IR from the hexacarbonyl clusters. The highest wavenumber absorption occurs at  $2039\text{ cm}^{-1}$  (rather than  $2068\text{ cm}^{-1}$  for the other clusters) and a bridging CO species is noted which was not observed for other NiCoFe clusters. The infrared spectrum as a KBr pellet the IR showed the  $\mu_2\text{-CO}$  species to be at a lower wavenumber than for the hexacarbonyl clusters;  $1827$  versus  $1858\text{--}1876\text{ cm}^{-1}$ . The X-ray structure of 1a confirmed the existence of a semibridging CO group.

Table 1. Infrared Data for Compounds 1-9

<u>Cluster</u>	<u><math>\nu(\text{CO})</math> Hexanes (<math>\text{cm}^{-1}</math>)</u>	<u><math>\nu(\text{CO})</math> KBr (<math>\text{cm}^{-1}</math>)</u>
$\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ (1)	2069s, 2024s, 2020s, 2006m, 1970m, 1965m	2067s, 2016s,br, 1971s
$\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$ (1a)	2039s, 1974s, 1969s, 1838vbr,w	2040s, 1965s,br, 1827m
$\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ (2)	2072s, 2043vs, 2010s, 2002s, 1885w,br	2072s, 2051s, 2046s, 2015s, 2008s, 1985s, 1865s
$\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ (3)	2072s, 2041s, 2015m, 2002s, 1989s, 1973m, 1884w,br	2069s, 2044s, 2000vs, 1983s, 1858s
$\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$ (4)	2065s, 2021vs, 2013s, 2004s, 1972m, 1968m	2064s, 2013vs,br, 1959m
$\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Et}_2)$ (5)	2068s, 2035vs, 2005s, 1980w, 1894w,br	2065s, 2027s, 1997s,br, 1874m,br

Table 1 (Cont'd)

<u>Cluster</u>	<u><math>\nu(\text{CO})</math> Hexanes (<math>\text{cm}^{-1}</math>)</u>	<u><math>\nu(\text{CO})</math> KBr (<math>\text{cm}^{-1}</math>)</u>
$\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Et}_2)$ (6)	2068s, 2036vs, 2008s, 2002s, 1981m, 1970m, 1874vw,br	2065s, 2031s, 1998s, 1967m,sh, 1875vw,br
$\text{CpNiCoFe}(\text{CO})_6(\text{MeC}_2\text{Ph})$ (7)	2067s, 2021vs, 2006s, 1980m, 1972m, 1967m	2053s, 2050s, 2029m,sh, 1996s,br
$\text{CpNiCoRu}(\text{CO})_6(\text{MeC}_2\text{Ph})$ (8)	2071s, 2039vs, 2008s, 1998s, 1887w,sl.br	2069s, 2033s, 1996s,br, 1876m,sl.br
$\text{CpNiCoOs}(\text{CO})_6(\text{MeC}_2\text{Ph})$ (9)	2069s, 2040vs, 2011m, 2002s, 1985m, 1972m, 1886w,br	2066s, 2033s, 1993s,br, 1871m,br



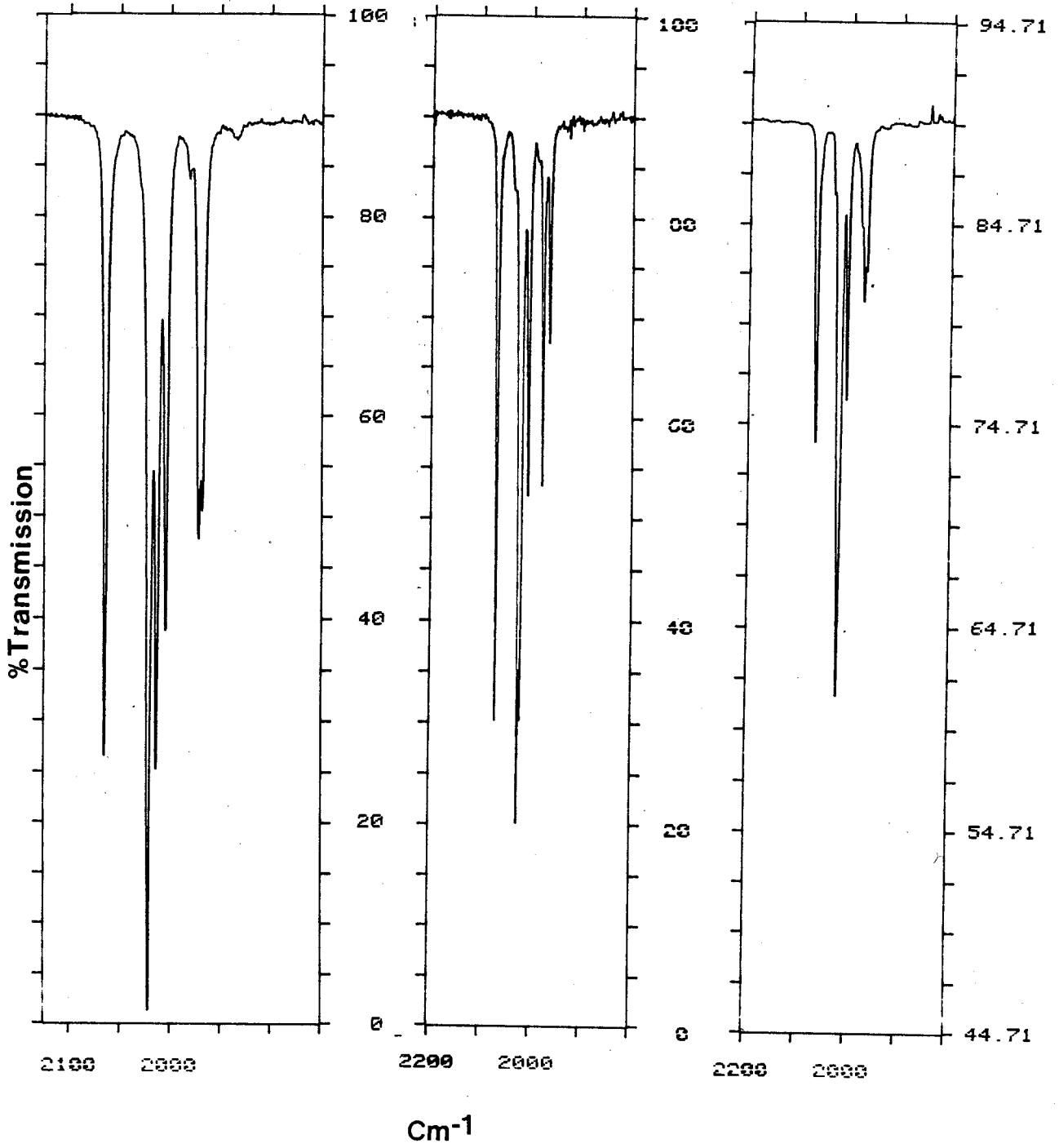
## FIGURE 1

Infrared Spectra (hexanes,  $\nu(\text{CO})$ ) of

$\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (1),

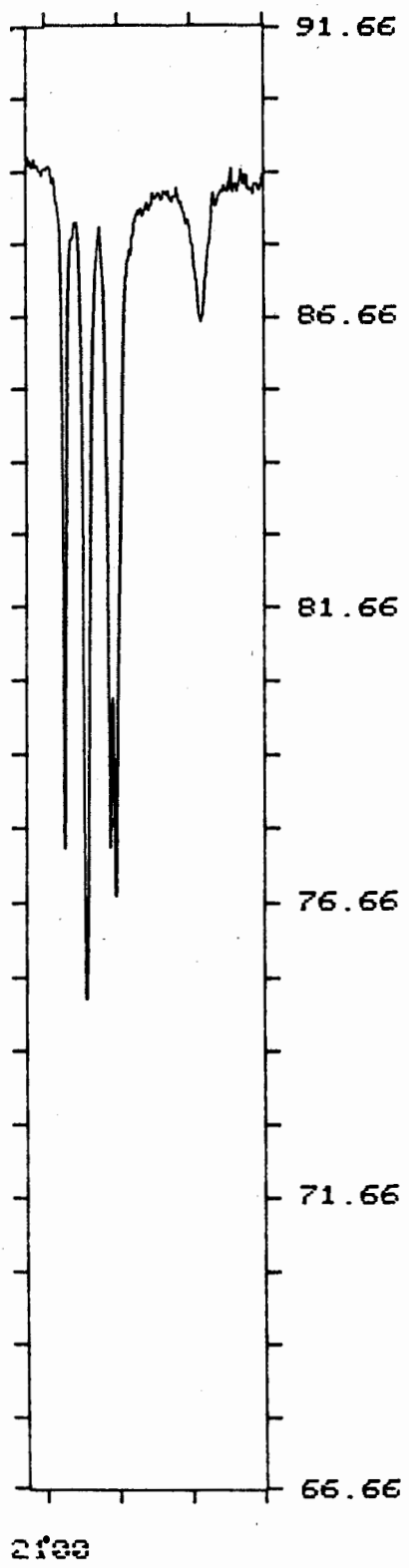
$\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (4),

$\text{CpNiCoFe}(\text{CO})_6(\text{MeC}_2\text{Ph})$  (7)

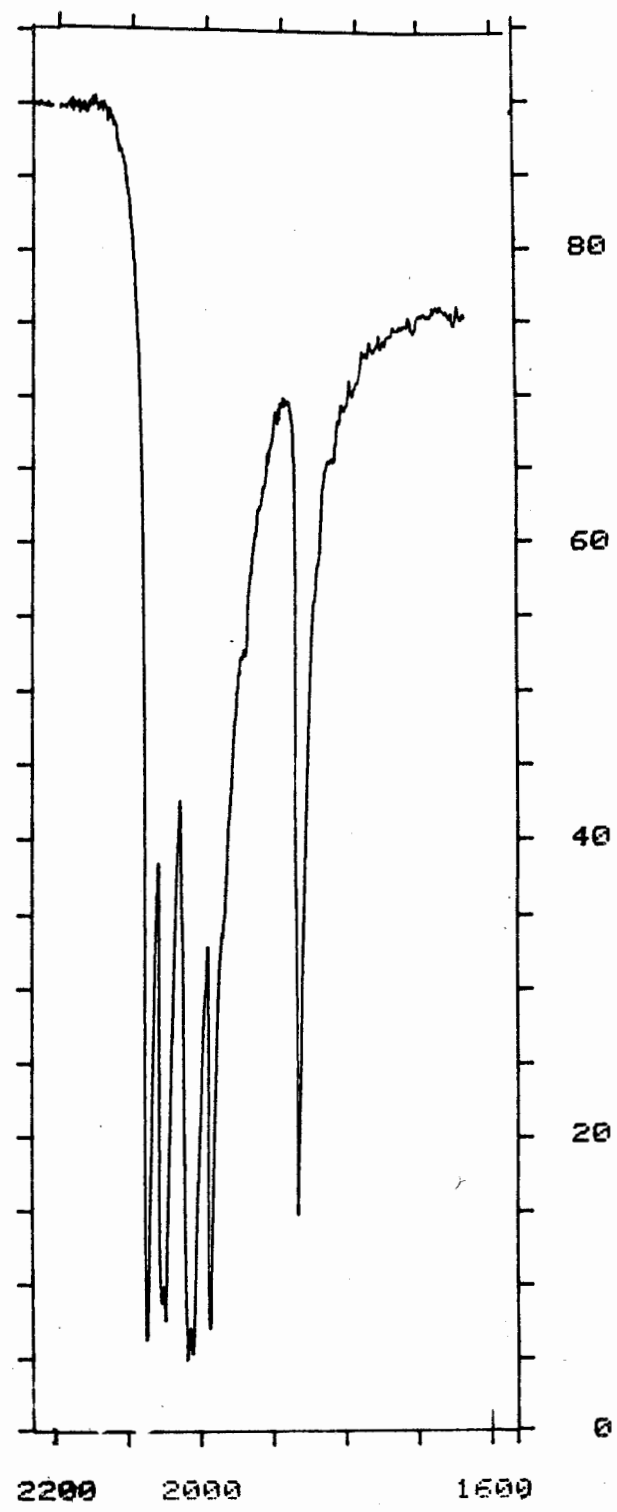
147

## FIGURE 2

Infrared Spectra of  $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$   
in hexanes and KBr



Hexanes



KBr

$\text{Cm}^{-1}$

### 1.3.2b Mass Spectra

As shown in Table 2 the mass spectrum of each hexacarbonyl heterotrimetallic cluster revealed the parent ion and the sequential loss of six carbonyl groups. 1a which contains a PPh<sub>3</sub> group also yielded a parent ion albeit in small intensity. Often clusters with phosphine ligands do not show parent ions due to a lack of stability. Other peaks, notably those belonging to CpNi (123 amu), CpNiCoFe (238 amu), CpNiCoRu (284 amu) and CpNiCoOs (374 amu) and the appropriate alkyne ligand; C<sub>2</sub>Et<sub>2</sub> (82 amu), MeC<sub>2</sub>Ph (116 amu) or C<sub>2</sub>Ph<sub>2</sub> (178 amu) were observed. In addition, 1a revealed the loss of its PPh<sub>3</sub> (262 amu) ligand.

The isotopic distribution of the parent ion can serve as a fingerprint for a given composition, particularly for the metal atoms, as some have a number of isotopes. All of the heterotrimetallic cluster alkynes 1-9, except 1a, had their parent ion distributions calculated and compared with the values obtained experimentally to confirm the presence of the three different metals. An example of the observed and calculated mass distribution found for the parent ion of 5 is shown in Figure 3 and clearly there is close agreement. All of the compounds 1-9 showed similar good agreement. Since 1a had a weak parent ion, the calculated mass distribution was performed with the P-CO<sup>+</sup> ion and this agreed satisfactorily with the observed values.

### 1.3.2c NMR Spectra

The <sup>1</sup>H NMR data for 1-9 are listed in Table 2. Compounds 1-3 all show a single sharp peak around δ 5.10 assignable to the Cp protons and a multiplet at δ 7.1-7.3 due to the protons of the diphenylacetylene ligand.

The diethylacetylene substituted compounds, 4-6, have an interesting <sup>1</sup>H

Table 2.  $^1\text{H}$  NMR and Mass Spectral Data for 1-9

<u>Cluster</u>	<u><math>^1\text{H}</math> NMR (<math>\delta</math>) (<math>\text{CDCl}_3</math>)</u>	<u>Mass Spectrum<sup>a</sup>, m/e</u> <u>(Relative Intensity)</u>	<u>Colour</u>
$\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Ph}_2)$	(1) 5.05s(5H)-Cp, 7.1-7.3m(10H)-Ph	584(10), 556(30), 528(25), 500(46), 472(85), 444(100), 416(65)	light brown
$\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$ (1a)	5.27s(5H)-Cp, 7.1-7.1m-Ph( $\text{C}_2\text{Ph}_2$ , $\text{PPh}_3$ )	846(0.5), 818(5), 790(7), 762(4), 734(10), 706(11), 444(30), 262(100)	dark brown
$\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$	(2) 5.13s(5H)-Cp, 7.1-7.3m(10H)-Ph	630(3), 602(12), 574(8), 546(4), 518(30), 490(15), 462(80)	dark brown
$\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$	(3) 5.07s(5H)-Cp, 7.1-7.4m(10H)-Cp	720(5), 692(60), 664(85), 636(12), 608(100), 580(18), 552(75)	dark brown
$\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$	(4) 1.41t, 1.63t(6H)- $\text{CH}_3$ , 2.47m, 2.60m, 2.82m, 3.04m(4H)- $\text{CH}_2$ , 5.20s(5H)-Cp	488(5), 460(20), 432(20), 404(15), 376(18), 348(100), 320(42)	green

Table 2 (Cont'd)

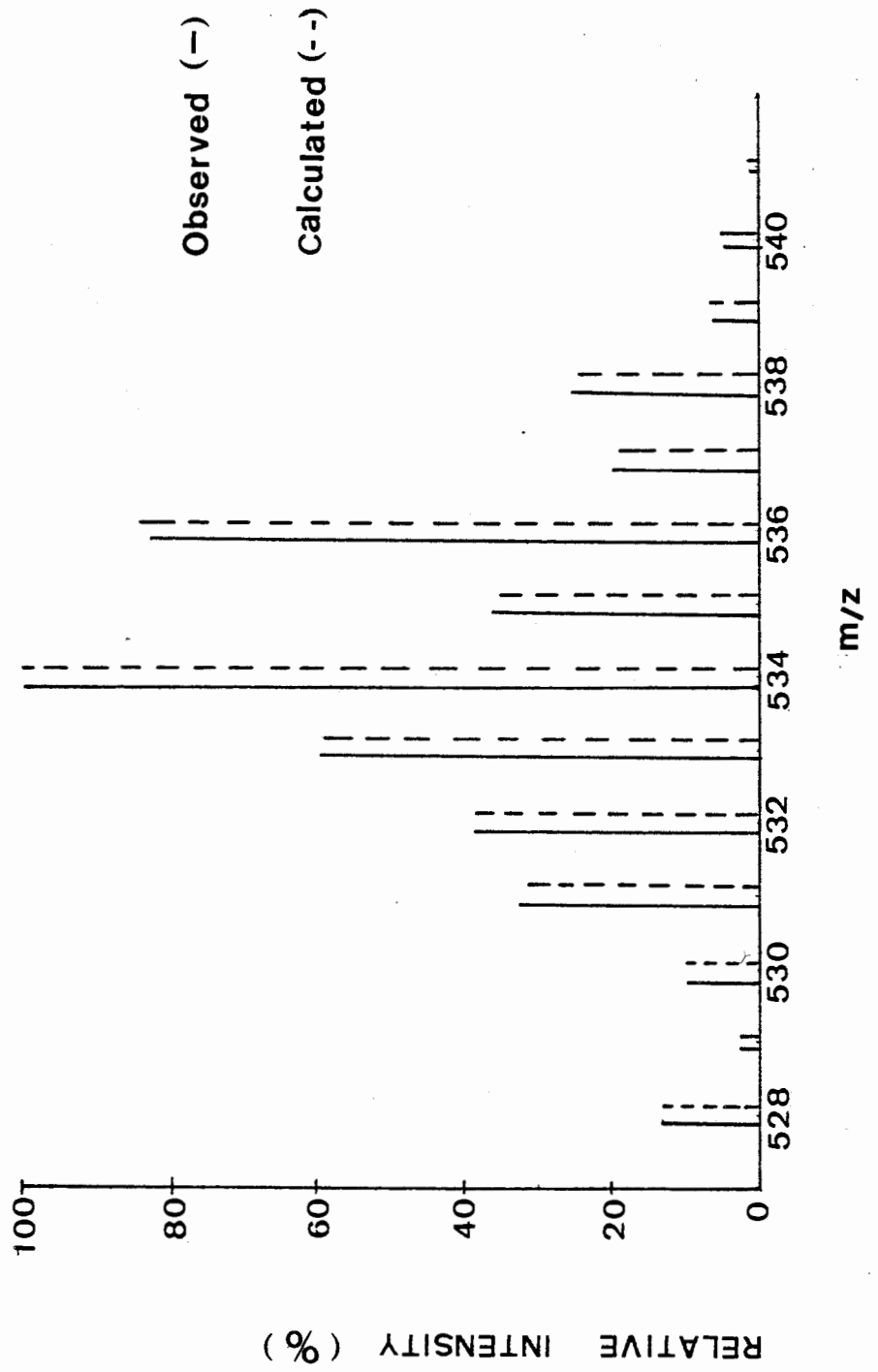
<u>Cluster</u>	<u><math>^1\text{H NMR } (\delta) \text{ (CDCl}_3\text{)}</math></u>	<u>Mass Spectrum<sup>a</sup>, m/e (Relative Intensity)</u>	<u>Colour</u>
$\text{CpNiCoRu(CO)}_6(\text{C}_2\text{Et}_2)$	(5) 1.24t, 1.37t(6H)-CH <sub>3</sub> , 2.20m, 2.50m, 2.78m(4H)-CH <sub>2</sub> , 5.25s(5H)-Cp	534(19), 506(22), 478(18), 450(40), 422(100), 394(100), 366(75)	dark brown
$\text{CpNiCoOs(CO)}_6(\text{C}_2\text{Et}_2)$	(6) 1.33t, 1.47t(6H)-CH <sub>3</sub> , 2.62m, 2.72m, 2.89m(4H)-CH <sub>2</sub> , 5.32s(5H)-Cp	624(14), 596(22), 568(26), 540(52), 512(46), 484(100), 456(65)	dark brown
$\text{CpNiCoFe(CO)}_6(\text{MeC}_2\text{Ph})$	(7) 2.62s, 2.68s(6H)-CH <sub>3</sub> , 4.94s, 5.22s(10H)-Cp, 7.1-7.3m(10H)-Ph	522(3), 494(19), 466(20), 438(18), 410(29), 382(93), 354(100)	green
$\text{CpNiCoRu(CO)}_6(\text{MeC}_2\text{Ph})$	(8) 2.31s, 2.64s(6H)-CH <sub>3</sub> , 4.90s, 5.22s(10H)-Cp, 7.1-7.4m(10H)-Ph	568(3), 540(10), 512(9), 484(9), 456(30), 428(29), 400(100)	brown
$\text{CpNiCoOs(CO)}_6(\text{MeC}_2\text{Ph})$	(9) 2.53s, 2.79s(6H)-CH <sub>3</sub> , 4.97s, 5.27s(10H)-Cp, 7.1-7.4m(10H)-Ph	658(2), 630(20), 602(20), 574(14), 546(100), 518(16), 488(85)	dark brown

<sup>a</sup> Mass position of the most intense peak in the isotopic envelope; Fe<sup>56</sup>, Co<sup>59</sup>, Ni<sup>58</sup>, Ru<sup>102</sup>, Os<sup>192</sup>.

## FIGURE 3

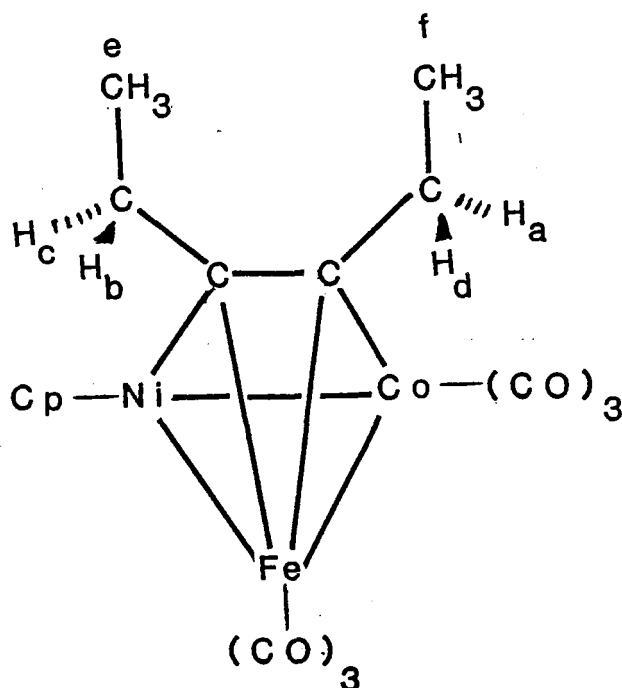
Comparison of the Observed and Calculated  
Isotopic Distribution of the Parent Ion  
of  $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Et}_2)$





NMR. One Cp resonance is observed for each compound but the methylene and methyl protons are more complex. Figures 4, 5 and 6 show the CH<sub>2</sub> and CH<sub>3</sub> region of 4, 5 and 6 respectively. In the spectrum of both 5 and 6 it is evident that two overlapping resonances contribute to the  $-\delta$  3.0 signal. The spectrum of 6 is deceptively similar to that of 5 but see later in the section. The methyl groups of 4 are triplets as expected because of coupling to the neighbouring CH<sub>2</sub> group. However, the protons on the methylene are clearly not the expected 1:3:3:1 pattern obtained from coupling with only the nearest CH<sub>3</sub> group.

From the representation of 4, CpNiCoFe(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>), below as I-10, it is clear that the cluster is chiral. The CH<sub>2</sub> protons and



methyl groups have been labelled such that H<sub>b</sub>, H<sub>c</sub> and H<sub>e</sub> are adjacent to the Ni atom and H<sub>a</sub>, H<sub>d</sub> and H<sub>f</sub> to the Co atom. Since the methylene protons are diastereotopic they should have distinct chemical shifts and the coupling

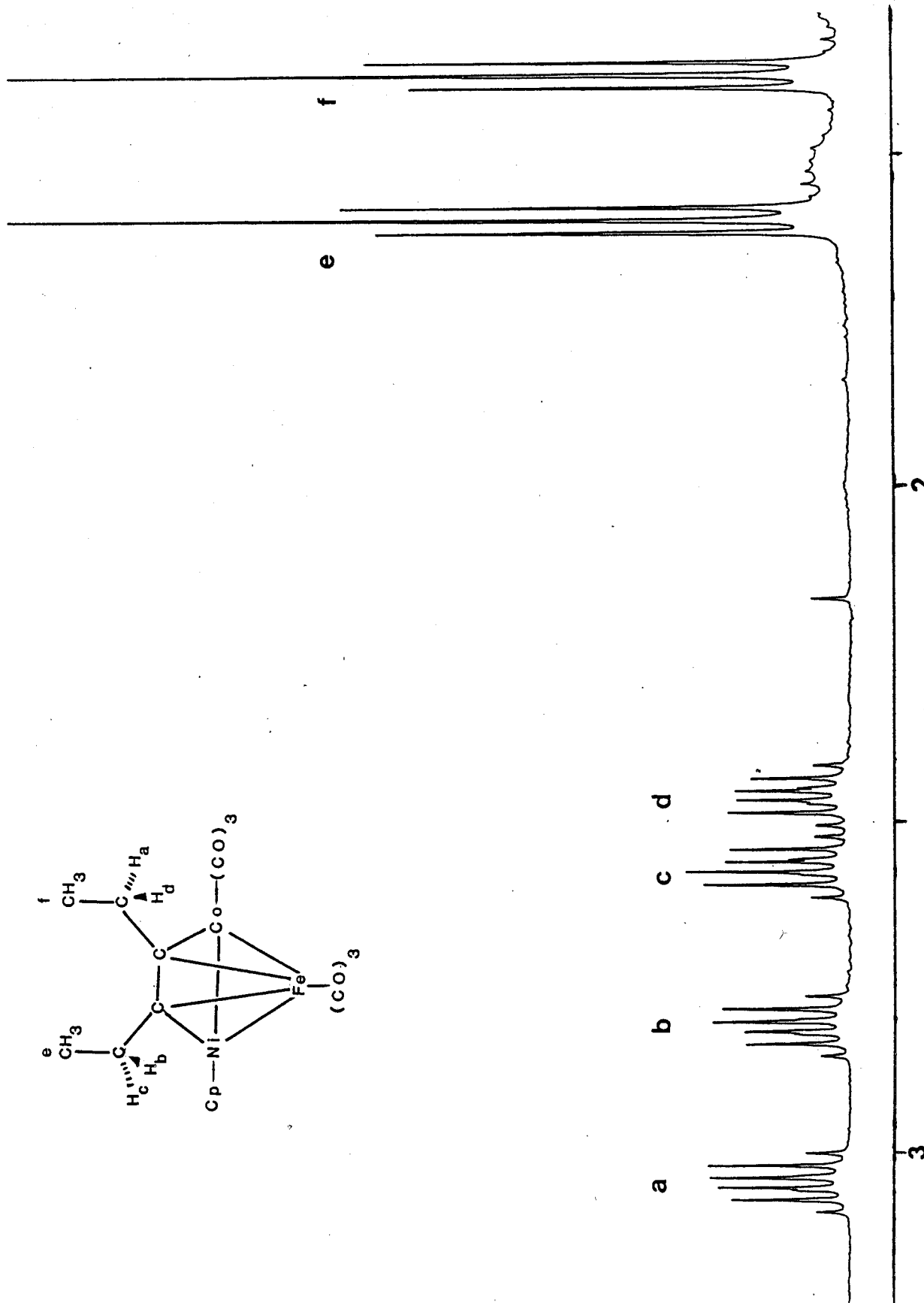
experienced by each methylene is now more complex. For example,  $H_d$  is split into a doublet by coupling with  $H_a$  and each line of the doublet is split into four lines of intensity 1:3:3:1 by the neighbouring  $CH_3$  group.

Ideally, a doublet of quartets should be observed for each methylene proton but some overlap has occurred, and the smaller peaks are sometimes obscured. Both 5 and 6, which are also chiral, have two methylene protons with nearly identical chemical shifts and show only 3 methylene multiplets with an intensity ratio of 2:1:1.

A decoupling experiment was performed to establish which methylene protons were  $\alpha$  to a particular methyl group. The methylene proton multiplets in 4, labelled  $H_{a-d}$  in Figure 4, were sequentially irradiated, and decoupling of either of the methyl groups ( $H_e$  or  $H_f$ ) was monitored. This revealed that the methylene protons  $H_a$  and  $H_d$  were  $\alpha$  to  $H_f$  and  $H_c$  and  $H_b$  were  $\alpha$  to  $H_e$ . Our labelling of the methylene protons in I-10 is consistent with this information, however we have not yet established which  $CH_2$  group is 'above' or adjacent to the CpNi fragment or which group is near the  $Co(CO)_3$  fragment. To attempt this a Nuclear Overhauser Enhancement (NOE) experiment was carried out. In this technique, one spectrum is accumulated while simultaneously saturating irradiation is applied at the transition of interest. From this spectrum is subtracted a second spectrum for which the irradiation frequency is moved to a region free from any transitions. The resultant difference spectrum shows a large negative peak corresponding to the first irradiation frequency and positive peaks for any resonances that are enhanced by the NOE. The magnitude of the indirect coupling decreases with the sixth power of the distance, so only nuclei in close proximity to the irradiated one are enhanced. Irradiation of the Cp resonance of 4 at room temperature or at 202K, unfortunately, did not result

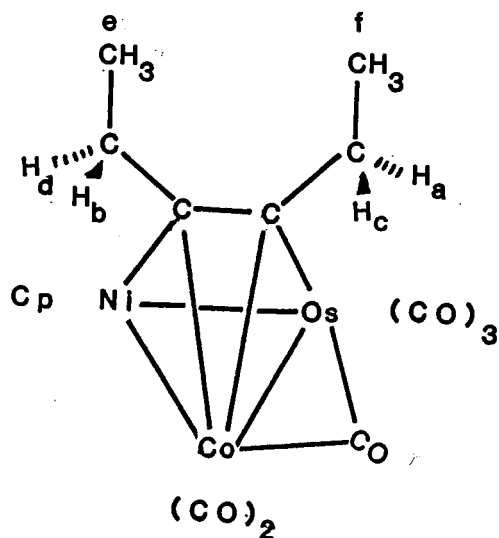
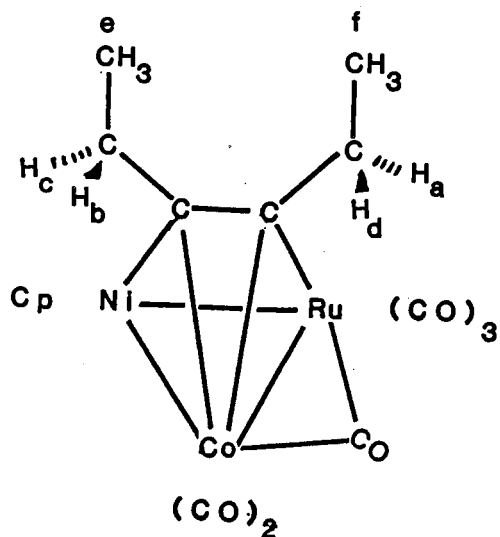
## FIGURE 4

$^1\text{H}$  NMR of the  $\text{CH}_2$  and  $\text{CH}_3$  region  
of  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (4)



in any observable enhancement of any of the  $\text{CH}_2$  or  $\text{CH}_3$  protons. The distance from the protons of the Cp ring to the  $\text{CH}_2$  or  $\text{CH}_2$  protons may be too large for this effect to be detected.

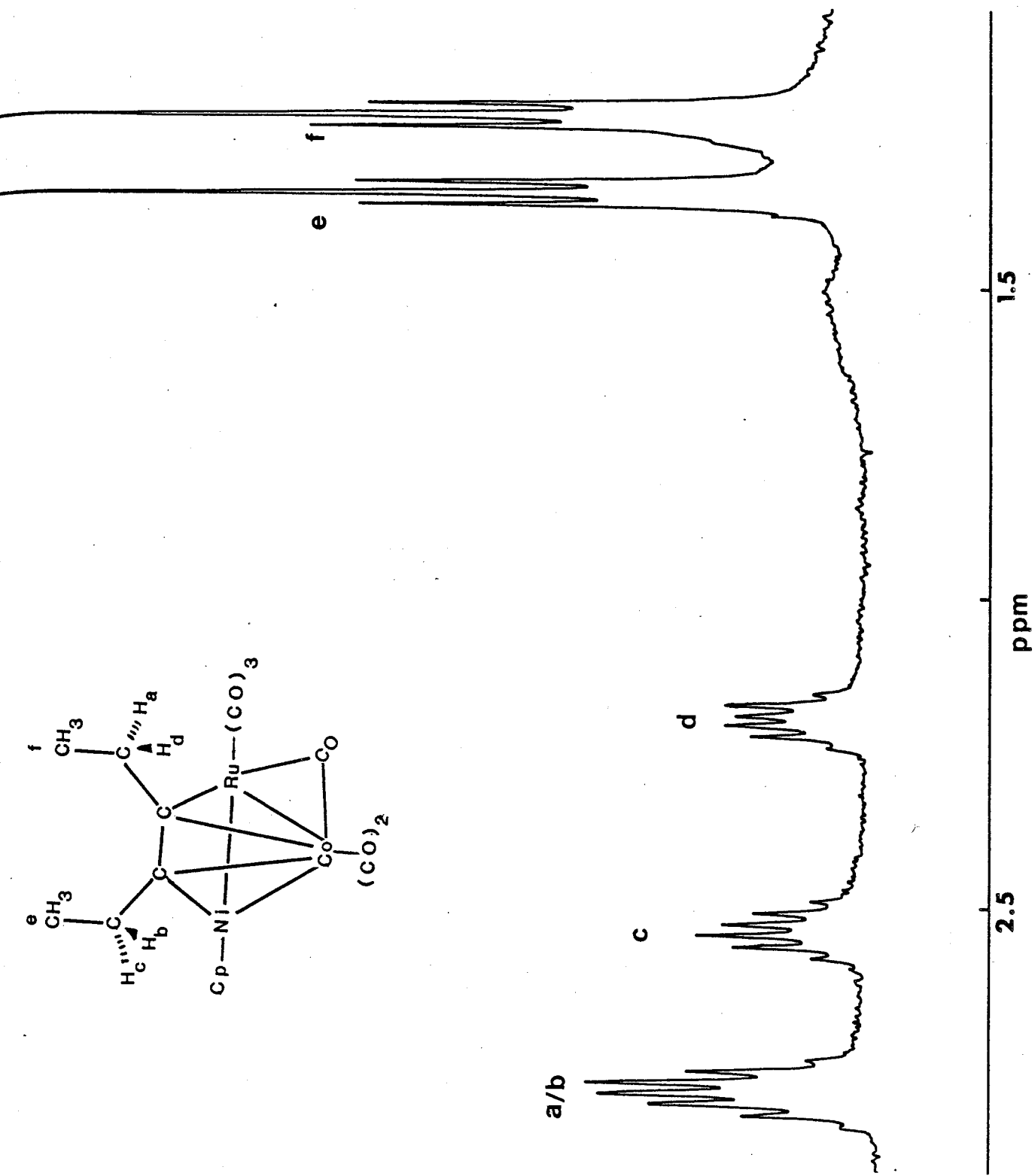
The methylene protons of 5 and 6 were also subjected to a decoupling experiment and are labelled in the NMR as shown in Figure 5. For 5, irradiation of the  $\text{H}_{a/b}$  multiplet caused both the  $\text{H}_e$  and  $\text{H}_f$  triplets to be decoupled, irradiation of  $\text{H}_c$  resulted in the decoupling of  $\text{H}_e$  and lastly irradiation of  $\text{H}_d$  decoupled the triplet  $\text{H}_f$ . This established that the methylene proton  $\text{H}_c$  and one of the two protons under the  $\text{H}_{a/b}$  envelope are  $\alpha$  to methyl group  $\text{H}_e$ , and  $\text{H}_d$  and the second proton under the  $\text{H}_{a/b}$  envelope is next to  $\text{H}_f$  as illustrated in I-11.



The decoupling experiment involving the  $\text{CH}_2$  hydrogens of 6 revealed a difference in the assignment of these peaks. The overall spectrum was very similar to 5 and the same labelling of the peaks is used in the NMR (Fig. 6). While irradiation of the  $\text{H}_{a/b}$  multiplet caused collapse of both  $\text{H}_e$  and

## FIGURE 5

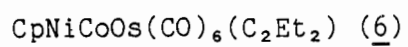
$^1\text{H}$  NMR of the  $\text{CH}_2$  and  $\text{CH}_3$  region  
of  $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (5)

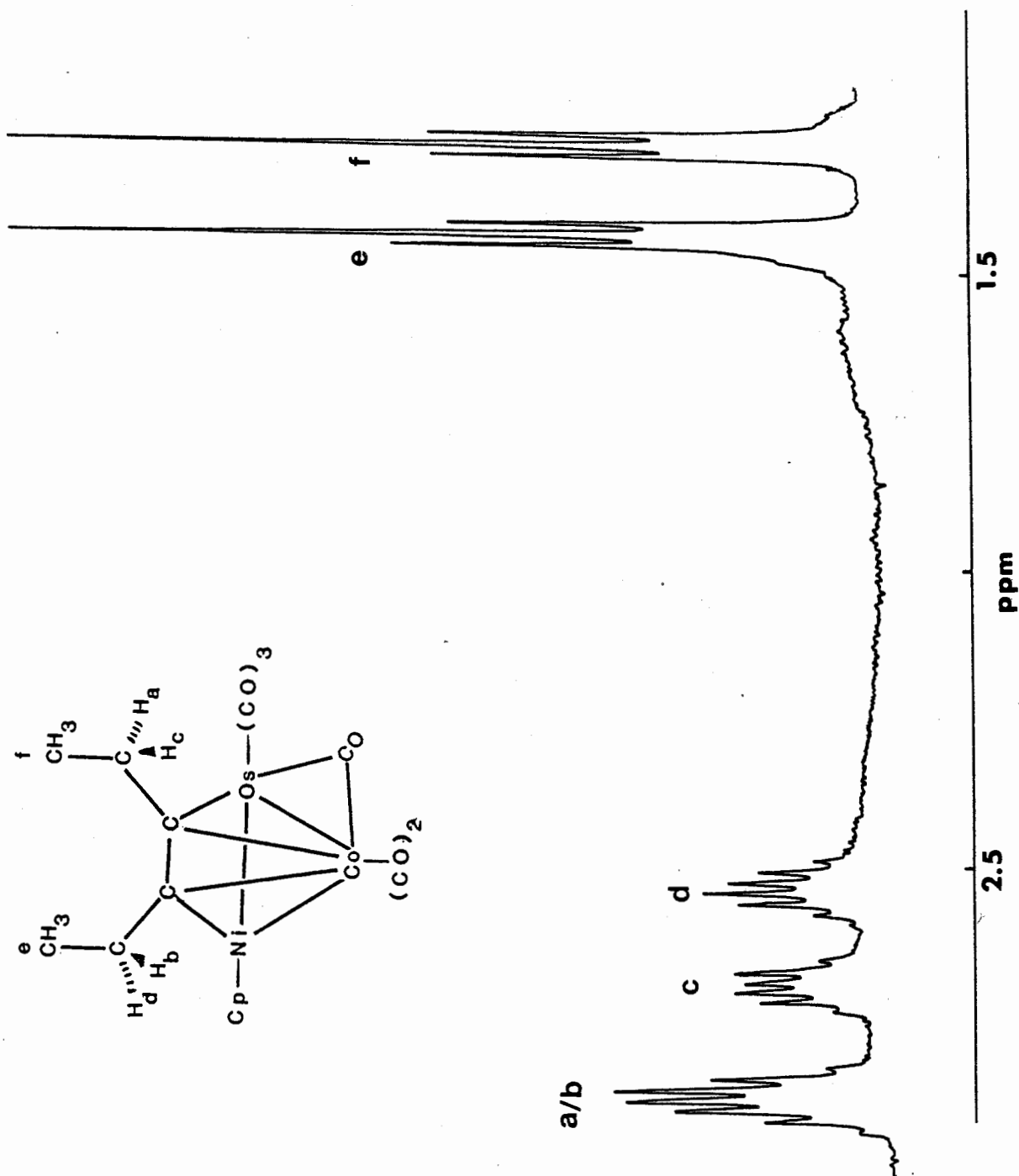




## FIGURE 6

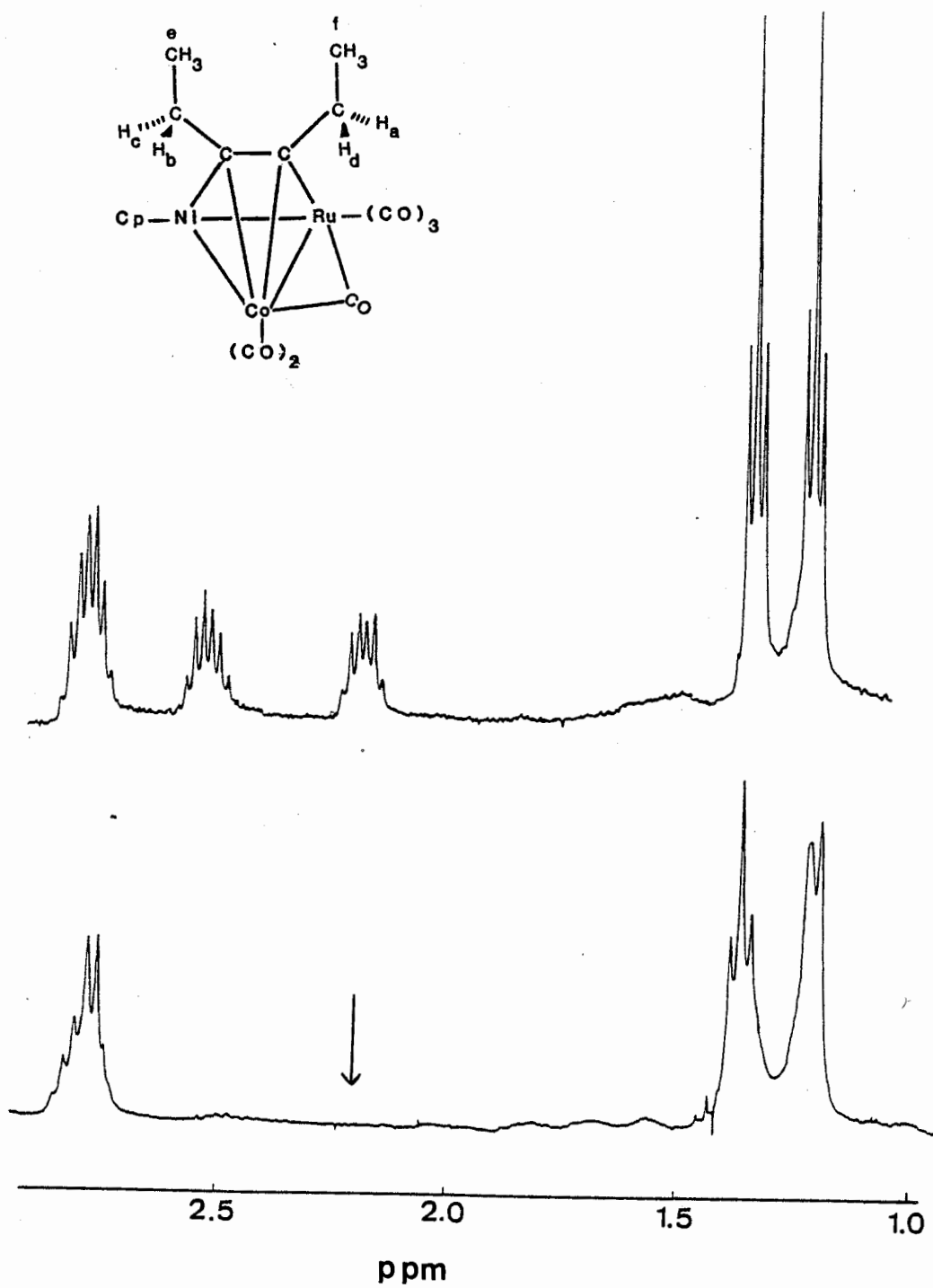
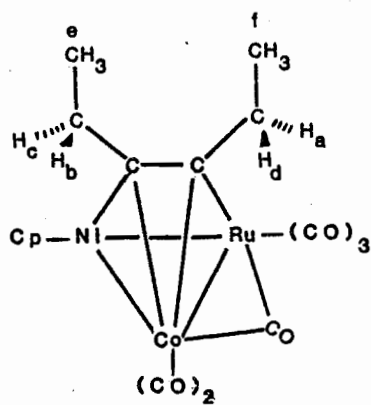
$^1\text{H}$  NMR of the  $\text{CH}_2$  and  $\text{CH}_3$  region





## FIGURE 7

Irradiation of a CH<sub>2</sub> Resonance  
in the <sup>1</sup>H NMR Spectrum of  
CpNiCoOs(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>) (5)



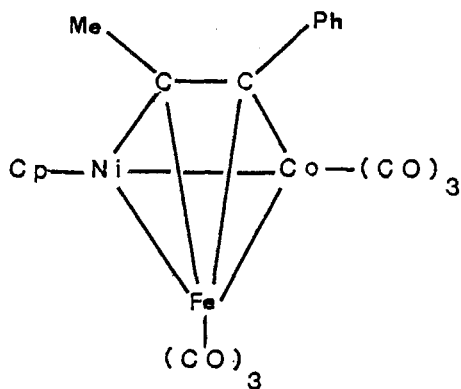
$H_f$ , irradiation of  $H_c$  collapsed  $H_f$  and irradiation of  $H_d$  collapsed  $H_e$ , which is the opposite to that encountered for 5. The decoupling experiment is consistent with the labelling scheme for 5 in I-12.

The decoupling experiments of 5 and 6 revealed further details. When methylene proton  $H_c$  of 5 was irradiated, in addition to decoupling the methyl resonance  $H_e$ , the multiplet due to proton  $H_d$  vanished. This is shown in Figure 7. The irradiation of  $H_d$  caused  $H_c$  to disappear but irradiation of the  $H_{a/b}$  envelope did not affect  $H_c$  or  $H_d$ . This is equivalent to a spin saturation labelling experiment. Saturation of the  $H_c$  resonance will cause saturation and hence disappearance of the  $H_d$  resonance only if  $H_c$  and  $H_d$  are exchanging and in addition, the exchange rate must be comparable to the relaxation rate at the two sites. In compound 6, irradiation of  $H_c$  caused  $H_d$  to vanish and vice-versa and, as in 5, irradiation of the  $H_{a/b}$  multiplet did not affect  $H_c$  or  $H_d$ .

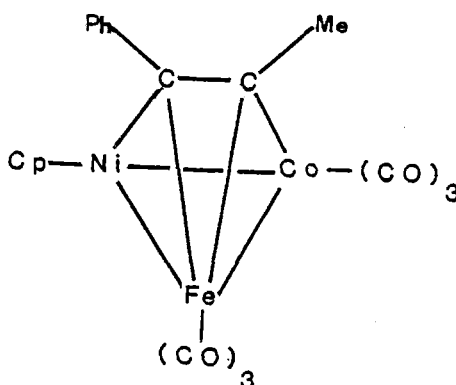
Saturation of the  $H_a$  methylene resonance in 4 did not result in transfer of the saturation to a different methylene hydrogen indicating that interchange of these hydrogens must be slower than in 5 or 6. In order to obtain a value for the activation energy of this interchange process the rate of saturation transfer between the methyl resonances in 4 at different temperatures was determined. This is the basis of the saturation transfer experiment detailed by Forsen and Hoffman<sup>49</sup> from which exchange kinetics are obtained. In their experiment one resonance is irradiated and the intensity of a second related resonance is determined as a function of time at varying temperatures. We pursued a slightly different approach from that which was suggested by Dahlquist<sup>50</sup> but one which yields the same information. A 180° pulse is applied to one resonance of interest in, for example, 4 causing the signal to be inverted. Spectra are recorded at varying delay times during

which a fraction of the saturation from the first peak is transferred to the second. Both peaks are also decaying in proportion to their lifetimes so the total intensity of the two peaks decreases with longer delay times. Measurement at five different temperatures (291, 300, 307, 314, 321K) gave an activation energy  $\Delta G^\ddagger = 74.5 \pm 1.3 \text{ kJ mol}^{-1}$ . An example of the calculation is provided in the Experimental Section. Unfortunately, for 5 and 6, there was enough overlap in the baseline of the methyl and methylene resonances to make it difficult to obtain accurate integrations because of some saturation transfer from one methyl group to the other even with very short delay times.

The  $^1\text{H}$  NMR of compounds 7-9 were relatively simple consisting of a multiplet for the aromatic protons, two Cp peaks and two singlets assignable to the methyl protons of the alkyne ligand. The  $^1\text{H}$  NMR of 7 is shown in Figure 8. The two sets of Cp, Me and Ph resonances occur because the alkyne ligand can be oriented on the heterotrimetallic framework in two different fashions, as shown below for 7 and labelled as I-13a and I-13b. The methyl group of the alkyne ligand can be adjacent to either the Ni or the Co atom



I-13a

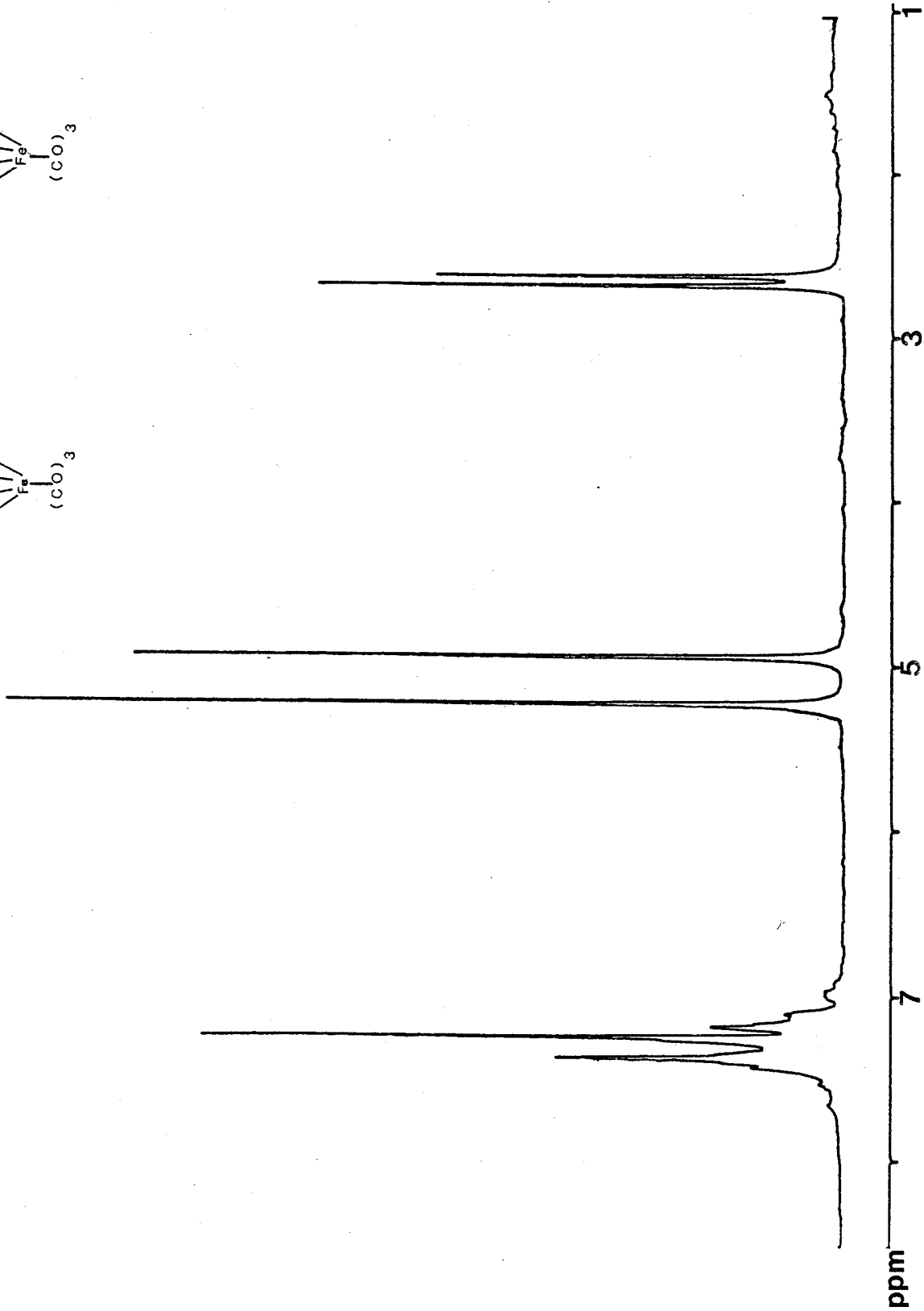
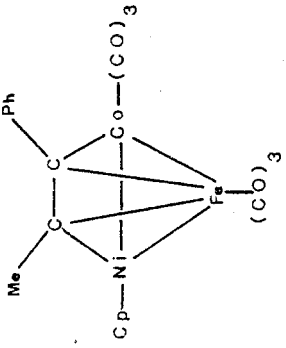
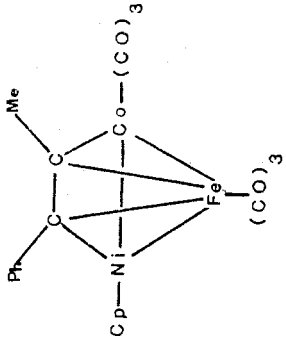


I-13b

and similarly for the phenyl ligand. Both 8 and 9 will also have two stereoisomers and the  $^1\text{H}$  NMR substantiates this.

FIGURE 8

 $^1\text{H}$  NMR of  $\text{CpNiCoFe}(\text{CO})_6(\text{MeC}_2\text{Ph})$  (7)





The ratio of the intensities of the Cp peaks (and the methyls) are very solvent dependent. This ratio and the chemical shifts are listed in Table 3 for three different solvents. From the varying intensity ratios it is easily established that the Cp and Me peak of lower  $\delta$  belong to the same isomer and the high  $\delta$  resonances belong to the second isomer, i.e., either I-13a or I-13b. An identical rationale can be applied to 8 and 9. To attempt to establish which set of resonances are attributable to a particular isomer a NOE experiment was carried out. The Cp resonances of 7, 8 and 9 were sequentially irradiated and the phenyl and methyl peaks of the spectrum were monitored to observe any enhancement of these hydrogens. An enhancement of the methyl hydrogens might be expected if the Me group is adjacent to the Ni atom putting the group in proximity of the Cp ring, i.e., I-13a. Unfortunately, no detectable enhancement could be observed (at 294 or 202K) for 7, 8 or 9 and consequently assignment of the isomers could not be proven.

In addition to a difference in the relative intensities of the two Cp and two Me resonances of 7, 8 and 9 there is a sizeable chemical shift when the solvent is changed. As the polarity of the solvent decreases, acetone- $d_6$  to  $CDCl_3$ , to toluene- $d_8$ , (Table 3) the Cp and Me resonances of 7, 8 and 9 shift quite significantly to higher field. The shifts are not equal, however. For example, while the Cp protons of 9 are separated by 0.35 ppm in acetone- $d_6$ , in toluene- $d_8$  the difference is only 0.09 ppm. Clearly, one of the stereoisomers of 7, 8 and 9 is being affected by the change in polarity to a greater extent than the other. This problem will be examined further in the Discussion Section.

From the proportions of the two isomers at room and lower temperatures the ground state energy difference,  $\Delta E$ , between the two stereoisomers of 7-

Table 3. Chemical Shifts and Intensity Ratios for the  
Cp and Me Resonances of 7, 8 and 9

Cluster	Solvent			
	CDCl <sub>3</sub>	Me <sub>2</sub> CO-d <sub>6</sub>	C <sub>6</sub> D <sub>5</sub> -CD <sub>3</sub>	
	Chemical Shift (Ratio)	Chemical Shift (Ratio)	Chemical Shift (Ratio)	
CpNiCoFe(CO) <sub>6</sub> (MeC <sub>2</sub> Ph) ( <u>7</u> )	Cp	5.22, 4.94 (1.24:1)	5.44, 5.11 (2.19:1)	4.78, 4.61 (1.47:1)
	Me	2.68, 2.62 (1.28:1)	2.76, 2.70 (2.20:1)	2.48, 2.41 (0.78:1)
CpNiCoRu(CO) <sub>6</sub> (MeC <sub>2</sub> Ph) ( <u>8</u> )	Cp	5.22, 4.90 (0.22:1)	5.45, 5.10 (0.42:1)	4.88, 4.66 (0.33:1)
	Me	2.64, 2.31 (0.19:1)	2.74, 2.43 (0.43:1)	2.39, 2.24 (0.28:1)
CpNiCoOs(CO) <sub>6</sub> (MeC <sub>2</sub> Ph) ( <u>9</u> )	Cp	5.27, 4.97 (0.43:1)	5.54, 5.19 (0.84:1)	4.93, 4.73 (0.63:1)
	Me	2.79, 2.53 (0.36:1)	2.93, 2.67 (0.76:1)	2.56, 2.47 (0.71:1)

Table 4. Ground State Energy Difference,  $\Delta E$ ,  
Between the Isomers of 7, 8 and 9

Cluster	CDCl <sub>3</sub>	$\Delta E$ kJ (mol <sup>-1</sup> ) Me <sub>2</sub> CO-d <sub>6</sub>	Toluene-d <sub>8</sub>
<u>7</u> , Cp	0.527	1.92	0.946
Me	0.607	1.95	0.611
average	0.567	1.93	0.778
<u>8</u> , Cp	3.72	2.13	2.72
Me	4.08	2.07	3.12
average	3.90	2.10	2.92
<u>9</u> , Cp	2.07	0.427	1.13
Me	2.51	0.640	0.840
average	2.28	0.536	0.987

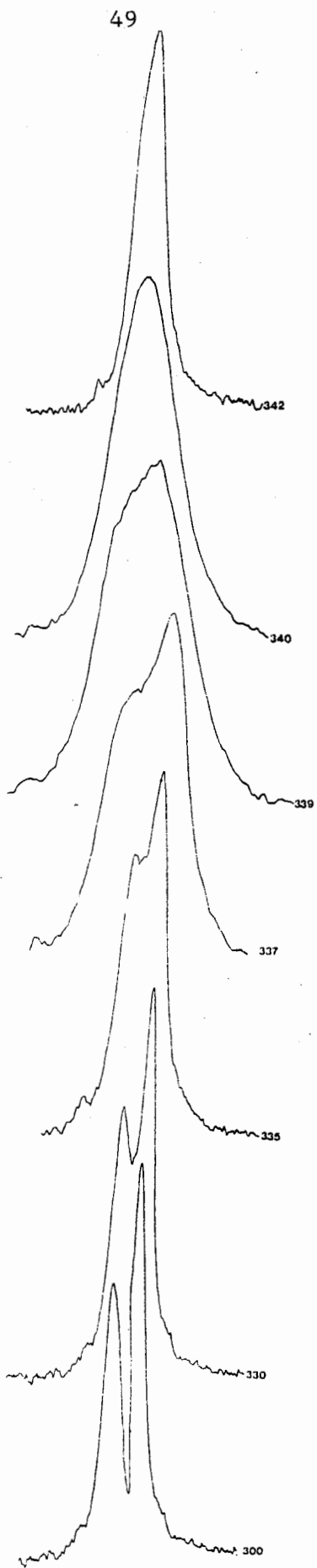
9 can be calculated. These  $\Delta E$  values for the clusters are listed in Table 4 for both the Cp and Me resonances and a sample calculation is included in the Experimental Section.

The activation energy  $\Delta G^\ddagger$ , was obtained for 7, 8 and 9 from variable temperature coalescence experiments on both the methyl and cyclopentadienyl resonances by using the Eyring equation, the coalescence temperature ( $T_c$ ), the rate constant ( $k_c$ ) and the separation of the two peaks at 243K,  $\Delta\nu_{243}$ . An example calculation is included in the Experimental Section. The collapse of the methyl peaks of 7 are shown in Figure 9. The coalescence temperature for peaks of unequal intensity was approximated by the method of Fraser<sup>51</sup>. The rate constant  $K_a$ , represents the rate constant for conversion of the isomer that resonates at higher  $\delta$  into the isomer which has the lower  $\delta$  resonances and  $k_b$  represents the rate constant for the reverse reaction. All these calculated values are summarized in Table 5.

Compound 7, (NiCoFe core), has the largest activation energy by approximately 4 kJ mol<sup>-1</sup> over 8 and 9 which are both very similar. In the C<sub>2</sub>Et<sub>2</sub> clusters, 4-6, the NiCoFe cluster also had the highest  $\delta G^\ddagger$  value since the methylene hydrogens did not exchange on the timescale of the saturation experiments but those of 8 and 9 did. However, this observation could not be quantified.

## FIGURE 9

Coalescence of the Methyl Resonances  
in the  $^1\text{H}$  NMR (Toluene- $d_8$ ) of  
 $\text{CpNiCoFe(CO)}_6(\text{MeC}_2\text{Ph})$  (7)



TEMPERATURE

2.8 2.5 PPM

Table 5. Coalescence Data for 7, 8 and 9

Cluster	$T_c$ (K) <sup>a</sup>	$\Delta_{2+3}K$	$\Delta G^\ddagger$ (kJ mol <sup>-1</sup> ) <sup>b</sup>	$K_{AB}$ (s <sup>-1</sup> ) <sup>c</sup>	$K_{BA}$ (s <sup>-1</sup> ) <sup>c</sup>
<u>7</u> , Me	339	11.18	74.3	30.8	18.9
<u>8</u> , Me	321	13.97	69.6	47.8	14.3
Cp	330	20.57	70.5	70.4	21.0
average			70.1		
<u>9</u> , Me	315	7.15	70.0	19.1	12.7
Cp	327	18.58	70.1	47.8	34.7
average			70.1		

<sup>a</sup> All experiments in toluene-d<sub>8</sub>;  $\pm 1$  K.

<sup>b</sup>  $\pm 1.3$  kJ mol<sup>-1</sup>

<sup>c</sup>  $\pm 1.1$  s<sup>-1</sup>

### 1.3.2d X-ray Structures

The X-ray structures were determined for four of the heterotrimetallic alkyne clusters, specifically 1a, 2, 3 and 4. The X-ray data collection parameters and refinement are in the Experimental Section along with the tables of bond lengths and angles, atom coordinates and temperature factors but a compilation of the more important bond lengths and angles are in Tables 6 and 7 for easy comparison. The labelled ORTEP and SNOOPI diagrams for all four cluster alkynes are shown in Figures 10, 11, 12 and 13 respectively.

All four crystal structures consist of discrete molecules with no unusually short inter- or intramolecular contacts. All the compounds have a bonded triangle of metal atoms (NiCoFe, NiCoRu or NiCoOs) which are capped by a diethyl or diphenylacetylene ligand bonded as a  $\mu_3-(\eta^2-|)$  group. A cyclopentadienyl ligand is  $\eta^5$  coordinated to the Ni atom in each structure. 1a, 2 and 3 have a semi-bridging carbonyl between the Co atom and the Fe, Ru or Os atom while 4 has only terminal carbonyl groups. In addition, 1a has a triphenylphosphine ligand bonded to the cobalt.

Following is a comparison (with some discussion) of the four clusters with respect to the metal triangle, alkyne and carbonyl coordination. A general overview of the structures with regard to other known structures and to the IR and NMR data is given in the Discussion Section.

#### Metal Triangles

As shown in Table 6, the Ni-M and Co-M bond length increases in the order 1a < 2 < 3, or Fe < Ru < Os, as expected although the difference between the Ru-M and Os-M lengths is small (0.019(2)Å). The Ni-Co distances



FIGURE 10

ORTEP Diagram of  
 $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$  (1a)

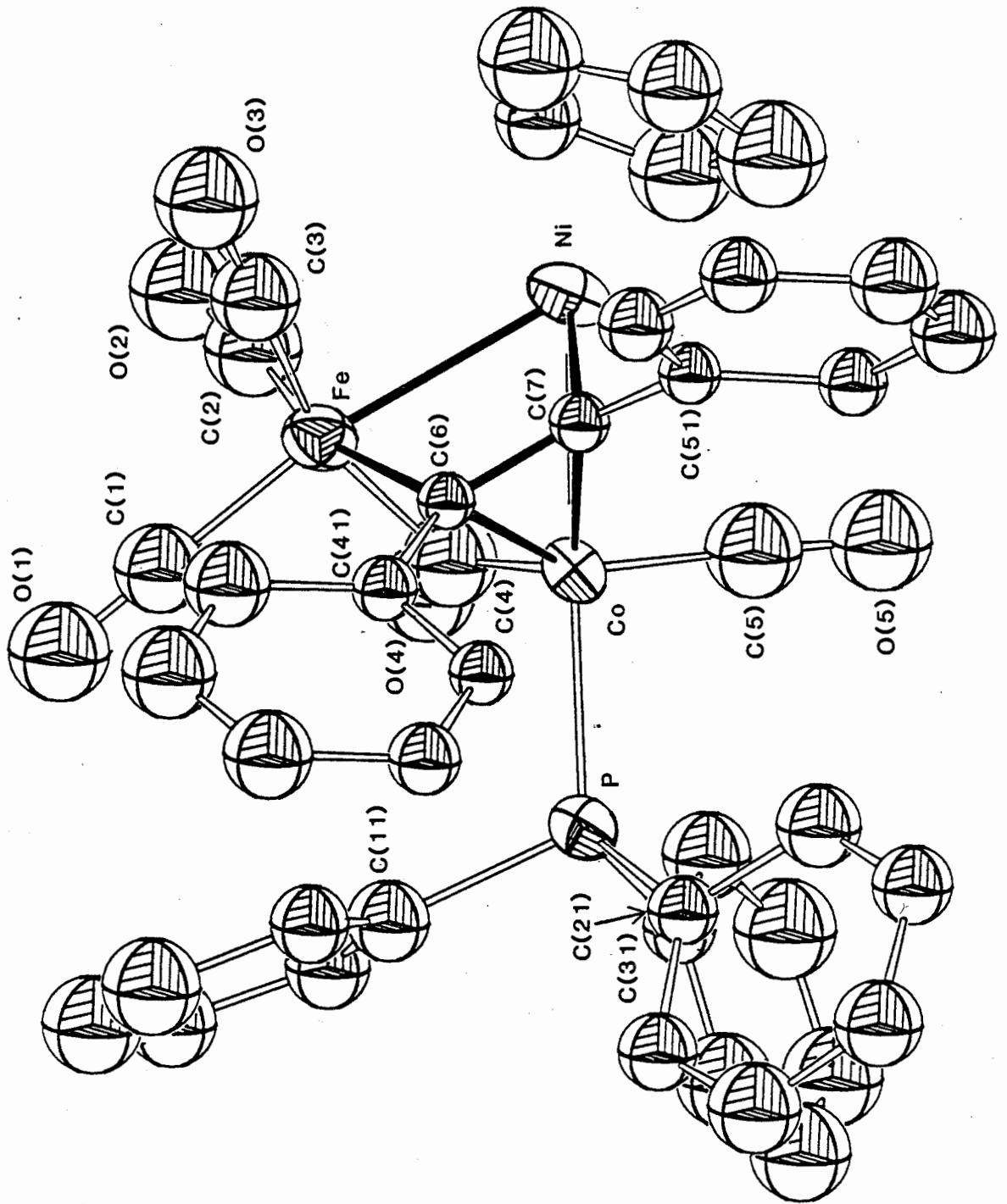


FIGURE 11

ORTEP Diagram of  
 $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (2)

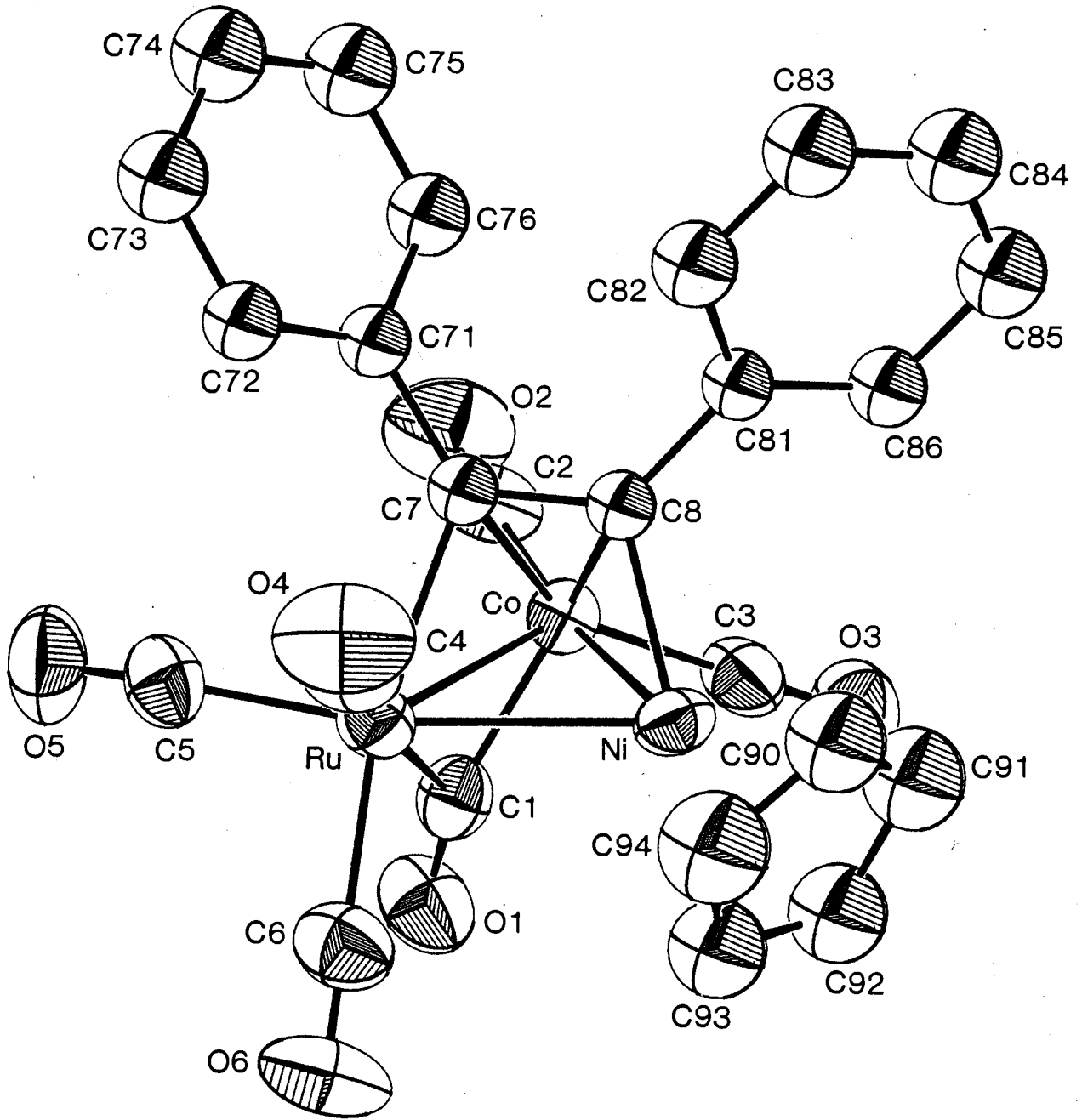


FIGURE 12

SNOOPI Diagram of  
 $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (3)

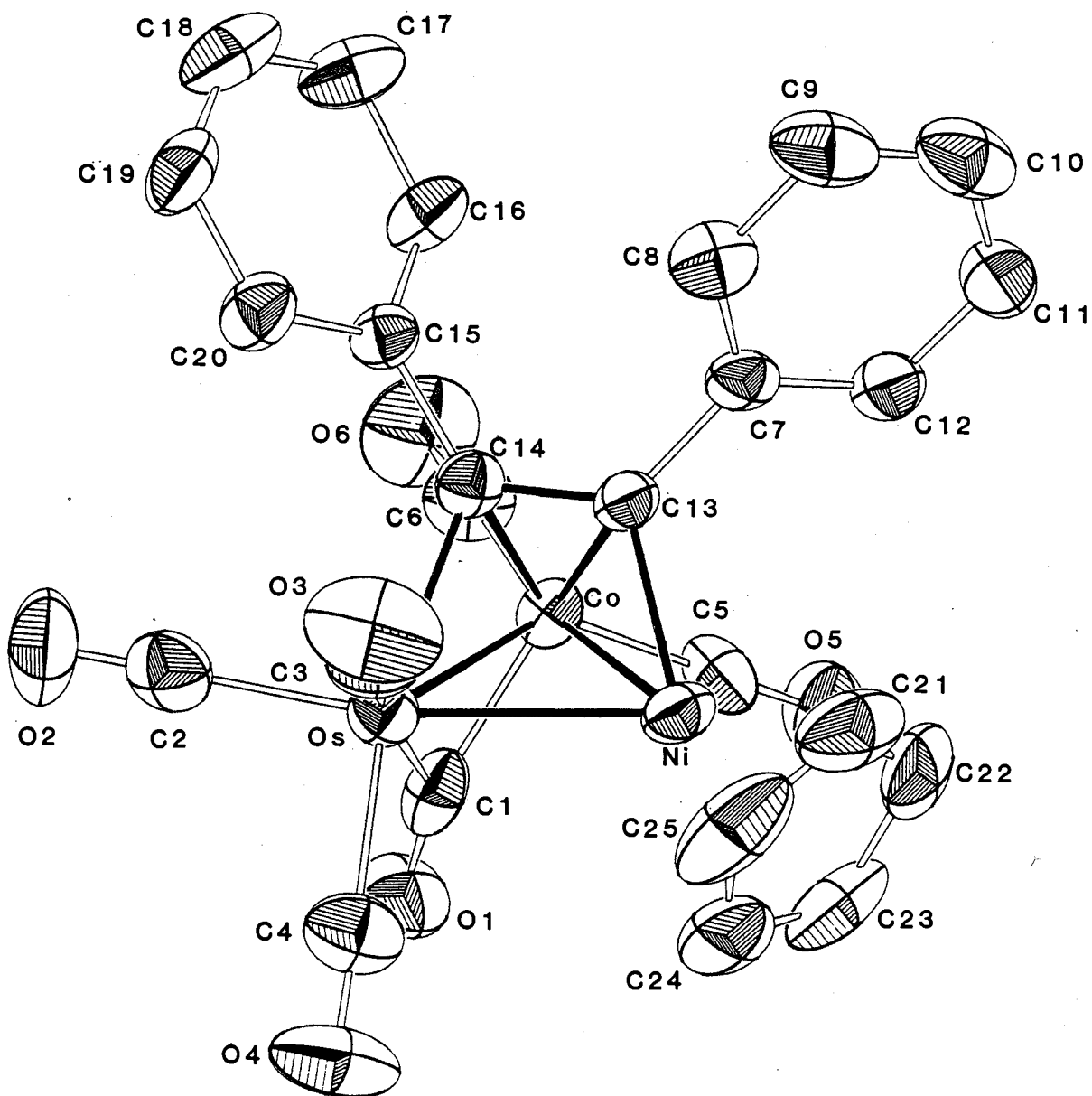
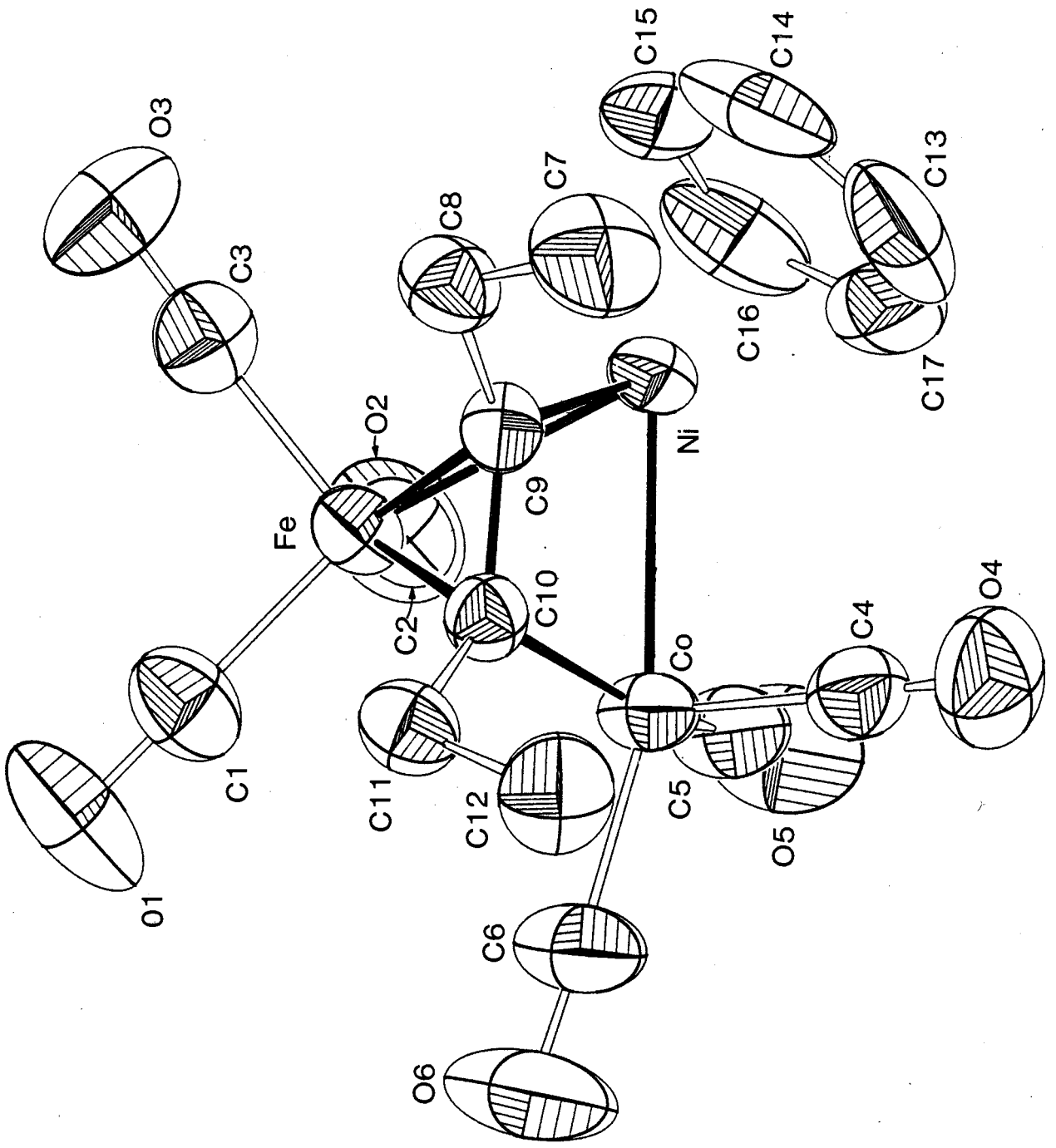


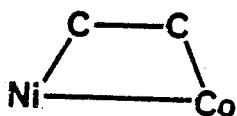
FIGURE 13

SNOOPI Diagram of  
 $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (4).

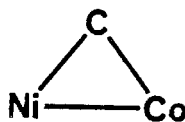




are very similar in 1a, 2 and 3 but they are all substantially shorter than in 4. While this may be the result of a different bridging alkyne ligand,  $C_2Et_2$  in 4 versus  $C_2Ph_2$  in 1a, 2 and 3, it is more likely a consequence of the different alkyne coordination with respect to the Ni-Co bond. In 1a-3 the alkyne is coordinated with respect to the Ni-Co bond as in I-14a but for 4 the orientation of the alkyne is represented by I-14b. Churchill<sup>32</sup> has



I-14a



I-14b

previously noted a difference in metal-metal bond lengths due to this dissimilarity in alkyne coordination, although our difference of 0.033-0.042(5)Å is considerably smaller than recorded (0.12-0.14(3)Å) by Churchill for  $Cp_2W_2Os(CO)_7[C_2(C_6H_4Me)_2]$ . This difference in alkyne coordination with respect to a particular metal-metal bond also explains the large difference in distances for the Ni-Fe bond lengths of 1a and 4 (2.486(4) versus 2.402(1)Å).

In 1a, a triphenylphosphine ligand has been substituted for one of the carbonyl groups on the Co atom. The Co-P bond is essentially colinear (177(1)°) with the Co-Ni bond. In all the other structures it is noted that there is a carbonyl on cobalt in a similar orientation and the other terminal CO groups are in similarly related positions in all structures. The presence of the  $PPh_3$  ligand has not disturbed the orientation of the remaining carbonyls, nor does there appear to be any effect on the Ni-Co bond length relative to 2 and 3.

Table 6. Comparison of Bond Lengths and Angles for 1a, 2, 3 and 4

Cluster	Bond Lengths (Å)				Bond Angles (deg)		
	<u>Ni-Co</u>	<u>Ni-M</u>	<u>Co-M</u>	<u>C≡C</u>	<u>Ni-Co-M</u>	<u>Ni-M-Co</u>	<u>M-Ni-Co</u>
CpNiCoFe(CO) <sub>5</sub> (PPh <sub>3</sub> )(C <sub>2</sub> Ph <sub>2</sub> ) ( <u>1a</u> )	2.390(4)	2.486(4)	2.467(4)	1.34(2)	61.6(1)	57.70(9)	60.7(1)
CpNiCoRu(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> ) ( <u>2</u> )	2.399(1)	2.560(1)	2.565(1)	1.377(9)	61.98(4)	55.82(4)	62.21(4)
CpNiCoOs(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> ) ( <u>3</u> )	2.394(1)	2.579(1)	1.570(1)	1.377(10)	62.49(4)	55.41(3)	62.10(4)
CpNiCoFe(CO) <sub>6</sub> (C <sub>2</sub> Et <sub>2</sub> ) ( <u>4</u> )	2.432(1)	2.402(1)	2.494(1)	1.371(7)	58.36(3)	59.52(3)	62.11(3)

The cyclopentadienyl ring is displaced on the same side of the NiCoM plane in all four structures. The dihedral angle between the cyclopentadienyl plane and the trimetallic plane is 71.8(5), 69.6(3), 70.3(3) and 68.8(4)° in 1a, 2, 3 and 4 respectively. This and the similar Ni-Cp(center) distance of 1.75(1), 1.747(1), 1.755(1) and 1.750(4)Å reveal a negligible effect of the different alkyne bonding modes or of different metals in the cluster on the Ni-Cp distance or on its disposition relative to the metal triangle.

### Alkyne Coordination

The interaction of the alkyne ligand with the three metal atoms occurs through a deformation of the alkyne fragment to an olefinic-like coordinating group forming two  $\sigma$ -bonds and one  $\pi$ -bond to the metal atoms. The formal 'triple-bond' of the alkyne has been found to range between 1.34(2)-1.377(10)Å in all 4 clusters, and is similar to the distance found for a double bond in an olefin. Similar lengthening has been noted in all (homo- and heterometallic) the  $\mu_3-(\eta^2-||)$  bonded clusters, and examination of our data and the lengths provided by Sappa<sup>24</sup> does not provide any rationalization of the degree of lengthening with respect to the different metals present, even when the comparison is restricted to diphenylacetylene complexes. For example,  $\text{Os}_3(\text{CO})_{10}(\text{C}_2\text{Ph}_2)$  was found to have a C≡C distance of 1.439(10)Å, that is the longest 'triple bond' distance found, but one of the shortest is also a triosmium cluster,  $\text{Os}_3(\text{CO})_7(\text{C}_2\text{Ph}_2)_2$ , having a C≡C distance of 1.33Å. It appears that factors more severe than just changing one metal atom as we have in 1a-3, are required to substantially change the length of the olefinic-type bond.

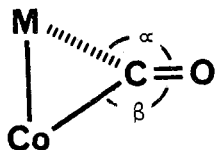
The alkyne has been found always to be symmetrically  $\pi$ -bonded to the

appropriate metal atom although the Fe-alkyne distances in 4 are slightly shorter than the Co-alkyne lengths found in 2 and 3. A comparison with 1a is inconclusive because of the large standard errors present in the bond lengths.

Although the alkyne is  $\pi$ -bonded to the Fe atom in 4 and to the cobalt in 1a, 2 and 3, the similarity of the alkyne bonding to the metal triangle is most revealing when the dihedral angle between the trimetal plane and the plane formed by the NiRuC(6)C(7) atoms in 2 and equivalent planes in 1a, 3 and 4 is considered. This dihedral angle is 57.8(3), 57.6(1), 57.3(1) and 57.1(1) $^\circ$  in 1a, 2, 3 and 4, respectively. Dahl<sup>31</sup> has recorded similar dihedral angles of 56.5(2) and 56.0(2) $^\circ$  for  $\text{Cp}_3\text{Rh}_3(\text{CO})(\text{C}_2\text{Ph}_2)$  and  $\text{Cp}_3\text{Rh}_3(\text{CO})(\text{C}_2(\text{C}_6\text{F}_5)_2)$ , compounds with only second row transition metals. The latter has a powerful electron withdrawing fluoro-alkyne ligand. This nearly constant dihedral or 'inclination' angle displayed for varying structures is probably a consequence of the alkyne ligand finding the optimum geometry for  $\pi$ -donation to, and  $\pi^*$ -acceptance from, the corresponding metal atom (see Discussion).

### Carbonyl Coordination

One of the more notable differences in these four structures is the presence (or absence) of a semi-bridging carbonyl ligand. Only compound 4 does not have this type of ligand. The relevant bond lengths and angles of the bridging species are summarized in Table 7 along with a recently reported and closely related structure,  $\text{Co}_2\text{Ru}(\text{CO})_9(\text{C}_2\text{Ph}_2)$ <sup>48</sup> which also has a semi-bridging carbonyl. The two angles in the table,  $\alpha$  and  $\beta$ , are defined as below as in I-15 and are consistent with those described by Cotton and



## I-15

Wilkinson<sup>52</sup>.

The M-C bond length increases for Fe < Ru < Os within our series of clusters 1a-3 but the Co<sub>2</sub>Ru cluster<sup>48</sup> has a substantially longer Ru-C distance. This longer M-C length would suggest that in this compound the carbonyl ligand is not as distorted from linearity as is the bridging carbonyl in 1a, 2 or 3. The listed  $\alpha$  angles substantiate this observation;  $\alpha$  is smaller for 2 and 3 than for the Co<sub>2</sub>Ru cluster (but comparison with 1a is of limited value because of the larger standard errors).

The Co-C bond lengths also increase as Fe < Ru < Os and both 2 and the Co<sub>2</sub>Ru cluster have similar values suggesting the difference noted for the Ru-C bond length is due to the replacement of the Co(CO)<sub>3</sub> group by the isoelectronic and isolobal CpNi unit. However, differences in M-C or any bond length associated with the semi-bridging carbonyl ligand are likely to involve very small energy changes comparable to other factors such as packing forces.

Table 7. Comparison of Bond Lengths and Angles for the Semi-bridging Carbonyl

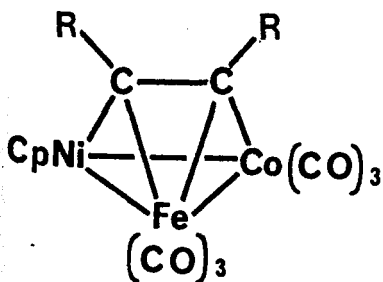
<u>Cluster</u>	<u>Bond Lengths (Å)</u>			<u>Bond Angles (deg)</u>	
	<u>Co-M</u>	<u>Co-C</u>	<u>M-C</u>	<u>α</u>	<u>β</u>
CpNiCoFe(CO) <sub>5</sub> (PPh <sub>3</sub> )(C <sub>2</sub> Ph <sub>2</sub> ) ( <u>1a</u> )	2.467(4)	1.74(2)	2.07(1)	148(1)	131(1)
CpNiCoRu(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> ) ( <u>2</u> )	2.565(1)	1.828(8)	2.251(8)	148.2(7)	134.5(6)
CpNiCoOs(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> ) ( <u>3</u> )	2.579(1)	1.853(8)	2.273(8)	147.3(7)	136.2(7)
Co <sub>2</sub> Ru(CO) <sub>9</sub> (C <sub>2</sub> Ph <sub>2</sub> ) <sup>48</sup>	2.5865(5) 2.6875(5)	1.820(3)	2.284(3)	149.9(1)	132.8(1)

## 1.4 Discussion

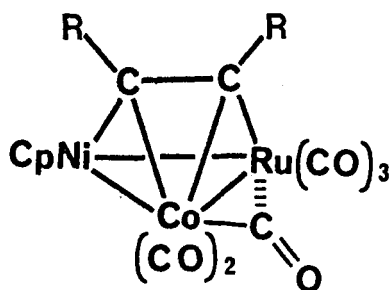
The synthesis of heterotrimetallic (MM'M") alkyne clusters of the formula  $\text{CpNiCoM}(\text{CO})_6(\text{RC}_2\text{R}')$  ( $\text{M} = \text{Fe}, \text{Ru}, \text{Os}; \text{R} = \text{R}' = \text{Ph}, \text{Et}; \text{R} = \text{Ph}, \text{R}' = \text{Me}$ ), 1-9 has been achieved. We are not aware of any reports on similar series where one metal of the trimetallic framework has been systematically varied down a given triad, and this has been accomplished for three different alkyne ligands. The X-ray structures of four compounds 1a, 2, 3 and 4 have been completed and these solid state structures will be discussed next.

### 1.4.1 Solid State Structures

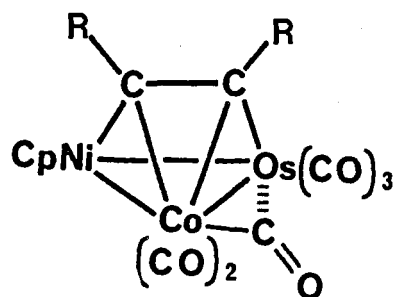
All of the clusters which have a Ru or Os atom in the trimetallic framework reveal the presence of a doubly bridging CO in the solid state IR, which is confirmed by the X-ray structures of 2 and 3. Where this semi-bridging CO is present, the alkyne ligand is found to  $\pi$ -bond with the Co atom and  $\sigma$ -bond with the Ni and Ru or Os atom. This structure is represented by I-17. Compounds where the Ru or Os has been replaced by an Fe atom all have their structures as shown by I-16 where the alkyne ligand



I-16



I-17

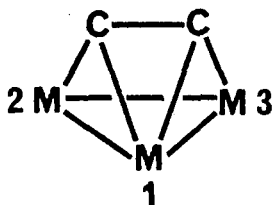


is  $\pi$ -bonded to the Fe atom and  $\sigma$ -bonded to the Ni and Co atoms. This is the arrangement found in the X-ray structure of 4. Only absorptions due to terminal carbonyl ligands are observed in both solid state and solution IR. The phosphine-substituted complex 1a, is an exception since only Ni, Co and Fe atoms are present but a bridging CO is found and this compound will be discussed later.

Since all the NiCoRu and NiCoOs clusters have the alkyne oriented as in I-17 regardless of the substituents on the alkyne (Ph, Me, Et), this suggests that minor changes in these substituents does not drastically alter the cluster-alkyne interaction. However, changing the Ru or Os for an Fe atom appears to affect the ground state structure and the alkyne has been rotated  $60^\circ$  on the trimetallic face.

Very recently McGlinchey et. al.,<sup>53</sup> have postulated that there should be a preference for certain orientations of an alkyne ligand in heterotrimetallic ( $M_2M'$  and  $MM'M''$ ) clusters. However, experimental evidence for this was largely based upon observed structures that are generally of the  $M_2M'$  type. There has not been a systematic variation of metal atoms within the trimetallic framework until this report.

Some of the relevant arguments and conclusions put forward by McGlinchey et. al. will be briefly reviewed here but the reader is referred to the original paper<sup>53</sup> for more detailed calculations. The following discussion will revolve around the heterotrimetallic core which is labelled as shown below, as I-18 with the orientation of the alkyne as indicated.



I-18



The alkyne ligand was shown to behave as an overall donor of electron density towards the trimetallic framework through its  $\pi$  orbital. Thus the ability of the  $M_3$  unit to accept electron density is crucial, and the positions labelled 2 and 3 were found to accept electron density more than the fragment in position 1. It is argued that the more electron-accepting fragments will prefer positions 2 and 3 since this will provide the most favourable interactions with the alkyne and thus form the stronger bonds. When the complex possesses different metal fragments, the acceptor orbitals are localized mainly on the more electronegative atoms(s). The preferred rotamer will then be the one with the more electronegative atom(s) occupying positions 2 and/or 3. In the solid state the ligands can be rearranged such that each metal obtains an 18 electron count. It should be stressed here that the authors did not consider varying atoms of the same triad in their trimetallic framework, but were generally concerned with only the 3d metals and the impact of different ligand (Cp, CO), and hence slightly different electronegativities, on the alkyne orientation. It is likely, however, that similar MO arguments will prevail with the 4d and 5d metals as with the 3d metals. This line of reasoning can be extended to help to explain our observed solid state structures.

Consideration of the Pauling electronegativities<sup>54</sup>, 2.20 for Ru and Os and 1.81 for Fe, 1.83 for Co and 1.91 for Ni, allows predictions to be made, based upon these arguments. When the trimetallic framework consists of NiCoRu or NiCoOs atoms the Ru or Os will prefer position 2 since these atoms are the most electronegative. The Ni atom has a Cp ligand bonded to it and is formally Ni(I), making it substantially more electronegative than the Co atom and it should occupy position 3 in the triangle. This would mean that the Co atom must then be in position 1. If all the CO groups are terminal,

this results in electron counts of 17(Os), 18(Ni) and 19(Co). To maintain an 18 electron count at each metal atom, a semi-bridging carbonyl ligand between the Co and Ru/Os is formed. This is the orientation of the metal atoms relative to the alkyne ligand that is indeed found in the X-ray structures of 2 and 3 and postulated for the other NiCoRu and NiCoOs clusters.

Where the trimetallic framework consists of NiCoFe atoms, the Ni atom should occupy either position 2 or 3. The Co atom has a slightly greater electronegativity than that of Fe and would be expected to occupy position 2 or 3 leaving position 1 for the Fe atom. However, it should be noted that the difference in electronegativities for Co and Fe is very small and other factors such as different M-C(alkyne) bond strengths may influence the choice of position of the metal atom on the triangle.

The structure of 1a (Figure 10) appears to be an exception as there is a semi-bridging CO present even though this is a NiCoFe cluster. However, there are two possible reasons for this anomaly. The first is that the PPh<sub>3</sub> ligand renders the cobalt atom less electronegative than the Fe atom (and of course, the Ni atom) and will then prefer position 1. This is the reason forwarded by McGlinchey et. al.<sup>53</sup>. However, it is also possible that the steric bulk of the PPh<sub>3</sub> ligand may be incompatible with having the phenyl group of the alkyne ligand in close proximity. This steric interference is relieved when the alkyne rotates by 60° and becomes  $\sigma$ -bonded to the iron atom. A terminal CO becomes semi-bridging to achieve an 18 electron count at each metal atom. It is not possible to state at this point which of these two factors is responsible for the observed orientation or indeed whether both factors are partly responsible.

Earlier, in the X-ray characterization section, various aspects of each

structure were examined and comparisons between them were made. The focus will now be on the similarities and differences between all the relevant structurally characterized heterotrimetallic clusters. Bond lengths of these clusters are given in Table 8.

Direct comparison between these clusters is often difficult because of the different metals involved; however, trends can often be established. The compounds  $\text{FeCo}_2(\text{CO})_9(\text{C}_2\text{Et}_2)^{55}$ ,  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  and  $\text{Cp}_2\text{Ni}_2\text{Fe}(\text{CO})_3(\text{C}_2\text{Ph}_2)^{46}$  provide a series where there has been stepwise replacement of a  $\text{Co}(\text{CO})_3$  by  $\text{CpNi}$ . All have the alkyne oriented as in I-16, shown earlier, with the Fe atom in position 1. As expected the bond length is found to decrease in the order;  $\text{Co-Co} > \text{Co-Ni} > \text{Ni-Ni}$ , (2.576(1), 2.432(1) and 2.404(4)Å), since the  $\text{Co}(0)$  is being replaced by a  $\text{Ni}(I)$  atom. The decrease is not orderly, however. The largest difference occurs upon replacement of one Co and a much smaller difference is noted when the second Co is replaced by a Ni atom.

Little difference was found for the  $\text{C}\equiv\text{C}$  lengths of the alkyne ligand for different structures. These values are not listed in Table 8 as they only varied between 1.34(2)Å and 1.40(3)Å with the majority centered around 1.37Å.

#### 1.4.2 Solution Structures

The  $^1\text{H}$  NMR of 1-3 exhibit only one Cp resonance and a multiplet due to the phenyl protons of the diphenylacetylene ligand. However information can be obtained regarding the dynamics of these heterotrimetallic alkyne systems by studying the diethyl- and phenylmethylacetylene substituted derivatives. In particular, irradiation of the resonance due to a particular methylene hydrogen revealed that exchange was occurring with a corresponding hydrogen

Table 8. Comparison of Bond Lengths in Related  
Heterotrimetallic Alkyne Clusters

Cluster	Bond Lengths			Reference
	Ni-Co	Ni-M	Co-M	
CpNiCoFe(CO) <sub>5</sub> (PPh <sub>3</sub> )(C <sub>2</sub> Ph <sub>2</sub> )*	2.390(4)	2.486(4)	2.467(4)	a, 56
CpNiCoRu(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> )*	2.399(1)	2.560(1)	2.565(1)	a
CpNiCoOs(CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> )*	2.394(1)	2.579(1)	2.570(1)	a
RuCo <sub>2</sub> (CO) <sub>9</sub> (C <sub>2</sub> Ph <sub>2</sub> )*	Co-Co 2.4535(5)		2.5685(5) 2.6875(5)	48
CpNiCoFe(CO) <sub>6</sub> (C <sub>2</sub> Et <sub>2</sub> )	2.422(1)	2.402(1)	2.494(1)	a
CpNiCoFe(CO) <sub>6</sub> (PhC <sub>2</sub> -i-CO <sub>2</sub> Pr)	2.422	2.423	2.487	25, 57
FeCo <sub>2</sub> (CO) <sub>9</sub> (C <sub>2</sub> Et <sub>2</sub> )	Co-Co 2.576(1)		2.479(1) 2.489(1)	55
Cp <sub>2</sub> Ni <sub>2</sub> Fe(CO) <sub>3</sub> (C <sub>2</sub> Ph <sub>2</sub> )	Ni-Ni 2.404(4)	2.388(4) 2.381(4)		58
CpNi <sub>2</sub> Ru(CO) <sub>3</sub> (C <sub>2</sub> Ph <sub>2</sub> )	Ni-Ni 2.418(4)	2.493(3) 2.496(3)		58
[CpNiFe <sub>2</sub> (CO) <sub>6</sub> (C <sub>2</sub> Ph <sub>2</sub> )] <sup>-</sup>	Fe-Fe 2.506(1)	2.453(1) 2.474(1)	Fe-Fe 2.506(1)	40

<sup>a</sup> This work.

\* Has bridging carbonyl.

of the second methylene group but none between CH protons on the same carbon atom. This was clearly seen at ambient temperatures for 5 and 6 but not for 4 as the exchange rate here was too slow at this temperature. This clearly rules out racemization of the clusters since this process equilibrates hydrogens on the same carbon atoms only, and demonstrates that no transfer of the alkyne from the front to the back face of the cluster can be occurring. For the phenylmethylacetylene substituted clusters 7-9 two methyl resonances were observed at room temperature in unequal intensities which coalesced upon warming the sample. This demonstrates exchange of groups on the alkyne ligand but does not reveal any information regarding front to back transfer of the alkyne ligand on the trimetallic face.

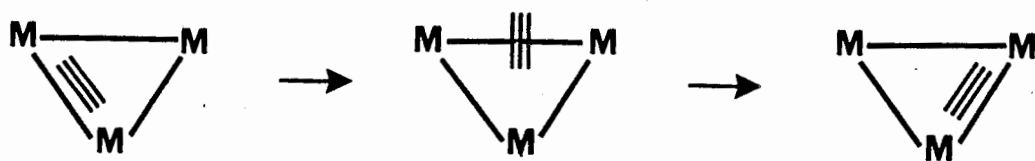
The observed fluxionality of the alkyne ligand is likely a consequence of this ligand rotating on one face of the heterotrimetallic core. It would account for the exchange of hydrogens on separate methylene groups for 4-6 and for the collapse of the methyl peaks in the NMR of 7-9. Rotation of an alkyne ligand about a trimetallic face has been noted before (see Table 9) but there are still not many examples known, or activation energies that have been determined.

In a study of  $\text{H}_2\text{Ru}_3(\text{CO})_9(\text{cyclooctyne})$  and the osmium analogue, Deeming<sup>61</sup> obtained activation energies of 59.8 and 70.7  $\text{kJ mol}^{-1}$ , respectively, and suggested that exchange of the alkyne ligand resulted from a pivotal arrangement (shown in Scheme I-4) with respect to the triangular metal face through the intermediacy of the closo configuration. Extending these 60° rotations around all three vertices of the triangle, the end result will be a rotation of the alkyne ligand, including exchange of diastereotopic methylene protons. Each metal is successively  $\sigma$ - or  $\pi$ -bonded to the alkyne ligand. The alternative approach, Scheme I-5, relates the fluxionality of

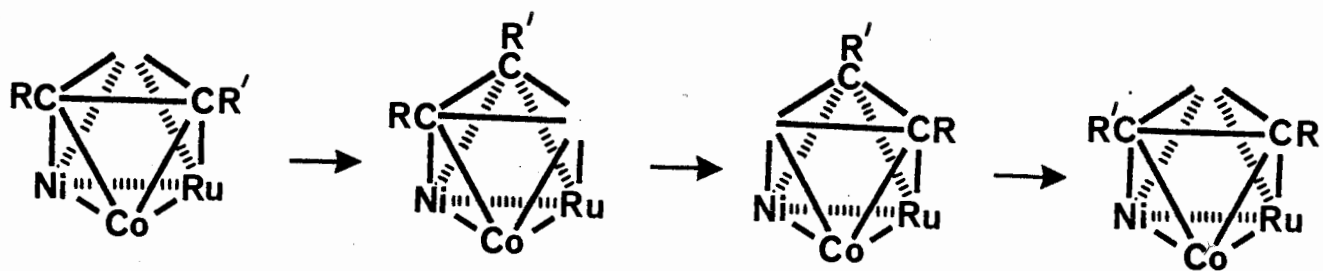
Table 9. Fluxional Trimetallic Alkyne Clusters

Cluster	$\Delta G^\ddagger$ (kJ mol <sup>-1</sup> )	Reference
$\text{Os}_3(\text{CO})_7(\text{PMe}_2)_2(\text{C}_6\text{H}_4)$	58.0	59
$\text{Os}_3(\text{CO})_7(\text{AsMe}_2)_2(\text{C}_6\text{H}_4)$	51.3	59
$\text{HOs}_3(\text{CO})_9(\text{PMe}_2)(\text{C}_6\text{H}_4)$	56.8	59
$\text{HOs}_3(\text{CO})_9(\text{AsMe}_2)(\text{C}_6\text{H}_4)$	58.0	59
$\text{Cp}_3\text{Rh}_3(\text{CO})(\text{C}_2\text{Ph}_2)$	-	60
$\text{Cp}_3\text{Rh}_3(\text{CO})(\text{C}_2(\text{C}_6\text{F}_5)_2)$	-	60
$\text{H}_2\text{Os}_3(\text{CO})_9(\mu_3\text{-indyne})$	60(2)	61
$\text{Os}_3(\text{CO})_7(\text{AsMe}_2)_2(\text{C}_6\text{H}_3\text{-iPr})$	-	62
$\text{Cp}_2\text{Ni}_2\text{Fe}(\text{CO})_3(\text{PhC}_2\text{-CO}_2\text{-iPr})$	63(2)	25,57
$\text{Cp}_2\text{Ni}_2\text{Fe}(\text{CO})_3(\text{PPh}_3)(\text{PhC}_2\text{-CO}_2\text{-iPr})$	73(2)	25,57
$\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$ ( <u>4</u> )	74(1)	a
$\text{CpNiCoFe}(\text{CO})_6(\text{PhC}_2\text{Me})$ ( <u>7</u> )	74(1)	a
$\text{CpNiCoRu}(\text{CO})_6(\text{PhC}_2\text{Me})$ ( <u>8</u> )	70(1)	a
$\text{CpNiCoOs}(\text{CO})_6(\text{PhC}_2\text{Me})$ ( <u>9</u> )	70(1)	a

<sup>a</sup> This work.



SCHEME I-4



SCHEME I-5

nido octahedra possessing one vacant site<sup>57</sup>. This view holds that the fluxionality is effectively the result of a migration of the vacant site over the octahedral surface. As shown in the sequence in Scheme I-5, a migration of the vacant site around one face of the octahedron by successive interchanges with a neighbouring vertex in the same sense equally results in the observed exchange of the alkyne substituents R and R'.

Both heating (to 30° above coalescence) and cooling (to 243K) compounds 7-9 did not result in any additional observable Cp or Me peaks which might have been attributable to intermediates of this rotational scheme. Similarly, cooling (to 220K) solutions of 4-6 did not result in extra Cp, CH<sub>2</sub> or CH<sub>3</sub> resonances. This implies that the intermediates have a low concentration or short lifetime. It is possible that this mechanism is not operable; if so, it is not clear what the nature of the actual mechanism could be. It is probably not dissociative, as no evidence of free alkyne was observed in any of the NMR solutions of the clusters nor was any exchange of alkynes noted between CpNiCoFe(CO)<sub>5</sub>(C<sub>2</sub>Ph<sub>2</sub>) and MeC<sub>2</sub>Ph. In addition, because the chirality of the methylene protons of 4-6 is maintained, the alkyne must rotate only on one face of the core as migration to the opposite face would result in exchange of protons on the same methylene group. As shown in the Characterization section, the NMR's of 4-6 do not show this exchange of protons. Thus any postulated mechanism must not involve free alkyne or interconvert enantiomers. The carbonyl ligands are also fluxional at room temperature. The <sup>13</sup>C{<sup>1</sup>H} NMR's of 1a and 5 showed only a single resonance attributable to the carbons of the CO. These were found at 209.3 and 193.2 ppm in these two compounds, respectively.

The ground state energy differences, ΔE, between the orientational isomers of 7-9 are the first to be calculated on this type of system.



Compound 7 (NiCoFe) shows the smallest  $\Delta E$ ,  $0.57 \text{ kJ mol}^{-1}$  and 8 (NiCoRu) the highest at  $3.90 \text{ kJ mol}^{-1}$ . All three compounds show similar solvent dependence in the preference for a particular conformation, but unfortunately the NOE experiment did not allow us to prove which set of signals belong to an individual orientation.

### 1.5 Conclusion

We have synthesized a series of heterotrimetallic alkyne clusters of general formula  $\text{CpNiCoM}(\text{CO})_6(\text{RC}_2\text{R}')$  ( $\text{M} = \text{Fe, Ru, Os}$ ;  $\text{R} = \text{R}' = \text{Ph, Et}$ ;  $\text{R}=\text{Ph}$ ,  $\text{R}'=\text{Me}$ ). This represents the largest complete series of heterotrimetallic clusters ( $\text{MM}'\text{M}''$ ) to be made. The alkyne ligand is shown by  $^1\text{H}$  NMR to be fluxional in solution and in addition, rotation is occurring only on one face of the heterotrimetallic core. The solid state IR revealed that clusters where  $\text{M} = \text{Fe}$  had only terminal carbonyls but when  $\text{M} = \text{Ru}$  or  $\text{Os}$  a bridging carbonyl was present. This was confirmed by the X-ray crystal structures of four clusters, 1a, 2, 3 and 4. The bridging CO arises as the result of a different preferred orientation of the alkyne ligand with respect to the trimetallic core in the solid state. This has been interpreted in the light of a recent theoretical paper as a preference for the atom with the greatest electronegativity to be  $\sigma$ -bonded to the alkyne.

## EXPERIMENTAL

$\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{RR}')(\text{R} = \text{R}' = \text{Ph, Et; R} = \text{Ph, R}' = \text{Me})$  and  $\text{Fe}_2(\text{CO})_9$  were prepared according to literature procedures<sup>38,63</sup>.  $\text{Ru}(\text{CO})_5$  was prepared by heating  $\text{Ru}_3(\text{CO})_{12}$  under 1500 psi of carbon monoxide.  $\text{Os}(\text{CO})_5$  was kindly supplied by Dr. R.K. Pomeroy of this department. Diphenylacetylene (Aldrich), 3-hexyne (Columbia Organic), phenylmethylacetylene (Aldrich) and  $\text{PPh}_3$  (Aldrich) were used as received.

All solvents were dried and distilled over Na or Na/benzophenone or  $\text{CaH}_2$  under nitrogen. All reactions were carried out in standard Schlenkware, connected to a switchable inert atmosphere/vacuum supply and were conducted under nitrogen.

Photochemical reactions were performed in a Pyrex vessel (~240 mL) equipped with a water cooled quartz finger. A 200 watt ultraviolet Hanovia high-pressure mercury source was placed inside the quartz finger. Nitrogen was passed through the reaction vessel prior to the addition of solvent and starting materials, and slow passage of nitrogen was maintained during the reaction.

Infrared spectra were recorded on a Perkin Elmer 983 spectrophotometer.  $^1\text{H}$  spectra were recorded on a Brüker WM-100 or a Brüker WM-400 instrument at 100 and 400 MHz respectively. All Nuclear Overhauser Enhancement and spin saturation transfer experiments were performed on the Brüker 400.  $^{13}\text{C}$  NMR were recorded on the Brüker WM-400 at 100 MHz. Mass spectral analyses were run by G. Owen on a Hewlett Packard 5985 system operating at 70eV. Microanalyses were performed by Mr. M.K. Yang of the microanalytical laboratory at Simon Fraser University.

The IR,  $^1\text{H}$  NMR and mass spectral data are not documented in the

experimental section but the reader is referred to the appropriate tables in the text.

### 1.6.1 Synthetic

#### Preparation of $\text{CpNiCoFe}(\text{CO})_6(\text{RC}_2\text{R}')$ (1,4,7)

A large excess, (10 fold), of  $\text{Fe}_2(\text{CO})_9$  and the corresponding heterobimetallic alkyne compound,  $\text{CpNiCo}(\text{CO})_3(\text{RC}_2\text{R}')$  ( $\text{R} = \text{R}' = \text{Ph, Et}$ ;  $\text{R} = \text{Ph, R}' = \text{Me}$ ) (typically in the range of 0.1-0.2 g) were stirred overnight at room temperature in approximately 30 mL hexane. The solvent was removed in vacuo, 6-8 mL of fresh hexane was added, and the solution chromatographed on a 1" x 18" Florisil column.  $\text{Fe}_3(\text{CO})_{12}$  and the starting heterobimetallic alkyne were eluted with hexanes. Using a 5:1 hexanes/methylene chloride eluant, a dark green or light brown band was recovered and identified as the appropriate heterotrimetallic alkyne,  $\text{CpNiCoFe}(\text{CO})_6(\text{RC}_2\text{R}')$ . Yields were variable but typically ranged from 30-60%.

The reaction of  $\text{Fe}_2(\text{CO})_9$  with  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  was used as a model. Warming the reaction mixture to 40-45°C and stirring overnight caused additional side-products to be formed. Some side products were made in a very small quantity to be eluted, and decomposed on the column. Two products formed in reasonable yields were  $\text{Cp}_2\text{Fe}_2(\text{CO})_4$  (IR, MS) and  $\text{Fe}_3(\text{CO})_9(\text{C}_2\text{Ph}_2)$  (IR, MS). The use of diethyl ether or THF rather than hexanes resulted in poorer yields of the desired heterotrimetallic complexes and with diethyl ether the formation of  $\text{Fe}_3(\text{CO})_9(\text{C}_2\text{Ph}_2)$  was especially favourable and often isolated in 30% yield. Additionally the amount of  $\text{Fe}_3(\text{CO})_{12}$  also increased with THF or diethyl ether as solvent.

Anal. Calcd. for  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (1): C, 51.34; H, 2.58. Found: C, 52.06; H, 3.62. Calcd. for  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (4): C, 41.77; H, 3.09.

Found: C, 41.87; H, 3.12.

Preparation of CpNiCoRu(CO)<sub>6</sub>(RC<sub>2</sub>R') (2, 5, 8)

In a typical reaction, CpNiCo(CO)<sub>3</sub>(RC<sub>2</sub>R') (R = R' = Ph, Et; R = Ph, R' = Me) and excess Ru(CO)<sub>5</sub> were added to 25 mL hexanes and stirred at room temperature until the bands in the infrared spectrum due to Ru(CO)<sub>5</sub> ( $\nu(\text{CO})$  2039, 2004 cm<sup>-1</sup>) disappeared. This was usually about 18 hours. The reaction mixture could be warmed (40-45°C) safely to decrease the time required for the reaction to go to completion, but more Ru<sub>3</sub>(CO)<sub>12</sub> was formed as a side product. After the reaction was complete, the solvent and any excess Ru(CO)<sub>5</sub> were removed by vacuum and the residue dissolved in the minimum amount of hexanes. Column chromatography with hexanes as eluant yielded Ru<sub>3</sub>(CO)<sub>12</sub> as a yellow band followed by the starting heterobimetallic, CpNiCo(CO)<sub>3</sub>(RC<sub>2</sub>R'). The desired product, CpNiCoRu(CO)<sub>6</sub>(RC<sub>2</sub>R') was then eluted as a brown band.

<sup>13</sup>C{<sup>1</sup>H} (5), (CDCl<sub>3</sub>, 293K); 193.2 (CO), 90.5 (Cp), 38.4, 36.9 (CH<sub>2</sub>), 17.1, 16.7 (CH<sub>3</sub>).

Anal. Calcd. for CpNiCoRu(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (2): C, 47.66; H, 2.40. Found: C, 47.82; H, 2.51. Calcd. for CpNiCoRu(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>) (5): C, 38.24; H, 2.83. Found: C, 38.18; H, 2.79.

Preparation of CpNiCoOs(CO)<sub>6</sub>(RC<sub>2</sub>R') (3, 6, 9)

CpNiCo(CO)<sub>3</sub>(RC<sub>2</sub>R') (R = R' = Ph, Et; R = Ph, R' = Me) and excess Os(CO)<sub>5</sub> were added to 25 mL hexanes and heated to 40-45°C for 10-12 hours. The solvent volume was reduced to 5 mL and the solution was poured onto a Florisil column and chromatographed. Variable, but usually small, amounts

of  $\text{Os}_3(\text{CO})_{12}$  were first eluted with hexanes followed by a green band identified as the starting heterobimetallic alkyne. Using a 5:1 hexane/chloroform mixture as eluant a dark brown band was next isolated and identified as  $\text{CpNiCoOs}(\text{CO})_6(\text{RC}_2\text{R}')$ .

Anal. Calcd. for  $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (3): C, 41.75; H, 2.10. Found: C, 41.79; H, 2.20. Calcd. for  $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (6): C, 32.77; H, 2.43. Found: C, 32.87; H, 2.49.

#### Preparation of $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$ (1a)

Compound 1 (0.12 g, 0.21 mmol) and  $\text{PPh}_3$  (0.07 g, 0.27 mmol) were added to 20 mL hexanes and stirred for 12 hours. The solvent volume was reduced to 4-5 mL and eluted on a Florisil column. Elution with methylene chloride/hexanes yielded a black band identified as 1a. No band due to 1 was visible although some decomposition had occurred.

Anal. Calcd for  $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$ : C, 61.6, H, 3.7, Found: C, 62.2, H, 4.4.  $^{13}\text{C}\{^1\text{H}\}$  ( $\text{CDCl}_3$ , 293K); 209.3 (CO), 156.2, 150.9 (acetylenic C), 133.8-125.4 (phenyl C), 92.8 (Cp).

#### Reaction of $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$ with $\text{M}_3(\text{CO})_{12}$ (M = Fe, Ru, Os)

$\text{Fe}_3(\text{CO})_{12}$  (0.21 g, 0.42 mmol) and  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  (0.19 g, 0.41 mmol) was added to 30 mL of hexanes and refluxed for 18 hours. The solvent volume was reduced by vacuum and the trap was noted to contain some  $\text{Fe}(\text{CO})_5$ . The residual hexanes was placed on a Florisil column and elution with hexanes yielded a small amount of  $\text{Fe}(\text{CO})_5$ , a red-brown band identified as  $\text{Co}_2(\text{CO})_6(\text{C}_2\text{Ph}_2)$  and the starting heterobimetallic as a green band. Elution with 4:1 hexanes/methylene chloride gave  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  as a brown

band, small and variable amounts of  $\text{Cp}_2\text{Ni}_2\text{Fe}_2(\text{CO})_6(\text{C}_2\text{Ph}_2)$  as a blue band ( $\nu(\text{CO}) = 2010, 1973 \text{ cm}^{-1}$ ),  $\text{Cp}_2\text{Ni}_2\text{Fe}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  as a green-brown band ( $\nu(\text{CO}) = 2030, 1979 \text{ and } 1963 \text{ cm}^{-1}$ ) and  $\text{Fe}_3(\text{CO})_9(\text{C}_2\text{Ph}_2)$  as a brown-black band. With only methylene chloride as the eluant some  $\text{Cp}_2\text{Fe}_2(\text{CO})_4$  was recovered.

$\text{Ru}_3(\text{CO})_{12}$  (0.143 g, 0.22 mmol) and  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  (0.10 g, 0.225 mmol) were added to 20 mL hexanes and heated at  $60^\circ\text{C}$  for 16 hours. The solvent was removed by vacuum and the residue dissolved in fresh hexanes and chromatographed on Florisil. Only small amounts of 5 could be isolated.

$\text{Os}_3(\text{CO})_{12}$  (0.20 g, 0.22 mmol) and  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  (0.2 g, 0.45 mmol) were added to 20 mL of hexanes and heated to  $60^\circ\text{C}$  for 16 hours. The solvent volume was reduced by vacuum and the residue chromatographed on a Florisil column. Only very small traces of 3 could be found.

#### Reaction of $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$ with $\text{Cr}(\text{CO})_5(\text{THF})$

$\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  (0.30 g, 0.68 mmol) was added to a THF solution of  $\text{Cr}(\text{CO})_5(\text{THF})$  (0.5 g, 2.2 mmol) which was freshly prepared from the photolysis of  $\text{Cr}(\text{CO})_6$  in THF. The reaction was monitored after 1 hour and since no new carbonyl bands appeared, it was stirred overnight at ambient temperature. An infrared spectrum revealed carbonyl bands due to  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  ( $\nu(\text{CO}) = 2060, 2014, 2008 \text{ cm}^{-1}$ ) and  $\text{Cr}(\text{CO})_6$ . The solvent was removed by vacuum and the residue dissolved in 5 mL of hexanes/methylene chloride and chromatographed on Florisil. The infrared spectrum of the fractions revealed only  $\text{Cr}(\text{CO})_6$  and  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  being eluted. No new material could be isolated.

Reaction of  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  with  $\text{CpRe}(\text{CO})_2(\text{THF})$

$\text{CpRe}(\text{CO})_2(\text{THF})$  (60 mg, 0.016 mmol) and  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  (70 mg, 0.16 mmol) were added to 25 mL hexanes and stirred at room temperature for 2 hours. An infrared spectrum of the solution revealed little remaining  $\text{CpRe}(\text{CO})_2(\text{THF})$  ( $\nu(\text{CO}) = 1925, 1857 \text{ cm}^{-1}$ ) but strong bands due to  $\text{CpRe}(\text{CO})_3$  ( $\nu(\text{CO}) = 2031, 1940 \text{ cm}^{-1}$ ). The solvent was removed under vacuum and the residue dissolved in hexanes/methylene chloride. Chromatography on a Florisil column with hexanes as eluant yielded  $\text{CpRe}(\text{CO})_3$ , then  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$ . With a stronger hexanes/methylene chloride eluant (3:1) a light yellow band was next recovered and identified as  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ . This compound and its properties are discussed in Chapter 4.

Reaction of  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  with  $\text{Fe}(\text{CO})_4(\text{PPh}_3)$

$\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  (51 mg, 0.113 mmol) and  $\text{Fe}(\text{CO})_4(\text{PPh}_3)$  (50 mg, 0.109 mmol) were added to 15 mL  $\text{CH}_3\text{CN}$  and stirred for 1 hour. As no reaction was apparent from monitoring the infrared spectrum,  $\text{Me}_3\text{NO}$  (20 mg) was added to the solution. After stirring for a further 1 hour, no new carbonyl bands in the IR were detected and the reaction was discontinued.

Reaction of  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  with  $\text{CpCo}(\text{CO})_2$

$\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  (0.1 g, 0.225 mmol) and excess  $\text{CpCo}(\text{CO})_2$  (60 mg, 0.333 mmol) were added to 15 mL of hexanes and refluxed for 12 hours. The solution was cooled and the reaction mixture chromatographed on a Florisil column. Only  $\text{CpCo}(\text{CO})_2$  ( $\nu(\text{CO}) = 2026, 1966 \text{ cm}^{-1}$ ) and  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  were identified. Some decomposition was indicated since material left at the top of the column could not be eluted with acetone or diethyl ether.

Reaction of CpNiCoFe(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) with MeC<sub>2</sub>Ph

CpNiCoFe(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (50 mg, 0.086 mmol) and excess methylphenylacetylene were added to 15 mL of hexanes and stirred for two hours. There was no apparent reaction as judged from an infrared spectrum and the reaction was discontinued.

Examples of the Calculation of  $\Delta G^\ddagger$  and  $\Delta E$

a) Calculation of  $\Delta G^\ddagger$  from Coalescence Temperature Experiments for 7

The  $\Delta\nu_\infty$  value was the separation of the two Cp or two Me peaks in Hz at 243K. From this the rate constant could be evaluated.

$$\Delta\nu_{2,3}(7) = 11.2 \pm 0.5 \text{ Hz}$$

$$K_c = \pi/\sqrt{2} \Delta\nu_{2,3} = 24.9 \pm 1.1 \text{ sec}^{-1}$$

By using the Eyring equation, the  $\delta G^\ddagger$  value can then be calculated.

$$K_c = KT/h[e^{-(\Delta G^\ddagger/RT_c)}]$$

$$K = \text{Boltzmann constant} = 1.3807 \times 10^{-16} \text{ ergs K}^{-1}$$

$$h = \text{Planck's constant} = 6.63 \times 10^{-27} \text{ ergs sec}$$

$$R = \text{gas constant} = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$T_c(7)(\text{methyl}) = 339 \pm 1\text{K}$$

$$\text{then } \Delta G^\ddagger = 74.5 \pm 1.3 \text{ kJ mol}^{-1}.$$

Additionally, from the fractional populations of the two sites  $\rho_A$  and  $\rho_B$ , the rate constants  $k_{AB}$  and  $k_{BA}$  could be calculated.

$$\rho_A(7)(\text{methyl}) = 0.62$$

$$\rho_B(7)(\text{methyl}) = 0.38$$



where A represents the high field resonance. Then,

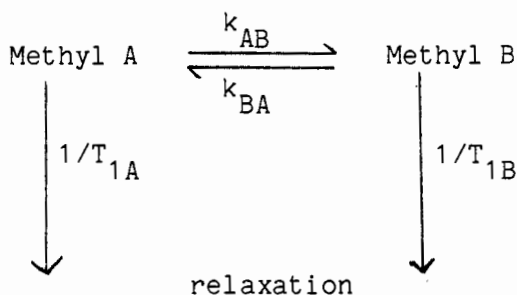
$$k_{AB} = 2(\pi/\sqrt{2})\Delta\nu_{\infty}(\rho_B) = 30.8 \text{ sec}^{-1}$$

and

$$k_{BA} = 2(\pi/\sqrt{2})\Delta\nu_{\infty}(\rho_A) = 18.9 \text{ sec}^{-1}$$

b) Calculation of  $\Delta G^\ddagger$  by Spin Saturation Transfer Experiments for 4

A brief outline will be given. A selective  $180^\circ$  pulse is applied to a methyl resonance followed by a variable delay time and then a normal  $90^\circ$  observation pulse applied. The  $180^\circ$  pulse inverts the methyl resonance and the second methyl resonance will then decrease in intensity since it is undergoing exchange with the first, and now inverted, methyl peak. The difference spectrum (spectrum 1 - spectrum 2) is plotted and this shows the effect of only the selective pulse. Varying the delay time between the  $180^\circ$  pulse and the observation pulse allows a rate constant to be established for one temperature. An example of varying delay times is shown in Figure 14. The results can be easily seen with a simple kinetic scheme. The chemical



rates,  $k_{AB}$  and  $k_{BA}$  are equal and the two relaxation rates,  $1/T_{1A}$  and  $1/T_{1B}$  are assumed to be equal or nearly so. If A and B represent the intensities of the two peaks, approximated by measuring the peak height of the center line of the triplet, then

$$(A + B) = A_0 \exp(-(1/T_{1A})t)$$

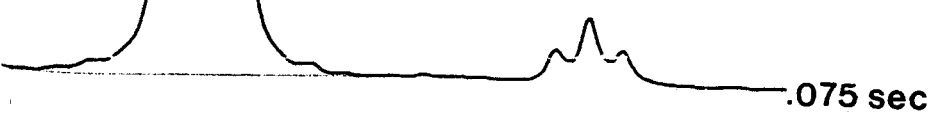
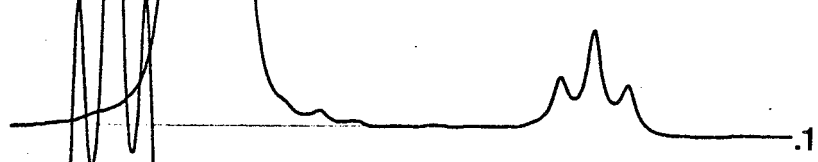
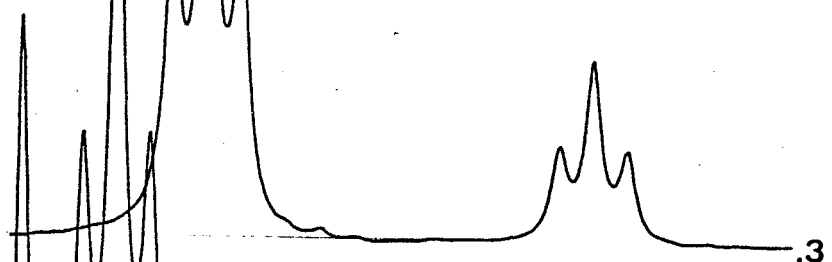
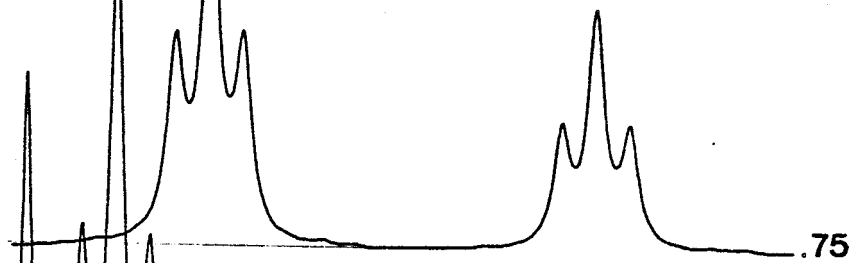
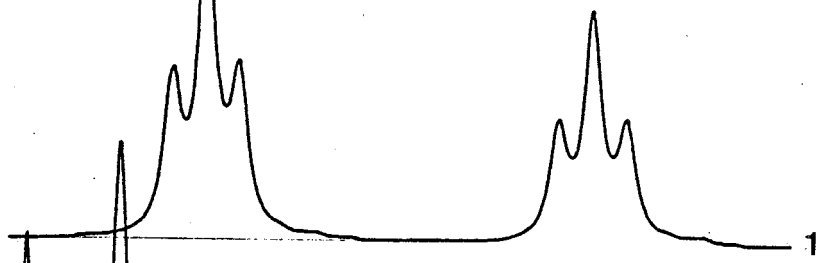
and

$$(A - B) = A_0 \exp(-(1/T_{1A} + k_{AB} + k_{BA})t)$$

## FIGURE 14

180° Pulse on the CH<sub>3</sub> Resonance  
of CpNiCoFe(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>) (4)  
with Varying Delay Times

86



T:314 K

where  $A_0$  is the intensity of the line A at time zero and  $t$  is the delay time. The sum  $(A + B)$  represents the total magnetization induced by the selective pulse and the difference  $(A - B)$  represents the equilibration of the excess magnetization between the two sites. The exchange rate constant between methyl group A and methyl group B is obtained by taking the difference between the sum and the difference totals. A plot of the  $k$  values obtained at different temperatures versus  $1/T$  (Figure 15) gave a straight line of slope  $-\Delta G^\ddagger/1.987$  from which the value of  $74.5 \pm 1.3 \text{ kJ mol}^{-1}$  was obtained.

c) Calculation of  $\Delta E$  for 7

From the population ratio of both isomers ( $^1\text{H NMR}$ ) and by using the Boltzmann distribution equation, the  $\Delta E$  value could be obtained.

$$\Delta E = -kT \ln(N_-/N_+) \times N$$

$$k = \text{Boltzmann constant} = 1.38 \times 10^{-28} \text{ kJ K}^{-1}$$

$$T = \text{Temperature (K)}$$

$$N = \text{Avogadro's number} = 6.023 \times 10^{23}$$

$$N_- = \text{integral of less dominant isomer}$$

$$N_+ = \text{integral of dominant isomer}$$

For 7 at  $T = 295\text{K}$ ,

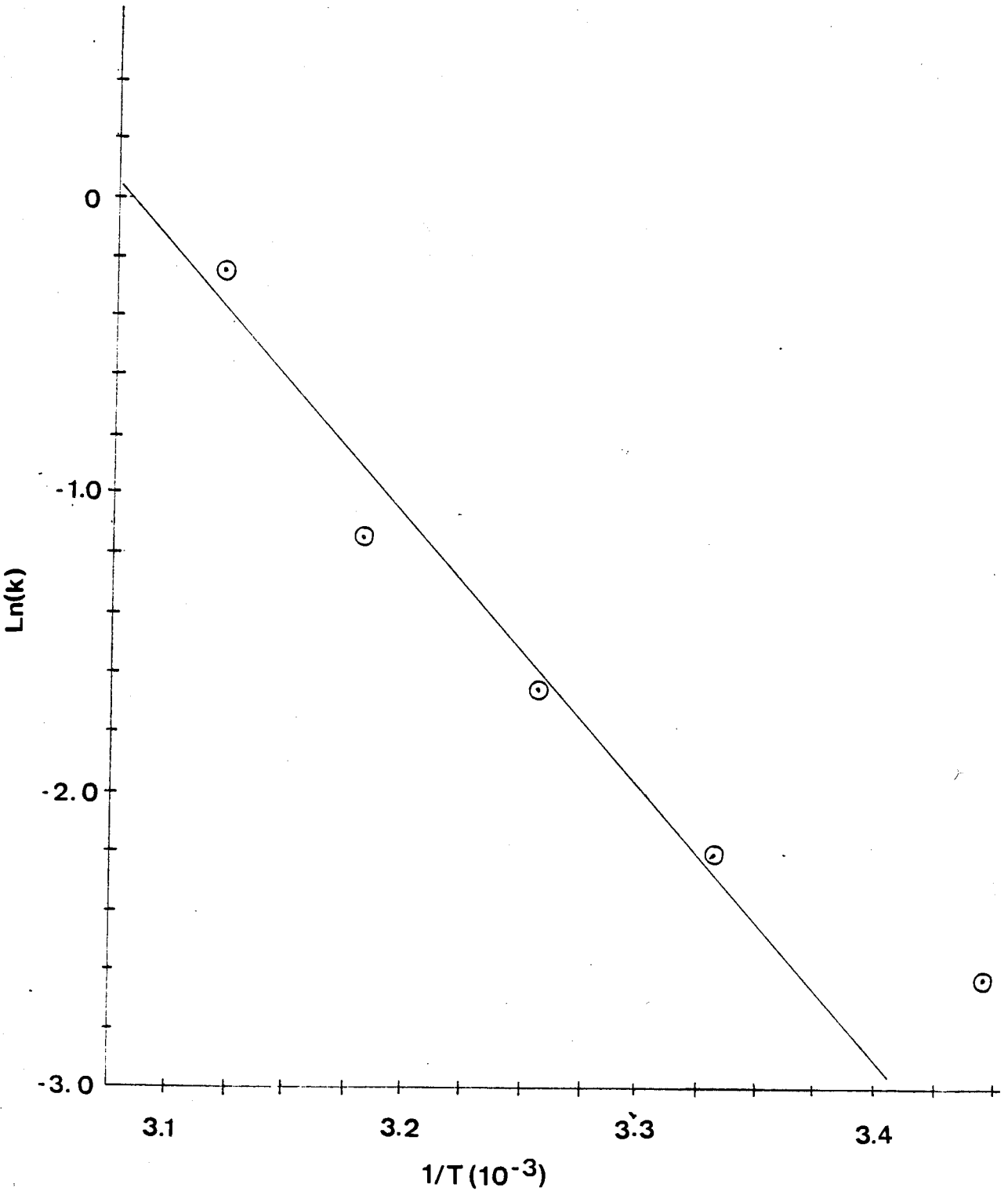
$$\begin{aligned} \Delta E &= -1.38 \times 10^{-28} \text{ kJ K}^{-1} T(295 \text{ K}) \ln(1.0/1.24) \times 6.023 \times 10^{23} \text{ mol}^{-1} \\ &= 0.53 \text{ kJ mol}^{-1} \end{aligned}$$

### 2.6.2 X-ray Structures

An overview of the X-ray data collection is given in Chapter 5. Many of the parameters and conditions used for the collection of data are listed in Table 10 but a brief introduction for each crystal structure is given

FIGURE 15

Graph of  $\ln K$  versus  $1/T$   
for Compound 4



under the appropriate heading.

1) X-ray Structure of  $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$  (1a)

A black crystal, 0.09 x 0.11 x 0.15 mm, obtained from a hexanes/chloroform solution was chosen for the diffraction studies. Preliminary Weissenberg and precession photos (Cu radiation) indicated the crystal to be monoclinic, and systematic absences  $0k0: k = 2n+1$ ,  $h0l: l = 2n+1$ , defined the space group as  $P2_1/c$ . Data were then collected on a Picker FACS-I diffractometer (Mo radiation) with a scan rate of  $1^\circ \text{ min}^{-1}$  for  $3^\circ \leq 2\theta \leq 35^\circ$  and a symmetrical scan width of  $(1.2 + 0.692\tan\theta)^\circ$ . There were very few reflections beyond  $35^\circ$  so the data collection was stopped at that point. Two standards were measured every 70 reflections and showed only random fluctuations.

Absorption corrections were applied via an analytical correction (transmission factors, 0.75-0.86). 1109 reflections out of a total of 2316 had  $I \geq 2.3\sigma(I)$  and were regarded as observed. The positions of the metal atoms and the phosphorus atom were located using MULTAN and the non-hydrogen atoms from a series of difference Fourier synthesis interspersed with least squares refinement. The hydrogen atom positions were calculated geometrically with a C-H bond length of 0.98Å and were given isotropic thermal parameters equal to the carbon atom to which they were attached and included as fixed contributions in further refinement. The metal atoms and the phosphorus atom were refined anisotropically. The largest peak in a final difference map  $0.43(20)\text{e}\text{\AA}^{-3}$  was located near C(61) of the aromatic ring. Computer programs are from reference 64. Bond distances and angles are in Table 11, atom coordinates are given in Table 15, calculated hydrogen atom coordinates in Table 19, anisotropic thermal parameters in Table 23 and

calculated meanplanes and dihedral angles in Table 27. An ORTEP diagram with the atom labelling scheme is shown in Figure 10 (Section 1.3.2d) and is grouped with all the heterotrimetallic alkyne structures.

2) X-ray Structure of  $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (2)

A dark brown crystal 0.11 x 0.19 x 0.24 mm, obtained from a hexane solution was chosen for the diffraction studies. Preliminary Weissenberg and precession photos (Cu radiation) indicated the crystal to be monoclinic and systematic absences  $0k0: k = 2n+1$ ,  $h0l: l = 2n+1$ , define the space group as  $P2_1/c$ . Data were then collected on a Picker FACS-I diffractometer (Mo radiation) with a scan range of  $2^\circ \text{ min}^{-1}$  for  $3^\circ \leq 2\theta \leq 45^\circ$  and a symmetrical scan width of  $(1.4 + 0.692\tan\theta)^\circ$ . Two standards were measured for intensity control after every 70 reflections but showed only random fluctuations.

Absorption corrections were applied via an analytical correction (transmission factors, 0.64-0.79). 1816 reflections out of a total of 3100 had  $I \geq 2.3\sigma(I)$  and were regarded as observed. The positions of the metal atoms were located using MULTAN and the non-hydrogen atoms were located from a series of difference Fourier synthesis and interspersed with least squares refinement. The hydrogen atom positions were calculated geometrically with a C-H bond length of 0.98Å and were given isotropic thermal parameters equal to the carbon atom to which they were attached and were included as fixed contributions in further refinement. The metal atoms, carbonyl ligands and the cyclopentadienyl ligand were refined anisotropically. The largest peak in a final difference map,  $0.53(22)\text{e}\text{\AA}^{-3}$  was located between the Ru atom and the nearest phenyl ring. Computer programs are from reference 64. Bond distances and angles are in Table 12, atomic coordinates are given in Table 16, calculated hydrogen atom coordinates in Table 20, anisotropic thermal



parameters in Table 24 and calculated meanplanes and dihedral angles in Table 28. An ORTEP diagram with the atom labelling scheme is shown in Figure 11 (Section 1.3.2d) and is grouped with the other relevant structures.

### 3) X-ray Structure of $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$ (3)

A triangular, platy dark brown crystal, 0.49 x 0.51 x 0.60 x 0.06 mm obtained from a hexane/methylene chloride solution at  $-20^\circ$  was chosen for the diffraction studies. Preliminary oscillation and a zero-level Weissenberg photo (Cu radiation) revealed the crystal to be of good quality. Data were collected on an Enraf Nonius CAD-4 diffractometer (Mo radiation) with a variable scan speed of  $0.8\text{-}5.0^\circ \text{ min}^{-1}$  for reflections with  $3.0^\circ \leq 2\theta \leq 45^\circ$ . The base scan width was  $0.7^\circ$  for  $3\text{-}40^\circ$ , and  $0.9^\circ$  for  $40^\circ\text{-}45^\circ$  plus dispersion correction. Two standards were measured after every 150 reflections but showed only random fluctuations.

Absorption corrections were applied via an analytical correction (transmission factors, 0.23-0.66). 2503 reflections out of a total of 3063 had  $I \geq 2.3\sigma(I)$  and were regarded as observed. The positions of the metal atoms were located using MULTAN, and non-hydrogen atoms from a series of difference Fourier synthesis interspersed with least squares refinement. The hydrogen atom positions were calculated geometrically with a C-H bond length of 0.98Å and assigned isotropic thermal parameters equal to the carbon atom to which they were bonded and included as fixed contributions in further refinement. All the non-hydrogen atoms were refined anisotropically. The largest peak in a final difference map,  $1.1(2)\text{e}\text{\AA}^{-3}$  was 1.2Å from the Os atom and 0.8Å from C(3) and of no chemical significance. Computer programs are from reference 65. Bond distances and angles are in Table 13, atomic coordinates are in Table 17, calculated hydrogen atom

coordinates in Table 21, anisotropic thermal parameters in Table 25 and calculated meanplanes and dihedral angles in Table 29. A labelled SNOOPI diagram is shown in Figure 12 (Section 1.3.2d) and is grouped with the other heterotrimetallic structures.

4) X-ray Structure of  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (4)

A dark green needle, 0.07 x 0.12 x 0.36 mm recrystallized from hexane/chloroform at 0°C was used for the diffraction studies. Oscillation photos (Cu radiation) revealed the crystal to be of good quality. Data were collected on an Enraf Nonius CAD-4 diffractometer (Mo radiation) with a variable scan speed of  $0.3\text{-}5.0^\circ \text{ min}^{-1}$  for  $3^\circ \leq 2\theta \leq 40^\circ$  and a symmetrical base scan width of  $0.7^\circ$  plus dispersion correction above and below the weighted  $K\alpha_1$  and  $K\alpha_2$ . Two standard reflections were monitored every 150 reflections and showed no indication of intensity loss.

Absorption corrections were applied via an empirical correction ( $\psi$  scans; transmission factors, 0.72-0.84). 1387 reflections from a total of 1748 had  $I \geq 2.3\sigma(I)$  and were regarded as observed. The positions of the metal atoms were located using MULTAN and the non-hydrogen atoms from a series of difference Fourier synthesis interspersed with least squares refinement. The hydrogen atom positions were calculated geometrically with a C-H bond length of 0.98Å and given isotropic thermal parameters equal to the carbon atom to which they were attached and were included as fixed contributions in further refinement. All non-hydrogen atoms were refined anisotropically. The largest peak in a final difference map,  $0.6(2)\text{e}\text{\AA}^{-3}$  was located between the Ni atom and the Cp ring. Computer programs are from reference 65. Bond distances and angles are in Table 14, atomic coordinates are given in Table 18, calculated hydrogen atom coordinates in Table 22,

anisotropic thermal parameters in Table 26 and calculated meanplanes and dihedral angles in Table 30. A SNOOPI diagram with the atom labelling scheme is shown in Figure 13 (Section 1.3.2d) and is grouped with the other relevant structures.

Table 10. Summary of Crystal Data for CpNiCoFe(CO)<sub>5</sub>(PPh<sub>3</sub>)(C<sub>2</sub>Ph<sub>2</sub>) (1a), CpNiCoRu(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (2), CpNiCoOs(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (3) and CpNiCoFe(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>) (4)

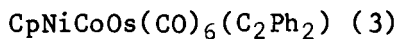
compound	<u>1a</u>	<u>2</u>	<u>3</u>	<u>4</u>
mol. wt.	819.15	630.1	719.3	488.8
space group	P2 <sub>1</sub> /c	P2 <sub>1</sub> /c	P2 <sub>1</sub> /c	P2 <sub>1</sub> /n
a, Å	11.698(2)	9.964(2)	9.950(2)	9.559(3)
b, Å	14.768(3)	9.5112(6)	9.505(2)	12.573(5)
c, Å	22.616(4)	24.835(2)	24.813(4)	16.175(4)
β, deg	111.82(1)	90.113(9)	90.10(3)	104.42(3)
Z	4	4	4	4
cell vol, Å <sup>3</sup>	3627	2353	2347	1883
ρ calcd, gcm <sup>-3</sup>	1.48	1.778	2.03	1.72
ρ obsd, gcm <sup>-3</sup>	1.51	1.80	2.11	1.74
μ (MoK <sub>α</sub> ) cm <sup>-1</sup>	14.7	21.7	69.4	26.4
total reflections, measured	2316	3100	3063	1748
no. reflections, I ≥ 2.3σ(I)	1109	1816	2503	1387
data limits, 2θ, deg	3, 35	3.5, 45	3, 45	3, 40
final R	0.057	0.036	0.033	0.029
final R <sub>w</sub>	0.070	0.042	0.039	0.033
GOF	1.32	1.37	1.52	1.23

Table 11. Bond Distances (Å) and Angles (deg) for  
 $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$  (1a)

Ni-Co	2.390(4)	Co-C(6)	2.05(2)
Ni-Fe	2.486(4)	Co-C(7)	2.04(2)
Fe-Co	2.467(4)	C(6)-C(7)	1.34(2)
Co-P	2.230(6)	C(6)-C(41)	1.50(2)
Fe-C(1)	1.70(2)	C(7)-C(51)	1.47(2)
Fe-C(2)	1.79(2)	C(1)-O(1)	1.20(3)
Fe-C(3)	1.69(2)	C(2)-O(2)	1.13(3)
Fe-C(4)	2.07(2)	C(3)-O(3)	1.21(3)
Co-C(4)	1.74(2)	C(4)-O(4)	1.19(3)
Co-C(5)	1.69(2)	C(5)-O(5)	1.15(3)
Fe-C(6)	2.03(2)	Ni-Cp (center)	1.75(1)
Ni-C(7)	1.93(2)	P-C(11)	1.86(2)
		P-C(21)	1.80(2)
		P-C(31)	1.79(2)
Ni-Co-Fe	61.55(12)	Co-C(5)-O(5)	175(2)
Co-Fe-Ni	57.71(11)	Ni-C(7)-Co	74.1(6)
Co-Ni-Fe	60.74(11)	Fe-C(6)-Co	74.3(6)
Fe-Co-P	119.4(2)	Ni-C(7)-C(6)	107(1)
Ni-Co-P	177.2(2)	Ni-C(7)-C(51)	117(1)
Fe-C(1)-O(1)	170(2)	Fe-C(6)-C(7)	107(1)
Fe-C(2)-O(2)	173(2)	Fe-C(6)-C(41)	125(1)
Fe-C(3)-O(3)	179(2)	C(6)-C(7)-C(51)	131(1)
Fe-C(4)-O(4)	131(2)	C(7)-C(6)-C(41)	126(1)
Co-C(4)-O(4)	148(2)		

Table 12. Bond Distances (Å) and Angles (deg) for CpNiCoRu(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (2)

Ni-Ru	2.560(1)	Ru-C(7)	2.126(7)
Ni-Co	2.399(1)	Ni-Cp(cnt r)	1.747(1)
Co-Ru	2.565(1)	Ni-C(8)	1.903(7)
Co-C(1)	1.828(8)	C(7)-C(71)	1.480(10)
Co-C(2)	1.791(9)	C(7)-C(8)	1.377(9)
Co-C(3)	1.758(8)	C(8)-C(81)	1.489(9)
Co-C(7)	2.085(7)	C(1)-O(1)	1.158(9)
Co-C(8)	2.071(7)	C(2)-O(2)	1.135(10)
Ru-C(1)	2.251(8)	C(3)-O(3)	1.129(10)
Ru-C(4)	1.913(9)	C(4)-O(4)	1.118(10)
Ru-C(5)	1.912(9)	C(5)-O(5)	1.135(10)
Ru-C(6)	1.943(8)	C(6)-O(6)	1.130(10)
Ni-Co-Ru	61.98(4)	Ru-C(5)-O(5)	179.1(8)
Ni-Ru-Co	55.82(4)	Ru-C(6)-O(6)	179.5(8)
Ru-Ni-Co	62.21(4)	Ru-C(7)-Co	75.0(2)
Co-C(1)-O(1)	148.2(7)	Ru-C(7)-C(8)	103.4(5)
Co-C(1)-Ru	77.2(3)	Ni-C(8)-Co	74.1(2)
Co-C(2)-O(2)	177.7(8)	Ni-C(8)-C(7)	110.6(5)
Co-C(3)-O(3)	175.7(8)	Co-Ru-C(1)	44.0(2)
Ru-C(1)-O(1)	134.5(6)	Ru-Co-C(1)	58.8(3)
Ru-C(4)-O(4)	178.1(8)		

Table 13. Bond Distances (Å) and Angles (deg) for

Ni-Os	2.579(1)	Co-C(14)	2.067(8)
Ni-Co	2.394(1)	Ni-Cp(ctr)	1.755(1)
Co-Os	2.570(1)	Ni-C(13)	1.900(7)
Os-C(1)	2.273(9)	C(7)-C(13)	1.485(10)
Os-C(2)	1.919(9)	C(13)-C(14)	1.377(10)
Os-C(3)	1.908(9)	C(14)-C(15)	1.477(10)
Os-C(4)	1.949(8)	C(1)-O(1)	1.148(10)
Os-C(14)	2.147(7)	C(2)-O(2)	1.126(11)
Co-C(1)	1.853(9)	C(3)-O(3)	1.141(11)
Co-C(5)	1.774(8)	C(4)-O(4)	1.122(10)
Co-C(6)	1.794(9)	C(5)-O(5)	1.122(10)
Co-C(13)	2.065(7)	C(6)-O(6)	1.128(10)
Ni-Co-Os	62.49(4)	Co-C(6)-O(6)	178.7(8)
Ni-Os-Co	55.41(3)	Os-C(1)-Co	76.3(3)
Os-Ni-Co	62.10(4)	Ni-C(13)-Co	74.2(2)
Os-C(1)-O(1)	136.2(8)	Ni-C(13)-C(14)	110.2(5)
Os-C(2)-O(2)	179.4(8)	Os-C(14)-Co	75.1(2)
Os-C(3)-O(3)	177.8(8)	Os-C(14)-C(13)	104.1(5)
Os-C(4)-O(4)	178.3(8)	Co-Os-C(1)	44.5(2)
Co-C(1)-O(1)	147.3(7)	Os-Co-C(1)	59.2(3)
Co-C(5)-O(5)	177.1(8)		

Table 14. Bond Distances (Å) and Angles (deg) for  
CpNiCoFe(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>) (4)

Ni-Co	2.432(1)	Ni-C(9)	1.898(5)
Ni-Fe	2.402(1)	C(1)-O(1)	1.144(7)
Co-Fe	2.494(1)	C(2)-O(2)	1.140(7)
Fe-C(1)	1.767(6)	C(3)-O(3)	1.144(7)
Fe-C(2)	1.786(6)	C(4)-O(4)	1.137(7)
Fe-C(3)	1.773(6)	C(5)-O(5)	1.133(7)
Fe-C(9)	2.036(5)	C(6)-O(6)	1.134(7)
Fe-C(10)	2.037(5)	C(7)-C(8)	1.492(9)
Co-C(4)	1.781(6)	C(8)-C(9)	1.510(7)
Co-C(5)	1.811(6)	C(9)-C(10)	1.371(7)
Co-C(6)	1.791(6)	C(10)-C(11)	1.522(7)
Co-C(10)	1.953(5)	C(11)-C(12)	1.507(8)
Ni-Cp(cntr)	1.750(4)		
Ni-Co-Fe	58.36(3)	Co-C(5)-O(5)	178.2(5)
Ni-Fe-Co	59.52(3)	Co-C(6)-O(6)	177.0(6)
Fe-Ni-Co	62.11(3)	Ni-C(9)-C(10)	108.0(4)
Fe-C(1)-O(1)	178.2(6)	Ni-C(9)-Fe	75.2(2)
Fe-C(2)-O(2)	178.2(6)	Co-C(10)-C(9)	103.9(4)
Fe-C(3)-O(3)	178.4(5)	Co-C(10)-Fe	77.3(2)
Co-C(4)-O(4)	178.3(5)		



Table 15. Atom Coordinates ( $\times 10^4$ ) for  $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$  (1a)

Atom	X	Y	Z	<u>B<sub>iso</sub></u>
Co	6468.4(26)	2282.0(20)	4855.4(13)	31(2)*
Fe	7936.3(28)	1045.6(20)	5340.0(14)	33(2)*
Ni	6941.6(26)	1930.5(19)	5952.7(12)	37(2)*
P	6075(5)	2673(4)	3846(3)	36(4)*
C(1)	7443(22)	37(16)	5515(11)	58(6)
C(2)	8533(20)	535(15)	4947(10)	49(6)
C(3)	9199(19)	1057(15)	6019(10)	48(6)
C(4)	6134(19)	1136(15)	4704(10)	54(6)
C(5)	5137(21)	2693(16)	4875(11)	56(6)
O(1)	7263(14)	-729(11)	5629(7)	67(4)
O(2)	8808(14)	166(10)	4434(7)	61(4)
O(3)	10120(14)	1060(11)	6498(7)	65(4)
O(4)	5410(13)	567(11)	4428(7)	59(4)
O(5)	4263(13)	3037(10)	4888(7)	63(4)
C(6)	8342(15)	2380(12)	5312(8)	20(4)
C(7)	7782(15)	2827(12)	5642(8)	22(4)
C(11)	6900(18)	1972(14)	3445(9)	41(5)
C(12)	6270(18)	1219(15)	3100(10)	45(6)
C(13)	7038(20)	678(16)	2865(10)	50(6)
C(14)	8168(21)	827(16)	2977(11)	60(7)
C(15)	8776(20)	1562(16)	3288(10)	59(6)
C(16)	8097(18)	2156(14)	3564(9)	38(5)
C(21)	4509(17)	2596(13)	3296(9)	31(5)

Table 15 (Cont'd)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>	
C(22)	4083(20)	2991(16)	2705(10)	56(6)
C(23)	2858(21)	2926(16)	2300(11)	61(6)
C(24)	2020(21)	2458(17)	2461(11)	61(6)
C(25)	2393(21)	2019(17)	3024(11)	63(7)
C(26)	3645(19)	2082(15)	3439(10)	47(6)
C(31)	6440(17)	3809(13)	3712(8)	29(5)
C(32)	6763(18)	4148(13)	3220(9)	37(5)
C(33)	6914(20)	5074(15)	3141(11)	52(6)
C(34)	6772(20)	5698(15)	3556(10)	46(6)
C(35)	6406(19)	5401(14)	4034(9)	39(5)
C(36)	6243(19)	4484(15)	4128(10)	43(6)
C(41)	9328(16)	2751(13)	5102(8)	26(5)
C(42)	9276(16)	3615(13)	4883(8)	25(5)
C(43)	10171(17)	3958(13)	4667(9)	32(5)
C(44)	11138(19)	3408(15)	4690(10)	49(6)
C(45)	11230(20)	2555(15)	4906(10)	53(6)
C(46)	10349(19)	2225(15)	5135(10)	45(6)
C(51)	8062(15)	3697(12)	5978(8)	23(4)
C(52)	7108(17)	4293(13)	5952(9)	28(5)
C(53)	7473(19)	5137(14)	6303(10)	42(6)
C(54)	8621(20)	5327(15)	6636(10)	46(6)
C(55)	9543(18)	4773(13)	6665(9)	35(5)
C(56)	9270(17)	3966(13)	6328(9)	33(5)

Table 15 (Cont'd)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>	
C(61)	7542(21)	1502(18)	6891(11)	67(7)
C(62)	6437(19)	979(14)	6487(10)	44(5)
C(62)	5481(22)	1538(17)	6266(11)	65(7)
C(64)	5826(22)	2376(17)	6482(12)	71(7)
C(65)	7081(19)	2324(16)	6857(10)	52(6)

\* Anisotropic temperature factors listed in Table 23.

Table 16. Atom Coordinates ( $\times 10^4$ ) and Equivalent Isotropic Temperature Factors ( $\text{\AA}^2 \times 10$ ) for  $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (2)

	<u>X</u>	<u>Y</u>	<u>Z</u>	<u>Biso</u>
Ni	3775.9(9)	2816.8(11)	913.4(4)	35.2(5)*
Co	1963.9(10)	4470.4(11)	969.6(4)	31.8(5)*
Ru	1484.3(7)	2061.6(7)	546.4(2)	36.7(3)*
C(1)	1639(8)	4291(8)	248(3)	40(4)*
O(1)	1579(6)	4819(6)	9829(2)	54(4)*
C(2)	609(9)	5620(9)	1144(3)	48(5)*
O(2)	-220(7)	6384(7)	1256(3)	78(5)*
C(3)	3272(9)	5708(9)	943(3)	44(5)*
O(3)	4053(7)	6567(7)	923(2)	66(4)*
C(4)	1738(9)	197(9)	818(3)	51(5)*
O(4)	1917(8)	-878(6)	985(3)	85(6)*
C(5)	-386(9)	1895(9)	386(3)	47(5)*
O(5)	-1496(6)	1806(7)	285(3)	68(4)*
C(6)	2040(10)	1644(9)	-186(3)	52(6)*
O(6)	2366(8)	1412(8)	-612(2)	78(5)*
C(7)	1304(7)	2669(7)	1367(3)	31(1)
C(8)	2606(7)	2967(8)	1516(3)	30(1)
C(71)	132(7)	2538(7)	1730(3)	31(2)
C(72)	-667(8)	1357(8)	1694(3)	40(2)
C(73)	-1716(9)	1125(10)	2062(3)	51(2)
C(74)	-1943(9)	2113(10)	2462(3)	55(2)

Table 16 (Cont'd)

	<u>X</u>	<u>Y</u>	<u>Z</u>	<u>Biso</u>
C(75)	-1156(9)	3281(9)	2504(3)	52(2)
C(76)	-115(8)	3514(8)	2136(3)	41(2)
C(81)	3194(7)	3043(8)	2066(3)	31(1)
C(82)	4167(8)	4011(8)	2196(3)	37(2)
C(83)	4754(9)	4043(9)	2714(3)	48(2)
C(84)	4308(9)	3117(9)	3091(3)	52(2)
C(85)	3363(9)	2152(10)	2981(3)	52(2)
C(86)	2772(8)	2077(9)	2462(3)	44(2)
C(91)	4951(9)	7196(12)	4712(3)	68(8)*
C(92)	4971(10)	6205(11)	4310(5)	86(9)*
C(93)	4461(9)	6806(12)	3846(4)	74(8)*
C(94)	4128(9)	8154(12)	3954(4)	71(8)*
C(95)	4434(9)	8392(10)	4508(4)	64(6)

\* Anisotropic thermal parameters listed in Table 24.

Table 17. Atom Coordinates ( $\times 10^4$ ) for  $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (3)

	<u>X</u>	<u>Y</u>	<u>Z</u>
Ni	1209.3(9)	2177.7(11)	917.1(4)
Co	3032.2(9)	538.3(11)	965.8(4)
Os	3510.3(3)	2956.7(3)	544.3(2)
C(1)	3379(7)	710(9)	235(3)
O(1)	3346(6)	163(6)	-176(2)
C(2)	5392(9)	3119(9)	388(3)
O(2)	6494(6)	3225(7)	295(3)
C(3)	3276(9)	4382(10)	802(3)
O(3)	3104(8)	5937(7)	967(3)
C(4)	2932(9)	3370(9)	-190(3)
O(4)	2624(8)	3586(8)	-617(2)
C(5)	1711(8)	-712(9)	937(3)
O(5)	915(7)	-1548(7)	920(3)
C(6)	4374(9)	-624(9)	1154(3)
O(6)	5210(7)	-1372(8)	1265(3)
C(7)	1805(7)	1934(8)	2071(3)
C(8)	2204(8)	2873(8)	2472(3)
C(9)	1641(10)	2831(10)	2987(3)
C(10)	675(9)	1849(10)	3098(3)
C(11)	246(8)	926(10)	2701(3)
C(12)	806(7)	954(9)	2201(3)
C(13)	2379(7)	2007(7)	1520(3)
C(14)	3679(7)	2314(8)	1371(3)

Table 17 (Cont'd)

	<u>X</u>	<u>Y</u>	<u>Z</u>
C(15)	4843(7)	2473(8)	1737(3)
C(16)	5118(8)	1518(9)	2141(3)
C(17)	6163(9)	1727(10)	2511(4)
C(18)	6949(9)	2890(10)	2468(4)
C(19)	6721(8)	3854(10)	2067(3)
C(20)	5683(8)	3636(9)	1707(3)
C(21)	-550(9)	3179(12)	1160(4)
C(22)	-906(8)	1830(12)	1046(4)
C(23)	-613(9)	1620(11)	489(4)
C(24)	-89(9)	2847(13)	303(4)
C(25)	-51(10)	3801(12)	713(5)

Table 18. Atom Coordinates ( $\times 10^4$ ) for  $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (4)

	<u>X</u>	<u>Y</u>	<u>Z</u>
Ni	1867.4(8)	2196.5(6)	937.7(5)
Co	3918.4(9)	3390.7(6)	1366.1(5)
Fe	3296.0(9)	2527.4(7)	-64.1(6)
C(1)	4719(7)	2952(6)	-505(4)
O(1)	5660(6)	3210(5)	-777(4)
C(2)	2036(7)	3553(5)	-495(4)
O(2)	1231(5)	4199(4)	-788(3)
C(3)	2593(7)	1559(5)	-849(4)
O(3)	2130(6)	918(4)	-1342(3)
C(4)	4010(7)	3027(5)	2444(4)
O(4)	4060(5)	2773(4)	3123(3)
C(5)	2692(7)	4508(5)	1199(4)
O(5)	1950(5)	5223(4)	1111(4)
C(6)	5556(7)	4116(5)	1432(5)
O(6)	6615(6)	4537(4)	1461(4)
C(7)	3759(8)	-336(5)	1648(5)
C(8)	3509(7)	202(5)	800(4)
C(9)	3512(6)	1400(4)	857(4)
C(10)	4687(6)	2060(4)	1045(3)
C(11)	6270(6)	1754(5)	1219(4)
C(12)	6989(7)	1638(6)	2154(4)
C(13)	605(8)	1833(8)	1778(5)
C(14)	291(8)	1128(6)	1122(6)



Table 18 (Cont'd)

	<u>X</u>	<u>Y</u>	<u>Z</u>
C(15)	-252(7)	1694(7)	388(5)
C(16)	-312(7)	2729(6)	575(5)
C(17)	219(8)	2813(6)	1443(6)

Table 19. Calculated Hydrogen Atom Coordinates for  
 $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$  (1a)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
H(12)	5325	1006	3034
H(13)	6703	107	2535
H(14)	8831	391	2864
H(15)	9785	1764	3354
H(16)	8530	2764	3857
H(22)	4701	3389	2539
H(23)	2499	3235	1824
H(24)	1042	2425	2155
H(25)	1726	1574	3168
H(26)	3926	1760	3908
H(32)	6881	3684	2872
H(33)	7232	5319	2771
H(34)	6856	6419	3490
H(35)	6330	5879	4393
H(36)	5849	4285	4487
H(42)	8339	4065	4880
H(43)	10084	4668	4499
H(44)	11823	3692	4486
H(45)	12062	2173	4975
H(46)	10446	1566	5323
H(52)	6118	4152	5675
H(53)	6786	5618	6293

Table 19 (Cont'd)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
H(54)	8885	5952	6900
H(55)	9468	4923	6948
H(56)	10	3507	6330
H(61)	8447	1203	7121
H(62)	6304	226	6374
H(63)	4506	1366	5921
H(64)	5257	2992	6402
H(65)	7634	2909	7100

Table 20. Calculated Hydrogen Atom Coordinates ( $\times 10^4$ ) for  
 $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (2)

	<u>X</u>	<u>Y</u>	<u>Z</u>
H(72)	-481	604	1374
H(73)	-2324	173	2052
H(74)	-2772	2043	2743
H(75)	-1315	4035	2839
H(76)	524	4422	2172
H(82)	4476	4760	1893
H(83)	5540	4809	2844
H(84)	4652	3142	3522
H(85)	3150	1295	3276
H(86)	1983	1291	2375
H(91)	5267	6983	5124
H(92)	5358	5092	4344
H(93)	4292	6230	3463
H(94)	3741	8815	3630
H(95)	4310	9364	4721

Table 21. Calculated Hydrogen Atom Coordinates ( $\times 10^4$ ) for  
 $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (3)

	<u>X</u>	<u>Y</u>	<u>Z</u>
H(8)	2898	3576	2391
H(9)	1932	3498	3265
H(10)	283	1799	3460
H(11)	-466	245	2782
H(12)	506	277	1928
H(16)	4565	668	2170
H(17)	6327	1039	2798
H(18)	7686	3038	2724
H(19)	7288	4694	2037
H(20)	5533	4329	1421
H(21)	-637	3631	1513
H(22)	-1290	1143	1296
H(23)	-762	756	281
H(24)	213	3019	-66
H(25)	280	4770	687

Table 22. Calculated Hydrogen Atom Coordinates ( $\times 10^4$ ) for  
 $\text{CpNiCoFe}(\text{CO})_6(\text{C}_2\text{Et}_2)$  (4)

	<u>X</u>	<u>Y</u>	<u>Z</u>
H(13)	944	1689	2244
H(14)	447	550	1152
H(15)	-653	1509	-250
H(16)	-654	3200	248
H(17)	289	3454	1783
H(71)	4647	9787	2065
H(72)	3752	8968	1538
H(73)	2876	9831	1935
H(81)	4244	10010	526
H(82)	2641	9943	421
H(111)	6820	12339	1001
H(112)	6340	11109	922
H(121)	6422	11222	2394
H(122)	8049	11335	2314
H(123)	7060	12339	2381

Table 23. Anisotropic Thermal Parameters ( $\text{\AA}^2 \times 10$ ) for  
 $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$  (1a)

<u>Atom</u>	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
Co	41.9(21)	41.0(20)	35.4(19)	5.8(17)	19.9(16)	-1.1(17)
Fe	50.6(22)	36.6(21)	37.4(19)	-3.2(20)	21.3(17)	-0.2(2)
Ni	58.3(21)	49.4(20)	32.9(18)	-6.9(18)	24.2(16)	-4.2(17)
P	56(5)	42(4)	38(4)	10(4)	26(4)	5(4)

Table 24. Anisotropic Thermal Parameters ( $\text{\AA}^2 \times 10$ ) for  
 $\text{CpNiCoRu(CO)}_6(\text{C}_2\text{Ph}_2)$  (2)

	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
Ni	38.3(6)	56.1(7)	39.3(6)	8.0(6)	6.8(4)	5.8(6)
Co	39.3(6)	38.4(6)	43.0(6)	0.8(5)	0.1(5)	1.8(5)
Ru	51.3(4)	44.7(4)	43.6(4)	0.0(4)	1.3(3)	-6.7(4)
C(1)	47(5)	45(5)	59(5)	-2.0(4)	-2.0(4)	-5.0(4)
O(1)	89(5)	61(4)	57(4)	-10(4)	-14(4)	11(4)
C(2)	78(6)	56(6)	46(5)	9(5)	-2(4)	3(4)
O(2)	105(6)	94(5)	97(5)	51(5)	6(4)	-10(4)
C(3)	76(6)	58(6)	34(5)	-17(5)	-3(4)	2(4)
O(3)	92(5)	84(5)	74(5)	-43(4)	-8(4)	8(4)
C(4)	78(7)	49(6)	66(6)	-2(5)	13(5)	-16(5)
O(4)	159(7)	51(4)	106(5)	27(5)	22(5)	14(4)
C(5)	67(6)	49(6)	63(5)	-3(5)	-9(5)	-6(5)
O(5)	53(4)	93(5)	114(5)	-1(4)	-18(4)	-20(4)
C(6)	83(7)	57(6)	56(5)	-6(5)	0(4)	-1(5)
O(6)	137(7)	107(6)	51(4)	24(5)	14(4)	-12(4)
C(91)	55(6)	151(10)	50(5)	-8(7)	20(5)	21(6)
C(92)	65(7)	71(7)	139(12)	-27(6)	48(7)	10(8)
C(93)	59(6)	139(10)	82(7)	-48(7)	13(5)	-41(7)
C(94)	37(5)	139(10)	95(7)	-2(6)	5(5)	51(7)
C(95)	50(6)	92(8)	102(7)	-14(5)	41(5)	-45(6)



Table 25. Anisotropic Thermal Parameters ( $\text{\AA}^2 \times 10$ ) for  
 $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$  (3)

	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
Ni	32.6(5)	49.6(7)	36.8(5)	7.0(5)	-8.6(4)	-4.9(5)
Co	30.5(5)	30.4(6)	41.6(6)	0.4(5)	-2.8(4)	-1.4(5)
Os	41.9(2)	38.3(2)	39.9(2)	-0.07(2)	-3.7(1)	6.7(2)
C(1)	28(4)	48(5)	69(6)	1(4)	-3(4)	4(5)
O(1)	79(4)	67(4)	43(3)	-6(4)	10(3)	-20(3)
C(2)	67(6)	45(5)	52(5)	-5(5)	-2(4)	12(4)
O(2)	46(4)	90(6)	109(5)	-13(4)	14(4)	24(4)
C(3)	66(6)	56(6)	50(5)	11(5)	-13(4)	3(5)
O(3)	153(7)	54(5)	83(5)	20(5)	-7(4)	-12(4)
C(4)	73(6)	53(6)	52(5)	19(5)	-4(4)	3(5)
O(4)	124(6)	98(6)	44(4)	30(5)	-19(4)	9(4)
C(5)	63(6)	47(5)	38(5)	-6(5)	12(4)	2(4)
O(5)	86(5)	78(5)	75(4)	-51(4)	3(4)	1(4)
C(6)	65(6)	47(6)	55(5)	7(5)	-4(4)	-1(4)
O(6)	79(5)	85(5)	114(6)	48(5)	-12(4)	9(5)
C(7)	34(4)	51(5)	34(4)	14(4)	-6(3)	3(4)
C(8)	57(5)	40(5)	50(5)	8(4)	-4(4)	-12(4)
C(9)	87(7)	65(7)	41(5)	21(6)	-3(4)	-16(5)
C(10)	71(6)	81(8)	43(5)	31(6)	9(4)	1(5)
C(11)	46(5)	68(6)	47(5)	4(5)	4(4)	10(4)
C(12)	40(5)	52(5)	40(5)	2(4)	-5(4)	-1(4)

Table 25 (Cont'd)

	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
C(13)	32(4)	35(5)	34(4)	-2(4)	-7(3)	-3(4)
C(14)	33(4)	36(4)	39(4)	7(4)	-2(3)	1(3)
C(15)	38(4)	32(4)	36(4)	-3(4)	-4(3)	-4(3)
C(16)	42(5)	48(5)	60(5)	-6(4)	-16(4)	-4(4)
C(17)	64(6)	59(6)	63(6)	15(5)	-28(5)	-5(5)
C(18)	50(5)	70(7)	74(6)	8(5)	-28(5)	-25(5)
C(19)	34(4)	68(6)	66(6)	-10(5)	-3(4)	-21(5)
C(20)	45(5)	43(5)	56(5)	-10(4)	-7(4)	-2(4)
C(21)	54(6)	119(10)	66(6)	36(6)	-16(5)	-27(6)
C(22)	26(5)	125(10)	93(7)	1(5)	-10(5)	45(7)
C(23)	44(5)	83(8)	94(7)	8(5)	-32(5)	-30(6)
C(24)	45(5)	133(11)	67(6)	21(6)	-18(5)	8(7)
C(25)	65(7)	66(7)	150(10)	10(6)	-53(7)	8(7)

Table 26. Anisotropic Thermal Parameters ( $\text{\AA}^2 \times 10$ ) for  
 $\text{CpNiCoFe(CO)}_6(\text{C}_2\text{Et}_2)$  (4)

	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
Ni	33.2(4)	48.1(5)	50.4(5)	-7.3(4)	12.9(4)	-4.3(4)
Co	37.3(5)	41.0(5)	59.9(6)	-5.4(4)	9.3(4)	-5.7(4)
Fe	39.3(5)	48.8(5)	44.7(5)	-1.1(4)	11.0(4)	7.2(4)
C(1)	62(5)	75(5)	65(5)	2(4)	18(4)	21(4)
O(1)	85(4)	150(5)	124(5)	-1(4)	61(4)	49(4)
C(2)	53(4)	69(5)	63(5)	3(4)	19(4)	13(4)
O(2)	72(3)	90(4)	102(4)	29(3)	22(3)	37(3)
C(3)	60(4)	65(5)	47(4)	7(4)	5(4)	11(4)
O(3)	111(4)	82(4)	71(4)	0(3)	-8(3)	-17(3)
C(4)	47(4)	62(4)	65(5)	-5(4)	13(4)	-14(4)
O(4)	85(4)	132(5)	56(3)	5(3)	12(3)	-1(3)
C(5)	54(4)	56(4)	83(5)	-8(4)	12(4)	-11(4)
O(5)	73(3)	63(3)	160(6)	21(3)	29(4)	3(3)
C(6)	56(5)	53(5)	109(6)	-12(4)	22(4)	-8(4)
O(6)	79(4)	96(4)	189(6)	-47(3)	42(4)	-12(4)
C(7)	109(7)	53(5)	96(6)	-2(4)	34(5)	21(4)
C(8)	58(4)	48(4)	63(5)	-2(3)	21(4)	4(4)
C(9)	47(4)	35(3)	38(4)	4(3)	13(3)	2(3)
C(10)	39(4)	43(4)	38(4)	1(3)	9(3)	0(3)
C(11)	40(4)	57(4)	61(5)	2(3)	14(3)	1(3)
C(12)	46(4)	100(6)	61(5)	11(4)	-2(4)	0(4)

Table 26 (Cont'd)

	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
C(13)	51(5)	193(9)	69(6)	-8(6)	28(4)	24(6)
C(14)	63(5)	59(5)	187(9)	-16(4)	66(6)	8(6)
C(15)	44(4)	150(8)	100(6)	-49(5)	28(4)	-56(6)
C(16)	33(4)	107(7)	128(8)	3(4)	21(4)	35(6)
C(17)	55(5)	108(7)	144(8)	-31(5)	55(5)	-69(6)

Table 27. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
 CpNiCoFe(CO)<sub>5</sub>(PPh<sub>3</sub>)(C<sub>2</sub>Ph<sub>2</sub>) (1a)

<u>Plane No.</u>	<u>Atoms Defining Plane<sup>b</sup></u>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D (Å)</u>	<u>χ<sup>2</sup></u>
1	Ni, Co, Fe	-0.7095	-0.6544	-0.2163	-7.3430	0.00
2 <sup>c</sup>	Cp ring: C61(0.2), C62(-1.1), C63(1.6), C64(-1.4), C65(0.7)	0.5630	0.1894	-0.8045	-9.5132	0.01
3 <sup>d</sup>	Ni(-2.4), Fe(2.3), C6(-4.3), C7(4.4)	-0.5472	0.1110	-0.8296	-8.4123	1.35
4	C6(-39), C7(41), C41(16), C51(-18)	-0.3994	0.4157	-0.8171	-9.7276	12.8

<sup>a</sup> Plane coordination is  $Ax + By + Cz - D = 0$ , with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviations of the atoms ( $\text{Å} \times 10^3$ ) from the meanplane are given in parentheses.

<sup>c</sup> Ni deviation 1.752 Å.

<sup>d</sup> Co deviation 1.766 Å.

Table 27 (Cont'd)

## Dihedral Angles (Between Normals to Meanplanes)

<u>Plane No.</u>	<u>Plane No.</u>	<u>Angle (deg)</u>
1	2	71.8
1	3	57.8
1	4	77.0
2	3	67.6
2	4	59.3
3	4	19.5

Table 28. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
CpNiCoRu(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (2)

<u>Plane No.</u>	<u>Atoms Defining Plane<sup>b</sup></u>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D</u>	<u>χ<sup>2</sup></u>
1	Ru, Ni, Co	0.2452	0.3614	-0.8996	-0.1513	0.00
2 <sup>c</sup>	Ru(-15), Ni(18), C7(30), C8(-32)	-0.1777	0.9472	-0.2670	1.2479	
3	Ph ring: C71(-.7), C72(-2), C73(-2), C74(10), C75(-12), C76(8)	-0.6270	0.4939	-0.6024	-1.4687	4.41
4	Ph ring: C81(-4), C82(-4), C83(11), C84(-11), C85(2), C86(5)	0.7067	-0.6441	-0.29286	-1.1222	4.32
5	Cp ring: C91(-3), C92(1), C93(2), C94(-3), C95(4)	0.9201	-0.2964	-0.2559	3.8286	0.42
6	Ru, Co, Cl	0.9788	-0.1321	-0.1565	0.9731	0.00
7	C7(-33), C8(33), C81(-14), C71(14)	-0.1596	0.9870	-0.0185	-2.268	

<sup>a</sup> Plane equation is  $Ax + By + Cz - D = 0$ , with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviations of the atoms ( $\text{Å} \times 10^3$ ) from the meanplane are given in parentheses.

<sup>c</sup> Co deviation 1.790 Å.

<sup>d</sup> Ni deviation 1.747 Å.

Table 28 (Cont'd)

## Dihedral Angles (Between Normals to Meanplanes)

<u>Plane No.</u>	<u>Plane No.</u>	<u>Angle (deg)</u>
1	2	57.4
1	3	55.5
1	4	78.3
1	5	69.6
1	6	70.6
1	7	70.5
2	3	42.3
2	4	131.1
2	5	112.1
2	6	104.9
3	4	125.8
3	5	124.7
3	6	125.8
4	5	23.6
4	6	34.7
5	6	11.5



Table 29. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
CpNiCoOs(CO)<sub>6</sub>(C<sub>2</sub>Ph<sub>2</sub>) (3)

<u>Plane No.</u>	<u>Atoms Defining Plane<sup>b</sup></u>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D</u>	<u>χ<sup>2</sup></u>
1	Ni, Co, Os	-0.2478	-0.3582	-0.9002	-3.0869	0.00
2 <sup>c</sup>	Os(-15), Ni(17), Cl3(-31), Cl4(29)	0.1804	-0.9445	-0.2746	-2.3806	
3	Ph ring: C7(4), C8(-3), C9(8), Cl0(-7), Cl1(2), Cl2(-3)	-0.6741	0.6710	-0.3089	-1.6728	3.21
4	Ph ring: Cl5(-9), Cl6(7), Cl7(-2), Cl8(-3), Cl9(2), Cl20(4)	0.6212	-0.4989	-0.6043	-0.7805	2.67
5 <sup>d</sup>	Cp ring: C21(2), C22(-3), C23(3), C24(-1), C25(-0.6)	-0.9184	0.3113	-0.2442	0.7425	0.28
6	Os, Co, Cl	-0.9766	0.1226	-0.1765	-3.3026	0.00
7	C7(15), Cl3(-28), Cl4(31), Cl5(-14)	-0.1738	0.9848	0.0088	-1.5291	22.0

<sup>a</sup> Plane equation is  $Ax + By + Cz - D = 0$ , with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviation of the atoms ( $\text{\AA} \times 10^3$ ) from the meanplane are given in parentheses.

<sup>c</sup> Co deviation 1.783  $\text{\AA}$ .

<sup>d</sup> Ni deviation 1.755  $\text{\AA}$ .

Table 29 (Cont'd)

## Dihedral Angles (Between Normals to Meanplanes)

<u>Plane No.</u>	<u>Plane No.</u>	<u>Angle (deg)</u>
1	2	57.3
1	3	78.2
1	4	55.3
1	5	70.3
1	6	69.1
1	7	71.5
2	3	132.1
2	4	41.5
2	5	112.9
2	6	104.1
3	4	104.7
3	5	25.4
3	6	37.3
4	5	125.3
4	6	124.1
5	6	12.0

Table 30. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
CpNiCoFe(CO)<sub>6</sub>(C<sub>2</sub>Et<sub>2</sub>) (4)

<u>Plane No.</u>	<u>Atoms Defining Plane<sup>b</sup></u>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D</u>	<u>χ<sup>2</sup></u>
1	Ni, Co, Fe	-0.5122	0.7746	-0.3710	0.8783	0.00
2 <sup>c</sup>	Cp ring: C13(8), C14(-8), C15(6), C16(-0.7), C17(-4)	0.9842	0.1234	-0.1267	-0.2096	2.84
3 <sup>d</sup>	Ni(17), Co(-16), C9(-30), C10(29)	0.0844	0.3013	-0.9498	-0.4617	
4	Fe, C9, C10	0.6173	-0.5950	-0.5147	0.1172	0.00
5	C8(-3), C9(6)  C10(-6), C11(3)	0.2090	0.0670	-0.9756	-0.5709	2.99

<sup>a</sup> Plane equation is  $Ax + By + Cz - D = 0$ , with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviations of the atoms ( $\text{\AA} \times 10^3$ ) from the meanplane are given in parentheses.

<sup>c</sup> Ni deviation 1.750  $\text{\AA}$ .

<sup>d</sup> Fe deviation 1.786  $\text{\AA}$ .

Table 30 (Cont'd)

## Dihedral Angles (Between Normals to Meanplanes)

<u>Plane No.</u>	<u>Plane No.</u>	<u>Angle (deg)</u>
1	2	68.8
1	3	57.2
1	4	125.9
1	5	72.2
2	3	76.1
2	4	53.2
2	5	70.3
3	4	68.8
3	5	15.3
4	5	53.8

## CHAPTER 2

X-RAY CRYSTAL STRUCTURE OF  $\text{CoMn(CO)}_6(\text{MeC}_2\text{Ph})_2$ 2.1 Introduction

The formation of  $\text{CoMn(CO)}_6(\text{RC}_2\text{R}')$  ( $\text{R}=\text{Ph}, \text{R}'=\text{Me}, \text{Ph}$ ) by the thermal reaction of  $\text{CoMn(CO)}_9$  with an equimolar amount of the appropriate alkyne was recently reported<sup>66</sup>. This is a facile reaction requiring conditions similar to that of the formation<sup>67</sup> of  $\text{Co}_2(\text{CO})_8(\text{RC}_2\text{R}')$  from  $\text{Co}_2(\text{CO})_8$  and  $\text{C}_2\text{Ph}_2$  but less severe than the reaction<sup>68</sup> of  $\text{Mn}_2(\text{CO})_{10}$  with  $\text{C}_2\text{Ph}_2$ . Subsequently, when two moles of alkyne are reacted with one mole of  $\text{CoMn(CO)}_9$ , compounds of the formula  $\text{CoMn(CO)}_6(\text{RC}_2\text{R}')_2$  are isolated as the major product. On the basis of the spectroscopic data<sup>69</sup> a 'ferrole' type structure was proposed.

Ferroles or tricarbonylferracyclopentadieneirontricarbonyl complexes are among the more common products of reactions of iron carbonyls with alkynes. Over eighty compounds have been reported and there are fourteen X-ray diffraction investigations on this structural type<sup>70</sup>. Other metals, notably rhodium<sup>71</sup>, iridium<sup>72,73</sup>, osmium<sup>74,75</sup> and cobalt<sup>76</sup> also form ferrole type structures but these compounds are far less common than those with iron. Structures where there are two different metals attached to the butadiene fragment are noticeably scarce, only  $(\text{CO})_3\text{Fe}(\text{CH}_3\text{C}_2\text{CH}_3)_2\text{Ni}(\text{C}_4(\text{CH}_3)_4)$  and  $\text{CpW(CO)}_2(\text{CF}_3\text{C}_2\text{CF}_3)_2\text{Co(CO)}_2$  being structurally characterized<sup>77,78</sup>. Very recently,  $\text{FeRu(CO)}_6(\text{PhC}_2\text{Ph})_2$  was synthesized but an X-ray structure was not reported<sup>30</sup>. Because of our interest in mixed metal alkyne chemistry and the lack of heterobimetallic ferrole type structures, the X-ray structure determination of  $\text{CoMn(CO)}_6(\text{MeC}_2\text{Ph})_2$  (10) was undertaken.

## 2.2 X-ray Data Collection and Refinement

A large hexagonal red-brown crystal obtained from  $\text{CHCl}_3$ -hexanes, having dimensions 0.35-0.40 mm in widths and 0.21 mm thick was used for the data collection. This crystal was supplied by P.J. Manning. Preliminary oscillation photos (Cu radiation) indicated the crystal to be of good quality. The crystal was transferred to an Enraf-Nonius diffractometer (Mo radiation) and the cell dimensions and space group were determined. Data were collected with a scan rate of  $1.0\text{-}6.0^\circ \text{ min}^{-1}$  for reflections with  $3^\circ \leq 2\theta \leq 45^\circ$  and a symmetrical scan width of  $\pm 0.45^\circ$  plus dispersion correction from the peak center. Four standard reflections were measured every 1 hour of X-ray exposure time, two to monitor intensity and two to monitor crystal movement. Neither pair showed any significant deviations. Absorption corrections were applied via an empirical correction (transmission factors 0.848-0.998). 2373 reflections out of a total of 2939 had  $I \geq 3.0\sigma(I)$  and were regarded as observed.

The positions of the metal atoms were located from a Patterson map and the non-hydrogen atoms were located from difference Fourier synthesis, interspersed with least squares refinement. All the hydrogen atoms were placed in geometrically calculated positions with  $d(\text{C-H})$  equal to 0.98Å. The phenyl hydrogens were given isotropic thermal parameters of  $0.06\text{\AA}^2$  and the methyl hydrogens,  $0.07\text{\AA}^2$ . All non-hydrogen atoms were refined anisotropically. The largest peak in a final difference map was  $0.21(12) \text{ e}\text{\AA}^{-3}$  and was located near the Mn atom. Computer programs are from reference 65. A summary of the crystal data is in Table 31, bond distances and angles in Table 32, atom coordinates in Table 33, anisotropic thermal parameters in Table 34, calculated hydrogen atom coordinates in Table 35 and calculated

Table 31. Summary of Crystal Data for  $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$  (10)

formula	$\text{CoMnC}_{24}\text{H}_{16}\text{O}_6$
mol. wt.	514.3
space group	$P2_1/c$
a, Å	16.308(3)
b, Å	8.875(2)
c, Å	15.574(4)
$\beta$ , deg	93.44(2)
Z	4
cell vol, Å <sup>3</sup>	2250
$\rho$ calcd, $\text{gcm}^{-3}$	1.516
$\rho$ obsd, $\text{gcm}^{-3}$	1.53
$\mu$ ( $\text{MoK}\alpha$ ) $\text{cm}^{-1}$	13.1
total reflections, measured	2939
no. reflections, $I \geq 3.0\sigma(I)$	2373
$2\theta$ limits, deg	3, 45
final R	3.34%
final $R_w$	4.28%
GOF	1.56

Table 32. Bond Distances (Å) and Angles (deg) for  
 $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$  (10)

Co-Mn	2.5488(8)	C(7)-C(21)	1.504(5)
Co-C(1)	1.823(4)	C(7)-C(8)	1.401(5)
Co-C(2)	1.797(4)	C(8)-C(81)	1.513(5)
Co-C(3)	1.829(4)	C(8)-C(9)	1.426(5)
Co-C(7)	1.977(3)	C(9)-C(10)	1.423(5)
Co-C(10)	1.977(3)	C(9)-C(91)	1.506(5)
Mn-C(4)	1.785(4)	C(10)-C(11)	1.486(5)
Mn-C(5)	1.798(4)	C(1)-O(1)	1.127(5)
Mn-C(6)	1.798(4)	C(2)-O(2)	1.132(5)
Mn-C(7)	2.100(3)	C(3)-O(3)	1.124(5)
Mn-C(8)	2.185(3)	C(4)-O(4)	1.152(5)
Mn-C(9)	2.170(3)	C(5)-O(5)	1.149(5)
Mn-C(10)	2.098(3)	C(6)-O(6)	1.145(5)
Co-C(1)-O(1)	179.5(4)	Mn-C(7)-C(8)	74.2(2)
Co-C(2)-O(2)	176.6(4)	Mn-C(7)-C(21)	134.2(2)
Co-C(3)-O(3)	176.0(4)	Mn-C(10)-C(9)	73.3(2)
Mn-C(4)-O(4)	178.6(4)	Mn-C(10)-C(11)	135.6(2)
Mn-C(5)-O(5)	177.7(4)	C(8)-C(7)-C(21)	120.6(3)
Mn-C(6)-O(6)	179.2(4)	C(81)-C(8)-C(7)	123.7(3)
Co-C(7)-C(8)	113.7(2)	C(81)-C(8)-C(9)	121.6(3)
Co-C(7)-C(21)	122.2(2)	C(91)-C(9)-C(8)	121.5(3)
Co-C(10)-C(9)	113.5(2)	C(91)-C(9)-C(10)	124.8(3)
Co-C(10)-C(11)	121.9(2)	C(9)-C(10)-C(11)	120.9(3)



Table 33. Atomic Coordinates ( $\times 10^4$ ) for  
 $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$  (10)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
Co	1851.5(3)	5311.1(5)	4384.2(3)
Mn	2933.1(3)	7311.8(6)	4765.0(3)
C(1)	863(2)	6219(4)	4145(2)
O(1)	254(2)	6786(4)	3993(2)
C(2)	1610(2)	3573(4)	4904(3)
O(2)	1496(2)	2480(3)	5252(2)
C(3)	1878(2)	4736(5)	3259(3)
O(3)	1846(2)	4371(5)	2566(2)
C(4)	3329(3)	8803(4)	5438(2)
O(4)	3583(3)	9746(4)	5887(2)
C(5)	3779(2)	7555(4)	4090(2)
O(5)	4335(2)	7731(3)	3682(2)
C(6)	2294(2)	8579(4)	4122(2)
O(6)	1895(2)	9393(4)	3708(2)
C(7)	1992(2)	6432(4)	5480(2)
C(8)	2727(2)	6101(4)	5954(2)
C(81)	2913(2)	6595(4)	6875(2)
C(9)	3318(2)	5315(4)	5485(2)
C(91)	4157(2)	4944(4)	5882(2)
C(10)	3041(2)	4970(4)	4625(2)
C(11)	3518(2)	3946(4)	4089(2)
C(12)	3707(2)	4281(4)	3251(2)

Table 33 (Cont'd)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
C(13)	4155(2)	3294(5)	2775(2)
C(14)	4422(3)	1948(5)	3133(3)
C(15)	4237(3)	1578(4)	3953(3)
C(16)	3782(3)	2571(4)	4430(2)
C(21)	1279(2)	7111(4)	5910(2)
C(22)	883(2)	6257(4)	6504(2)
C(23)	209(2)	6827(5)	6900(2)
C(24)	-79(2)	8232(5)	6699(3)
C(25)	302(3)	9093(5)	6119(3)
C(26)	984(2)	8544(4)	5271(2)

Table 34. Anisotropic Thermal Parameters ( $\text{\AA}^2 \times 10$ ) for  
 $\text{CoMn(CO)}_6(\text{MeC}_2\text{Ph})_2$  (10)

<u>Atom</u>	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
Co	33.3(3)	41.0(3)	36.9(3)	-0.2(2)	1.0(2)	-5.0(2)
Mn	38.3(3)	32.4(3)	31.1(3)	-0.2(2)	4.9(2)	1.1(2)
C(1)	45(2)	54(3)	48(2)	6(2)	-3(2)	-15(2)
O(1)	55(2)	88(2)	88(2)	26(2)	-17(2)	-22(2)
C(2)	47(2)	54(3)	60(3)	-5(2)	6(2)	-12(2)
O(2)	99(3)	60(2)	115(3)	-16(2)	21(2)	20(2)
C(3)	36(2)	82(3)	54(2)	3(2)	0(2)	-16(2)
O(3)	70(2)	185(4)	49(2)	5(2)	-5(2)	-43(2)
C(4)	85(3)	45(3)	41(2)	-9(2)	6(2)	7(2)
O(4)	169(4)	60(2)	65(2)	-41(2)	-7(2)	-13(2)
C(5)	50(2)	28(2)	51(2)	1(2)	5(2)	-1(2)
O(5)	59(2)	55(2)	86(2)	-6(1)	38(2)	1(2)
C(6)	54(2)	53(3)	42(2)	2(2)	17(2)	6(2)
O(6)	80(2)	83(2)	68(2)	36(2)	10(2)	34(2)
C(7)	39(2)	33(2)	34(2)	-1(2)	7(2)	2(2)
C(8)	43(2)	34(2)	29(2)	2(2)	5(2)	6(2)
C(81)	51(2)	62(3)	31(3)	3(2)	2(2)	-5(2)
C(9)	38(2)	35(2)	33(2)	-2(2)	4(2)	6(2)
C(91)	40(2)	64(3)	41(2)	9(2)	-3(2)	3(2)
C(10)	34(2)	31(2)	37(2)	-3(1)	8(1)	1(2)
C(11)	34(2)	33(2)	40(2)	1(2)	2(2)	-4(2)

Table 34 (Cont'd)

<u>Atom</u>	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
C(12)	38(2)	40(2)	38(2)	4(2)	0(2)	-3(2)
C(13)	47(2)	62(3)	37(2)	1(2)	8(2)	-13(2)
C(14)	59(3)	51(3)	65(3)	13(2)	10(2)	-20(2)
C(15)	76(3)	29(2)	78(3)	13(2)	7(2)	-1(2)
C(16)	62(3)	35(2)	53(2)	1(2)	11(2)	1(2)
C(21)	37(2)	38(2)	35(2)	3(2)	5(2)	-4(2)
C(22)	47(2)	43(2)	51(2)	0(2)	14(2)	3(2)
C(23)	50(2)	61(3)	56(3)	-2(2)	24(2)	3(2)
C(24)	44(2)	73(3)	51(2)	9(2)	15(2)	-12(2)
C(25)	66(3)	54(3)	56(3)	21(2)	9(2)	-5(2)
C(26)	58(2)	48(2)	45(2)	10(2)	18(2)	3(2)

Table 35. Calculated Hydrogen Atom Coordinates ( $\times 10^4$ ) for  
 $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$  (10)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
H(12)	3518	5236	2989
H(13)	4281	3553	2185
H(14)	4745	1261	2795
H(15)	4420	611	4193
H(16)	3648	2289	5009
H(22)	1085	5245	6656
H(23)	- 60	6209	7325
H(24)	-563	8624	6980
H(25)	97	10104	5983
H(26)	1251	9181	5308
H(811)	3404	7212	6850
H(812)	3005	5723	7256
H(813)	2452	7175	7088
H(911)	4366	5924	6097
H(912)	4513	4519	5462
H(913)	4127	4221	6357

Table 36. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
 $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$  (10)

Plane No.	Atoms Defining Plane <sup>b</sup>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D</u>	<u><math>\chi^2</math></u>
1 <sup>c</sup>	C7(4.7), C8(-8.7) C9(8.8), C10(-4.7)	0.4001	0.8662	-0.3225	3.487	10.5
2	C11(6.3), C12(-3.0), C13(-3.2), C14(-5.8), C15(-0.2), C16(-7.3)	0.8284	0.4356	0.3023	8.202	6.58
3	C21(-1.7), C22(-2.1), C23(6.3), C24(-4.9), C25(-0.8), C26(3.6)	0.5606	0.3791	0.7017	10.02	3.28
4	Mn(0.2), Co(0.0), C1(3.3), O1(-1.5)	-0.2611	-0.0668	0.9768	5.561	0.7
5	Co(-0.1), Mn(0.0), C2(11.3), O2(-5.6)	-0.6810	0.4788	0.5935	4.253	7.4
6	Co(-0.1), Mn(0.0), C3(19.3), O3(-12.3)	0.6496	-0.7171	0.2135	0.415	19.1
7	Co(0.0), Mn(0.1), C4(-9.9), O4(6.2)	0.4933	-0.6669	0.5282	1.956	5.3
8	Co(0.0), Mn(0.1), C5(-9.1), O5(3.6)	0.4963	-0.6684	0.5236	1.926	5.8
9	Co(0.0), Mn(0.0), C6(5.2), O6(-2.2)	-0.5354	0.2528	0.8362	5.286	1.8

<sup>a</sup> Plane equation is  $Ax + By + Cz = D$  with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviations of the atoms ( $\text{\AA} \times 10^3$ ) from the meanplane are given in parentheses.

<sup>c</sup> Mn deviation 1.654(2)  $\text{\AA}$ .

meanplanes and dihedral angles in Table 36. A labelled SNOOPI diagram is shown in Figure 16 and a sideview of the molecule in Figure 17.

### 2.3 Description and Discussion of the Structure

The crystal structure consists of the packing of discrete molecules with no unusual inter- or intramolecular contacts. The atom labelling scheme is presented in Figure 16 and a sideview of the molecule, without the substituents on the cobaltacyclopentadiene ring, is illustrated in Figure 17. Hydrogen atoms have been omitted in all diagrams.

The gross features of the structure consist of two methylphenylacetylene ligands linked in a head to head fashion with  $\sigma$ -bonding from the end carbons C(7) and C(10), to the  $\text{Co}(\text{CO})_3$  group forming a cobaltacyclopentadienyl ring. This ring is inclined towards the  $\text{Mn}(\text{CO})_3$  moiety, as shown clearly in Figure 17 and allows the Mn to complete its coordination by  $\pi$ -bonding with the delocalized electron density of the butadiene residue of the cobaltacyclopentadiene ring.

There are few structures with accurate Co-Mn bond lengths in the literature and the distance in 10 of 2.5488(8)Å, while shorter than the covalent radii<sup>79</sup> of 2.65Å, is similar to that found<sup>80</sup> in  $(\text{MeCp})\text{MnCo}_2(\text{Me}_5\text{C}_5)_2(\text{CO})_4$  of 2.561Å. The Co-Mn distance in 10 is slightly longer than the closely related ferrole<sup>81</sup>  $(\text{CO})_3\text{Fe}(\text{PhC}_2\text{Ph})_2\text{Fe}(\text{CO})_3$ , 2.505(1)Å.

The manganese has three terminal carbonyl groups with similar Mn-C bond lengths of 1.785(4) to 1.798(4)Å. The cobalt atom also has three terminal carbonyls but Co-C(2) is significantly ( $4.6\sigma$ ) shorter than the Co-C(1) or Co-C(3) distances; 1.797(4) versus 1.823(4) and 1.829(4)Å. As shown in Figure 17 this carbonyl ligand is almost perpendicular to the

FIGURE 16

SNOOPI Diagram of  
 $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$  (10)



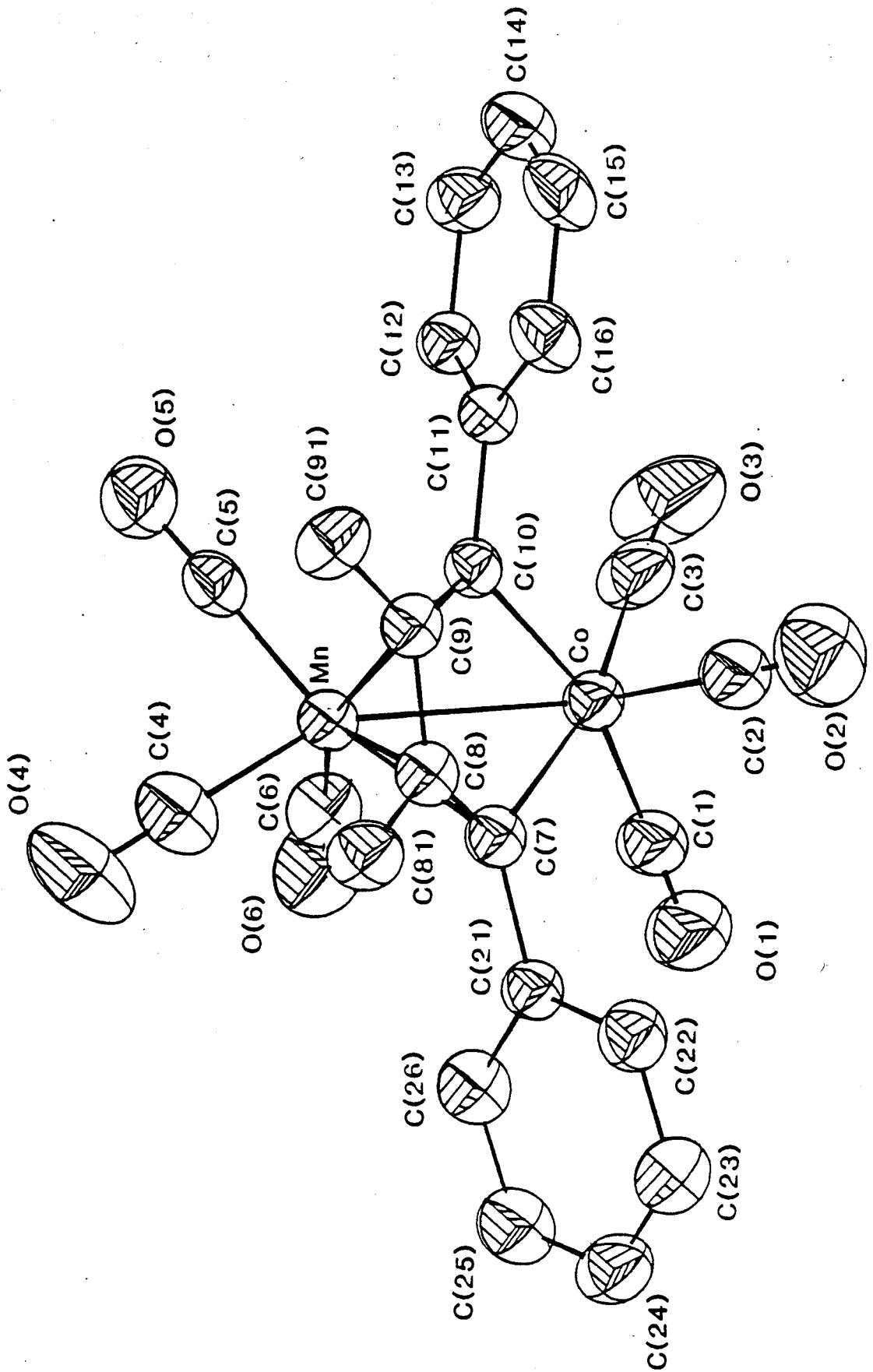
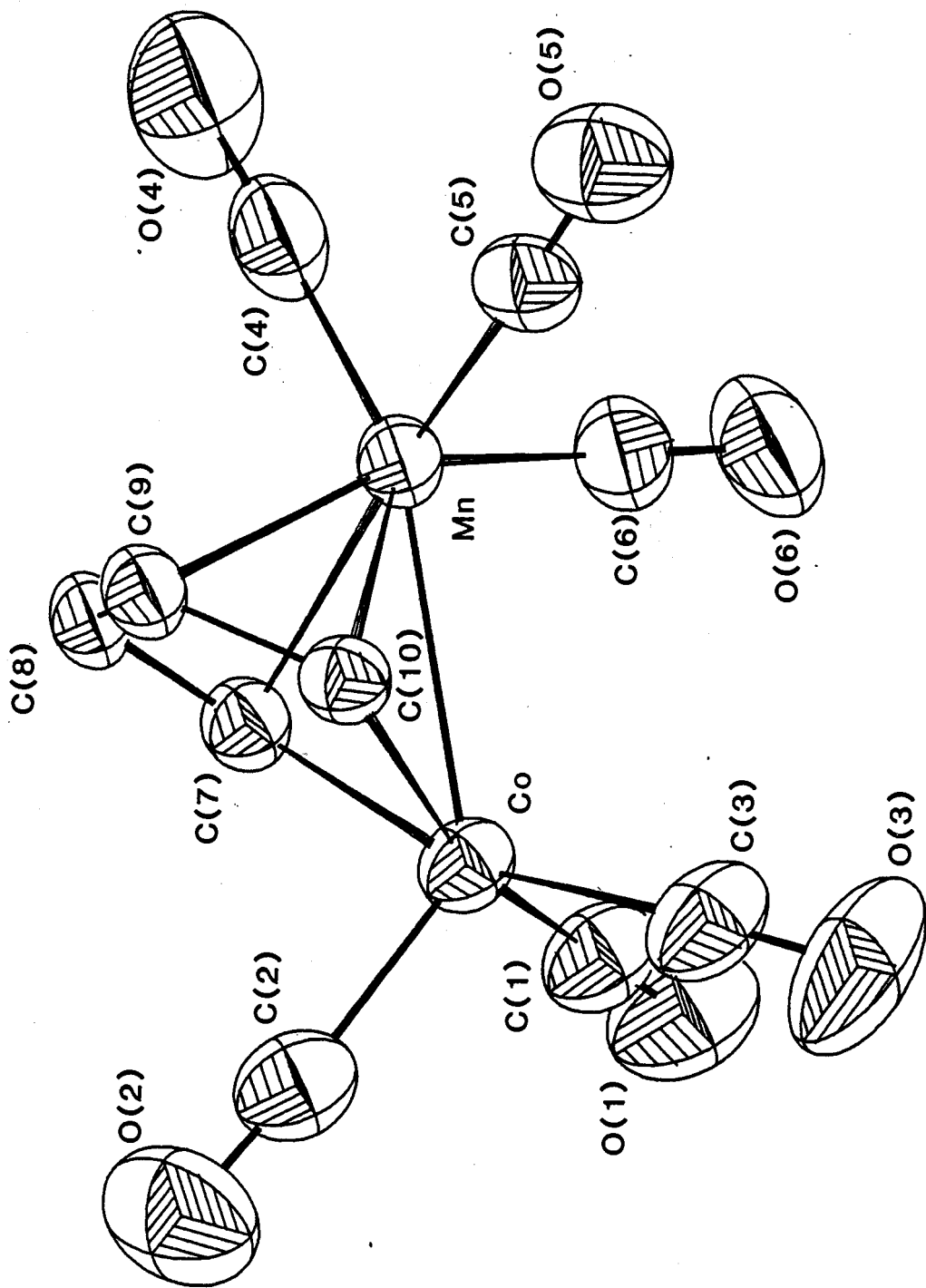
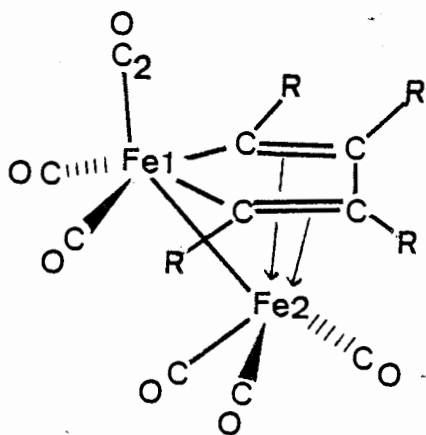


FIGURE 17

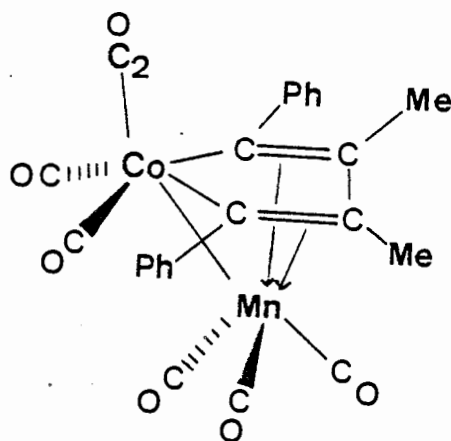
Side-view of  
 $\text{CoMn}(\text{CO})_6(\text{MeC}_2\text{Ph})_2$  (10)



cobaltacyclopentadiene ring and similar shortening effects have been noted for the identically positioned carbonyl in the isoelectronic ferrole series<sup>82,83</sup>. However when there are two iron atoms present in the structure, as shown in II-1 a dative Fe(2)→Fe(1) bond must be invoked to allow each metal to achieve an 18 electron count. Fe(1) can relieve this excess charge by increased back donation to the carbonyl groups of which C(2) is in a unique position to accept more electron density than the other



II-1

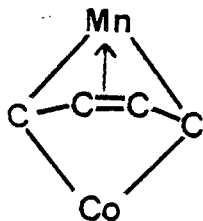


II-2

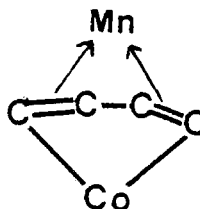
two. Alternatively, it was suggested<sup>84</sup> that the shortening of this Fe(1)-C(2) bond is due to an increased electron donation from this unique carbonyl to Fe(1), thus helping to balance the electron deficiency of the metal. This latter view is supported by the replacement of C(2) by PPh<sub>3</sub>, a stronger  $\sigma$ -donor<sup>83</sup>. In the structure of 10 shown as II-2 a semi-bridging carbonyl is not observed and a dative bond does not have to be invoked since both metals have an 18 electron count. Yet the Co-C(2) bond is still shorter than the other two cobalt-carbon distances. Cobalt is slightly more electronegative than Mn (1.88 versus 1.55)<sup>54</sup> which lends support to increased  $\sigma$ -donation from C(2) to the Co atom as the reason for the observed

small shortening effect.

The cobaltacyclopentadienyl ring donates four electrons to the  $\text{Mn}(\text{CO})_3$  group and two canonical forms of the structure can be written as II-3a and II-3b. Both satisfy the requirement of donating two electrons to the cobalt (2  $\sigma$ -bonds) and four electrons to the Mn (2  $\sigma$  and 1  $\pi$ -bond or 2  $\pi$ -bonds).



II-3a



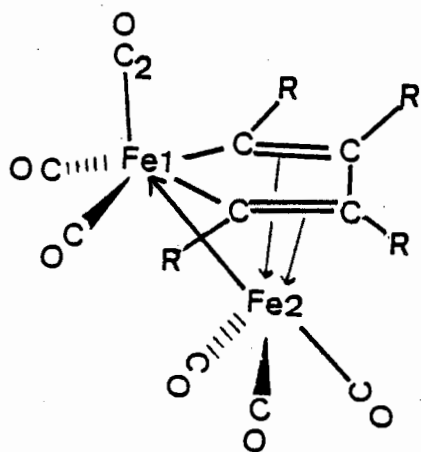
II-3b

Which, if any, of these two extreme canonical forms represents the molecular structure will be determined by the relevant C-C bond lengths of the butadiene fragment. However, the observed distances in 10, C(7)-C(8), 1.401(5); C(8)-C(9), 1.426(5); C(9)-C(10), 1.423(5)Å; reveal that neither canonical form alone is correct. The true molecular structure has almost equal C-C bond distances. This 'bond levelling' effect is likely a result of back donation from the Mn into the lowest  $\pi^*$  orbital of the diene<sup>84</sup>. This orbital has nodal planes between the doubly bonded C(7)-C(8) and C(9)-C(10) atoms, thus back donation from the Mn would be expected to reduce the C-C bond order making all the distances similar. Structurally characterized ferrole structures<sup>83</sup> have also shown nearly identical butadiene bond lengths.

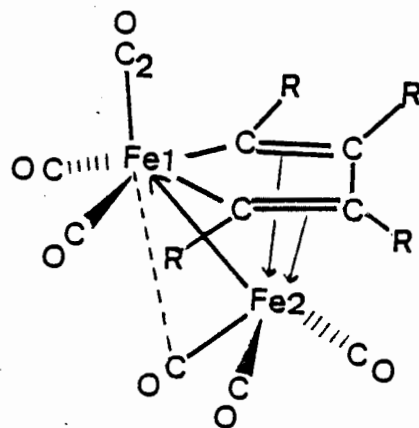
The manganese-carbon bond lengths to the cobaltacyclopentadienyl ring are not equivalent with Mn-C(7) and Mn-C(10) being shorter than Mn-C(8) and Mn-C(9), 2.099(4) versus 2.177(4)Å. Differing bond lengths between the butadiene fragment and the  $\pi$ -bound metal have also been noted in ferrole

structures<sup>84</sup> and in the Fe-Ni ferrole type compound<sup>77</sup>, although the differences in M-C bond lengths were smaller than 10. However the pattern is always similar with the shortest M-C bond to the outer two carbons, C(7) and C(10) and the longest bond to the inner two carbons, C(8) and C(9).

The ferrole structures possess in the solid state either a 'sawhorse' geometry (carbonyls eclipsed) or a 'non-sawhorse' geometry (carbonyls staggered)<sup>84</sup>. These two geometries are shown as II-4a and II-4b with the former having the eclipsing carbonyls and the latter with staggered carbonyls. These structures are interconverted simply by a 60° rotation of the Fe(CO)<sub>3</sub> group which is π-bonded to the butadiene fragment. This



II-4a



II-4b

rotation positions one carbonyl in a semi-bridging position. In the structure of 10, the carbonyls do not adopt either the sawhorse or non-sawhorse geometry but are positioned intermediate between the two extremes. The relevant dihedral angles between the carbonyls on opposite metals are 20 and 26°, while angles of 0° or 60° would be expected for II-4a or II-4b, respectively. This 'semi-sawhorse' geometry is a result of two factors.

While adoption of a non-sawhorse geometry is favoured by most ferroles<sup>84</sup> this places one of the carbonyls on Fe(2) in a semi-bridging position approximately 2.4-2.5Å from Fe(1). The formation of a semi-bridging CO ligand in 10 is not necessary since both the Co and Mn already have 18 electrons. However the sawhorse geometry was calculated by Hoffmann<sup>84</sup> to be approximately 0.2eV higher in energy than the non-sawhorse configuration making the sawhorse geometry for 10 also unfavourable. To minimize both these higher energy interactions, a configuration of the carbonyl ligands approximately halfway (26°), between the eclipsed (0°) and staggered (60°) geometries was found to be adopted.

## CHAPTER 3

SYNTHETIC AND STRUCTURAL STUDIES ON  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{R}_2)$  and  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{R}_2)$ 3.1 Introduction

This chapter describes the structural investigation of two related types of mononuclear rhenium complexes; the alkyne complex  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ , and the diazene complex  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ . The  $\text{CpRe}(\text{CO})_2$  fragment has a varied and interesting chemistry<sup>85</sup>. It can bind to a wide variety of ligands such as  $\text{PR}_3$ ,  $\text{CO}$ ,  $\text{N}_2\text{R}^+$ ,  $\text{NO}^+$ , olefins, etc. However it is not as common to find this fragment bound to another metal atom and certainly not to 2 or 3 metal atoms in a cluster arrangement. Because of our interest in heterotrimetallic (MM'M") alkyne clusters of the Ni, Co and Fe triads, we wished to extend this chemistry to include the Re group analogs in a similar arrangement such as  $\text{Cp}_2\text{NiCoRe}(\text{CO})_4(\text{RC}_2\text{R}')$ . In such an attempt, using  $\text{CpRe}(\text{CO})_2(\text{THF})$  and  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$ , a trimetallic cluster was not obtained but the mononuclear alkyne complex,  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  was isolated in low yield. This compound had been synthesized in 1976 by Nesmeyanov<sup>86</sup> and some properties very briefly reported at that time. We know of no further studies on it. Consequently we have synthesized the compound directly and have carried out a complete characterization including an X-ray structure determination. The analogous complex  $\text{CpRe}(\text{CO})_2(\text{MeC}_2\text{Ph})$  has also been synthesized.

The synthesis and study of transition metal carbonyl complexes containing coordinated diazenes is becoming an area of active investigation. The two nitrogen lone pairs and the  $\text{N}=\text{N}$   $\pi$  bond provide for a rich variety of



coordination modes in contrast to the alkyne ligand. In connection with the structure of the alkyne complex  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  and continuing interest in the structures and properties of the related diazenes  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{R}_2)$ <sup>87</sup> in this laboratory, an investigation of some of the latter was undertaken.

To provide a brief background for the subsequent discussion of these two types of compounds, this introduction continues with an outline of some of the features of alkyne and diazene complexes.

### 3.1.1 Mononuclear Alkyne Complexes

Many aspects of the synthesis of transition metal alkyne compounds have been dealt with in Chapter 1 and only a brief review will be given here.

#### 3.1.1a Synthesis

The synthesis of mononuclear transition metal complexes involves addition of the appropriate alkyne to a solution of the transition metal compound. Substitution of one of the ligands on the metal by the alkyne yielding the mononuclear alkyne complex is very common and the synthesis of  $\text{Cp}_2\text{Ti}(\text{CO})(\text{C}_2\text{Ph}_2)$  from  $\text{Cp}_2\text{Ti}(\text{CO})_2$  and  $\text{C}_2\text{Ph}_2$  is illustrative<sup>88</sup>.

It is possible for two alkyne ligands to bond to one metal atom.  $\text{Pt}(\text{C}_2\text{Ph}_2)_2$  can be easily synthesized from  $\text{Pt}(\text{COD})_2$  and two moles of alkyne<sup>89</sup>. Alternatively, one alkyne ligand has been known to displace two  $e^-$  electron donors on a transition metal<sup>90</sup> and a representative reaction is shown below.



The reduction of Re(V) to Re(III) has also been found to occur upon

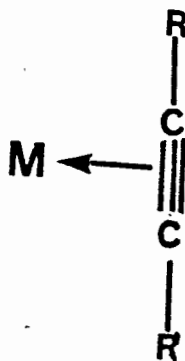
addition of an alkyne ligand<sup>90</sup> and one reaction is shown below. The tri-



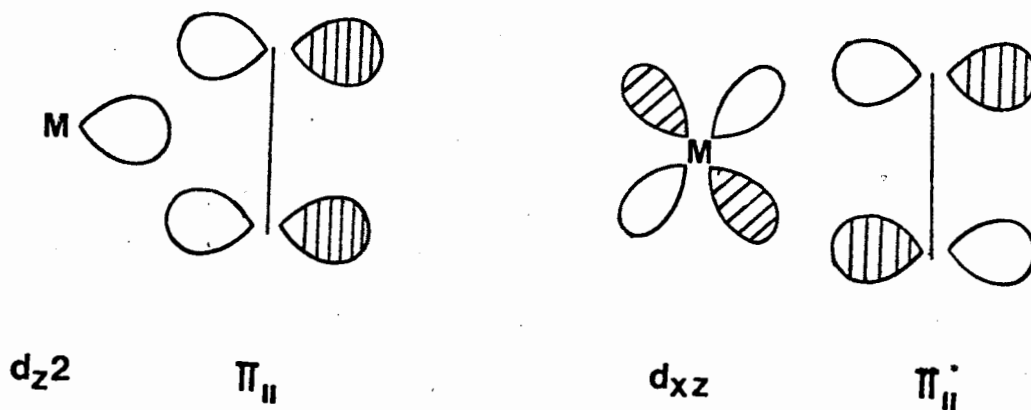
phenylphosphine oxide does not appear in the absence of the alkyne.

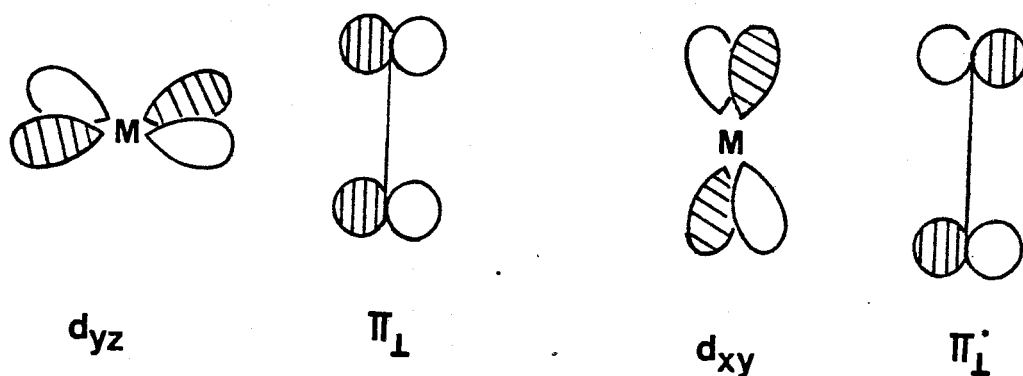
### 3.1.1b Bonding

A schematic of the transition metal alkyne interaction is shown below and superficially resembles the olefin-transition metal interaction forwarded



by Dewar<sup>26</sup> and Chatt and Duncanson<sup>27</sup>. However additional interactions with the alkyne ligand are possible and are discussed further. The alkyne has a pair of  $\pi$  and  $\pi^*$  orbitals (the  $\pi_{||}$  and  $\pi_{||}^*$  below) which are in the MCC plane and a pair of  $\pi$  and  $\pi^*$  molecular orbitals ( $\pi_{\perp}$  and  $\pi_{\perp}^*$  below)





perpendicular to the MCC plane. Both the  $\pi_{\parallel}$  and  $\pi_{\perp}$  have the ability to donate two electrons to the metal atom making the alkyne a potential four electron donor. The majority of mononuclear compounds have the alkyne as a two electron donor but four electron behaviour is certainly known, for example,  $\text{CpMo}(\text{CO})(\text{CH}_3(\text{RC}\equiv\text{CR}))^{91}$ . Back donation from the metal  $d_{xz}$  to the  $\pi^*_{\parallel}$  orbital is similar to that described by Dewar, Chatt and Duncanson for ethylene complexes. An additional interaction which is not present in the ethylene compounds is the possibility of back donation from the  $d_{xy}$  orbital of the metal to the  $\pi^*_{\perp}$  orbital of the alkyne.

### 3.1.2 Mononuclear Diazenes Complexes

While we are interested in diazenes from a coordination point of view, specifically to the  $\text{CpRe}(\text{CO})_2$  unit, it is worthwhile to note that much of the current interest in diazenes results from their postulated intermediacy in biological nitrogen fixation. Schrauzer<sup>92</sup> has invoked the parent diazene ( $\text{N}_2\text{H}_2$ ) as a crucial intermediate in a variety of nitrogen fixation reactions

and diazene has also been considered as an intermediate in the nonenzymatic conversion of coordinated dinitrogen<sup>93</sup>. Interest was also sparked into the coordination chemistry of diazenes when it was discovered that the introduction of metal ions into azo dyes resulted in modification of some of their properties<sup>94</sup>. Very recently, azobenzenes have been found to react with chromium carbenes to yield mixtures of 1,2- and 1,3-diazetidiones<sup>95</sup>.

In this chapter we shall not be concerned with the parent diazene,  $\text{HN}=\text{NH}$ , but only with the disubstituted derivatives and in particular the aromatic acyclic derivatives,  $\text{ArN}=\text{NAr}$ .

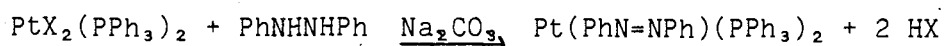
### 3.1.2a Preparation of Transition Metal Diazene Complexes

#### i) Insertion of Diazonium Salts into Metal Hydride Bonds

This particular synthetic route was first discovered by Parshall<sup>96</sup> when he isolated  $[\text{Pt}(\text{NHNAr})\text{Cl}(\text{PEt}_3)_2]^+$  from  $\text{Pt}(\text{H})\text{Cl}(\text{PEt}_3)_2$  and  $\text{ArN}_2^+$ . Many other workers have utilized this method for other metals<sup>97,98</sup>. Two major problems occur with this synthetic route: 1) treatment of some metal hydrides with diazonium salts decomposes the diazonium salt to dinitrogen and 2) facile orthometallation of the aryl group is common. These drawbacks make this method less popular than it might have otherwise become.

#### ii) Reactions with Hydrazines

The reaction of substituted hydrazines with halide complexes of the transition metals can yield diazene complexes. An example<sup>99</sup> is the formation of a platinum azobenzene complex from the platinum (II) dihalide and diphenylhydrazine.



iii) Reaction with Diazenes

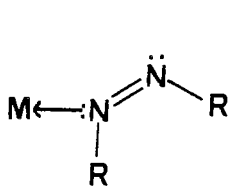
The instability of the parent diazene, HNNH, often precludes the direct reaction of this species with transition metal complexes. However, this route is the most common one for the synthesis of transition metal disubstituted diazene compounds. An example<sup>100</sup> is the reaction of Ni(cyclooctadiene)<sub>2</sub> with azobenzene or azotoluene in the presence of PPh<sub>3</sub> to yield Ni(N<sub>2</sub>Ar<sub>2</sub>)(PPh<sub>3</sub>)<sub>2</sub>. Generally the transition metal complex must have a labile ligand for the azo ligand to be able to displace it.

iv) Reaction of Coordinated Dinitrogen

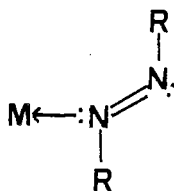
Sellmann<sup>101</sup> has found that phenyllithium will attack the dinitrogen ligand in CpMn(CO)<sub>2</sub>(N<sub>2</sub>) and when followed by protonation yields the diazene compound CpMn(CO)<sub>2</sub>(PhN=NH). This compound was postulated to have a σ-bonded diazene with the bound nitrogen containing the phenyl group. However, Leigh<sup>102</sup> suggests the product may be a hydrazido (2-) species and not a diazene. Confirmation of the structure awaits further information.

3.1.2b Structure and Bonding

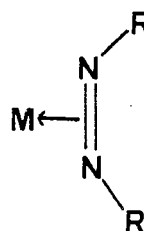
Upon coordination to a metal atom, the diazene or azo ligand can adopt one of three potential configurations shown below as III-1, III-2 and III-3.



III-1



III-2



III-3

The trans structure III-2 has not been authenticated to date although Vrieze<sup>103</sup> has prepared several platinum complexes of formula  $(RNNR)PtCl_2L$  ( $L = PPh_3, AsEt_3, C_2H_4$ ) and suggested from the  $\nu(N=N)$  vibration in the infrared that the azo ligand is in a trans configuration. The cis form III-1 has been observed<sup>104,105</sup> in  $[ReCl_2(HN_2C_6H_4F)(PEt_3)_2]^-$  and  $[RuCl(CO)_2(HN_2Ph)(PPh_3)_2]^+$ . The sideways bonded structure has been documented by X-ray crystallography<sup>106,107</sup> for  $L_2Ni(PhNNPh)$  ( $L = Me_3CNC, P(p-MeC_6H_4)_3$  and  $Cp_2Ti(PhNNPh)$ <sup>108</sup>. In the two nickel complexes the phenyl groups on the azobenzene ligand are oriented in a trans arrangement but adopt a cis geometry in the titanium compound.

### 3.2 Synthesis and Characterization

#### 3.2.1 Synthesis

$\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11) was initially observed to be formed in low yield as yellow crystals in the reaction of  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  with  $\text{CpRe}(\text{CO})_2(\text{THF})$ . Other than this, only unidentified decomposition products and  $\text{CpRe}(\text{CO})_3$  were observed to be formed in this reaction. Nesmeyanov<sup>86</sup> reported the synthesis of 11 in 20% yield by stirring a solution of  $\text{CpRe}(\text{CO})_2(\text{THF})$  (formed in situ in THF by irradiation of  $\text{CpRe}(\text{CO})_3$  at 10°C for 30 minutes) and  $\text{C}_2\text{Ph}_2$  for a further 15 minutes at 20°C with no further irradiation. In our hands, this procedure was ineffective at 20°C and furthermore, it gave 11 in only ca. 5% yield at 45°C, along with extensive decomposition. The yield improved to ca. 15% when the reaction of excess  $\text{C}_2\text{Ph}_2$  with a concentrated THF solution of  $\text{CpRe}(\text{CO})_2(\text{THF})$  was conducted in hexanes at 20°C. Finally, 11 was obtained in 30% yield by reacting a large excess of  $\text{C}_2\text{Ph}_2$  with crystalline  $\text{CpRe}(\text{CO})_2(\text{THF})$  entirely in hexanes at 20°C. The analogous complex  $\text{CpRe}(\text{CO})_2(\text{MeC}_2\text{Ph})$  (12) was similarly synthesized. The in situ irradiation of excess  $\text{C}_2\text{Ph}_2$  and  $\text{CpRe}(\text{CO})_3$  in hexanes produced 11 in rather low yield.

$\text{CpRe}(\text{CO})(\text{N}_2\text{Ph}_2)$  (13),  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$  (14) and  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$  (15) were synthesized by addition of the appropriate azobenzene ligand with  $\text{CpRe}(\text{CO})_2(\text{THF})$  in hexanes. Yields were seldom greater than 25%. As with 11, very little product could be isolated when THF was used as the solvent although acetone could be substituted for hexanes with no apparent loss in % yield.

### 3.2.2 Properties

$\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11) is a pale yellow crystalline solid soluble in hexane and polar organic solvents. Its solutions show remarkable thermal stability. It was recovered unchanged after 8 hours in refluxing iso-octane ( $99^\circ\text{C}$ ). The solid and its hexane solution are also stable in air for periods exceeding one week and 12 hours, respectively. 12 was obtained as a pale yellow low melting solid which appeared to be less stable and exhibited decomposition within a few hours when exposed to air.

Compounds 13, 14 and 15 were isolated as orange or light brown crystals. The solids were reasonably stable in air but in THF solution 13 decomposes within 30-45 minutes at room temperature, and 15 was found to slowly decompose in  $\text{CHCl}_3$  or acetone solutions.

In attempts to substitute either the  $\text{C}_2\text{Ph}_2$  or  $\text{N}_2\text{Ph}_2$  ligand or a carbonyl group in these compounds we have found them to be remarkably inert. With 11, no reaction occurred with triphenylphosphine in either hexanes, octane or acetonitrile at or somewhat above room temperature, and in the case of acetonitrile, no substitution by the solvent occurred. Similarly, no reaction of 11 in hexanes was observed with CO at either 1 atm or 1200 psi pressure, and no decomposition of 11 occurred. This is in contrast to the behaviour of  $\text{Cp}_2\text{Ti}(\text{CO})(\text{C}_2\text{Ph}_2)$ <sup>88</sup> in which the diphenylacetylene group is readily substituted by CO to give  $\text{Cp}_2\text{Ti}(\text{CO})_2$  and the decomposition of  $\text{Cp}_2\text{Ti}(\text{CO})(\text{C}_2\text{Ph}_2)$  is rapid above ca.  $30^\circ\text{C}$ .  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13) also did not substitute the azobenzene ligand for a CO at 1 atm.

The reaction of  $\text{I}_2$  with either 11 or 13 formed trans- $\text{CpRe}(\text{CO})_2\text{I}_2$  when the reaction was conducted in  $\text{CH}_2\text{Cl}_2$ . However when performed in hexanes, cis- $\text{CpRe}(\text{CO})_2\text{I}_2$ <sup>109,110</sup> was formed immediately rather than the corresponding



trans compound. A control reaction revealed that cis- $\text{CpRe}(\text{CO})_2\text{I}_2$  in  $\text{CH}_2\text{Cl}_2$  is converted to trans- $\text{CpRe}(\text{CO})_2\text{I}_2$  in less than 10 minutes.

11 and excess  $\text{N}_2\text{Ph}_2$  were stirred for 2 hours in hexanes but no exchange of alkyne for the azo ligand was noted as monitored by IR. A similar lack of reaction was also observed for the reverse reaction, that is  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  with excess  $\text{C}_2\text{Ph}_2$ .

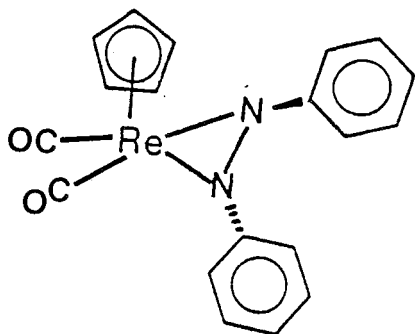
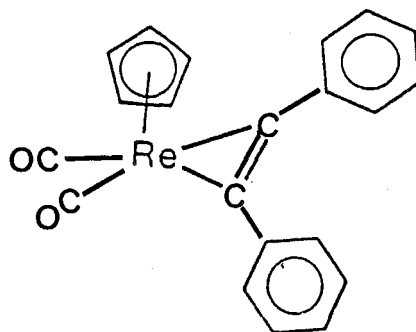
The spectroscopic properties of these compounds will be discussed after the solid state structures of 11 and 13 have been described.

### 3.3 X-ray Structures of $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ (11) and $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ (13)

The analysis of the X-ray data and a listing of all the bond lengths and angles, atom coordinates, etc., are in the Experimental section.

Both structures are composed of discrete mononuclear molecules with no unusually short inter- or intramolecular contacts.  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11) crystallizes in the space group  $\text{P}2_1/n$  and the asymmetric unit contains two independent molecules, referred to as molecules 1 and 2. Molecule 1 is illustrated in Figure 18, which also provides the numbering scheme. The differences between these crystallographically distinct molecules are only minor and are discussed below. When quoting bond lengths and angles in the discussion only molecule 1 will be mentioned unless there is a significant difference between molecules 1 and 2. Compound 13 crystallized in a non centrosymmetric space group  $\text{P}2_1$ , and a SNOOPI diagram of the molecule with atom labelling is shown in Figure 19.

In both structures the overall geometry of the rhenium coordination sphere conforms to that of a four legged 'piano stool' (shown below for 11 and 13) with the Cp ring as the seat and the four carbons in 11 [C(11),

1311

C(12), C(13), C(14)], or the two nitrogen atoms and C(1) and C(2) in 13, as

## FIGURE 18

SNOOPI Diagram of  
 $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11)

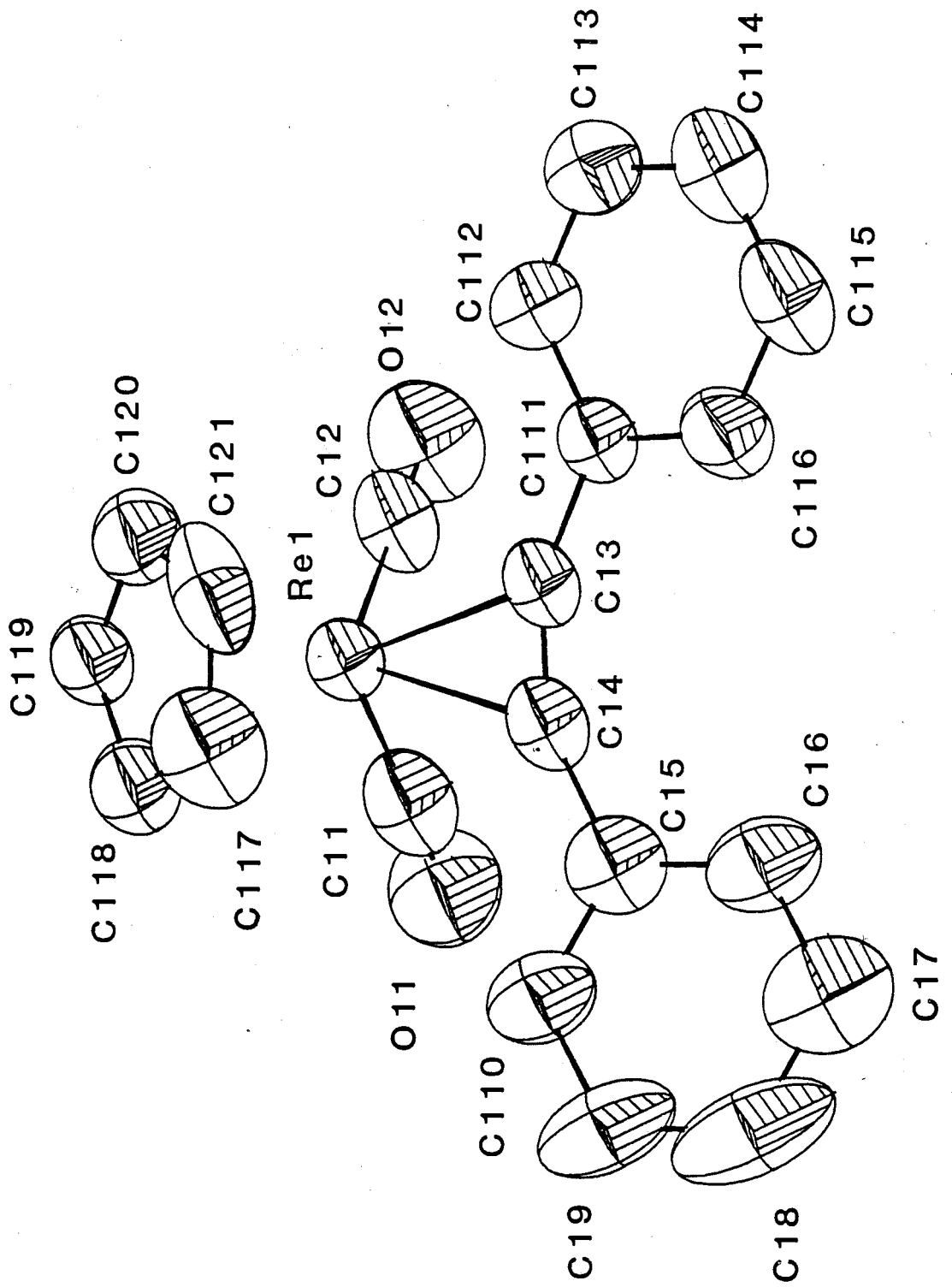
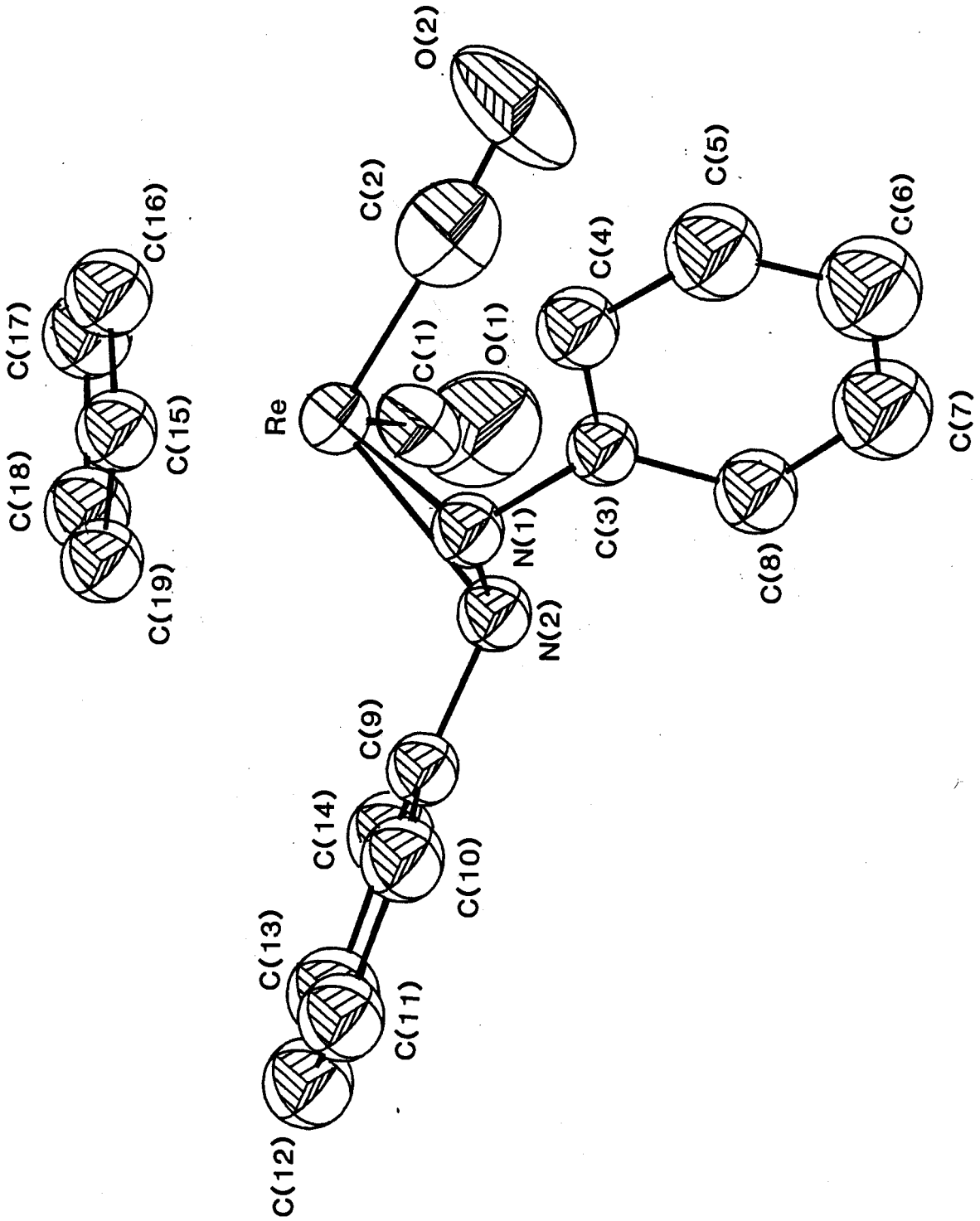


FIGURE 19

SNOOPI Diagram of  
 $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13)



the base. However these four basal atoms in 13 do not form a well defined plane ( $\chi^2 = 154$ ) unlike that of the four basal atoms in 11 ( $\chi^2 = 1.8$ ) Also the dihedral angle between the Cp ring and the four basal atoms is larger in 13 at  $9(1)^\circ$  than in 11 at  $2.3(8)^\circ$  showing the greater tilt of this ring.

Both  $N_2Ph_2$  and  $C_2Ph_2$  are bound side-on to the rhenium atom. The  $C_2Ph_2$  ligand is symmetrically bound with Re-C(13), 2.188(10)Å, Re-C(14), 2.182(9)Å and C(13)-Re-C(14),  $33.5^\circ$ . In addition, the coordinated diphenylacetylene is no longer linear but displays a cis geometry with the phenyl rings bent at angles of  $149.9(9)$ - $152.6(10)^\circ$  from the C≡C bond axis. This cis bending of the  $C_2Ph_2$  ligand has been observed in all previous mononuclear structures but it is noteworthy that the deformation angles observed here for 11 are among the largest that have been found; the reported range being  $135$ - $154^\circ$ . A summary of the observed angles and C≡C bond lengths for diphenylacetylene compounds is provided in Table 37. As shown in this table all the C≡C bond lengths are elongated with respect to free diphenylacetylene but those of 11 are among the shortest observed.

The geometries of the inner coordination spheres of 11 and 13 are shown in Figure 20. Unlike the  $C_2Ph_2$  ligand in 11, the  $N_2Ph_2$  ligand in 13 is not symmetrically bound to the rhenium atom. One Re-N bond length is 2.048(12)Å and the second 2.136(11)Å. The Re-N-N angles are also not equal and the N-Re-N angle (the 'bite' angle) is  $39.5(4)^\circ$ , significantly larger than observed in 11 which was  $33.5(4)^\circ$ . These small bite angles preclude any expectation of approximate four fold symmetry. The N(1)-N(2) distance of 1.415(17)Å appears to be one of the longest found for a metal bound azobenzene, but as shown in Table 38, which lists the relevant N=N distances for the structurally determined azobenzene compounds, the errors in the bond lengths of the N=N bonds are too large for accurate comparisons.

TABLE 37. The C≡C Bond Lengths (Å), *cis*-Deformation Angles (deg) and  $\nu(\text{C}\equiv\text{C})$  Values ( $\text{cm}^{-1}$ ) in Some Mononuclear  $\eta^2$ - Diphenylacetylene Complexes.

Complex	$d(\text{C}\equiv\text{C})$	Ph-C≡C angle	$\nu(\text{C}\equiv\text{C})$	Ref.
PhC≡CPh	1.198(3)	178.1(2)	2223	111,112
CpRe(CO) <sub>2</sub> (C <sub>2</sub> Ph <sub>2</sub> )	1.261(14) 1.232(13)	152.6(10), 151.1(10) 151.6(9), 149.9(9)	1848	a
[Co(C <sub>2</sub> Ph <sub>2</sub> )(PMe <sub>3</sub> ) <sub>3</sub> ] <sup>+</sup>	1.265(7)	143.6(5), 137.5(5)	-	113
[Co(MeCN)(C <sub>2</sub> Ph <sub>2</sub> )PMe <sub>3</sub> ] <sub>3</sub> <sup>+</sup>	1.267(7)	147.0(5), 148.8(5)	1780	113
Cp <sub>2</sub> Mo(C <sub>2</sub> Ph <sub>2</sub> )	1.269(7)	-	-	114
CpNb(CO)(C <sub>4</sub> Ph <sub>4</sub> )(C <sub>2</sub> Ph <sub>2</sub> )	1.26(4)	142(3)	1780	115
Mo(CNBu) <sub>2</sub> (SBu) <sub>2</sub> (C <sub>2</sub> Ph <sub>2</sub> )	1.28(2)	139(1)	-	116
Pt(C <sub>2</sub> Ph <sub>2</sub> ) <sub>2</sub>	1.291(5)	152.4(5), 154.1(5)	1881	117
Ni(C <sub>2</sub> Ph <sub>2</sub> )(CNBu-t) <sub>2</sub> <sup>b</sup>	1.276(12) 1.291(16)	148.2(10), 147.5(11) 147.2(14), 151.3(14)	1810	118
Cp <sub>2</sub> Ti(CO)(C <sub>2</sub> Ph <sub>2</sub> )	1.285(10)	138.8(7), 145.8(7)	1780	88
[CpMo(CO)(PPh <sub>3</sub> )(C <sub>2</sub> Ph <sub>2</sub> )] <sup>+</sup>	1.29(2)	135(1), 141(1)	1645	119
CpW(O)Ph(C <sub>2</sub> Ph <sub>2</sub> )	1.29(3)	143.5(4)	1748	120
Cp <sub>2</sub> Nb(O <sub>2</sub> CBu-t)(C <sub>2</sub> Ph <sub>2</sub> )	1.29(1)	141.2(7), 146.4(8)	1800	121
W(CO)(C <sub>2</sub> Ph <sub>2</sub> ) <sub>3</sub>	1.30	140	1702	122
Mo(S <sub>2</sub> CNEt <sub>2</sub> ) <sub>2</sub> (CO)(C <sub>2</sub> Ph <sub>2</sub> )	1.313(4)	140.3(3)	-	116
Pt(C <sub>2</sub> Ph <sub>2</sub> )(PPh <sub>3</sub> ) <sub>2</sub>	1.32(9)	140	1740	123
[pyH][TaCl <sub>4</sub> (py)(C <sub>2</sub> Ph <sub>2</sub> )]	1.325(12)	138.2(8), 141.2(8)	-	124
Mo(Porph)(C <sub>2</sub> Ph <sub>2</sub> )	1.324(5)	136.4(4), 145(5)	-	114
Mo(TTP)(C <sub>2</sub> Ph <sub>2</sub> )C <sub>6</sub> H <sub>5</sub> CH <sub>3</sub>	1.324(5)	136.4(4), 145.4(4)	-	114



Table 37 (cont'd)

Complex	<u>d</u> (C≡C)	Ph-C≡C angle	$\nu$ (C≡C)	Ref.
Fe(C <sub>2</sub> Ph <sub>2</sub> )[P(OMe <sub>3</sub> ) <sub>3</sub> ]	1.332(10)	-	-	125
Cp*TaCl <sub>2</sub> (C <sub>2</sub> Ph <sub>2</sub> )	1.35(2)	138.5(5), 140.3(5)	-	126
CpNb(C <sub>2</sub> Ph <sub>2</sub> ) <sub>2</sub> (CO)	1.35(2)	138(4)	-	127

<sup>a</sup> This work. Values are for independent molecules 1 and 2, respectively.

<sup>b</sup> Two independent molecules occur in the asymmetric unit.

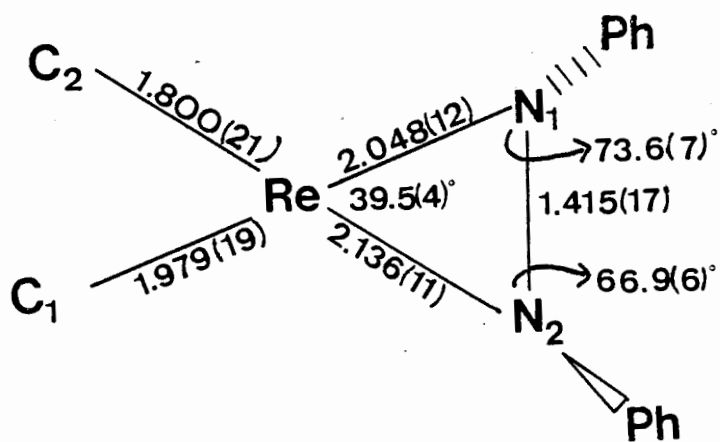
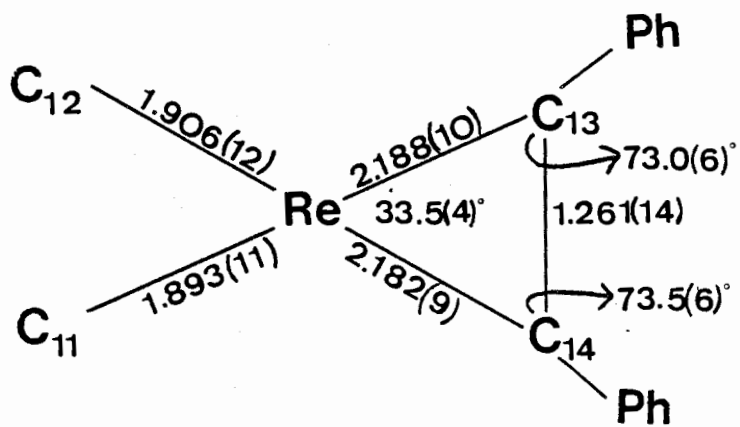
Table 38. Structurally Characterised Azobenzene and Transition  
Metal Azobenzene Compounds.

Compound	C-N	N=N	<C-N=N	M-N	<N-M-N	Ref.
<u>trans</u> -azobenzene	1.434(3)	1.243(3)	113.3(3)	-	-	128
<u>cis</u> -azobenzene	1.449(4)	1.253(4)	121.9(3)	-	-	129
<u>trans</u> -p,p'-azotoluene	1.433(4)	1.244(4)	113.8(3)	-	-	130
Cp <sub>2</sub> Ti(N <sub>2</sub> Ph <sub>2</sub> )	1.397(9) 1.393(9)	1.339(8)	122.5(3) 123.7(3)	1.965(6) 1.971(5)	39.8(5)	108
(CpTiCl) <sub>2</sub> (μ-NPh)(μ-N <sub>2</sub> Ph <sub>2</sub> )	1.438(6)	1.422(6)	116.1(4)	1.920(4)	-	108
[P(C <sub>6</sub> H <sub>4</sub> CH <sub>3</sub> ) <sub>3</sub> ] <sub>2</sub> Ni(N <sub>2</sub> Ph <sub>2</sub> )	1.427(8) 1.423(8)	1.371(6)	111.6(5) 112.2(5)	1.932(5) 1.927(5)	41.6(2)	111
(t-BuNC) <sub>2</sub> Ni(N <sub>2</sub> Ph <sub>2</sub> )	1.421(5) 1.424(5)	1.385(5)	111.6(3) 111.9(3)	1.897(4) 1.899(5)	42.8(2)	106
CpRe(CO)(N <sub>2</sub> Ph <sub>2</sub> )	1.406(20) 1.424(18)	1.415(17)	113.7(10) 114.4(10)	2.048(12) 2.136(11)	39.5(4)	a

<sup>a</sup> This work.

## FIGURE 20

View of the Inner Coordination Sphere of  
 $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11) and  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13)



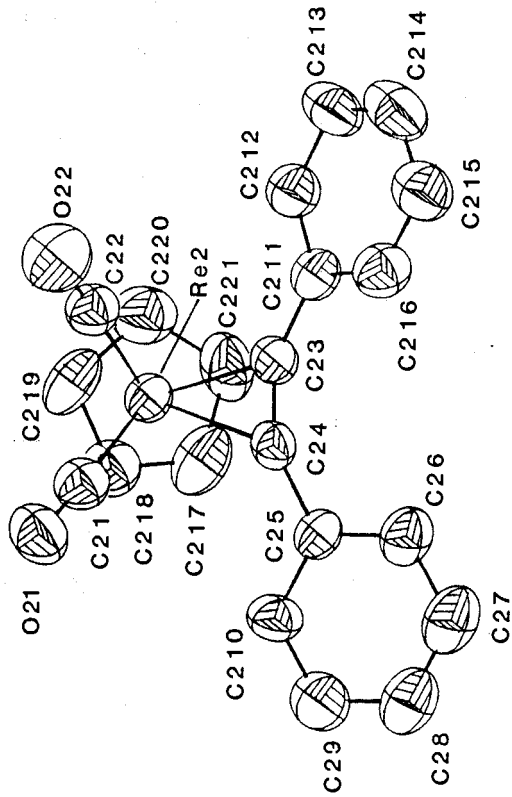
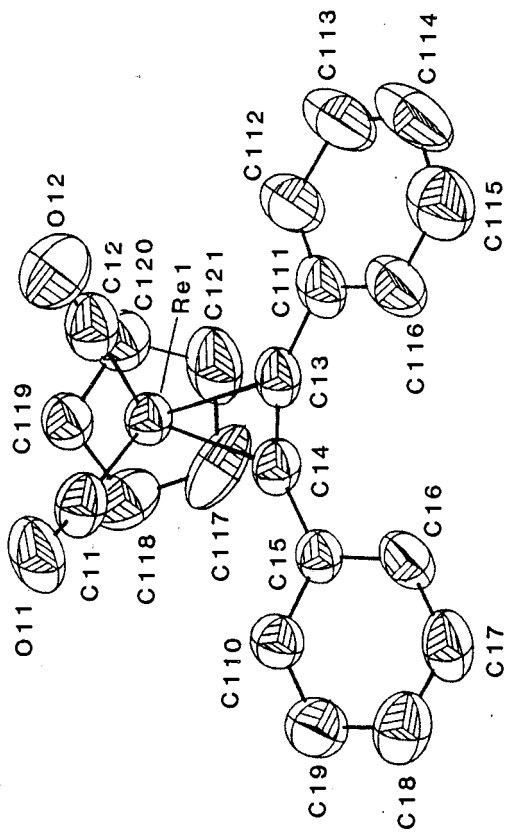
The Re-C(carbonyl) bond lengths for both molecules of 11 are very similar, with distances ranging from 1.893(11) to 1.916(11)Å. However in 13 these lengths, as shown in Figure 20, are not equal with the Re-C(carbonyl) bonds being 1.80(2)Å and 1.98(2)Å.

The most obvious dissimilarity between the independent molecules 1 and 2, which are compared in Figure 21 projected onto the plane of the Cp ring, is in the relative orientation of the two phenyl rings in each case. This is most evident by comparing the dihedral angles between the planes of these phenyl rings. In molecule 1 this angle is 16.2(9)°, and in molecule 2 it is 28.2(9)°. In the determination of the X-ray structure of Ni(C<sub>2</sub>Ph<sub>2</sub>)(CNBu-t)<sub>2</sub>, which also crystallizes with two independent molecules in the asymmetric unit, Dickson and Ibers<sup>106</sup> observed an analogous difference in these dihedral angles to be the main dissimilarity in the molecules. They noted that two contributing features may be calculated: i) δ - the dihedral angle between an individual phenyl ring plane and its adjacent C≡C-C(phenyl) skeletal plane, ii) - a 'tilt' angle which is the angle between the two distinguishable adjacent C≡C(phenyl) skeletal planes in the diphenylacetylene ligand. These values are listed in Table 37.

In 13 the phenyl groups are bent back away from the rhenium atom and the extent of this deviation is shown by the dihedral angle of 44(1)° between the N(1)N(2)C(3) and N(1)N(2)C(9) planes. Ibers<sup>106</sup> has suggested that this angle is an indication of the extent of metal-azo interaction and the larger the angle, the greater the interaction. In [P(C<sub>6</sub>H<sub>4</sub>CH<sub>3</sub>)<sub>3</sub>]<sub>2</sub>Ni(N<sub>2</sub>Ph<sub>2</sub>)<sup>107</sup> and [(CH<sub>3</sub>)<sub>3</sub>CNC]<sub>2</sub>Ni(N<sub>2</sub>Ph<sub>2</sub>)<sup>106</sup> dihedral angles of 23.5(4) and 26.8(4)° were noted which suggests that a smaller metal-azo interaction is present in these compounds than in 13. This is substantiated by the longer N=N bond length found in 13 than in the two nickel compounds.

## FIGURE 21

SNOOPI Diagram of the Two Molecules in the Asymmetric  
Unit of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11)



The ortho hydrogens in 13 may be causing a small amount of steric repulsion. The angles  $N(1)C(3)C(4) = 117.6(13)^\circ$  and  $N(1)C(3)C(8) = 123.6(14)^\circ$  for the  $N(1)-C(3)$  bond and  $N(2)C(9)C(10) = 122.0(11)^\circ$  and  $N(2)C(9)C(14) = 116.8(13)^\circ$  for the  $N(2)-C(9)$  bond should be equal for each individual C-N bond. But possibly because of a non-bonded interaction of the ortho hydrogens with the opposite nitrogen atom, a small amount of steric repulsion is introduced resulting in an observed slight twist ( $5-6^\circ$ ) of the individual phenyl groups.

### 3.4 Solution Structures

#### 3.4.1 $CpRe(CO)_2(C_2Ph_2)$ (11) and $CpRe(CO)_2(MeC_2Ph)$ (12)

The solution structure of the alkyne compounds 11 and 12 appear to be comparable to that observed for 11 in the solid state. The IR spectra (hexanes) of 11 and 12 display the expected 2  $\nu(CO)$  absorptions with those of the  $MeC_2Ph$  derivative at slightly lower wavenumber ( $1976, 1899\text{ cm}^{-1}$ ) than observed for 11 ( $1980, 1904\text{ cm}^{-1}$ ). The  $^1H$  NMR spectra of both compounds show a single peak due to the Cp protons and a multiplet assignable to the phenyl protons. In addition 12 has a single methyl resonance at  $\delta 2.75$ . The  $^{13}C$  and  $^{13}C\{^1H\}$  spectra allow all the carbon resonances of 11 and 12 to be assigned (see Experimental Section) and at room temperature both exhibit single carbonyl resonances. In 11 this could result from a rigid structure with equivalent carbonyl groups like that observed in the crystal structure. This cannot be true for 12 which contains the unsymmetric alkyne  $MeC_2Ph$ , and suggests in each case the probable rotation of the alkyne ligand about the Re-alkyne axis. In 12 the  $^{13}C$  spectra at  $-60^\circ C$  exhibit two equal intensity carbonyl carbon resonances at  $\delta 206.5$  and  $\delta 206.0$ , but the resonances for



all other carbon atoms remain essentially unperturbed. This is consistent with a single frozen-out structure and this must have the  $\text{PhC}_2\text{Me}$  ligand oriented parallel, not perpendicular, to the plane of the Cp ring as is observed for the  $\text{PhC}_2\text{Ph}$  ligand in the X-ray structure of 11.

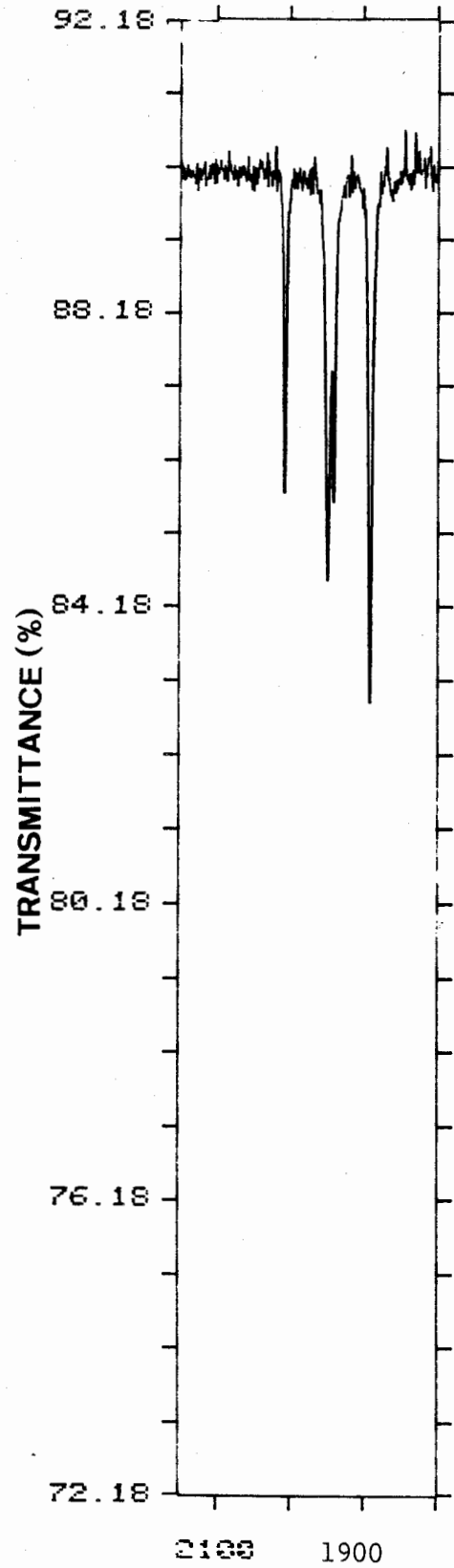
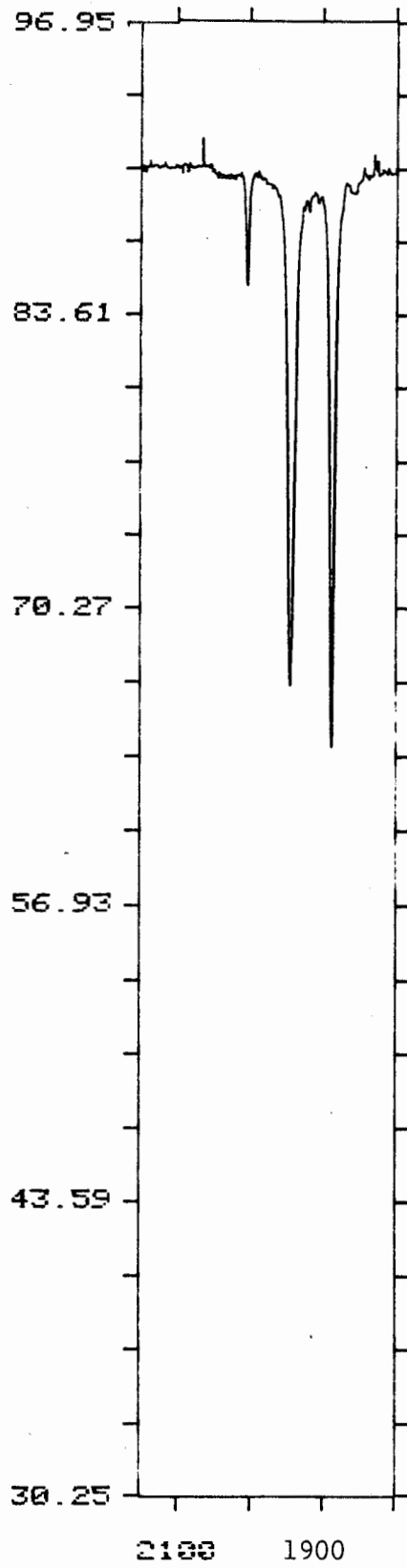
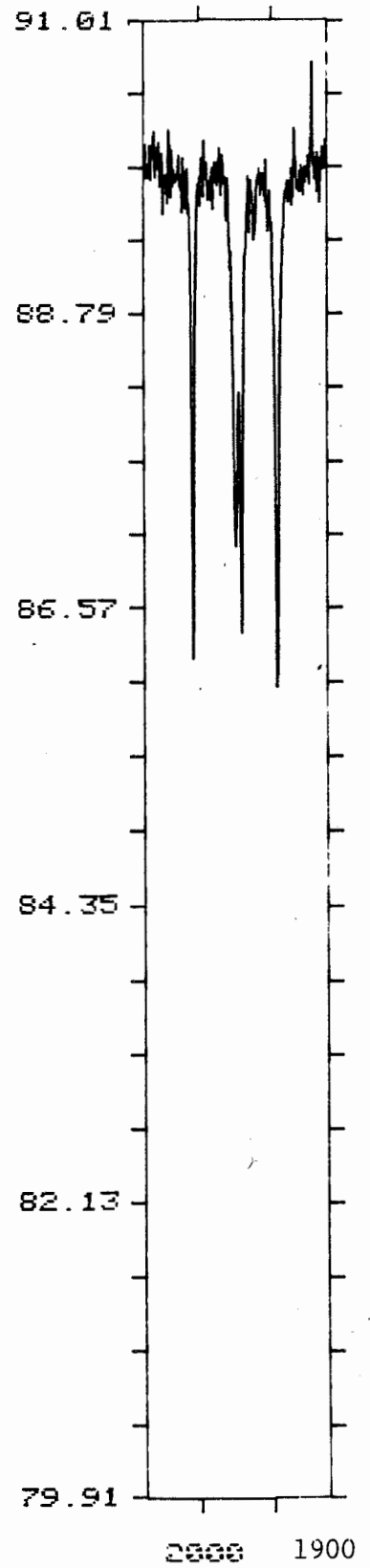
An NOE experiment was performed by irradiation of the Cp resonance and observing the increase in the resonances of the phenyl protons. The para protons were not observed to be enhanced but the meta protons were partially enhanced and the ortho protons of the phenyl ring were fully enhanced.

### 3.4.2 $\text{CpRe}(\text{CO})_2(\text{N}_2\text{R}_2)$ [R = Ph (13), $\text{C}_6\text{H}_4\text{Me-p}$ (14), $\text{C}_6\text{H}_4\text{F-p}$ (15)]

The IR spectra of 13, 14 and 15 show more carbonyl bands than are to be expected if only the structure determined for 13 is present in solution. All three spectra (hexanes) are shown in Figure 22 for comparison. Compound 13 has four sharp bands at 2006, 1947, 1939 and 1890  $\text{cm}^{-1}$  as has 15 at 2008, 1950, 1941 and 1893  $\text{cm}^{-1}$ . However, for 14 there are only three discernible absorptions at 2004, 1945 and 1887  $\text{cm}^{-1}$  with the latter two very strong in comparison with the higher wavenumber band. Most probably the 1945  $\text{cm}^{-1}$  peak consists of two nearly coincident bands. A decrease in intensity of the  $\nu(\text{CO})$  bands in different solvents, allows assignment of the spectrum as two pairs of bands, each pair attributable to an individual isomer. As an example, for 13, in  $\text{CH}_2\text{Cl}_2$  the bands at 2006 and 1939  $\text{cm}^{-1}$  are reduced to approximately one-third intensity clearly showing these bands belong to one isomer. Removal of the  $\text{CH}_2\text{Cl}_2$  and redissolving in hexanes restores the original spectrum. Table 39 lists the  $\nu(\text{CO})$  bands of each isomer of 13, 14 and 15 and those of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  and cis- $\text{CpRe}(\text{CO})_2\text{I}_2$  for comparison. Isomer A has arbitrarily been assigned the higher wavenumber  $\nu(\text{CO})$  bands and isomer B the lower  $\nu(\text{CO})$  absorptions.

## FIGURE 22

IR Spectra (hexanes) of 13, 14 and 15

131415

WAVENUMBER

Table 39.  $\nu(\text{CO})$  Bands for Mononuclear Rhenium Compounds

Compound	$\nu(\text{CO}) \text{ cm}^{-1}$ (hexanes)
$\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ (Isomer A) ( <u>13</u> )	2006, 1939
$\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ (Isomer B) ( <u>13</u> )	1947, 1898
$\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$ (Isomer A) ( <u>14</u> )	2004, (ca. 1945)
$\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$ (Isomer B) ( <u>14</u> )	1945, 1887
$\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$ (Isomer A) ( <u>15</u> )	2008, 1941
$\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$ (Isomer B) ( <u>15</u> )	1950, 1893
$\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ ( <u>11</u> )	1980, 1904
<u>cis</u> - $\text{CpRe}(\text{CO})_2\text{I}_2$	2040, 1977

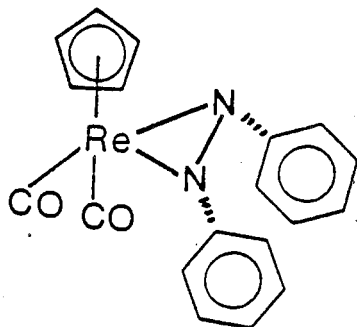
The infrared spectrum of 13, 14 and 15 changes as the solvent polarity (CHCl<sub>3</sub>, acetone) is increased. The lower wavenumber absorptions i.e., isomer B, become dominant and the bands arising from isomer A are reduced to approximately one-quarter their original intensity in hexanes. In addition the  $\nu(\text{CO})$  absorptions are much broader in the more polar solvents and as a consequence the 'inner two' bands, e.g., 1947 and 1939 cm<sup>-1</sup> for 13, merge into one absorption.

From the infrared data in Table 39 it is possible to deduce some information regarding the probable geometry of these two isomers. The  $\nu(\text{CO})$  bands of isomer A are more closely related to a Re(III) compound, for example, cis-CpRe(CO)<sub>2</sub>I<sub>2</sub>, than are the bands of isomer B. In addition, N<sub>2</sub>Ph<sub>2</sub> appears to be a better  $\pi$ -acceptor than is C<sub>2</sub>Ph<sub>2</sub><sup>131</sup> so an increase in wavenumber for the  $\nu(\text{CO})$  bands (by comparison with those for CpRe(CO)<sub>2</sub>C<sub>2</sub>Ph<sub>2</sub>) would be expected for the isomer in which the azobenzene is bound in the same fashion (side-on) as C<sub>2</sub>Ph<sub>2</sub> is in CpRe(CO)<sub>2</sub>(C<sub>2</sub>Ph<sub>2</sub>). Since side-on configuration of the azo ligand was confirmed in the X-ray crystal structure, solution isomer A is proposed to have the N<sub>2</sub>Ph<sub>2</sub> ligand bound side-on also; this is therefore present as the minor isomer in solution. Though no information is available from the IR spectrum regarding the stereochemistry of the two phenyl groups, they are for the moment regarded to be the same as in the crystal, i.e., trans (structure A-1).

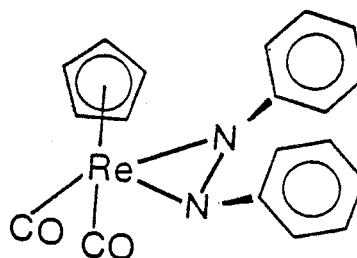
The KBr spectrum of powdered CpRe(CO)<sub>2</sub>(N<sub>2</sub>Ph<sub>2</sub>) shows three  $\nu(\text{CO})$  bands at 2001, 1928 and 1900 cm<sup>-1</sup> rather than the two anticipated from the structure. Using only a small number of crystals from the same sample that furnished the crystal used for X-ray analysis, a KBr spectrum was obtained identical to that obtained from the powder. It proved impossible to grow crystals from just a polar solvent (CHCl<sub>3</sub>, CH<sub>2</sub>Cl<sub>2</sub>, acetone), without the

addition of a non-polar solvent such as hexanes or pentanes. It was hoped that use of only a polar solvent would favour the second isomer for crystal growth. Similarly, X-ray quality crystals could not be grown for 14, which appears to exist in solution as mainly one isomer. Thus, it was not possible to obtain a satisfactory correlation between the solid state IR and X-ray structure to use in further support of the solution IR assignments of isomers A and B.

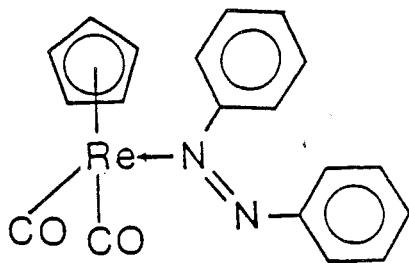
There are a number of possibilities for the geometry of isomer B. These are shown below. A-2 and A-3 have the azo ligand also bound in an  $\eta^2$  fashion, but with the aromatic rings oriented in a cis geometry. The rings could be either cis to the carbonyl ligands as in A-2 or cis to the Cp



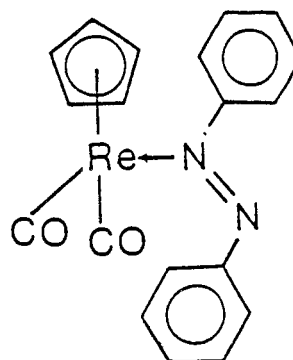
A-2



A-3



B-1



B-2

ligand as in A-3. Alternatively, the azo ligand could be bound to the rhenium in an  $\eta^1$  fashion, donating the two electrons of a nitrogen lone pair. Again the aromatic rings could be in two different possible orientations, cis (B-1) or trans (B-2). Because A-2 and A-3 resemble a Re(III) species similar to isomer A-1, (13), these would be expected to have  $\nu(\text{CO})$  values similar to A-1\*. But B-1 and B-2 would be expected to have lower  $\nu(\text{CO})$  values than isomer A-1 since they are more effectively Re(I) compounds with the  $\text{N}_2\text{Ph}_2$  ligand bound as a Lewis base. Thus on the basis of the observed IR values isomer B is assigned the  $\sigma$ -bonded structure which is represented as B-1 or B-2. In an effort to obtain more information relating to the solution structures of isomers A and B, attention was turned to NMR spectroscopy.

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\* The results of variable temperature NMR spectroscopy (see below) indicate that structures A-2 and A-3 contribute to the fluxionality of isomer A in solution.

3.4.2a CpRe(CO)<sub>2</sub>(N<sub>2</sub>Ph<sub>2</sub>) (13)

The proton spectra were obtained for solutions in acetone-d<sub>6</sub> or CDCl<sub>3</sub>, which, from the IR data, will favour the presence of isomer B. Since at the present stage there is no evidence on which to distinguish between possibilities B-1 or B-2 as the correct structure for the major isomer B, any reference to the  $\sigma$ -bonded isomer will be made as "isomer B". As shown in Figure 23, in addition to resonances of the major isomer B, the Cp resonance due to isomer A-1 is also visible and is slightly broader than the Cp peak for the major isomer. The phenyl region of 13 only shows the ortho, meta and para protons due to the major isomer B, and none for the minor isomer A-1.

The <sup>13</sup>C and <sup>13</sup>C{<sup>1</sup>H} room temperature spectra of 13 essentially confirm the proton spectra. Two Cp peaks due to the major and minor isomers, and resonances due to two distinct phenyl rings of the major isomer are observed. The Cp resonance of the minor isomer is slightly broader than the Cp peak of the major isomer, similar to the <sup>1</sup>H NMR data. Only one peak was observed in the carbonyl carbon region at room and lower temperatures.

We have suggested from the IR data that the minor isomer in solution has the side-on structure that was characterized by the X-ray structure determination. The identity of the major isomer is unknown but was tentatively assigned as either B-1 or B-2 because of the decrease in  $\nu(\text{CO})$  values that a coordinated Lewis base would be expected to cause as compared to a  $\pi$ -bound acid ligand. From the <sup>1</sup>H and <sup>13</sup>C data both A-2 and A-3 can again be eliminated as possible candidates for the structure of the major isomer, since resonances for two distinct phenyl rings are observed in the <sup>1</sup>H and <sup>13</sup>C NMR, and A-2 and A-3 have their phenyl rings in equivalent



## FIGURE 23

$^1\text{H}$  NMR (acetone- $d_6$ ) of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13)

Cp B

Cp A

OB

pB

mB

3

4

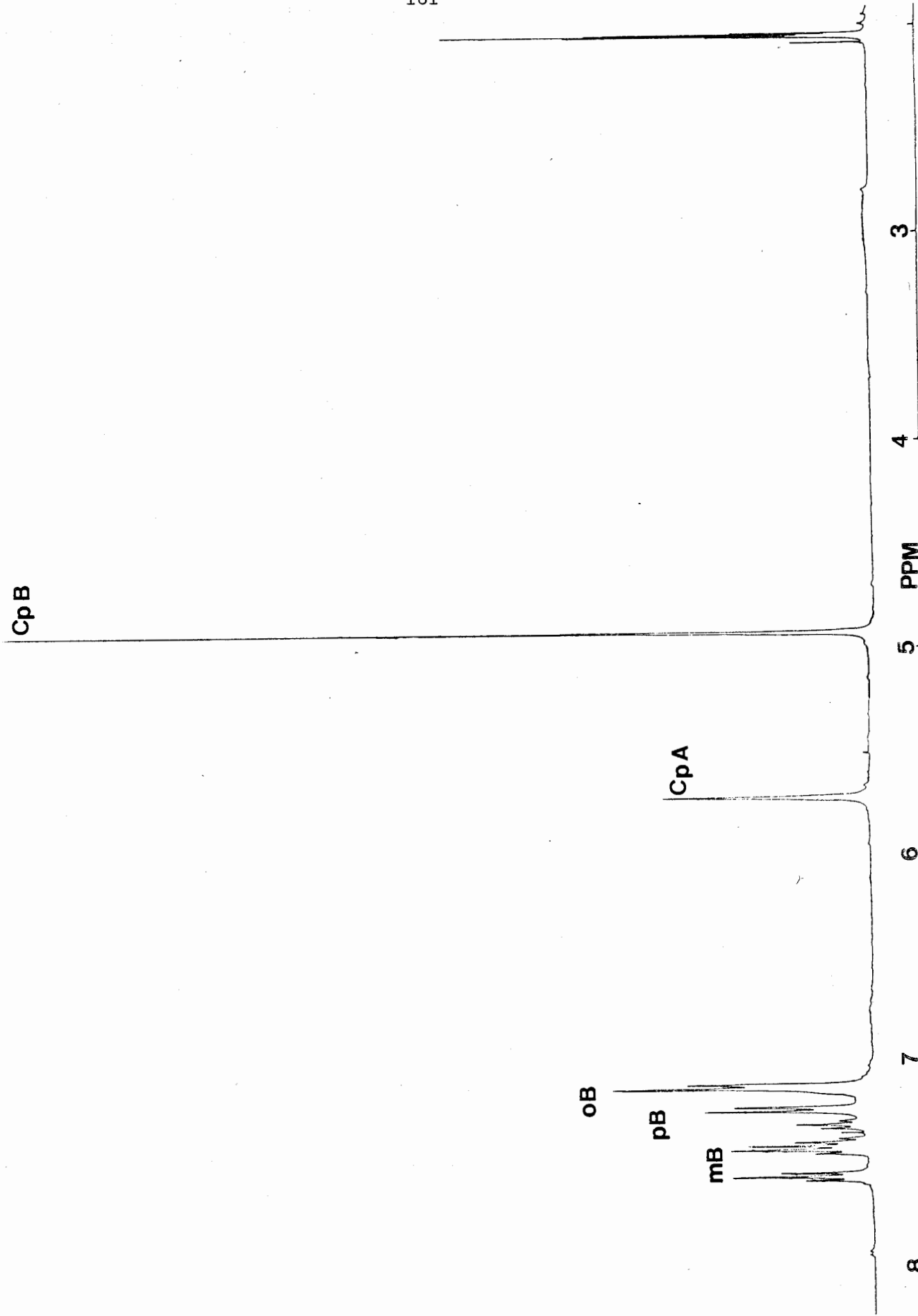
PPM

5

6

7

8



environments and would only show one set of phenyl resonances in the NMR. A 50:50 mixture of A-2 and A-3 is also not possible as this would produce two equal intensity Cp peaks in the NMR and a total of six  $\nu(\text{CO})$  bands (4 for the major isomers and 2 for the minor isomer) in the IR, neither of which is observed. Both of the alternative structures B-1 and B-2, have the phenyl groups of the azobenzene ligand in non-identical environments and either structure would result in the observance of distinct phenyl proton or carbon resonances. Rotation about the Re-N bond in B-1 or B-2 would result in only one carbonyl carbon resonance being observed in the  $^{13}\text{C}$  NMR.

To help establish or confirm the geometry of the minor isomer A in solution as A-1 an NOE experiment was performed. The Cp resonance of the minor isomer ( $\delta$  5.72) was irradiated in an effort to selectively enhance the resonances of the phenyl protons of the minor isomer which may be obscured in the  $^1\text{H}$  NMR by the phenyl protons of the major isomer. If the geometry of the minor isomer is as observed in the X-ray structure, enhancement of the protons on the phenyl ring which is closest to the Cp ring would be expected. The phenyl rings of the alkyne ligand in  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11) are further away from the Cp ligand than the phenyl ring in  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13) and, as mentioned previously (sec. 3.4.1), enhancement of the ortho and meta protons is observed for 11. Upon irradiation of the Cp resonance of the minor isomer of 13 an  $^1\text{H}$  NMR spectrum identical to that shown in Figure 24 was observed. Saturation transfer from the minor Cp peak to the major Cp resonance has occurred and then subsequently, an NOE transfer from the protons of the Cp ring of the major isomer to the protons of both phenyl rings. The transfer between the Cp peaks of the major and minor isomers firmly establishes that reversible interconversion is occurring. Irradiation of the Cp resonance of the major isomer also resulted in transfer

## FIGURE 24

NOE Experiment on  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$

a) Expansion of the Phenyl Region

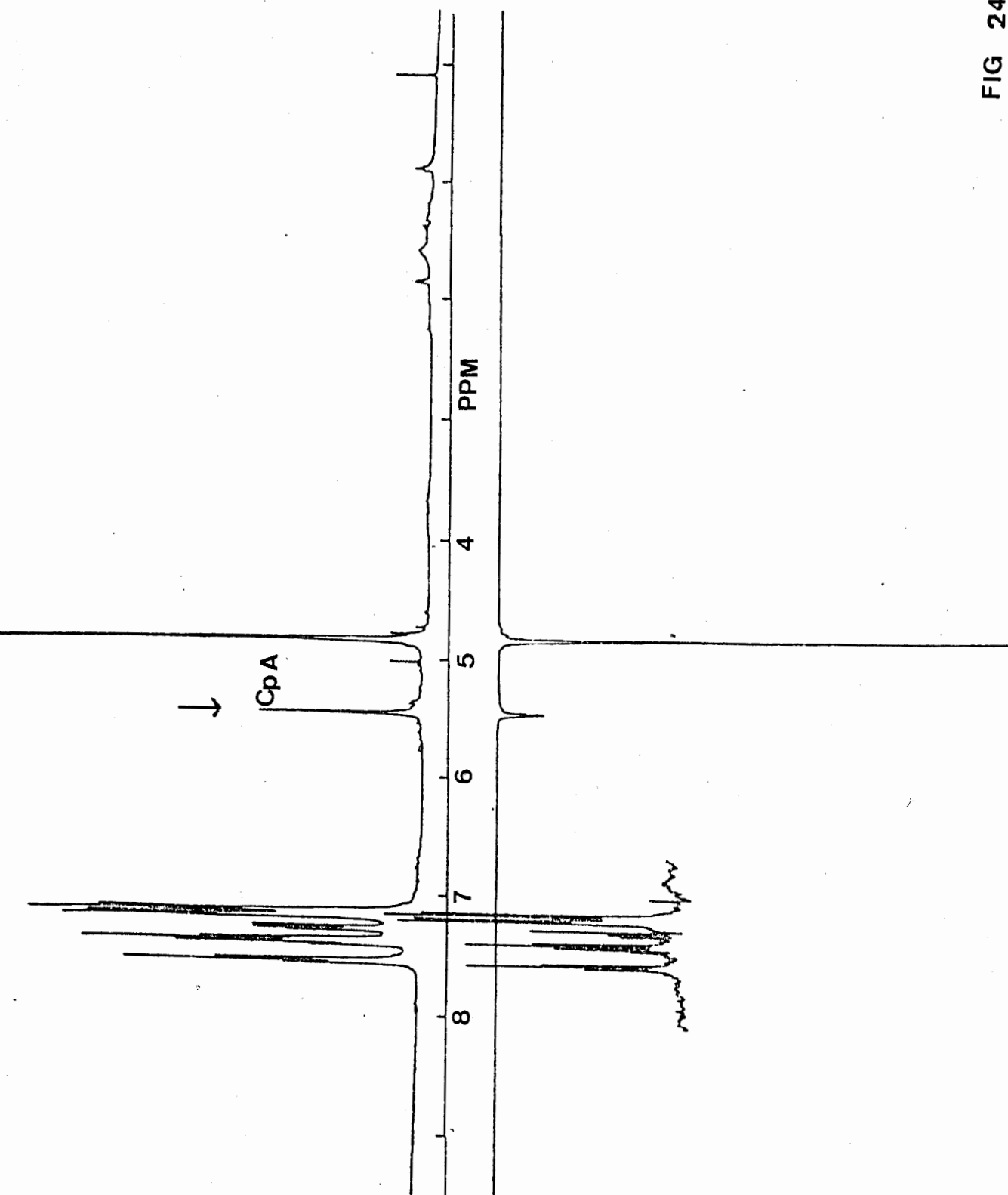


FIG 24

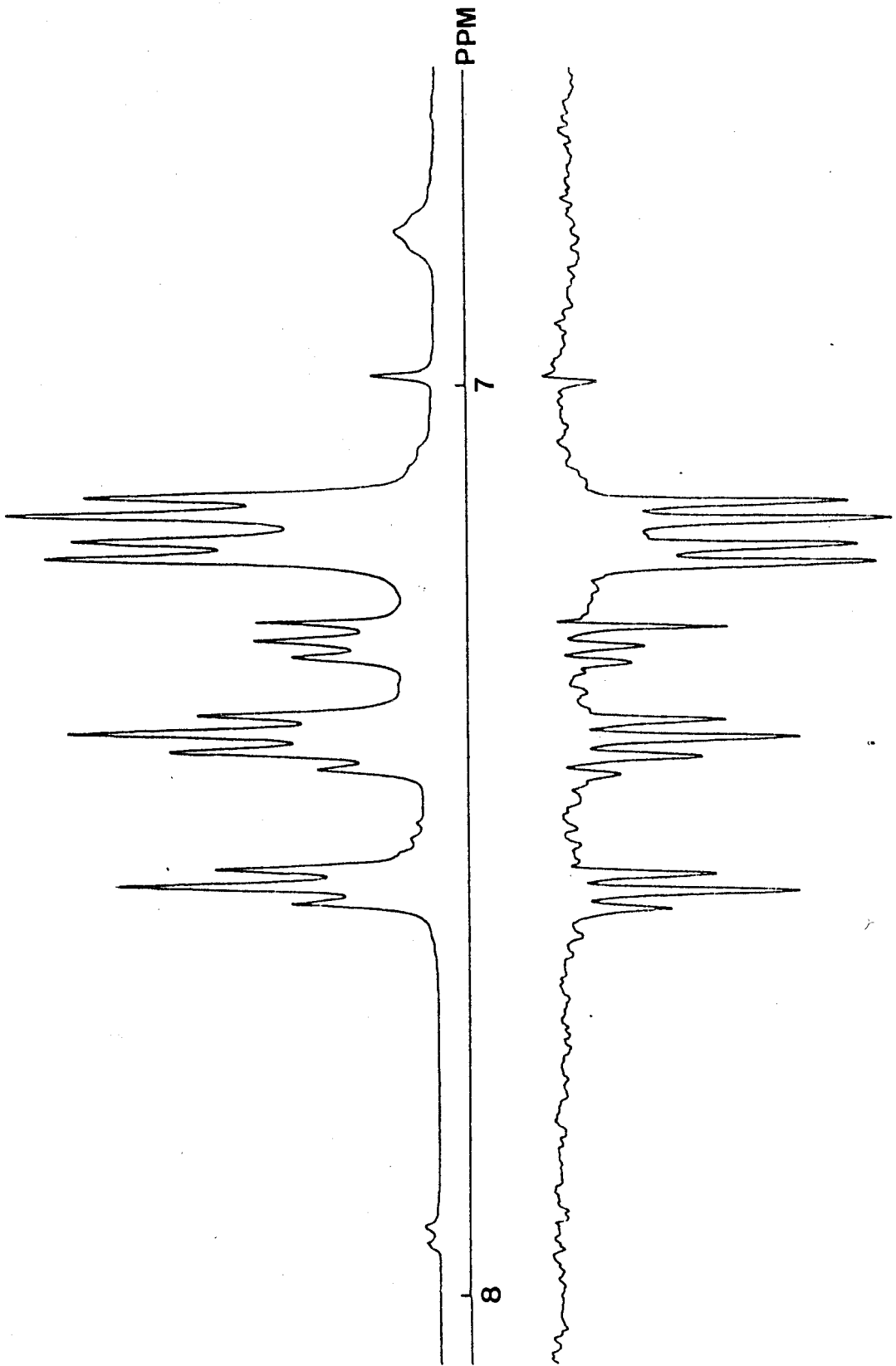
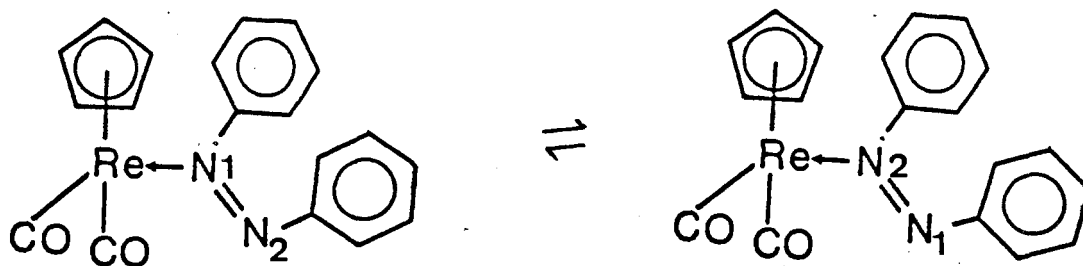


FIG 24 a

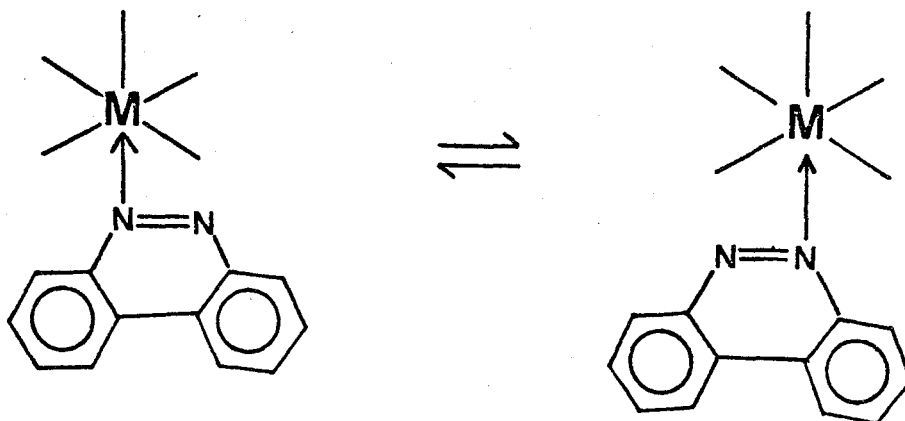
to the Cp peak of the minor isomer and enhancement of the phenyl protons of the major isomer. This transfer was observed even with using shorter NOE buildup times (5, 3, 1, 0.3, 0.1 sec) than used originally (10 sec). Shorter buildup times result in less NOE transfer per pulse and poorer quality spectra, hence buildup times less than 0.1 sec were not pursued.

Since the phenyl protons of both phenyl groups of the major isomer are enhanced by NOE transfer, both of these groups must approach the Cp ring. This can occur for the major isomer as shown below. The phenyl ring on N<sub>1</sub>



initially is close to the Cp ring and coordination site exchange with the N<sub>2</sub> of the azobenzene ligand would put the second phenyl group in closer proximity to the Cp group. However this exchange is not fast enough to average the chemical shifts of the protons of the phenyl ligands.

There is precedent for coordination site exchange between the nitrogens of an azo group in an  $\eta^1$  bonded complex. Nixon<sup>132</sup> and Herberhold<sup>133</sup> have recorded this type of exchange in  $M(\text{CO})_5\text{L}$  ( $M = \text{Cr}, \text{Mo}, \text{W}$ ;  $\text{L} =$  benzo[*c*]cinnoline) and  $\text{ArCr}(\text{CO})_2(2,3\text{-diazabicyclo-[2.2.1]hept-2-ene})$  ( $\text{Ar} =$  substituted benzene rings) respectively. The exchange for the  $M(\text{CO})_5(\text{benzo[}c\text{]cinnoline})$  complex is shown below. Exchange between the two



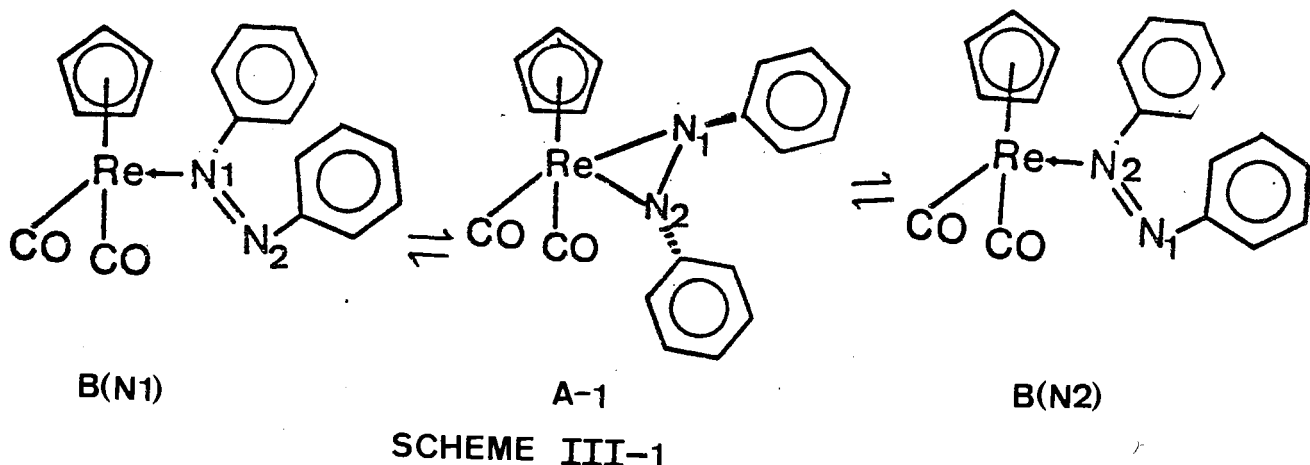
nitrogens on the benzo[c]cinnoline was found to occur and is an example of a  $\eta^1-\eta^1$  shift. However the benzo[c]cinnoline and the 2,3-diazabicyclo-[2.2.1]hept-2-ene are fixed into a cis geometry and cannot form products where the groups on the nitrogen are in a trans position. Acyclic cis azo ligands also undergo similar coordination site exchange.  $M(\text{CO})_5\text{L}$  ( $M = \text{Cr}, \text{Mo}, \text{W}$ ;  $\text{L} = \text{cis-dimethyldiazene}$  (c-DMD)) were shown recently by Ackerman<sup>134,135</sup> to be fluxional and undergo coordination site exchange, but in addition Ackermann notes that for the complexes where  $\text{L} = \text{trans-dimethyldiazene}$  (t-DMD) there is no observable exchange. Unfortunately, a direct analogy between Ackermann's and our work is not possible since Ackermann did not make use of any aromatic azobenzene ligands.

For the  $\eta^1-\eta^1$  coordination site shift in  $\text{ArCr}(\text{CO})_2(\text{azo})$ <sup>133</sup>,  $\Delta G^\ddagger$  values of 71.1-84.5  $\text{kJ mol}^{-1}$  were recorded depending upon the substituent on the benzene ring. Similarly  $\Delta G^\ddagger$  values of 62.8, 48.1 and 62.8 ( $\pm 0.4$ )  $\text{kJ mol}^{-1}$  were recorded<sup>134</sup> for  $M(\text{CO})_5(\text{c-DIPD})$  with  $M = \text{Cr}, \text{Mo}$  and  $\text{W}$  respectively. Coalescence experiments (Brüker WM-100) on the Cp peaks of the major and minor isomer of 13 established a  $\Delta G^\ddagger$  value of  $67.4 \pm 0.4 \text{ kJ mol}^{-1}$ . This is the value for the  $\eta^1-\eta^2$  shift between isomer A-1 and isomer B and cannot be directly compared to the values obtained by Ackermann<sup>134</sup> for  $M(\text{CO})_5(\text{c-DIPD})$  since the latter involve an  $\eta^1-\eta^1$  shift. There have not been any reports in



the literature of an  $\eta^1$ - $\eta^2$  shift of an azo ligand for comparison. Only very recently<sup>136,137</sup> was an  $\eta^1$ - $\eta^2$  shift established for any ligand, specifically a phosphalkene ligand ( $R-P=CR_2$ ). This ligand exhibits a coordination shift between the phosphorus lone pair and the phosphorus-carbon double bond in the complex  $Pt(PPh_3)_2(P(\text{mesityl})=CPh_2)$ . However,  $\Delta G^\ddagger$  values were not given for this process.

We have established that a reversible interconversion is occurring between isomer A-1 and isomer B, and that exchange of the phenyl group positions in isomer B is occurring. This process is probably represented by Scheme III-1. The major isomer B is interconverting with the minor isomer A-1 by a  $\sigma$ - $\pi$  or an  $\eta^1$ - $\eta^2$  shift. The three membered ring of A-1 can open to



reform B or to form B', these two compounds differing only in which particular nitrogen is bonded to the rhenium atom.

In the room temperature  $^{13}C\{^1H\}$  NMR spectrum of 13 (Figure 25), the Cp resonance of the minor isomer, Cp A, is much broader than the Cp resonance of the major isomer, Cp B. The phenyl carbon resonances of the minor isomer are very broad and are seen as large humps centered about  $\delta$  129, 123 and 120. The broadness encountered for these carbon resonances is indicative of a second fluxional process occurring and that the minor isomer is undergoing

exchange with further species. In an attempt to obtain more information on this process a number of lower temperature  $^{13}\text{C}\{^1\text{H}\}$  spectra were recorded.

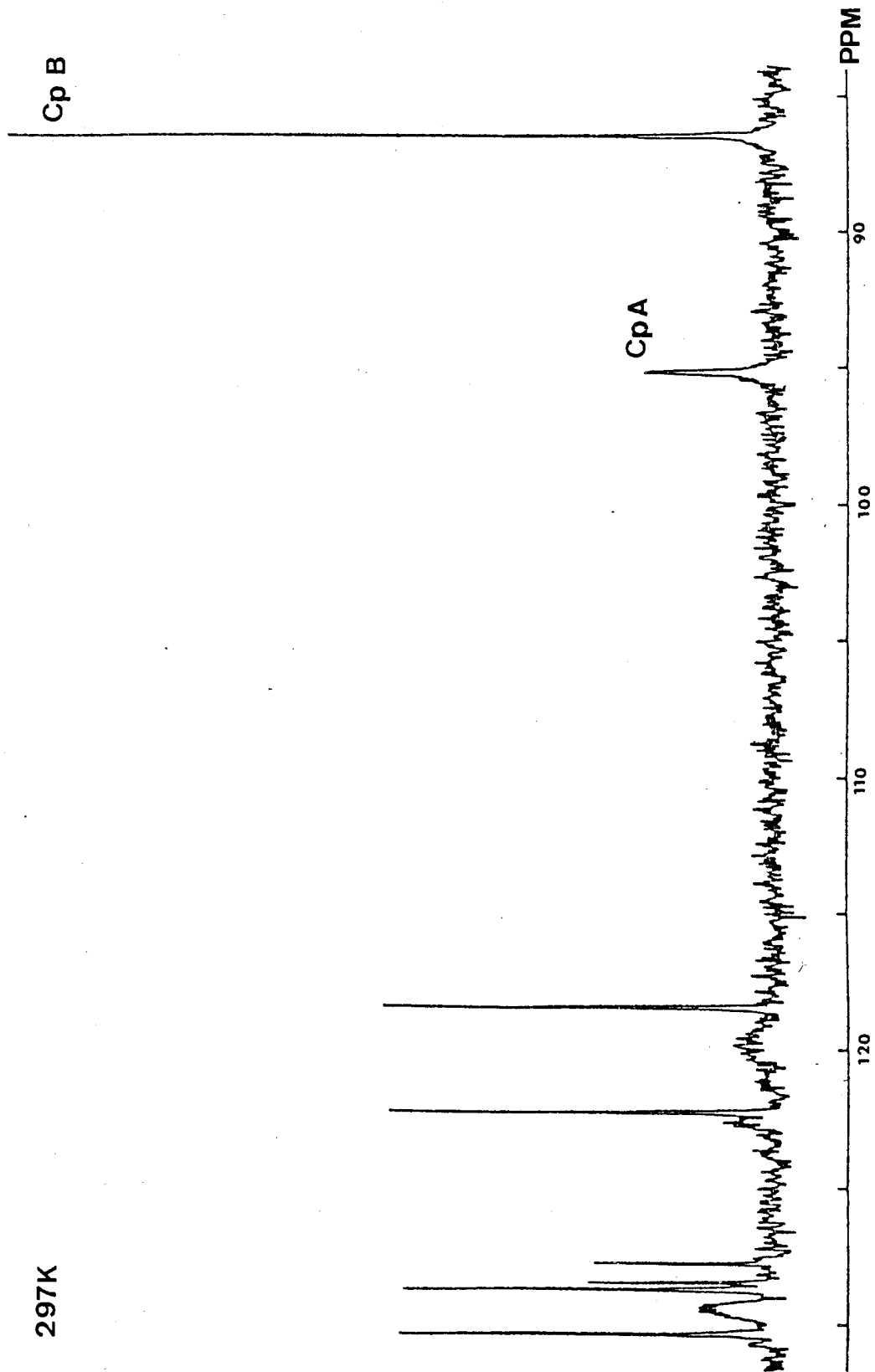
Figure 26 shows the aromatic region in the  $^{13}\text{C}\{^1\text{H}\}$  NMR of 13 at 221 K. Two of the broad resonances in Figure 25 have sharpened into two lines and the third broad resonance at  $\sim 119.5$  ppm has disappeared. However subsequent cooling to 200 K (Figure 27) and to 189 K (Figure 28) results in further splitting of the lines until the limiting spectrum is obtained. A total of 11 and possibly 12 lines (one may be under the 122 ppm peak of the major isomer) are observed. While it is difficult to follow the emergence of new peaks precisely, a labelling scheme such as that shown in Figure 28 is a possibility. If the ordering of aromatic carbons is similar to that in the major isomer, then the resonances labelled A and B would be due to the ortho protons of the minor isomer(s) and the labels C-E would be assigned to the meta and para carbons. Regardless of the actual assignment of the aromatic protons of the minor species, there are a minimum of 11 lines, and possibly 12, present. This cannot be attributed to only one minor isomer in solution, since the maximum number of phenyl carbon resonances attributable to one isomer is six (not including the ipso carbons), but must be due to two or more minor species. It is not clear whether further splitting would occur below 189 K, or whether some of the broadening of the resonances is due to the viscosity increase of the solvent. However, very little broadening was observed in the resonances of the major isomer which would tend to refute the viscosity argument.

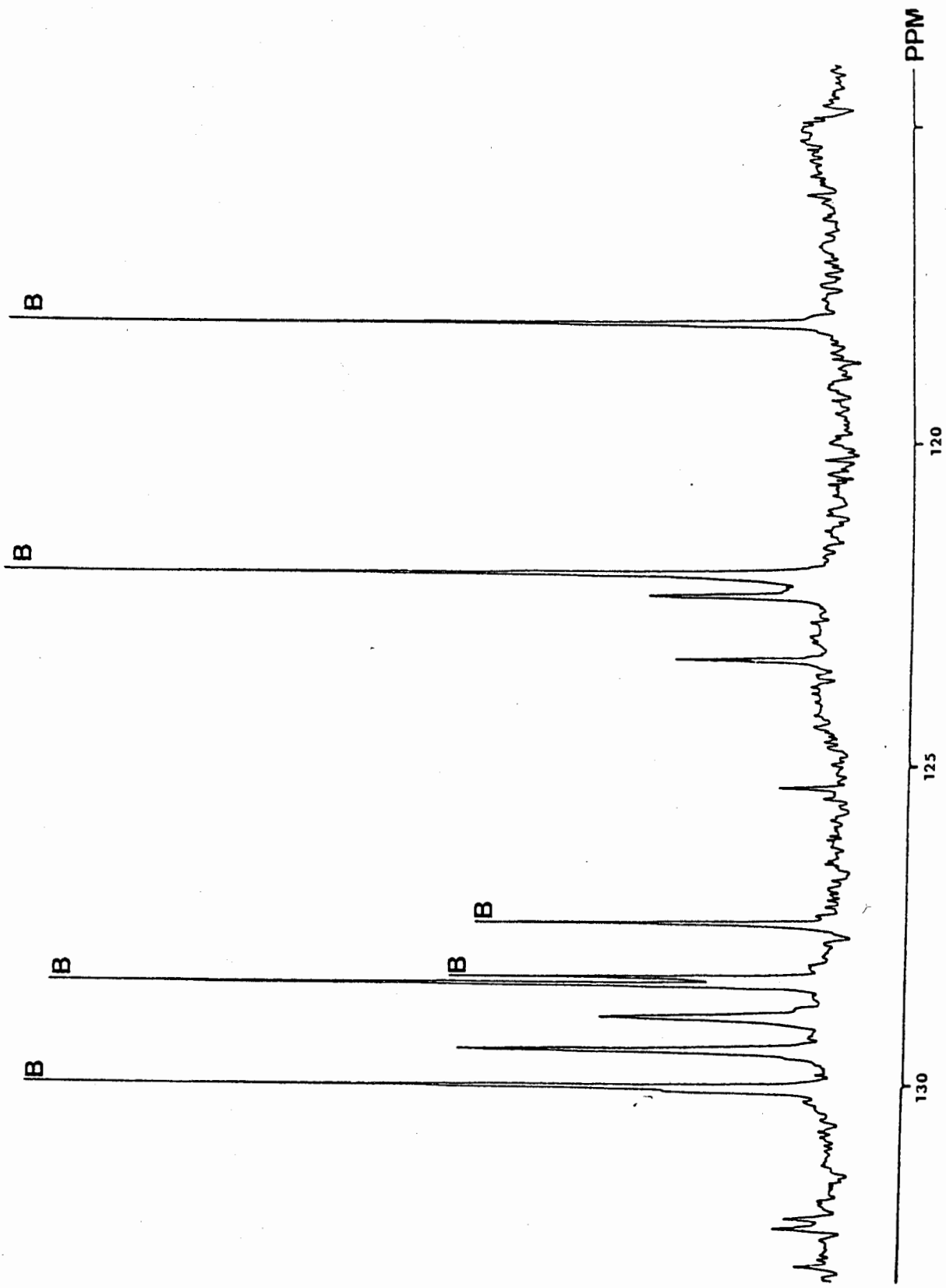
While only the aromatic region was shown in Figure 25 to 28, the other regions were examined. However little change could be discerned. Figure 29 shows the carbonyl carbon, ipso carbon and Cp carbon region of the limiting spectrum (189 K). Only one sharp carbon resonance in the CO

## FIGURE

25.  $^{13}\text{C}\{^1\text{H}\}$  NMR (acetone- $d_6$ ) of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  at 297 K
26.  $^{13}\text{C}\{^1\text{H}\}$  NMR (acetone- $d_6$ ) of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  at 221 K
27.  $^{13}\text{C}\{^1\text{H}\}$  NMR (acetone- $d_6$ ) of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  at 200 K
28.  $^{13}\text{C}\{^1\text{H}\}$  NMR (acetone- $d_6$ ) of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  at 189 K

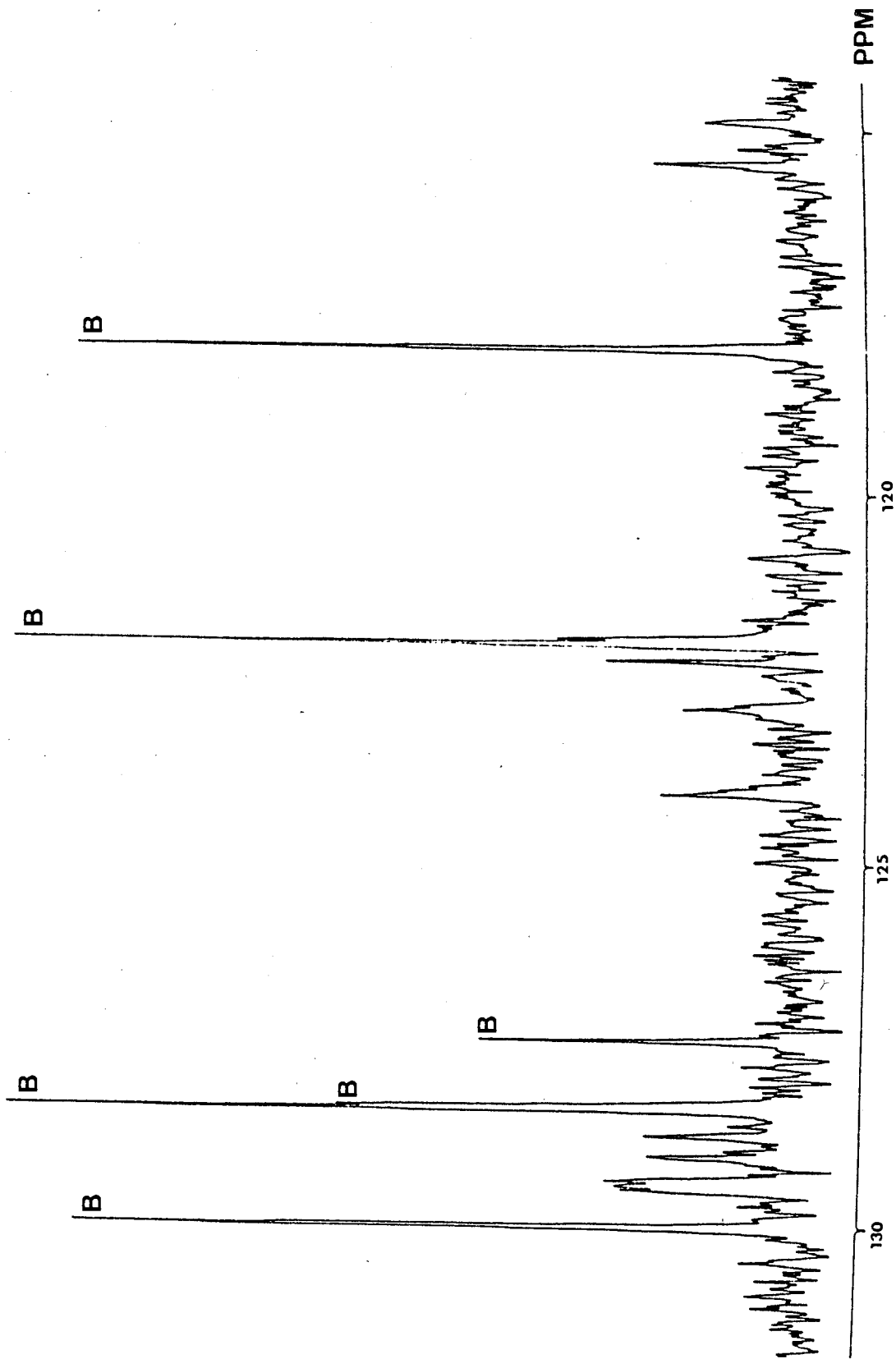
297K





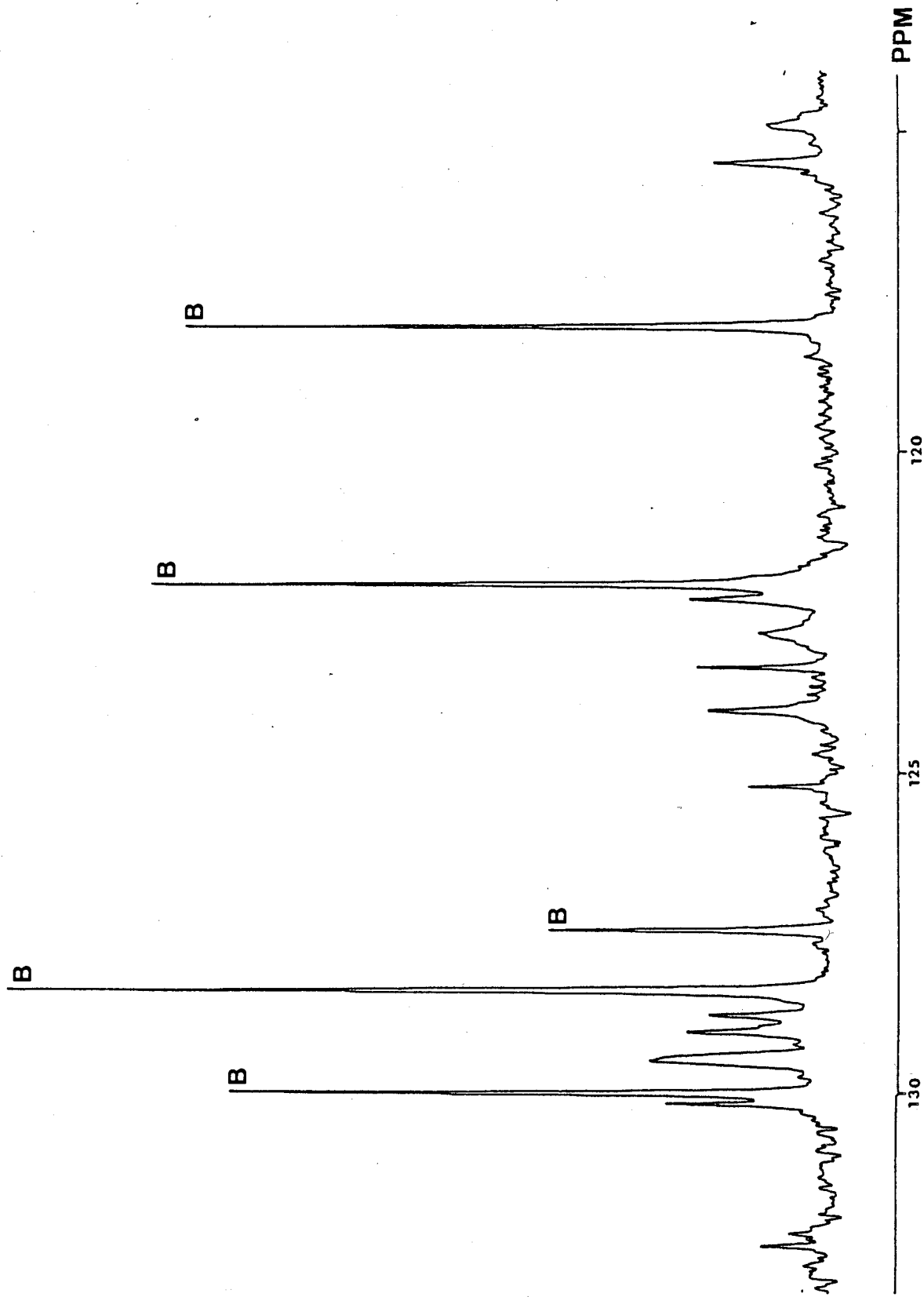
221K

200 K



189K

194



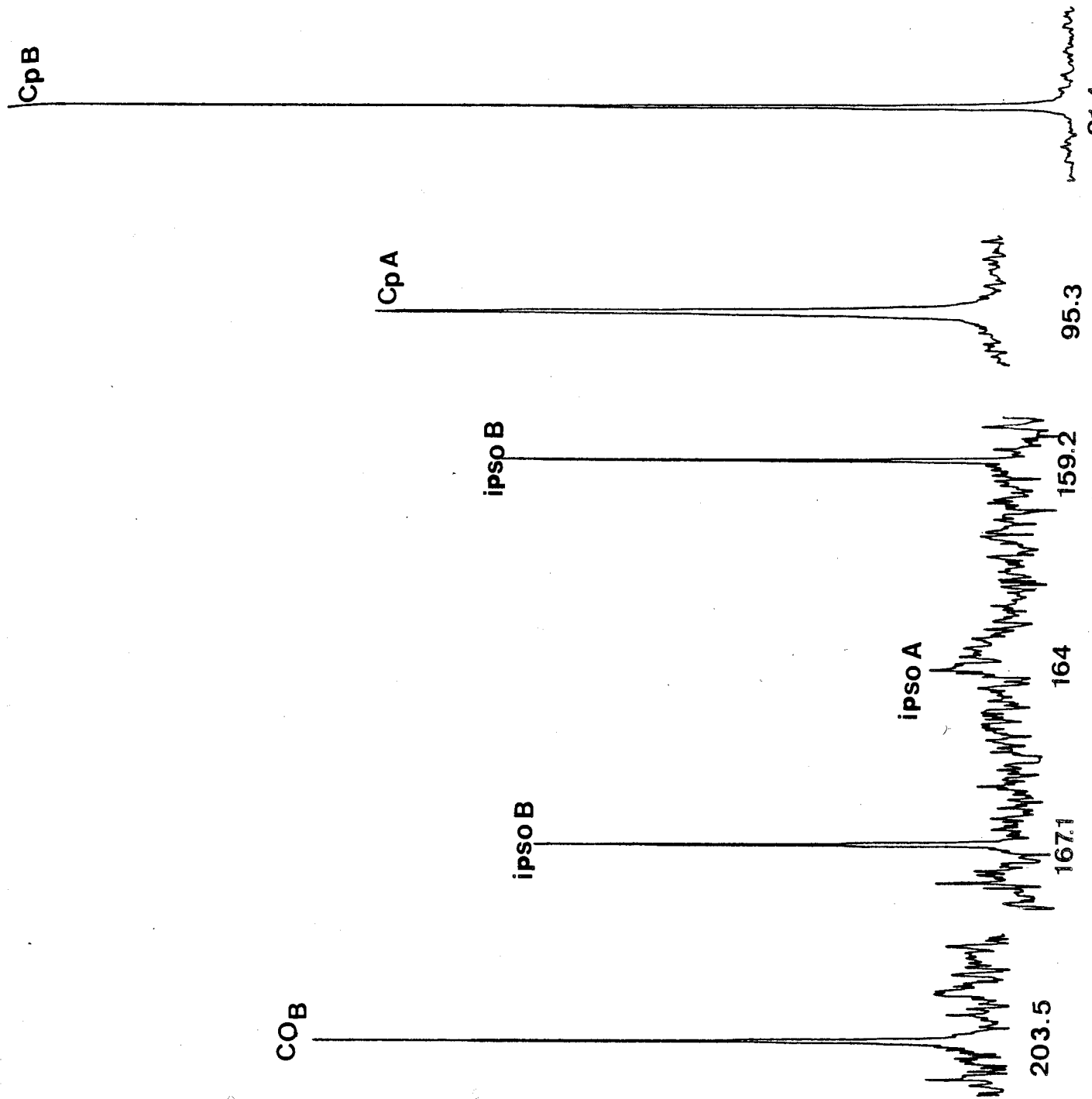
region was observed which was due to the major isomer, substantiating that the major isomer is the  $\sigma$ -bonded azobenzene complex with free rotation about the Re-N bond. Resonances due to the carbonyl carbons of the minor isomers were not observed. The ipso carbon region shows two sharp peaks due to the major isomer (167.2 and 159.2 ppm) and a small broad resonance at approximately 163 ppm. This broad resonance could certainly be due to the ipso carbons of the minor isomers which have not been resolved. They would appear to be more similar in chemical shift than the ipso carbons of the major isomer. A  $\pi$  or  $\eta^2$ -bonded arrangement of the azobenzene ligand to the rhenium could account for this observation. It would not be expected that two ipso carbons of the azobenzene ligand, that is bound to the rhenium atom through both nitrogens so that the ipso carbons differ only in the spatial orientation of the phenyl ligand, would be as magnetically different as the two ipso carbons of an azobenzene ligand in which one nitrogen donates its lone pair to the Re and the second is unattached. The Cp resonance of the minor isomer(s) appears at 95.3 ppm. This peak is substantially broader at half height than the Cp resonance due to the major isomer (81.4 ppm), but separation into the two or more peaks which would be required for 2 or more different minor species, was not observed. However, the broadness of this resonance is certainly suggestive of an exchange process occurring within the  $\eta^2$  bound form which does not involve the  $\eta^1$ - $\eta^2$  shift.

The interconversions between the  $\eta^1$ - $\eta^2$  species and within the  $\eta^2$  bound isomers may be summarized in Scheme III-2. The major isomer in solution is isomer B (only the cis form B-1, is drawn) which can undergo a  $\sigma$ - $\pi$  or  $\eta^1$ - $\eta^2$  shift to form A-1 or A-1'. These two enantiomers differ only in which phenyl ligand on N<sub>1</sub> or N<sub>2</sub> is 'close to' the Cp ring or the carbonyl ligands, and are indistinguishable to us in the IR or NMR spectrum. The nitrogen

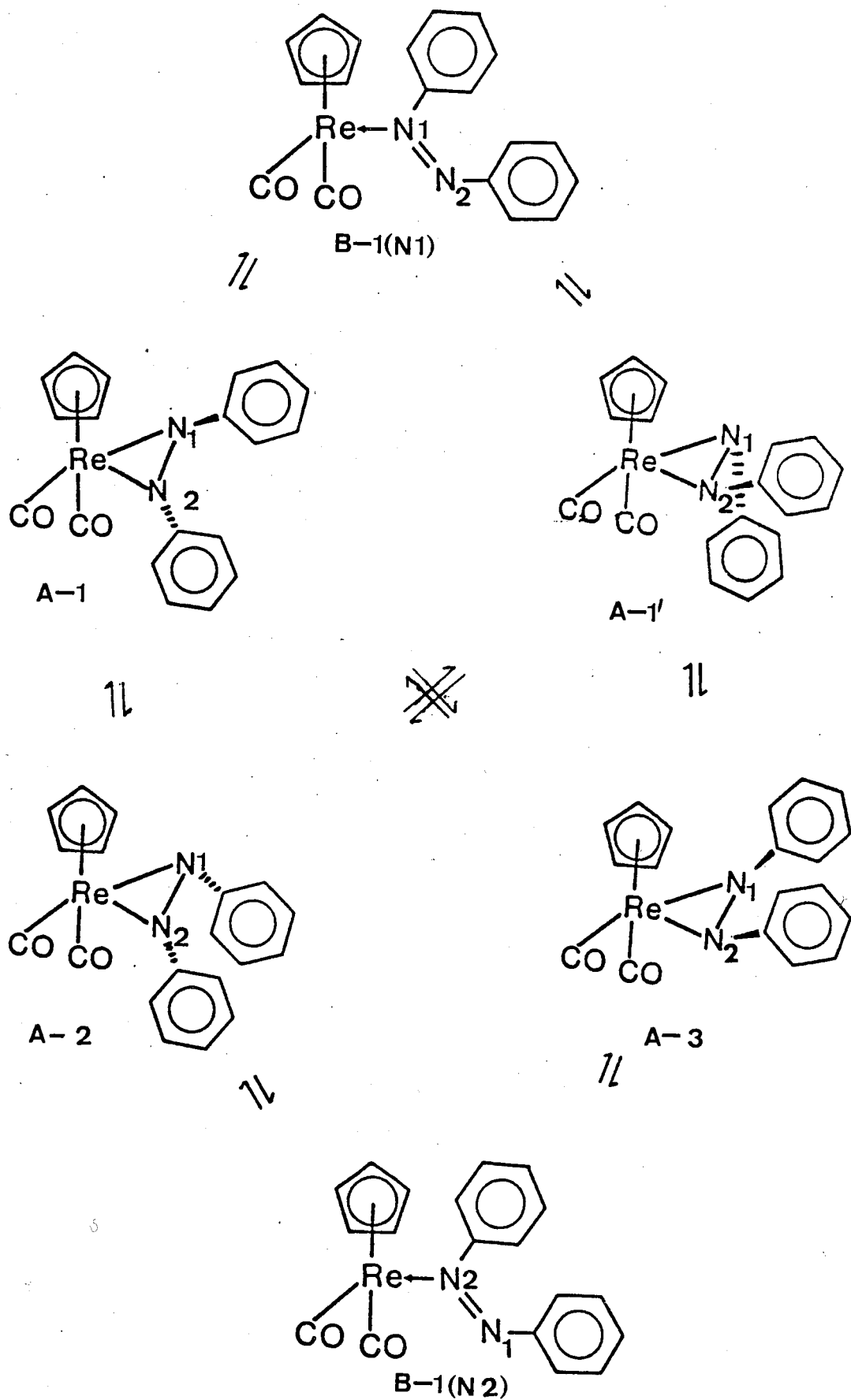


## FIGURE 29

$^{13}\text{C}\{^1\text{H}\}$  NMR (acetone- $d_6$ ) of the CO, ipso and Cp region  
of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13) at 189 K



## SCHEME III-2



labelled  $N_1$  on isomer A-1 may then undergo a nitrogen inversion to yield A-2, where both phenyl ligands are cis and 'towards' the carbonyl ligand, or the nitrogen labelled  $N_2$  may invert to form A-3 which has the phenyl groups cis and 'towards' the Cp ring. Similar arguments hold for A-1'. A  $\pi$ - $\sigma$  or  $\eta^2$ - $\eta^1$  shift of the azobenzene ligand in A-2 or A-3 could reform B-1 with  $N_2$  acting as the Lewis base.

This scheme accounts for the reversible  $\eta^1$ - $\eta^2$  shift between B-1 and A-1 (or A-1') and also the large number of aromatic resonances observed in the  $^{13}\text{C}\{^1\text{H}\}$  spectrum of 13 at 189 K (Figure 27) for the minor isomers. The compounds A-1, A-2 and A-3 have 12 different phenyl carbon resonances (not including ipso carbons) and we observe at least 10 or 11 of them. However differing populations of the A-1, A-2 or A-3 isomers may account for the absence of, and varying intensities of the resonances.

Nitrogen inversion e.g., A-1 to A-2, is expected to have a lower barrier than an  $\eta^1$ - $\eta^2$  shift (B-1 to A-1) of the azobenzene ligand which involves substantial rehybridization of the orbitals. Thus it is not surprising that the low energy interconversions in this system appear to involve inversion at the nitrogen atoms. However, it is noted that diaziridines<sup>138,139</sup>, which have a CNN ring versus our ReNN ring, can have a high barrier to inversion (111-115  $\text{kJ mol}^{-1}$ ) depending upon the substituents on the nitrogens and the carbon atom. In scheme III-2, only interconversion between stereoisomers by a consecutive nitrogen inversion is assumed and two synchronous inversions of the nitrogens i.e., A-1 to A-1' or A-2 to A-3, is not considered. This assumption is supported by calculations by Mannschreck<sup>138</sup> which indicate that the mono-planar transition state for inversion at one nitrogen atom is much lower in energy than the bi-planar transition state for two synchronous inversions. The latter transition

state will suffer from the high strain energy caused by the two planar  $sp^2$  hybridized nitrogen atoms. While our system is very different from the diaziridine system used for the calculations, since we have introduced a transition metal atom, there is no evidence to support a synchronous inversion in this case and it is presumed not to be occurring.

Ackermann<sup>134,135</sup> had only observed coordination site exchange (an  $\eta^1-\eta^1$  shift) between nitrogens of a cis diazene ligand and not with the corresponding trans diazenes, and did not observe an  $\eta^1-\eta^2$  shift with either a cis or trans azo ligand. However, an  $\eta^2$  transition state was postulated in the  $\eta^1-\eta^1$  shift of the cis azo ligand. Since our  $\eta^1-\eta^2$  shift had been noted with trans-azobenzene as the starting material, in order to test whether there would be a further, different isomer formed that would display different site exchange behaviour in our system, the reaction of cis-azobenzene and  $CpRe(CO)_2(THF)$  was carried out. After chromatography, a brown compound with an identical IR and  $^1H$  NMR spectrum to 13 was isolated. No other compounds were isolated. This supports the interconversions outlined in Scheme III-2. Both cis and trans azobenzene will yield the same mixture of products upon reacting with  $CpRe(CO)_2(THF)$  because both the cis and trans azo ligands go through common intermediates when coordinated in the  $\eta^2$  form.

### 3.4.2b $CpRe(CO)_2(N_2(C_6H_4Me-p)_2)$ (14) and $CpRe(CO)_2(N_2(C_6H_4F-p)_2)$ (15)

The  $\eta^2$ -bonded azo structure has been shown by Albini and Kisch<sup>140</sup> to be a better  $\pi$ -acceptor than the  $\eta^1$ -form. Any substituent on the phenyl ring of the azo ligand that increases the  $\pi$  acidity should increase the proportion of the  $\eta^2$  isomer. Conversely, an electron donating substituent

on the phenyl rings should disfavour the  $\eta^2$  form and favour the  $\eta^1$  structure in which the azo ligand functions as a Lewis base. To substantiate this conjecture the p,p'-methylazobenzene and p,p'-fluoroazobenzene were synthesized and reacted with  $\text{CpRe}(\text{CO})_2(\text{THF})$  to form  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$  (14) and  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$  (15), respectively. The IR's of 14 and 15 have been mentioned briefly previously (Section 3.4.2) and are shown in Figure 22. The use of a methyl group in the para position, compound 14, causes the higher wavenumber  $\nu(\text{CO})$  bands to almost disappear. With fluorine in the para position, 15, the proportion of the higher wavenumber  $\nu(\text{CO})$  bands increases, although not to the same extent as the decrease noted for 14. These changes in the intensities of the  $\nu(\text{CO})$  bands for 14 and 15 are consistent with the higher wavenumber  $\nu(\text{CO})$  absorptions being due to the  $\eta^2$  or  $\pi$ -bound isomer and the lower wavenumber  $\nu(\text{CO})$  bands to the  $\eta^1$  or  $\sigma$ -bound azo ligand, and so further support the assignments already adopted.

The  $^1\text{H}$  and  $^{13}\text{C}\{^1\text{H}\}$  NMR of  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$  (14) showed resonances due to two non-identical aromatic phenyl rings and two methyl groups. The Cp resonance due to the minor isomer A-1 in  $\text{CDCl}_3$  was broad and only slightly above the baseline. For  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$  (15) the Cp resonance due to the minor isomer was broad but of reasonable intensity; however, other peaks in the aromatic region were not observed.

Unfortunately, 15 was not stable enough in solution for long term NMR studies to be carried out. However, a relatively quick coalescence experiment was performed on 15, with a consequent large error in the temperature (see Experimental) and yielded a  $\Delta G^\ddagger$  of  $63.6 \pm 1.2 \text{ kJ mol}^{-1}$ . This is slightly smaller than the  $\Delta G^\ddagger$  value of  $67.4 \pm 0.4 \text{ kJ mol}^{-1}$  obtained for  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ . This suggests that the  $\eta^1$ - $\eta^2$  shift has a lower energy barrier when the  $\eta^2$  form has electron-withdrawing groups on the phenyl

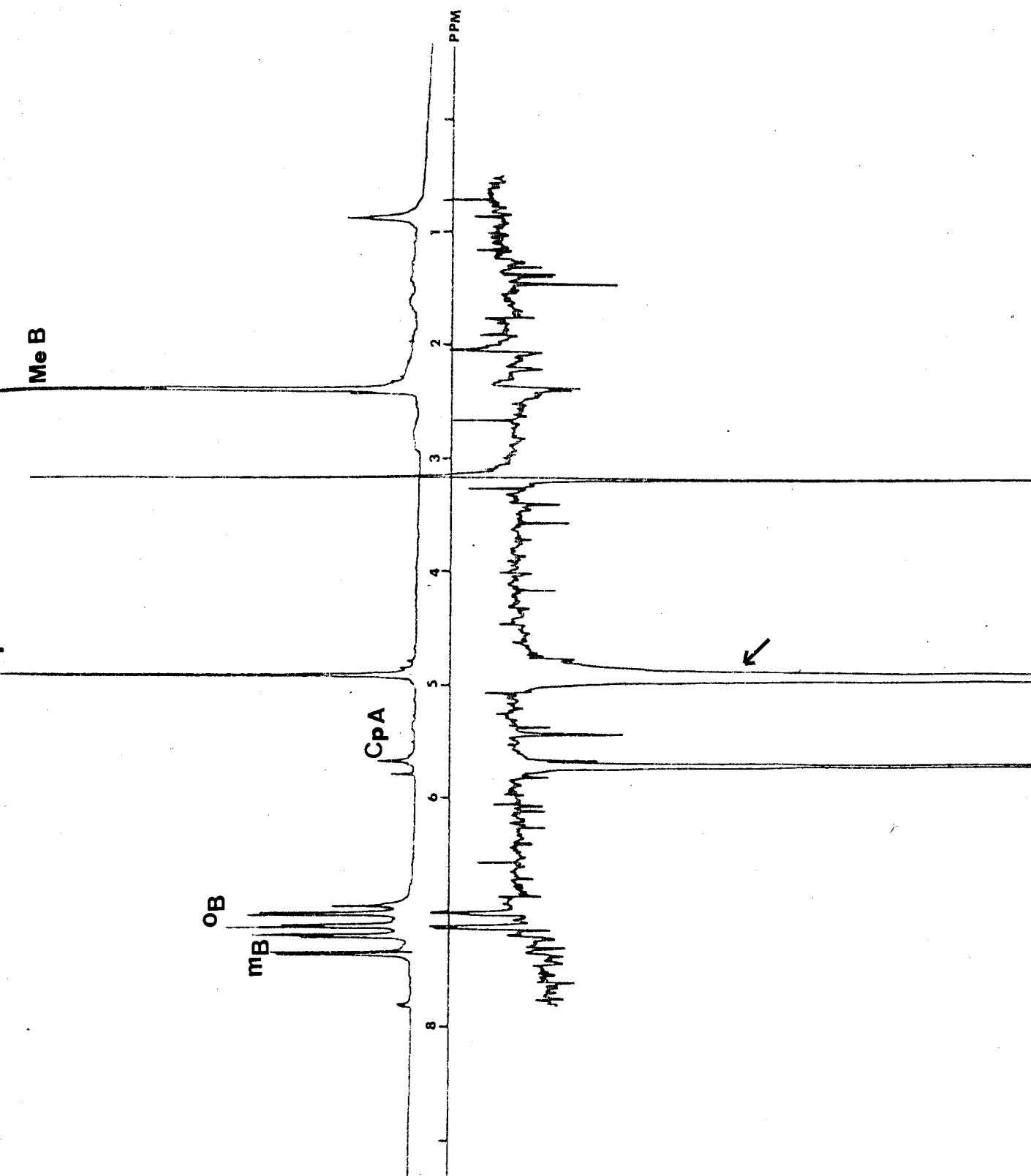
ligands which help to favour the  $\eta^2$  structure (isomer A).

An NOE experiment was performed on  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$  (14). The Cp resonance of the major isomer (CpB) was irradiated (250 K) and the results of this experiment are shown in Figure 29. The ortho protons of both phenyl rings of the azotoluene ligand were found to be enhanced but not the meta protons. No enhancement of the para methyl hydrogens was found. Saturation transfer from the Cp resonance of the major isomer CpB, to the Cp resonance of the minor isomer CpA was observed, similar to the results found for  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13), establishing the interconversion process between isomer A and isomer B.

FIGURE 30

NOE Experiment on  
 $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$  (14)





### Summary

The synthesis of  $\text{CpRe}(\text{CO})_2\text{L}$  complexes, where L is a two electron donor such as  $\text{C}_2\text{Ph}_2$ ,  $\text{MeC}_2\text{Ph}$ ,  $\text{N}_2\text{Ph}_2$ ,  $\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2$  and  $\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2$  was accomplished. The X-ray crystal structures of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11) and  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13) reveal both the diphenylacetylene and the azobenzene ligand to be bound to the rhenium in an  $\eta^2$  fashion and the overall geometry of the complex approximates that of a 'four legged piano stool'. Fluxionality of the alkyne ligands was restricted to rotation about the Re-alkyne vector and for  $\text{CpRe}(\text{CO})(\text{MeC}_2\text{Ph})$  (12), this rotation could be stopped at  $-60^\circ\text{C}$ .

For the azobenzene complex  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13), a different type of fluxionality was present. The azobenzene ligand exhibited an  $\eta^1$ - $\eta^2$  shift with the  $\eta^1$  (or  $\sigma$ ) form being predominant in solution and the  $\eta^2$  (or  $\pi$ ) structure present as the minor isomer. The  $^1\text{H}$ ,  $^{13}\text{C}\{^1\text{H}\}$  NMR and IR data support the  $\eta^1$  isomer being the major one in solution and saturation experiments show a reversible equilibrium occurring between the major and minor isomer. At room temperature the  $^{13}\text{C}\{^1\text{H}\}$  NMR spectrum had only very broad peaks in the phenyl carbon region attributable to the minor isomer. Cooling the solution to 189 K resulted in at least 11 distinct resonances being noted. This temperature dependence and larger number of minor isomers is possible because the nitrogens in the  $\eta^2$  isomer can undergo nitrogen inversion. This is a low energy process causing the phenyl carbons to coalesce at room temperature but at lower temperatures the inversion is slow enough to observe at least three different minor isomers.

Coalescence of the Cp resonances of the major and minor isomers of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13) gave a value of  $67.4 \pm 0.4 \text{ kJ mol}^{-1}$  for the  $\eta^1$ - $\eta^2$  shift of the azobenzene ligand. This is the first observed  $\eta^1$ - $\eta^2$  shift of

an azobenzene ligand and only the second example of this type of shift for any ligand<sup>136</sup>.

Ackermann<sup>134</sup> has observed coordination site exchange or an  $\eta^1$ - $\eta^1$  shift for the azo ligand in  $M(\text{CO})_5(\underline{c}\text{-DIPD})$  ( $M = \text{Cr, Mo, W}$ ) and  $\text{W}(\text{CO})_5(\underline{c}\text{-DMD})$ . Although an  $\eta^2$  or  $\pi$  transition state was proposed as the intermediate in the  $\eta^1$ - $\eta^1$  shift, no experimental evidence was forwarded. Coordination site exchange was not found for trans azo ligands, specifically  $M(\text{CO})_5(\underline{t}\text{-DIPD})$  ( $M = \text{Cr, Mo, W}$ ). However, we observe an  $\eta^1$ - $\eta^2$  shift of the azobenzene ligand in  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  and in the  $\eta^2$  isomer, as confirmed by X-ray crystallography, the azobenzene ligand has a trans geometry with the phenyl rings in nonequivalent environments.

$\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$  (15) has a fluorine in the para position and the  $\nu(\text{CO})$  bands in the solution IR which are attributable to the  $\eta^2$  isomer increased in intensity. When a methyl group was substituted in the para position, as in  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{Me-p})_2)$ , the  $\eta^1$  isomer was found to be heavily favoured. These observations are consistent with an  $\eta^1$ - $\eta^2$  shift, since the use of an electron withdrawing group should favour the  $\eta^2$  form relative to the  $\eta^1$  isomer because in the  $\eta^2$  geometry the azo ligand is a better  $\pi$  acid than in the  $\eta^1$  form<sup>140</sup>. Conversely, an electron donating group should favour the  $\eta^1$  isomer because in this geometry the azo ligand is acting as a Lewis base donating the lone pair of electrons of a nitrogen to the rhenium.

## 3.6

## EXPERIMENTAL SECTION

3.6.1 Synthetic

$^1\text{H}$  and  $^{13}\text{C}$  NMR spectra were recorded on a Brüker WM-400 spectrometer at 400 MHz and 100 MHz, respectively. Infrared spectra were measured on a Perkin-Elmer 983 spectrophotometer, and Raman spectra using a Cary Model 81 with a Spectra-Physics Model 125 HeNe laser. Solvents were dried and distilled under nitrogen and all operations were conducted under nitrogen.  $\text{CpRe}(\text{CO})_2(\text{THF})$  was synthesized and isolated as previously described<sup>138</sup>. Azobenzene was purchased from Eastman Kodak, azotoluene and p,p'-difluoroazobenzene were made by the reduction of nitrotoluene and p-fluoronitrotoluene respectively by the method of Alper<sup>139</sup>. Cis-azobenzene was made by photolysis of the trans isomer and the two compounds separated by column chromatography.

Reaction of  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  with  $\text{CpRe}(\text{CO})_2(\text{THF})$ 

$\text{CpRe}(\text{CO})_2(\text{THF})$  (60 mg, 0.158 mmol) and  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  (70 mg, 0.157 mmol) were added to hexane (25 mL) and stirred at room temperature for 2 hours. An infrared spectrum of the solution revealed little remaining  $\text{CpRe}(\text{CO})_2(\text{THF})$  ( $\nu(\text{CO})$  1925, 1857  $\text{cm}^{-1}$ ) but strong bands due to  $\text{CpRe}(\text{CO})_3$  ( $\nu(\text{CO})$  2031, 1940  $\text{cm}^{-1}$ ). The solvent was removed under vacuum and the residue dissolved in hexane- $\text{CH}_2\text{Cl}_2$ . Chromatography on a Florisil column with hexanes as eluant yielded  $\text{CpRe}(\text{CO})_3$ , then  $\text{CpNiCo}(\text{CO})_3(\text{C}_2\text{Ph}_2)$  ( $\nu(\text{CO})$  2060, 2008  $\text{cm}^{-1}$ ) as a green solution. With hexanes: $\text{CH}_2\text{Cl}_2$  (3:1) a light yellow band was next recovered, and identified as  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ , with physical properties identical to those listed below.

Preparation of CpRe(CO)<sub>2</sub>(C<sub>2</sub>Ph<sub>2</sub>) (11)Method 1

CpRe(CO)<sub>2</sub>(THF) (51 mg, 0.132 mmol) and a large excess of C<sub>2</sub>Ph<sub>2</sub> (115 mg, 0.645 mmol) was added to 12 mL of hexanes and stirred vigorously for 4 hours. Chromatography on Florisil with hexane as an eluant gave yellow CpRe(CO)<sub>2</sub>(C<sub>2</sub>Ph<sub>2</sub>) (19 mg, 30% yield). IR (hexane):  $\nu(\text{CO})$ : 1980 s, 1904 s  $\text{cm}^{-1}$ . Raman (solid): 1848  $\text{cm}^{-1}$  ( $\nu(\text{C}=\text{C})$ ). <sup>1</sup>H NMR (CDCl<sub>3</sub>):  $\delta$  7.85-7.32 m (Ph), 5.55 s (Cp). <sup>13</sup>C{<sup>1</sup>H} NMR (CDCl<sub>3</sub>):  $\delta$  204.5 (CO), 131.8, 131.6, 128.6, 127.6 (Ph), 88.2 (Cp), 81.7 (acetylenic carbons). Mass spectrum: m/e (based on Re<sup>185</sup> and Re<sup>187</sup>) 484, 486 (M<sup>+</sup>), 456, 458 (M-CO<sup>+</sup>), 428, 430 (M-2CO<sup>+</sup>). Anal. calcd.: C, 51.95; H, 3.11. Found: C, 52.21; H, 3.26%.

Method 2

CpRe(CO)<sub>3</sub> (130 mg, 0.336 mmol) and C<sub>2</sub>Ph<sub>2</sub> (220 mg, 1.23 mmol) were added to 30 mL hexanes in a quartz Carius tube under vacuum and irradiated with a 200 W Hanovia high pressure mercury source. Chromatography on Florisil yielded CpRe(CO)<sub>3</sub> (colorless) and CpRe(CO)<sub>2</sub>(C<sub>2</sub>Ph<sub>2</sub>) (yellow) (18 mg, 11% yield).

Preparation of CpRe(CO)<sub>2</sub>(PhC<sub>2</sub>Me) (12)

This compound was synthesized as in Method 1 (above) using CpRe(CO)<sub>2</sub>(THF) (100 mg, 0.259 mmol) and PhC<sub>2</sub>Me (0.15 mL, ~1.2 mmol). CpRe(CO)<sub>2</sub>(PhC<sub>2</sub>Me) was isolated, after chromatography, as an oily, yellow solid and identified by spectroscopy. IR (hexane):  $\nu(\text{CO})$ : 1976 s, 1899 s  $\text{cm}^{-1}$ . <sup>1</sup>H NMR (CDCl<sub>3</sub>):  $\delta$  204.9 (CO), 131.4, 128.3, 126.8 (Ph), 87.9 (Cp), 78.2 (CMe), 69.4 (CPh), 15.2 (Me). (in the <sup>1</sup>H-coupled spectrum:  $\delta$  78.2 q

( $^2J=9$  Hz), 69.4 s, 15.2 q ( $^1J=130$  Hz). At  $-60^\circ\text{C}$   $^{13}\text{C}\{^1\text{H}\}$  coupled spectrum:  $\delta$  206.5 s, 206.0 s (CO), 132.6-126.2 complex (Ph), 88.3 ( $^1J=180$  Hz, Cp), 78.9 q ( $^2J=9$  Hz, CMe), 69.3 s (CPh), 15.9 q ( $^1J=130$  Hz, Me) Mass spectrum: m/e 422, 424 ( $\text{M}^+$ ), 394, 396 ( $\text{M}-\text{CO}^+$ ), 366, 368 ( $\text{M}-2\text{CO}^+$ ).

#### Oxidative Stability of $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ (11)

The color of solid samples of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (yellow) remained unchanged after standing in air for one week and they could then be dissolved to give essentially clear IR and NMR spectra. A solution of  $\text{CpRe}(\text{CO})(\text{C}_2\text{Ph}_2)$  in hexanes was left open to the air and periodically monitored by IR. No decomposition was noted in 12 hours.

#### Thermal Stability of $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$

A solution of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  in iso-octane was refluxed ( $99^\circ\text{C}$ ) for 8 hours and no decomposition or loss of CO was noted by IR. A solution of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  in THF was refluxed ( $66^\circ\text{C}$ ) for 1 hour, after which no starting material remained and no carbonyl containing products were apparent from the IR spectrum.

#### Reaction of $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ with Carbon Monoxide

$\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (15 mg, 0.039 mmol) was dissolved in 10 mL hexanes in a Paar bomb and pressurized with CO (1200 psi, 82 atm). After 1 hour, the pressure was released and an IR spectrum of the yellow solution revealed only CO bands due to 11 and none for  $\text{CpRe}(\text{CO})_3$ , ( $2031, 1940\text{ cm}^{-1}$ ).

#### Reaction of $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ with $\text{PPh}_3$

To a solution of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (15 mg, 0.039 mmol) in 6 mL hexanes,

octane or acetonitrile, excess  $\text{PPh}_3$  (20 mg, 0.076 mmol) was added. As monitored by IR, no reaction was observed after 2 hours.  $\text{Me}_3\text{NO}$  (6 mg, 0.080 mmol) was then added to the acetonitrile solution but no reaction was noted after a further 2 hours. The  $\text{CH}_3\text{CN}$  solution was refluxed for 5 hours but no loss of CO ligand was observed.

#### Reaction of $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ with $\text{I}_2$

$\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (10 mg, 0.026 mmol) was dissolved in 5 mL hexanes and  $\text{I}_2$  (10 mg, 0.039 mmol) in 5 mL  $\text{CH}_2\text{Cl}_2$  was added slowly to the solution. After 2 hours only bands for trans- $\text{CpRe}(\text{CO})_2\text{I}_2$  ( $\nu(\text{CO})$  2048 m, 1987 s  $\text{cm}^{-1}$ ) were observed in the IR.

#### Reaction of $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ with Azobenzene

$\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (75 mg, 0.195 mmol) and excess  $\text{N}_2\text{Ph}_2$  (90 mg, 0.494 mmol) were added to 10-15 mL hexanes and vigorously stirred at room temperature. After 2 hours, monitoring of the IR revealed only 11 present and no evidence for  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  ( $\nu(\text{CO})$  2007 m, 1948 s, 1939 m, 1890 s  $\text{cm}^{-1}$ ) were noted. The solution was warmed to 55°C for three hours but still the IR showed that  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  was not formed.

#### Reaction of $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$ with $\text{H}_2$

To a solution of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (50 mg, 0.130 mmol) in 8 mL hexanes,  $\text{H}_2$  was bubbled through at a constant rate. No new  $\nu(\text{CO})$  bands in the infrared spectrum were observed and the experiment was discontinued after 1 hour.

Reaction of CpRe(CO)<sub>3</sub> with C<sub>2</sub>Ph<sub>2</sub>

CpRe(CO)<sub>3</sub> (50 mg, 0.149 mmol) and C<sub>2</sub>Ph<sub>2</sub> (31 mg, 0.174 mmol) were added to 15 mL hexanes-CH<sub>2</sub>Cl<sub>2</sub> (13:2) and stirred overnight at room temperature. Inspection by IR revealed only bands due to CpRe(CO)<sub>3</sub> ( $\nu(\text{CO})$  2031, 1940 cm<sup>-1</sup>) and none for 11. Me<sub>3</sub>NO (8 mg, 0.107 mmol) was added to the solution and stirred for a further two hours. As there were no new  $\nu(\text{CO})$  bands the solution was heated to 55°C for an additional two hours at which point the reaction was discontinued as no additional  $\nu(\text{CO})$  bands were apparent.

Preparation of CpRe(CO)<sub>2</sub>(N<sub>2</sub>Ph<sub>2</sub>) (13)

In a typical reaction, excess azobenzene was added to a hexane solution of CpRe(CO)<sub>2</sub>(N<sub>2</sub>Ph<sub>2</sub>) (100 mg, 0.26 mmol) and stirred vigorously for 6 hours. Chromatography on Florisil with hexanes yielded excess azobenzene and with a 1:1 hexane-CH<sub>2</sub>Cl<sub>2</sub> as eluant a brown band identified as CpRe(CO)<sub>2</sub>(N<sub>2</sub>Ph<sub>2</sub>) was isolated in 25% yield. IR (hexane): 2006 m, 1947 s, 1939 m, 1890 s cm<sup>-1</sup>  $\nu(\text{CO})$ , IR (KBr): 2001 s, sl.br., 1928 s, 1900 s cm<sup>-1</sup>,  $\nu(\text{CO})$ . <sup>1</sup>H NMR (Acetone-d<sub>6</sub>):  $\delta$  7.59, 7.57, 7.55 (Ph, meta), 7.45, 7.44, 7.42, 7.40 (Ph, meta and para), 7.34, 7.32, 7.30 (Ph, para), 7.25, 7.23 (Ph, ortho), 7.14, 7.12 (Ph, ortho), 5.72 s (Cp, minor isomer), 4.95 s (Cp, major isomer). <sup>13</sup>C{<sup>1</sup>H} NMR (CDCl<sub>3</sub>):  $\delta$  201.1 (CO), 166.8, 158.6 (Ph, ipso) 129.4, 127.8 (Ph, ortho), 127.7, 127.0 (Ph, meta), 121.3, 117.7 (Ph, para), 93.4 (Cp, minor isomer), 85.8 (Cp, major isomer). Mass spectrum: m/e (based on Re<sup>185</sup> and Re<sup>187</sup>) 489, 491 (M<sup>+</sup>), 461, 463 (M-CO<sup>+</sup>), 433, 435 (M-2CO<sup>+</sup>). Anal. calcd.: C, 46.62, H, 3.09, N, 5.72; found: C, 46.81, H, 3.11 N, 5.80%.

Preparation of CpRe(CO)<sub>2</sub>(N<sub>2</sub>(C<sub>6</sub>H<sub>4</sub>Me-p)<sub>2</sub>) (14)

CpRe(CO)<sub>2</sub>(N<sub>2</sub>(C<sub>6</sub>H<sub>4</sub>Me-p)<sub>2</sub>) was prepared in an identical fashion to 13 in



20% yield using azotoluene rather than azobenzene. IR (hexane): 2004 w, 1945 s, 1887 s  $\text{cm}^{-1}$ . IR (KBr): 1997 w, 1932 s, 1868 s  $\text{cm}^{-1}$   $\nu(\text{CO})$ .  $^1\text{H}$  NMR (Acetone- $d_6$ ):  $\delta$  7.30, 7.28 (Ph, meta), 7.22, 7.20 (Ph, meta), 7.14, 7.12 (Ph, ortho), 7.04, 7.02 (Ph, ortho), 5.68 s (Cp, minor isomer), 4.92 s (Cp, major isomer), 2.40, 2.38 (Me).  $^{13}\text{C}\{^1\text{H}\}$  NMR ( $\text{CDCl}_3$ ):  $\delta$  202.4 (CO), 156.7, 156.1 (Ph, ipso), 141.1, 138.9 (Ph, para), 129.9, 129.7, 128.2, 122.7, 121.2, 117.8 (Ph), 85.6 (Cp), 29.7 ( $\text{CH}_3$ ), 21.4 ( $\text{CH}_3$ ). Mass spectrum: m/e, 516, 518 ( $\text{M}^+$ ), 488, 490 ( $\text{M}-\text{CO}^+$ ), 460, 462 ( $\text{M}-2\text{CO}^+$ ). Anal. calcd.: C, 48.73, H, 3.70, N, 5.41; found: C, 48.84, H, 3.80, N, 5.48%.

#### Preparation of $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$ (15)

$\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$  was prepared in an identical fashion to 13 but using p,p'-difluoroazobenzene rather than azobenzene. IR (hexanes): 2008 s, 1950 s, 1941 s, 1893 s  $\text{cm}^{-1}$ .  $^1\text{H}$  NMR ( $\text{CDCl}_3$ ):  $\delta$  6.8-7.4 (Ph), 5.83 s (Cp, minor isomer), 5.05 s (Cp, major isomer).  $^{13}\text{C}\{^1\text{H}\}$  ( $\text{CDCl}_3$ ):  $\delta$  202.2 (CO), 162.4, 160.0 (Ph, para,  $^1\text{J}_{\text{C-F}}=246$  Hz), 154.07 (Ph, ipso), 131.1, 128.9 (Ph, para,  $^1\text{J}_{\text{C-F}}=223$  Hz), 123.12, 123.04 (Ph, ortho,  $^3\text{J}_{\text{C-F}}=8.4$  Hz), 119.53, 119.44 (Ph, ortho,  $^3\text{J}_{\text{C-F}}=8.5$  Hz), 116.54, 116.32 (Ph, meta,  $^2\text{J}_{\text{C-F}}=22.6$  Hz), 114.64, 114.41 (Ph, meta,  $^2\text{J}_{\text{C-F}}=22.7$  Hz), 93.4 s (Cp, minor isomer), 86.8 s (Cp, major isomer). Mass spectrum: m/e, 524, 526 ( $\text{M}^+$ ), 496, 498 ( $\text{M}-\text{CO}^+$ ), 468, 470 ( $\text{M}-2\text{CO}^+$ ).

#### Reaction of $\text{CpRe}(\text{CO})_2(\text{THF})$ with cis-azobenzene

Excess cis-azobenzene was added to a hexane (8 mL) solution of  $\text{CpRe}(\text{CO})_2(\text{THF})$  (20 mg, 0.041 mmol) and stirred vigorously overnight. Chromatography on Florisil with hexane- $\text{CH}_2\text{Cl}_2$  (1:2) yielded a dark brown compound with an identical IR and NMR to that found for  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ .

( $\nu(\text{CO})$  (hexane) 2006 m, 1947 s, 1939 m, 1890 s  $\text{cm}^{-1}$ ).

Reaction of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  with  $\text{C}_2\text{Ph}_2$

$\text{C}_2\text{Ph}_2$  (16 mg, 0.089 mmol) and  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (20 mg, 0.041 mmol) were added to 6-7 mL of hexanes and 0.5 mL of  $\text{CH}_2\text{Cl}_2$ . The reaction was stirred at room temperature for 2 hours and the IR spectrum showed only 13 and no  $\nu(\text{CO})$  bands due to  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  ( $\nu(\text{CO})$  1980, 1904  $\text{cm}^{-1}$ ). The hexanes- $\text{CH}_2\text{Cl}_2$  was removed and 6 mL of benzene was added. After stirring at ambient temperature for 1 hour, the solution was warmed to 40°C where after 1 hour no formation of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  was detected, although some decomposition was noted.

Reaction of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  with  $\text{I}_2$

To a solution of 13 (920 mg, 0.041 mmol) in 5-6 mL hexanes an  $\text{I}_2$ /hexanes solution was added dropwise until the IR revealed the  $\nu(\text{CO})$  bands due to 13 had disappeared. The hexanes were removed and the residue dissolved in  $\text{CH}_2\text{Cl}_2$ . An IR spectrum revealed two carbonyl absorptions (2041 s, 1977 m  $\text{cm}^{-1}$ ) corresponding to cis- $\text{CpRe}(\text{CO})_2\text{I}_2$ . When the identical experiment was conducted using benzene as solvent, only trans- $\text{CpRe}(\text{CO})_2\text{I}_2$  was formed ( $\nu(\text{CO})$  ( $\text{CH}_2\text{Cl}_2$ ): 2047 m, 1988 s  $\text{cm}^{-1}$ ). A control experiment revealed that pure cis- $\text{CpRe}(\text{CO})_2\text{I}_2$  was fully converted into trans- $\text{CpRe}(\text{CO})_2\text{I}_2$  in benzene in 10 minutes at ambient temperature.

Reaction of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  with Carbon Monoxide

A hexane-benzene solution of  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (15 mg, 0.031 mmol) was purged with CO gas. As monitored by IR, no formation of  $\text{CpRe}(\text{CO})_3$  ( $\nu(\text{CO})$ ; 2031, 1940  $\text{cm}^{-1}$ ) was noted, nor that of any other carbonyl containing

species.

Reaction of  $\text{CpRe}(\text{CO})_3$  with  $\text{N}_2\text{Ph}_2$

$\text{CpRe}(\text{CO})_3$  (25 mg, 0.075 mmol) and azobenzene (28 mg, 0.15 mmol) were added to 12 mL hexanes- $\text{CH}_2\text{Cl}_2$  (10:2) and stirred overnight. Inspection by IR revealed bands due to  $\text{CpRe}(\text{CO})_3$  only and none for  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ . Addition of  $\text{Me}_3\text{NO}$  or iodobenzene to the solution and further stirring for 2 hours did not yield any 13 as monitored by IR and the reaction was discontinued.

Photolysis of  $\text{CpRe}(\text{CO})_3$  with  $\text{N}_2\text{Ph}_2$

Photolysis of  $\text{CpRe}(\text{CO})_3$  (20 mg, 0.060 mmol) and azobenzene (27 mg, 0.14 mmol) in 15 mL hexanes using pyrex-filtered UV radiation yielded only small amounts of 13 as monitored by IR. The reaction was discontinued after 2 hours because of extensive darkening of the solution.

Calculation of  $\Delta G^\ddagger$  for  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (toluene- $d_8$ )

The variable temperature behaviour of the Cp resonances for isomers A and B was observed, with coalescence at  $332 \pm 1$  K.

$$T_c = 332 \pm 1 \text{ K}$$

$$\Delta\nu_\infty = 77.1 \pm 0.5 \text{ Hz}$$

$$K_c = \pi/\sqrt{2} \Delta\nu_\infty = 171.3 \pm 1.1 \text{ sec}^{-1}$$

and since,

$$\Delta G^\ddagger = -RT_c \ln(K_c h / KT_c)$$

where,

$$K = \text{Boltzmann constant} = 1.3807 \times 10^{-16} \text{ ergs K}^{-1}$$

$$h = \text{Plank's constant} = 6.63 \times 10^{-27} \text{ erg sec}$$

$$R = \text{gas constant} = 1.987 \text{ cal K}^{-1} \text{ mol}^{-1}$$

thus

$$\Delta G^\ddagger = 67.4 \pm 0.4 \text{ kJ mol}^{-1}$$

Calculation of  $\Delta G^\ddagger$  for  $\text{CpRe}(\text{CO})_2(\text{N}_2(\text{C}_6\text{H}_4\text{F-p})_2)$  (toluene- $d_8$ )

Coalescence of the Cp resonances of the A and B isomers was again monitored.

$$T_c = 315 \pm 3 \text{ K}$$

$$\Delta\nu_\infty = 82.6 \pm 0.5 \text{ Hz}$$

$$K_c = \pi/\sqrt{2} \pi\nu_\infty = 183.3 \pm 1.1 \text{ sec}^{-1}$$

and thus,

$$\Delta G^\ddagger = 63.6 \pm 1.2 \text{ kJ mol}^{-1}$$

### 3.6.2 X-ray Structures

X-ray Structure of  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11)

A pale yellow needle-like crystal, 0.06 x 0.11 x 0.30 mm, obtained from a hexanes solution, was chosen for the diffraction studies. Preliminary oscillation photos (Copper radiation) indicated the crystal to be of good quality. Data were collected on an Enraf-Nonius diffractometer (Mo radiation) with a scan speed of 0.8-4.0° min<sup>-1</sup> for 2.5° ≤ 2θ ≤ 45° and a symmetrical scan width of 0.9° plus dispersion correction. Two standards were measured for intensity control and two standards were measured for orientation control every 150 reflections. Neither pair showed any significant deviation. Absorption corrections were applied via an empirical

correction ( $\psi$  scans, transmission factors 0.447-0.993). 3008 reflections out of a total of 4545 had  $I \geq 3.0\sigma(I)$  and were regarded as observed.

The coordinates of the Re atoms were derived from a three dimensional Patterson map and non-hydrogen atoms from successive difference Fourier syntheses following least squares refinement. The hydrogen atom positions were calculated geometrically with  $d(C-H)$  of 0.95Å and assigned isotropic thermal parameters equal to their parent carbon atoms. All non-hydrogen atoms were refined anisotropically. The largest peaks in a final difference map,  $0.56(12)e\text{Å}^{-3}$ , were located near the Re atoms. Computer programs are from reference 65. A summary of the crystal data is in Table 40, bond distances and angles are in Table 41, atom coordinates in Table 43, anisotropic thermal parameters in Table 45, calculated hydrogen atom coordinates in Table 47 and calculated meanplanes and dihedral angles in Table 49. A labelled SNOOPI diagram is shown in Figure 18 and a view of the two independent molecules in the unit cell is shown in Figure 20.

#### X-ray Structure of $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$ (13)

A small orange crystal, 0.08 x 0.11 x 0.13 mm, obtained from hexanes/ $\text{CHCl}_3$ , was chosen for the diffraction studies. Data were collected on an Enraf Nonius diffractometer (Mo radiation) with a scan speed of 0.5-4.0°  $\text{min}^{-1}$  for  $3^\circ \leq 2\theta \leq 45^\circ$  and a symmetrical scan width of 0.6° plus dispersion correction. Two standards were measured for intensity control after every 3600 seconds of X-ray beam time and showed only random variations. Absorption corrections were applied via an empirical method ( $\psi$  scans, transmission factors 0.492-0.992). 991 reflections out of a total of 1180 had  $I \geq 3.0\sigma(I)$  and were regarded as observed.

The Re atom coordinates were derived from a three dimensional

Patterson map and non-hydrogen atoms from successive difference Fourier synthesis following least squares refinement. The hydrogen atom positions were calculated geometrically with  $d(\text{C-H})$  of 0.98Å and assigned isotropic temperature factors of  $0.075\text{Å}^2$  and included as fixed contributions in further refinement. The Re atom, carbonyl groups and nitrogen atoms were refined anisotropically and all other non-hydrogen atoms isotropically. An extinction correction was included and refined to  $0.473 \times 10^{-6}$ . To ascertain that the correct enantiomorph had been chosen, the coordinates of all the atoms were inverted and the model refined until there was no further change. This gave values of  $R_{\text{F}} = 0.0248$  and  $R_{\text{WF}} = 0.0279$  and  $\text{GOF} = 1.60$  and showed the original model to be correct. The coordinates were set back to their original values and a final difference map showed the highest peak to be  $0.55(21)\text{eÅ}^{-3}$  and located near the cyclopentadienyl ring. Computer programs are from reference 65. A summary of the crystal data is in Table 40, bond distances and angles in Table 42, atom coordinates in Table 44, anisotropic thermal parameters in Table 46, calculated hydrogen atom coordinates in Table 48 and calculated meanplanes and dihedral angles in Table 50. An SNOOPI diagram showing the labelling scheme is presented in Figure 19.

Table 40. Summary of Crystal Data and Intensity Collection for  
 $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11) and  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13)

compd	$\text{C}_{21}\text{H}_{15}\text{O}_2\text{Re}$	$\text{C}_{19}\text{H}_{15}\text{N}_2\text{O}_2\text{Re}$
mol wt	485.55	489.52
$\underline{a}$ , Å	10.841(1)	8.315(1)
$\underline{b}$ , Å	13.961(2)	11.214(2)
$\underline{c}$ , Å	23.121(2)	9.511(2)
$\beta$ , deg	91.360(8)	105.75(2)
cell vol, Å <sup>3</sup>	3498	853.6
$\underline{Z}$	8	2
$\rho$ (calcd) g cm <sup>-3</sup>	1.490	1.905
$\rho$ (obsd) g cm <sup>-1</sup>	1.53(1)	2.02(1)
space group	$\underline{P2}_1\underline{n}$	$\underline{P2}_1$
cryst dimens, mm	0.06x0.11x0.30	0.08x0.11x0.13
t, °C	21	21
radiation	$\text{MoK}_\alpha$	$\text{MoK}_\alpha$
linear abs coeff, cm <sup>-1</sup>	70.5	72.3
transmission factors	0.447-0.993(emperical)	0.492-0.992(emperical)
scan speed, deg/min	0.8-4.0 ( $\omega$ -2 $\theta$ scan)	0.5-4.0 ( $\omega$ -2 $\theta$ scan)
2 $\theta$ limits, deg	2.5, 45.0	3.0, 45.0
data collected	4545	1180
unique data, $I \geq 3.0\sigma(I)$	3008	991
$\underline{R}_f$ , %	2.95	2.44
$\underline{R}_{\text{wf}}$ , %	3.40	2.71
goodness of fit	1.17	1.55
extinction	-	$0.473 \times 10^{-6}$

Table 41. Bond Distances (Å) and Angles (deg) for  
 $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11).

Metal Atom Coordination

Bond Lengths

Re(1)-C(11)	1.893(11)	Re(2)-C(21)	1.916(11)
Re(1)-C(12)	1.906(12)	Re(2)-C(22)	1.902(10)
Re(1)-C(13)	2.188(10)	Re(2)-C(23)	2.190(9)
Re(1)-C(14)	2.182(9)	Re(2)-C(24)	2.194(9)
Re(1)-Cp(center)	2.039(12)	Re(2)-Cp(center)	1.968(12)

Bond Angles

C(13)-Re(1)-C(14)	33.5(4)	C(23)-Re(2)-C(24)	32.7(3)
Re(1)-C(13)-C(14)	73.0(6)	Re(2)-C(23)-C(24)	73.8(6)
Re(1)-C(14)-C(13)	73.5(6)	Re(2)-C(24)-C(23)	73.5(6)
Re(1)-C(13)-C(111)	134.4(7)	Re(2)-C(23)-C(211)	134.5(7)
Re(1)-C(14)-C(15)	135.3(7)	Re(2)-C(24)-C(25)	136.4(7)
Re(1)-C(11)-O(11)	177.2(9)	Re(2)-C(21)-O(21)	177.2(9)
Re(1)-C(12)-O(12)	177.6(10)	Re(2)-C(22)-O(22)	178.4(9)

Acetylene Coordination

Bond Lengths

C(13)-C(14)	1.261(14)	C(23)-C(24)	1.232(13)
C(13)-C(111)	1.45(1)	C(23)-C(211)	1.44(1)
C(14)-C(15)	1.43(1)	C(24)-C(25)	1.46(1)

Bond Angles

C(14)-C(13)-C(111)	152.6(10)	C(24)-C(23)-C(211)	151.6(9)
C(13)-C(14)-C(15)	151.1(10)	C(23)-C(24)-C(25)	149.9(9)



Table 42. Bond Distances (Å) and Angles (deg) for  
 $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13)

## Bond Lengths

Re-N(1)	2.048(12)	N(1)-N(2)	1.415(17)
Re-N(2)	2.136(11)	N(1)-C(3)	1.406(20)
Re-C(1)	1.979(18)	N(2)-C(9)	1.424(18)
Re-C(2)	1.800(21)	C(1)-O(1)	1.150(23)
Re-Cp(center)	1.988(15)	C(2)-O(2)	1.189(26)

## Bond Angles

Re-N(1)-N(2)	73.6(7)	N(2)-Re-C(1)	81.3(5)
Re-N(2)-N(1)	66.9(6)	N(1)-N(2)-C(9)	113.7(10)
Re-N(1)-C(3)	119.7(9)	N(2)-N(1)-C(3)	114.4(10)
Re-N(2)-C(9)	115.0(8)	Re-C(1)-O(1)	175.6(14)
N(1)-Re-N(2)	39.5(4)	Re-C(2)-O(2)	174.6(19)
N(1)-Re-C(2)	94.7(7)	C(1)-Re-C(2)	81.2(4)

Table 43. Atom Coordinates for  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11)

<u>Atom</u>	<u>x</u>	<u>y</u>	<u>z</u>
<u>Molecule 1</u>			
Re(1)	0.07979(4)	0.16622(3)	0.09831(2)
C(11)	0.1422(9)	0.0563(7)	0.0610(5)
C(12)	0.0553(10)	0.0788(8)	0.1603(5)
O(11)	0.1774(8)	-0.0137(6)	0.0395(4)
O(12)	0.0370(8)	0.0255(6)	0.1965(4)
C(13)	0.2242(9)	0.2255(7)	0.1555(5)
C(14)	0.2693(8)	0.2183(7)	0.1059(4)
C(15)	0.3679(8)	0.2329(7)	0.0660(4)
C(16)	0.4779(10)	0.2744(8)	0.0863(5)
C(17)	0.5750(9)	0.2861(18)	0.0495(6)
C(18)	0.5658(10)	0.2562(8)	-0.0069(6)
C(19)	0.4602(10)	0.2156(9)	-0.0280(5)
C(110)	0.3638(10)	0.2053(7)	0.0092(4)
C(111)	0.2309(10)	0.2519(7)	0.2163(5)
C(112)	0.1275(12)	0.2616(9)	0.2487(5)
C(113)	0.1369(15)	0.2938(9)	0.3050(5)
C(114)	0.2509(14)	0.3126(9)	0.3315(6)
C(115)	0.3541(14)	0.3003(9)	0.2977(6)
C(116)	0.3460(12)	0.2708(8)	0.2417(5)
C(117)	0.0379(10)	0.3047(9)	0.0482(6)
C(118)	-0.0177(10)	0.2293(9)	0.0183(5)
C(119)	-0.1063(9)	0.1897(7)	0.0533(5)
C(120)	-0.1075(10)	0.2397(8)	0.1042(5)
C(121)	-0.0206(12)	0.3105(8)	0.1030(6)
<u>Molecule 2</u>			
Re(2)	0.27088(4)	0.87942(3)	0.66720(2)
C(21)	0.3365(9)	0.7683(7)	0.6302(4)
C(22)	0.2314(9)	0.7893(8)	0.7256(4)
O(21)	0.3708(7)	0.7013(5)	0.6075(3)

Table 43 (Cont'd)

<u>Atom</u>	<u>x</u>	<u>y</u>	<u>z</u>
O(22)	0.2087(8)	0.7331(6)	0.7601(4)
C(23)	0.4086(8)	0.9317(7)	0.7305(4)
C(24)	0.4616(8)	0.9261(7)	0.6840(4)
C(25)	0.5705(8)	0.9435(7)	0.6495(4)
C(26)	0.6709(9)	0.9905(8)	0.6736(5)
C(27)	0.7723(10)	1.0108(9)	0.6407(6)
C(28)	0.7751(10)	0.9848(9)	0.5845(5)
C(29)	0.6765(10)	0.9361(8)	0.5593(5)
C(210)	0.5733(9)	0.9166(8)	0.5917(4)
C(211)	0.4077(9)	0.9529(7)	0.7916(4)
C(212)	0.2979(9)	0.9704(8)	0.8215(4)
C(213)	0.2971(11)	0.9903(8)	0.8801(4)
C(214)	0.4072(12)	0.9927(8)	0.9099(5)
C(215)	0.5170(11)	0.9792(8)	0.8825(5)
C(216)	0.5162(10)	0.9584(8)	0.8245(5)
C(217)	0.2413(10)	1.0081(8)	0.6069(5)
C(218)	0.1720(10)	0.9366(8)	0.5859(4)
C(219)	0.0846(9)	0.9130(8)	0.6249(5)
C(220)	0.0978(11)	0.9747(10)	0.6737(5)
C(221)	0.1989(11)	1.0358(8)	0.6610(5)

Table 44. Atom Coordinates ( $\times 10^4$ ) for  $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>	<u>Biso</u>
Re	429.8(6)	2500	2932.8(5)	2.63(6)
N(1)	1616(1)	1306(1)	1927(1)	3.2(6)
N(2)	2834(1)	2157(1)	2615(1)	2.9(6)
C(1)	1826(2)	3245(1)	4740(2)	4.2(8)
O(1)	2698(2)	3613(1)	5805(1)	8.3(8)
C(2)	71(2)	1441(2)	4237(2)	5.4(10)
O(2)	-300(2)	740(1)	5032(2)	8.3(9)
C(3)	1936(2)	143(1)	2484(2)	3.0(3)
C(4)	614(2)	-637(2)	2225(2)	4.4(3)
C(5)	900(2)	-1826(2)	2705(2)	5.6(4)
C(6)	2445(2)	-2225(2)	3466(2)	5.7(4)
C(7)	3731(2)	-1404(2)	3773(2)	5.8(4)
C(8)	3508(2)	-254(2)	3269(2)	4.0(3)
C(9)	3310(2)	2941(1)	1622(1)	3.4(3)
C(10)	3287(1)	2590(1)	215(1)	3.8(2)
C(11)	3907(2)	3397(2)	-652(2)	4.2(3)
C(12)	4559(2)	4500(2)	-102(2)	5.0(4)
C(13)	4611(2)	4830(2)	1308(2)	5.1(4)
C(14)	3974(2)	4060(2)	2177(2)	4.7(4)
C(15)	8313(1)	2736(2)	779(1)	3.5(3)
C(16)	7635(2)	2791(2)	1917(1)	4.3(3)
C(17)	8249(2)	3826(2)	2716(2)	4.7(4)
C(18)	9466(2)	4375(2)	2013(2)	4.5(3)
C(19)	9374(2)	3658(1)	822(2)	3.8(3)

Table 45. Anisotropic Thermal Parameters ( $\times 100$ ) for  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11)

<u>Atom</u>	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
<u>Molecule 1</u>						
Re(1)	4.91(2)	5.09(3)	6.17(3)	0.01(2)	-0.027(2)	-0.33(2)
C(11)	5.3(7)	6.1(8)	9.7(9)	-0.8(6)	-1.0(6)	-0.1(7)
C(12)	6.6(8)	6.4(8)	11.0(10)	0.7(6)	1.4(7)	-1.7(7)
O(11)	10.6(7)	6.5(6)	14.0(8)	0.5(5)	1.3(6)	-4.2(5)
O(12)	11.5(7)	9.2(7)	11.2(7)	1.2(5)	3.2(6)	3.0(5)
C(13)	5.4(6)	4.4(7)	8.4(8)	0.5(5)	-0.7(6)	-1.3(6)
C(14)	4.6(6)	5.6(7)	7.5(8)	0.2(5)	-0.5(5)	-0.7(6)
C(15)	5.1(6)	4.8(7)	6.6(7)	0.1(5)	-0.6(5)	-0.7(5)
C(16)	6.3(7)	8.9(9)	8.3(9)	0.2(7)	-0.3(6)	-3.9(7)
C(17)	5.4(7)	7.2(8)	13.4(11)	-1.3(6)	0.1(7)	-3.6(8)
C(18)	6.5(8)	7.7(9)	12.3(11)	-2.4(7)	3.1(7)	-0.2(8)
C(19)	8.3(9)	8.9(9)	7.5(8)	-1.6(7)	2.1(7)	-0.5(7)
C(110)	7.6(8)	5.4(7)	6.4(7)	-1.0(6)	-2.1(6)	0.0(7)
C(111)	8.8(8)	3.9(7)	8.7(8)	0.4(6)	-1.4(7)	-1.6(6)
C(112)	11.2(10)	8.4(9)	6.7(8)	-0.7(8)	-0.3(7)	-1.4(7)
C(113)	18.7(15)	8.3(9)	6.1(8)	-1.7(10)	2.0(9)	-1.5(7)
C(114)	27.9(21)	6.4(9)	5.9(9)	-1.8(11)	-3.4(11)	-1.5(7)
C(115)	15.9(14)	7.1(9)	10.6(11)	0.4(9)	-6.3(10)	-0.5(8)
C(116)	11.6(10)	6.9(9)	7.6(9)	0.4(7)	-2.2(8)	-1.8(7)
C(117)	5.3(7)	8.2(9)	15.9(13)	-1.3(6)	-3.8(8)	6.2(9)
C(118)	7.0(8)	10.7(10)	7.4(9)	2.1(7)	-0.1(6)	2.6(7)
C(119)	5.5(6)	6.4(8)	7.4(8)	0.1(5)	-1.5(6)	-1.2(6)
C(120)	6.9(8)	7.8(9)	8.1(8)	1.7(6)	-1.6(6)	-0.3(7)
C(121)	10.5(10)	5.3(8)	13.7(12)	3.4(7)	-6.2(9)	-2.6(8)
<u>Molecule 2</u>						
Re(2)	4.96(2)	5.15(3)	5.36(2)	-0.36(2)	-0.31(2)	0.41(2)
C(21)	7.2(7)	6.6(8)	5.1(7)	-1.0(6)	-0.1(5)	1.1(6)
C(22)	6.4(7)	6.7(7)	5.9(7)	0.3(6)	0.2(6)	0.2(6)

Table 45 (Cont'd)

<u>Atom</u>	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
O(21)	10.3(6)	6.9(5)	7.8(6)	0.3(5)	0.8(5)	-1.2(4)
O(22)	13.0(7)	7.9(6)	9.7(6)	-0.7(5)	2.4(6)	2.7(5)
C(23)	5.5(6)	5.4(7)	5.2(6)	0.0(5)	-0.9(5)	-0.3(5)
C(24)	5.8(6)	4.2(6)	5.2(6)	-0.4(5)	-0.6(5)	-0.3(5)
C(25)	5.1(6)	4.7(6)	5.5(6)	0.9(5)	0.0(5)	1.1(5)
C(26)	6.7(7)	8.0(8)	4.2(6)	-0.2(6)	1.3(5)	0.1(6)
C(27)	8.6(9)	8.1(9)	7.5(8)	-0.6(7)	3.3(7)	-0.4(7)
C(28)	6.0(7)	7.9(9)	11.0(10)	0.1(6)	2.0(7)	1.0(7)
C(29)	4.7(7)	8.8(9)	12.0(10)	0.1(6)	0.3(7)	0.2(8)
C(210)	5.0(6)	7.2(8)	8.5(8)	0.1(6)	-0.9(6)	-0.1(7)
C(211)	5.5(6)	5.2(7)	6.4(7)	0.7(5)	0.0(5)	0.7(5)
C(212)	6.9(7)	8.0(9)	6.8(8)	1.7(6)	-0.6(6)	0.8(6)
C(213)	11.0(10)	7.2(9)	7.2(8)	0.6(7)	-2.9(7)	0.0(7)
C(214)	15.1(12)	6.2(8)	6.1(8)	-1.5(8)	-2.6(8)	-0.4(6)
C(215)	11.1(10)	7.9(9)	5.4(7)	-1.1(7)	3.0(7)	-1.0(6)
C(216)	6.4(7)	7.2(8)	6.7(7)	-0.3(6)	0.2(6)	-0.1(6)
C(217)	7.0(8)	6.9(8)	11.1(10)	-0.5(6)	-1.6(7)	2.7(7)
C(218)	8.6(8)	7.5(8)	6.1(7)	-0.4(7)	-1.7(6)	1.1(6)
C(219)	4.3(6)	7.5(8)	11.0(9)	0.1(6)	-2.0(6)	2.9(7)
C(220)	8.0(8)	12.0(11)	7.1(8)	4.2(8)	-0.7(7)	-0.1(8)
C(221)	9.8(9)	5.2(8)	11.5(10)	2.8(7)	-5.2(8)	-2.3(7)

Table 46. Anisotropic Thermal Parameters ( $\times 100$ ) for  
 $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13)

<u>Atom</u>	<u><math>U_{11}</math></u>	<u><math>U_{22}</math></u>	<u><math>U_{33}</math></u>	<u><math>U_{12}</math></u>	<u><math>U_{13}</math></u>	<u><math>U_{23}</math></u>
Re	3.63(2)	3.16(2)	3.16(2)	-0.55(1)	0.86(2)	-0.43(1)
N(1)	4.9(7)	3.9(8)	4.0(7)	0.5(6)	1.9(6)	-0.5(6)
N(2)	3.9(6)	3.1(10)	3.7(6)	0.2(5)	0.5(5)	-0.3(5)
C(1)	6.3(10)	4.3(10)	5.0(9)	2.3(9)	0.8(8)	0.0(8)
O(1)	10.2(10)	13.2(14)	5.9(8)	2.8(10)	-1.8(7)	-2.3(9)
C(2)	9.0(10)	6.6(13)	6.0(11)	-1.3(11)	3.5(11)	-1.8(10)
O(2)	19.0(15)	6.5(9)	9.1(9)	1.3(10)	9.2(10)	2.7(8)

Table 47. Calculated Hydrogen Atom Coordinates ( $\times 10^2$ )for  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11)

<u>Atom</u>	<u>x</u>	<u>y</u>	<u>z</u>
H(16)	486	295	127
H(17)	652	316	64
H(18)	636	264	-32
H(19)	453	194	-68
H(110)	287	176	-6
H(112)	46	246	232
H(113)	62	304	327
H(114)	258	333	372
H(115)	436	314	315
H(116)	421	263	219
H(117)	104	346	34
H(118)	2	208	-21
H(119)	-159	135	43
H(120)	-162	227	136
H(121)	-2	357	134
H(26)	670	1010	714
H(27)	843	1044	658
H(28)	847	1000	561
H(29)	679	916	519
H(210)	502	884	574
H(212)	219	968	800
H(213)	219	1002	900
H(214)	408	1004	952
H(215)	595	984	904
H(216)	595	947	806
H(217)	312	1037	587
H(218)	182	906	548
H(219)	23	862	620
H(220)	48	975	709
H(221)	232	1087	685



Table 48. Calculated Hydrogen Atom Coordinates ( $\times 10^4$ ) for  
 $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$  (13)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
H(4)	-507	-374	1703
H(5)	-41	-2385	2493
H(6)	2633	-3063	3781
H(7)	4837	-1649	4371
H(8)	4455	293	3462
H(10)	2854	1807	-165
H(11)	3878	3173	-1654
H(12)	4988	5044	-723
H(13)	5081	5603	1697
H(14)	3989	4291	3175
H(15)	8067	2112	28
H(16)	6857	2216	2146
H(17)	7946	4133	3573
H(18)	10163	5085	2316
H(19)	10000	3795	101

Table 49. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
CpRe(CO)<sub>2</sub>(C<sub>2</sub>Ph<sub>2</sub>) (11)

<u>Plane No.</u>	<u>Atoms Defining Plane<sup>b</sup></u>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D (Å)</u>	<u>χ<sup>2</sup></u>
<u>Molecule 1</u>						
1 <sup>c</sup>	C11(-4), C12(4), C13(-9), C14(8)	-0.7014	0.6010	-0.3831	-1.1215	1.85
2	Ph ring: C15(-3), C16(-2), C17(5), C18(-4), C19(-1), C110(4)	-0.3401	0.9025	-0.2642	1.1866	0.64
3	Ph ring: C111(9), C112(-16), C113(13), C114(-3), C115(-4), C116(1)	-0.0652	0.9507	-0.3031	1.6644	3.49
4 <sup>d</sup>	Cp ring: C117(1), C118(3), C119(-3), C120(2), C121(0)	-0.6724	0.6280	-0.3919	1.9777	0.18
5 <sup>e</sup>	C13(5), C14(-5), C15(2), C111(-2)	-0.3567	0.9050	-0.2318	1.1749	0.50
6	C13, C14, C15	-0.3722	0.8949	-0.2461	1.0601	0.00
7	C13, C14, C111	-0.3338	0.9141	-0.2304	1.2669	0.00
8	Re1, C13, C14	-0.3092	0.9251	-0.2203	1.3954	0.00
<u>Molecule 2</u>						
9 <sup>f</sup>	C21(-6), C22(6), C23(-12), C24(12)	-0.6430	0.6359	-0.4269	-1.5150	3.82

Table 49 (Cont'd)

Plane No.	Atoms Defining Plane <sup>b</sup>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D (Å)</u>	<u>χ<sup>2</sup></u>
10	Ph ring: C25(1), C26(4), C27(1), C28(7), C29(10), C210(6)	-0.4006	0.8760	-0.2685	5.1724	1.55
11	Ph ring: C211(-8), C212(7), C213(5), C214(-15), C215(13), C216(-2)	0.0716	0.9773	-0.1994	9.647	4.49
12 <sup>g</sup>	Cp ring: C217(-7), C218(7), C219(-3), C220(-1), C221(5)	-0.6175	0.6555	-0.4349	1.7220	1.04
13 <sup>h</sup>	C23(14), C24(-14), C25(5), C211(-5)	-0.3157	0.9237	-0.2169	7.0664	4.85
14	C23, C24, C25	-0.3573	0.8957	-0.2647	5.7421	0.00
15	C23, C24, C211	-0.2456	0.9472	-0.2060	7.8525	0.00
16	Re2, C23, C24	-0.2582	0.9424	-0.2127	7.6255	0.00

<sup>a</sup> Plane equation is  $Ax + By + Cz - D = 0$ , with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviations of the atoms ( $\text{Å} \times 10^3$ ) from the meanplane are given in parentheses. The e.s.d.'s in these values are  $0.01\text{Å}$ .

<sup>c</sup> Rel deviation  $1.077\text{Å}$ .

<sup>d</sup> Other deviations ( $\text{Å}$ ): Rel  $-1.956$ , C11  $-3.05$ , C12  $-3.08$ , C13  $-2.99$ , C14  $-2.95$ , C15  $-3.20$ , C111  $-3.33$ .

<sup>e</sup> Rel deviation  $0.109\text{Å}$ .

<sup>f</sup> Re2 deviation  $1.086\text{Å}$ .

<sup>g</sup> Other deviations ( $\text{Å}$ ): Re2  $-1.968$ , C21  $-3.07$ , C22  $-3.10$ , C23  $-3.03$ , C24  $-2.98$ , C25  $-3.22$ , C211  $-3.42$ .

<sup>h</sup> Re2 deviation  $0.118\text{Å}$ .

Table 49 (Cont'd)

## Dihedral Angles (Between Normals to Meanplanes)

<u>Plane No.</u>	<u>Plane No.</u>	<u>Angle (deg)</u>
<u>Molecule 1</u>		
1	4	2.3
1	5	28.0
2	3	16.2
2	5	2.1
2	6	2.2
3	5	17.5
3	7	16.1
5	8	3.0
6	7	2.6
<u>Molecule 2</u>		
9	12	1.9
9	13	28.0
10	11	28.2
10	13	6.3
10	14	2.7
11	13	22.6
11	15	18.3
13	16	3.5
14	15	7.8

Table 50. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
CpRe(CO)<sub>2</sub>(N<sub>2</sub>Ph<sub>2</sub>) (13)

<u>Plane No.</u>	<u>Atoms Defining Plane<sup>b</sup></u>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D</u>	<u>χ<sup>2</sup></u>
1 <sup>c</sup>	Cp ring: C15(-3), C16(16), C17(-22), C18(18), C19(-9)	0.6194	-0.5492	0.3719	2.875	2.0
2	Ph ring: C9(-7), C10(7), C11(-8), C12(-5), C13(11), C14(-2)	0.8463	-0.4099	0.0978	1.135	0.8
3	Ph ring: C3(-9), C4(18), C5(-9), C6(-17), C7(23), C8(-5)	-0.4467	0.2640	0.9440	1.562	2.3
4	N1(-108), N2(98), C1(-137), C2(178)	0.48986	-0.6174	0.4595	0.7054	153.8
5	Re, N1, N2	0.0125	-0.5587	0.7948	0.6552	0.0
6	Re, C1, C2	-0.8415	0.5338	0.3087	2.057	0.0

<sup>a</sup> Plane equation is  $Ax + By + Cz = D$ , with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviations of the atoms ( $\text{Å} \times 10^3$ ) from the meanplane are given in parentheses.

<sup>c</sup> Re deviation  $\text{Å}$ .

Table 50 (Cont'd)

## Dihedral Angles (Between Normals to Meanplanes)

<u>Plane No.</u>	<u>Plane No.</u>	<u>Angle (deg)</u>
1	2	19.9
1	3	86.7
1	4	9.0
1	5	38.8
1	6	140.1
2	3	101.3
2	4	28.7
2	5	58.5
2	6	154.5
3	4	81.7
3	5	56.2
3	6	54.0
4	5	30.5
4	6	133.7
5	6	103.9

## CHAPTER 4

X-RAY STRUCTURES OF cis-Cp\*Re(CO)<sub>2</sub>I<sub>2</sub> and trans-Cp\*Re(CO)<sub>2</sub>Br<sub>2</sub>4.1 Introduction

In view of the current activity in cyclopentadienyl and pentamethylcyclopentadienyl carbonyl chemistry, it is surprising that several members of the series of simple rhenium dicarbonyl dihalides CpRe(CO)<sub>2</sub>X<sub>2</sub> and Cp\*Re(CO)<sub>2</sub>X<sub>2</sub> (Cp = η<sup>5</sup>-C<sub>5</sub>H<sub>5</sub>, Cp\* = η<sup>5</sup>-C<sub>5</sub>Me<sub>5</sub>, X = Cl, Br, I) were unknown and reports of the others were fragmentary. CpRe(CO)<sub>2</sub>Br<sub>2</sub> and CpRe(CO)<sub>2</sub>I<sub>2</sub> (cis and trans) have been synthesized<sup>109,110,144,145,146</sup> but CpRe(CO)<sub>2</sub>Cl<sub>2</sub> was unknown. Cp\*Re(CO)<sub>2</sub>I<sub>2</sub> was briefly mentioned<sup>147</sup> but without details of the synthesis or stereochemistry; the dibromides and dichlorides were unknown.

It was hoped preferably to determine the X-ray structures of the cis and trans isomers of a single halide for better comparison, but unfortunately no single pair could be found which gave acceptable crystals for both isomers. However the structure determinations of cis-Cp\*Re(CO)<sub>2</sub>I<sub>2</sub> (16) and trans-Cp\*Re(CO)<sub>2</sub>Br<sub>2</sub> (17) were carried out to confirm the geometries deduced from the infrared intensities and to provide typical metrical data for this class of compounds, none of which has been structurally investigated before.

#### 4.2 X-ray Data Collection and Refinement

A needle-like orange-red crystal of 16 and a small red crystal of 17, both obtained from methylene chloride/hexanes, was supplied by A.H. Klahn and used for the diffraction studies. The major morphological axis was aligned with the glass fibre and preliminary oscillation photos (Cu radiation) showed the crystals to be of an acceptable quality. The crystal was transferred to an Enraf-Nonius diffractometer, and the cell dimensions and space group were determined. The crystallographic data and scan conditions are listed in Table 51. The data collection was monitored by measuring two standard reflections every 1 hour of X-ray exposure time and showed only random fluctuations. The data were then corrected for Lorentz, polarization and absorption errors ( $\psi$  scans) and only the reflections with  $I \geq 3.0\sigma(I)$  were regarded as observed.

The heavy atoms were located from a Patterson map and all other non-hydrogen atoms from a series of difference Fourier syntheses interspersed with least squares refinement. All hydrogen atoms were placed in geometrically calculated positions with  $d(C-H) = 0.98\text{\AA}$  and thermal parameters of  $U = 0.09\text{\AA}^2$  for 16 and  $U = 0.08\text{\AA}^2$  for 17, and included as fixed contributions in further refinement. The disposition of the hydrogen atoms of the methyl groups was set identical to that found experimentally by Dahl<sup>148</sup> for  $\text{Cp}^*\text{Co}(\text{CO})_2$ . The Re atom and halide ligands were refined anisotropically and all other atoms isotropically. For the final cycles the weighting scheme was of the form  $\omega = 1/(\sigma(\underline{F}_o)^2 + n(\underline{F}_o)^2)$ . The value of the parameter  $n$  ( $= 0.0005$  for 16 and  $0.0004$  for 17) was that for which the variation of the averaged  $\sum\{w|\underline{F}_o| - |\underline{F}_c|\}^2$  as a function of  $|\underline{F}_o|$  and  $(\sin\theta/\lambda)$  was kept to a minimum. Final residuals for 16 were  $R_f = 0.046$ ,  $R_{wf} = 0.058$  and for 17,  $R_f = 0.043$ ,  $R_{wf} = 0.054$ . A small extinction



correction was necessary for 17. The largest peaks in a final difference map for 16 were  $1.7(2)$  and  $1.2(2)e\text{\AA}^{-3}$  located near the C(1)O(1) and C(2)O(2) ligands respectively, and both  $2.62\text{\AA}$  from the Re atom. The location of the peaks was suggestive of a small fractional disorder of I(2) with the carbonyl group C(2)O(2) and I(1) with C(1)O(1). The two peaks were refined as iodine atoms with variable occupancy and a fixed isotropic temperature factor of  $0.05\text{\AA}^2$ . After several cycles, the peaks had refined to occupancies of 2.2% and 1.4%, respectively, with Re-I bond lengths of  $2.65\text{\AA}$ . The residuals were  $R_f = 0.044$ ,  $R_{wf} = 0.056$ . However, because of the small occupancy of the partial iodine atoms this model was not pursued. A final difference map for 17 revealed numerous peaks and troughs of  $\pm 0.8-0.9(2)e\text{\AA}^{-3}$  around the Re and Br atoms and a single large peak of  $1.6(2)e\text{\AA}^{-3}$  located  $1.5\text{\AA}$  from the Re and  $1.2\text{\AA}$  from C(1) having no apparent chemical significance.

Computer programs are from reference 65. A summary of the crystal data is in Table 51, bond lengths and angles in Table 52 for 16 and in Table 53 for 17, atom coordinates for 16 in Table 54 and for 17 in Table 55, anisotropic thermal parameters for 16 and 17 in Table 56, calculated hydrogen atom coordinates in Table 57 for 16 and in Table 58 for 17 and calculated meanplanes and dihedral angles in Table 59 for 16 and Table 60 for 17. Labelled SNOOPI diagrams are presented in Figures 31 and 32 for 16 and 17, respectively.

#### 4.3 Results and Discussion

Both compounds exist as discrete molecular species in the solid state with no unusually short inter- or intramolecular contacts. SNOOPI diagrams and the atom labelling schemes for 16 and 17 are presented in Figures 31 and

FIGURE 31

SNOOPI Diagram of cis-Cp\*Re(CO)<sub>2</sub>I<sub>2</sub>

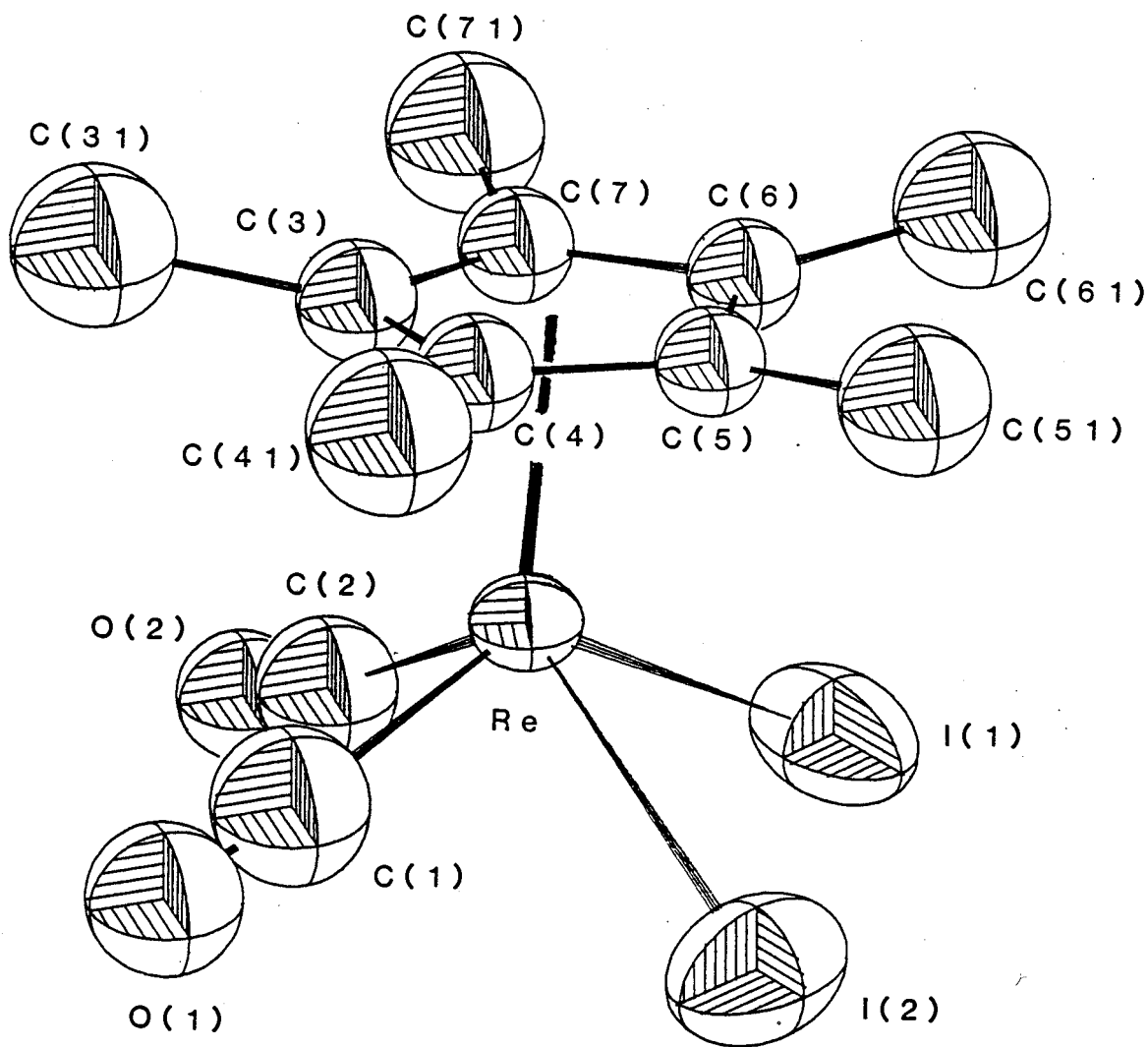


FIGURE 32

SNOOPI Diagram of trans-Cp\*Re(CO)<sub>2</sub>Br<sub>2</sub>

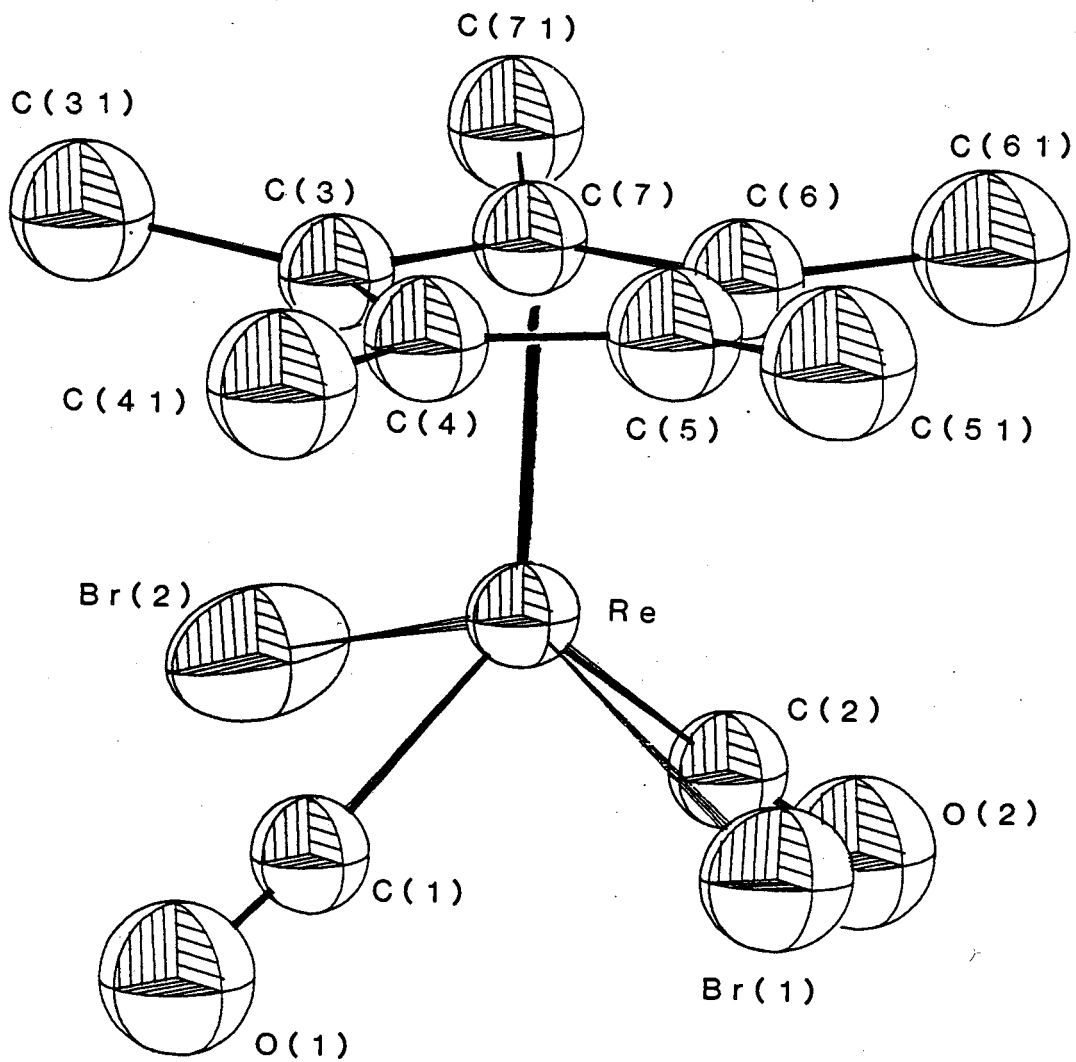


Table 51. Summary of Crystal Data and Data Collection Conditions.  
Collection Conditions for 16 and 17.

compound	<u>cis</u> -(C <sub>5</sub> Me <sub>5</sub> )Re(CO) <sub>2</sub> I <sub>2</sub>	<u>trans</u> -(C <sub>5</sub> Me <sub>5</sub> )Re(CO) <sub>2</sub> Br <sub>2</sub>
formula	C <sub>12</sub> H <sub>15</sub> I <sub>2</sub> O <sub>2</sub> Re	C <sub>12</sub> H <sub>15</sub> Br <sub>2</sub> O <sub>2</sub> Re
mol. wt.	631.3	537.3
space group	P2 <sub>1</sub> /n	P2 <sub>1</sub> /c
a, Å	7.201(3)	8.512(2)
b, Å	27.254(5)	12.610(4)
c, Å	8.682(2)	13.794(6)
β, deg	110.04(4)	102.40(3)
Z	4	4
cell vol., Å <sup>3</sup>	1600.2(2)	1446.1(3)
d <sub>calcd</sub> g cm <sup>-3</sup>	2.62	2.47
d <sub>meas</sub> g cm <sup>-3</sup>	2.64(1)	2.43(1)
cryst dimension, mm	0.05 × 0.08 × 0.25	0.09 × 0.12 × 0.15
radiation	Mo	Mo
μ, cm <sup>-1</sup>	115.08	139.82
transmission factors	0.565-0.996	0.679-0.988
scan speeds, deg min <sup>-1</sup>	1.0-4.0	0.5-4.0
scan width, deg	0.70 + dispersion	0.75 + dispersion
data limits, 2θ deg	3, 45	3, 46
measured reflections	2081	2008
observed reflections I > 3.0σ(I)	1510	1509
no. variables	84	85
R <sub>F</sub>	0.046	0.043
R <sub>wF</sub>	0.058	0.054
GOF	1.83	2.05
extinction	-	1.28 × 10 <sup>-7</sup>

Table 52. Bond Distances (Å) and Angles (deg) for cis-(C<sub>5</sub>Me<sub>5</sub>)Re(CO)<sub>2</sub>I<sub>2</sub>

Re-I(1)	2.773(2)	C(3)-C(4)	1.40(3)
Re-I(2)	2.761(2)	C(4)-C(5)	1.40(3)
Re-C(1)	1.92(3)	C(5)-C(6)	1.38(2)
Re-C(2)	2.03(3)	C(6)-C(7)	1.46(3)
Re-C(3)	2.20(2)	C(7)-C(3)	1.41(3)
Re-C(4)	2.26(2)	C(3)-C(31)	1.56(3)
Re-C(5)	2.37(2)	C(4)-C(41)	1.54(3)
Re-C(6)	2.41(2)	C(5)-C(51)	1.56(3)
Re-C(7)	2.26(2)	C(6)-C(61)	1.52(3)
C(1)-O(1)	1.03(3)	C(7)-C(71)	1.48(3)
C(2)-O(2)	0.85(3)	Re-Cp*(center)	1.951(9)
I(1)-Re-I(2)	83.00(7)	I(2)-Re-C(1)	75.3(8)
C(1)-Re-C(2)	78.5(10)	I(2)-Re-C(2)	122.4(6)
I(1)-Re-C(1)	129.5(7)	Re-C(1)-O(1)	177(2)
I(1)-Re-C(2)	75.4(7)	Re-C(2)-O(2)	173(2)

Table 53. Bond Distances (Å) and Angles (deg) for trans-(C<sub>5</sub>Me<sub>5</sub>)Re(CO)<sub>2</sub>Br<sub>2</sub>

Re-Br(1)	2.579(2)	C(3)-C(4)	1.46(2)
Re-Br(2)	2.579(2)	C(4)-C(5)	1.42(2)
Re-C(1)	1.93(2)	C(5)-C(6)	1.39(2)
Re-C(2)	1.92(2)	C(6)-C(7)	1.43(2)
Re-C(3)	2.25(2)	C(7)-C(3)	1.41(2)
Re-C(4)	2.31(2)	C(3)-C(31)	1.52(2)
Re-C(5)	2.36(2)	C(4)-C(41)	1.49(2)
Re-C(6)	2.31(2)	C(5)-C(51)	1.56(3)
Re-C(7)	2.33(2)	C(6)-C(61)	1.49(3)
C(1)-O(1)	1.10(2)	C(7)-C(71)	1.50(2)
C(2)-O(2)	1.15(2)	Re-Cp*(center)	1.945(7)
Br(1)-Re-Br(2)	138.97(7)	Br(2)-Re-C(2)	78.1(4)
Br(1)-Re-C(1)	76.9(4)	C(1)-Re-C(2)	104.3(6)
Br(1)-Re-C(2)	77.7(4)	Re-C(1)-O(1)	174(1)
Br(2)-Re-C(1)	77.6(4)	Re-C(2)-O(2)	176(1)



Table 54. Atom Coordinates ( $\times 10^4$ ) for Re and I Atoms and  $\times 10^3$  for Others) and Equivalent Isotropic Temperature Factors ( $\text{\AA}^2 \times 10$ ) for cis-(C<sub>5</sub>Me<sub>5</sub>)Re(CO)<sub>2</sub>I<sub>2</sub>

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>	<u>Biso</u>
Re	748(1)	1315.5(3)	2286.7(8)	30.7(4)*
I(1)	3104(2)	525.7(7)	3791(2)	70.2(11)*
I(2)	2206(3)	1113.6(8)	-200(2)	79.2(14)*
C(1)	151(4)	194(1)	169(3)	74(6)
O(1)	185(3)	227.9(7)	130(2)	70(4)
C(2)	250(4)	157.9(9)	450(3)	59(5)
O(2)	311(2)	168.2(5)	546(2)	49(3)
C(3)	-205(3)	160.6(7)	242(2)	43(4)
C(31)	-236(4)	211.0(10)	314(3)	78(6)
C(4)	-240(3)	150.5(7)	76(2)	41(4)
C(41)	-307(4)	185.4(10)	-72(3)	79(7)
C(5)	-232(3)	99.7(7)	57(2)	34(3)
C(51)	-286(4)	70.2(9)	-107(3)	66(6)
C(6)	-196(3)	76.9(7)	207(2)	36(4)
C(61)	-196(4)	22.3(10)	242(3)	70(6)
C(7)	-170(3)	115.6(7)	328(2)	39(4)
C(71)	-152(4)	108.5(10)	502(3)	75(6)

Table 55. Atom Coordinates ( $\times 10^4$  for Re and Br Atoms and  $\times 10^3$  for Others) and Equivalent Isotropic Temperature Factors ( $\text{\AA}^2 \times 10$ ) for trans-(C<sub>5</sub>Me<sub>5</sub>)Re(CO)<sub>2</sub>Br<sub>2</sub>

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>	<u>Biso</u>
Re	7392.4(6)	7800.5(5)	7093.1(4)	31.7(3)*
Br(1)	10184(2)	7859(2)	6649(1)	57.6(11)*
Br(2)	4403(2)	7696(2)	6207(1)	68.3(116)*
C(1)	735(2)	659(1)	624(1)	38(3)
C(2)	719(2)	899(1)	621(1)	39(3)
O(1)	734(1)	584(1)	582(1)	65(3)
O(2)	708(2)	974(1)	573(1)	65(3)
C(3)	650(2)	724(1)	842(1)	46(3)
C(4)	819(2)	695(1)	860(1)	42(3)
C(5)	906(2)	792(1)	869(1)	50(3)
C(6)	801(2)	878(1)	855(1)	39(3)
C(7)	742(2)	836(1)	839(1)	36(3)
C(31)	514(2)	647(1)	846(1)	55(4)
C(41)	884(2)	585(2)	873(1)	62(4)
C(51)	1093(2)	795(2)	902(1)	68(5)
C(61)	850(2)	991(2)	863(1)	67(4)
C(71)	494(2)	899(1)	841(1)	52(3)

Table 56. Anisotropic Thermal Parameters ( $\text{\AA}^2 \times 10^3$ ) forcis-Cp\*Re(CO)<sub>2</sub>I<sub>2</sub> (16)

<u>Atom</u>	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
Re	33.0(4)	47.3(5)	36.3(4)	-3.6(4)	9.1(3)	2.3(4)
I(1)	59.2(9)	78.7(13)	129(2)	16.8(9)	14.9(9)	27.9(10)
I(2)	76.3(11)	140(2)	84.3(11)	-14.6(11)	48.2(9)	-28.2(11)

trans-Cp\*Re(CO)<sub>2</sub>Br<sub>2</sub> (17)

<u>Atom</u>	<u>U<sub>11</sub></u>	<u>U<sub>22</sub></u>	<u>U<sub>33</sub></u>	<u>U<sub>12</sub></u>	<u>U<sub>13</sub></u>	<u>U<sub>23</sub></u>
Re	30.9(4)	51.2(4)	38.5(4)	-0.2(3)	8.1(2)	-0.4(3)
Br(1)	51.7(10)	93.3(14)	73.7(11)	-2.1(9)	24.8(8)	-2.6(10)
Br(2)	45.5(10)	152(2)	62.7(11)	-7.3(11)	1.7(8)	3.6(11)

Table 57. Calculated Hydrogen Atom Coordinates ( $\times 10^3$ ) for  
cis-(C<sub>5</sub>Me<sub>5</sub>)Re(CO)<sub>2</sub>I<sub>2</sub> (16)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
H(31)	-376	213	302
H(32)	-136	208	423
H(33)	-185	233	247
H(41)	-449	184	-120
H(42)	-238	215	- 16
H(43)	-224	171	-131
H(51)	-286	70	-107
H(52)	-187	44	- 67
H(53)	-231	94	-168
H(61)	-332	12	218
H(62)	-126	11	167
H(63)	-101	22	352
H(71)	-281	101	506
H(72)	- 49	83	532
H(73)	- 87	140	547

Table 58. Calculated Hydrogen Atom Coordinates ( $\times 10^3$ ) for  
trans-(C<sub>5</sub>Me<sub>5</sub>)Re(CO)<sub>2</sub>Br<sub>2</sub> (17)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
H(31)	507	639	916
H(32)	550	587	811
H(33)	426	683	800
H(41)	884	578	944
H(42)	997	579	862
H(43)	820	532	831
H(51)	1119	798	974
H(52)	1112	851	860
H(53)	1112	726	871
H(61)	867	1011	934
H(62)	753	1021	818
H(63)	937	990	829
H(71)	487	908	911
H(72)	414	852	799
H(73)	513	959	802

Table 59. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
cis-(C<sub>5</sub>Me<sub>5</sub>)Re(CO)<sub>2</sub>I<sub>2</sub> (16)

<u>Plane No.</u>	<u>Atoms Defining Plane<sup>b</sup></u>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D</u>	<u>χ<sup>2</sup></u>
1	Re, I1, I2	0.6629	0.6219	0.1639	2.912	0
2	Re, C1, C2	-0.898	0.419	0.433	1.88	0
3	I1(0.0), I2(0.0), C1(-15), C2(12)	0.945	0.203	-0.083	2.13	32.1
4 <sup>c</sup>	C3(-1.5), C4(0.1) C5(1.0), C6(-1.7), C7(1.9)	0.986	0.055	-0.194	-1.61	1.71

<sup>a</sup> Plane equation is  $Ax + By + Cz = D$ , with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviation of the atoms ( $\text{\AA} \times 10^2$ ) from the meanplane are given in parentheses

<sup>c</sup> Deviations: Re, 1.951(9); C31, 0.28(4); C41, 0.19(4); C51, 0.14(3); C61, 0.16(3); C71, 0.15(4) \text{\AA}.

Table 59 (Cont'd)

## Dihedral Angles (Between Normals to Meanplanes)

<u>Plane No.</u>	<u>Plane No.</u>	<u>Angle (Deg)</u>
1	2	106.2(7)
1	3	30.7(4)
1	4	41.2(6)
2	3	136.8(8)
2	4	147.4(9)
3	4	10.6(7)

Table 60. Calculated Meanplanes<sup>a</sup> and Dihedral Angles for  
trans-(C<sub>5</sub>Me<sub>5</sub>)Re(CO)<sub>2</sub>Br<sub>2</sub> (17)

<u>Plane No.</u>	<u>Atoms Defining Plane<sup>b</sup></u>	<u>A</u>	<u>B</u>	<u>C</u>	<u>D</u>	<u>χ<sup>2</sup></u>
1	Re, Br1, Br2	-0.0383	0.9986	-0.0278	9.310	0
2	Re, C1, C2	0.9904	0.0439	-0.0848	5.835	0
3	Br1(0.4), Br2(0.6), C1(-28), C2(-25)	-0.1247	0.0200	0.9956	8.244	380
4 <sup>c</sup>	C3(0.6), C4(-1.3), C5(1.5), C6(-0.7), C7(0.1)	-0.1767	0.0240	0.9990	10.842	1.14

<sup>a</sup> Plane equation is  $Ax + By + Cz = D$ , with  $x, y, z$  in orthorhombic coordinates.

<sup>b</sup> Deviation of the atoms ( $\text{Å} \times 10^2$ ) from the meanplane are given in parentheses.

<sup>c</sup> Deviations: Re, 1.944(7); C31, 0.24(3); C41, 0.04(3); C51, 0.19(3); C61, 0.08(3); C71, 0.28(3) Å.



Table 60 (Cont'd)

## Dihedral Angles (Between Normals to Meanplanes)

<u>Plane No.</u>	<u>Plane No.</u>	<u>Angle (Deg)</u>
1	2	89.9(4)
1	3	90.7(4)
1	4	90.3(5)
2	3	89.6(6)
2	4	92.6(7)
3	4	30.0(5)

32, respectively.

The overall molecular configuration consists of a rhenium atom bonded to a Cp\* ligand and to two CO ligands and two halide ligands in a cis arrangement for 16 and a trans arrangement for 17. This is in agreement with the  $\nu(\text{CO})$  infrared intensity patterns observed by Klahn<sup>149</sup>. Neither molecule has any crystallographically imposed symmetry in the lattice; however both molecules (including the hydrogen atoms) possess an approximate mirror plane that is easily seen in Figure 33 which shows the molecules side by side projected onto the plane of the Cp\* ring. In 16 this bisects the C<sub>5</sub>Me<sub>5</sub> ligand at C(3) and C(31) and bisects the angles I(1)-Re-I(2) and C(1)-Re-C(2). In 17 it bisects C(5) and C(51) and both of the Re-Br bonds.

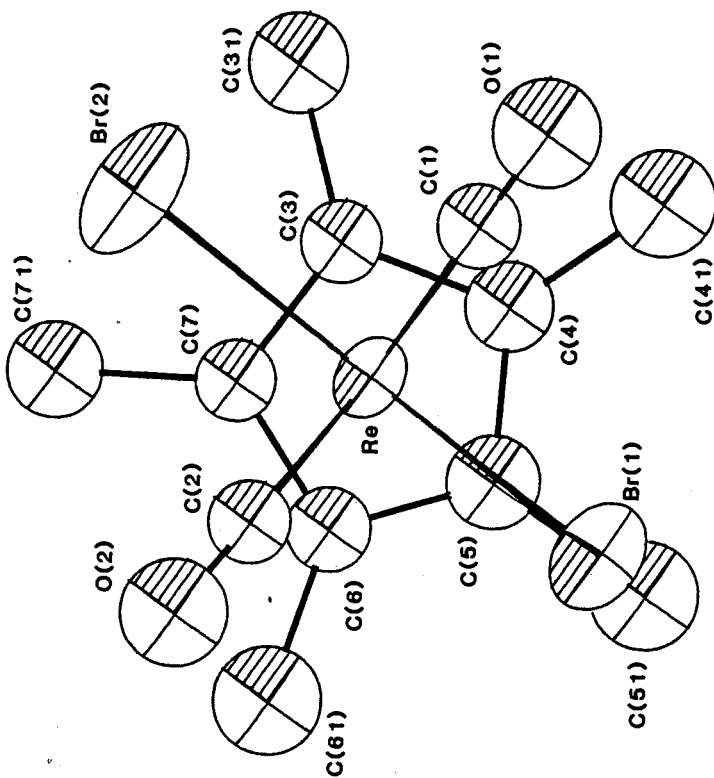
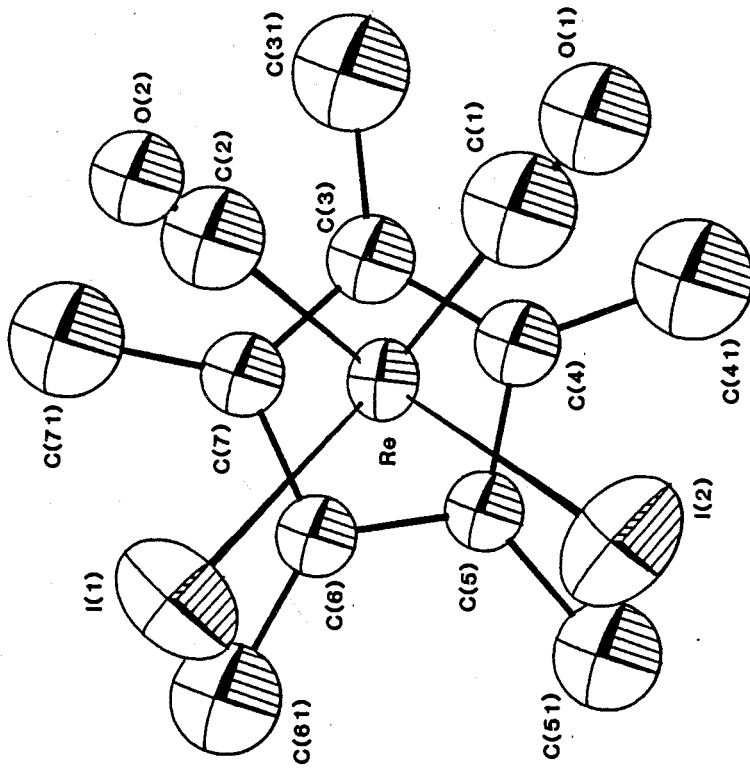
#### 4.3.1 Rhenium Coordination

The rhenium (III) atoms are seven coordinate if the Cp\* group is considered as a three coordinate monoanion and the overall geometry is that of a typical four-legged piano stool Cp\*ML<sub>4</sub> complex<sup>150</sup>. The Re-C(CO) bond lengths for both compounds are in the range 1.9-2.0Å that we have observed in CpRe(CO)<sub>2</sub>( $\eta^2$ -C<sub>2</sub>Ph<sub>2</sub>) and CpRe(CO)<sub>2</sub>( $\eta^2$ -N<sub>2</sub>Ph<sub>2</sub>) (Chapter 3). The Re-Br and Re-I bond lengths are unexceptional.

The Re-Cp\* (centroid) distances of 1.945(7)Å and 1.951(9)Å in 16 and 17 are identical and evidently insensitive to the differing cis- and trans-arrangements of the carbonyl and halide basal ligands in the two structures. Examination of the individual Re-C bond lengths to the Cp\* ligand reveals, however, differences in 16 and 17. The individual Re-C bonds in 17 vary from 2.25(2) to 2.36(2)Å in no evidently regular manner, but in 16 they vary systematically in a way that is chemically sensible even though only marginally insignificant at the level of precision. The bonds Re-C(5) and

FIGURE 33

Projection onto the  $Cp^*$  plane of (16) and (17)



Re-C(6) are long at 2.37(2) and 2.41(2)Å, Re-C(4) and Re-C(7) intermediate at 2.26(2)Å and Re-C(3) the shortest at 2.20(2)Å. With reference to the projection in Figure 33, it can be seen that the long Re-C bonds are 'opposite' or pseudotrans to the two CO groups, while the shortest is 'opposite' the iodide ligands. The asymmetry therefore follows a pattern that would be expected considering the  $\pi$ -acceptor properties possessed by CO ligands but not by iodide. The CO ligands withdraw  $\pi$ -electron density and diminish the  $\pi$ -contribution to the Re-C(Cp\*) bonds trans to them thereby resulting in lengthening. The asymmetry pattern also can be interpreted as a tilt of the Cp\* ring away from the bulky iodide ligands. Hoffmann<sup>150</sup> has suggested that a degree of tilting of the Cp ring observable in some CpML<sub>n</sub> compounds is accountable by interactions involving  $\delta$  orbitals of the metal and Cp ligand. In essence, it appears that (in projection) the eclipsing of a M-C(Cp) bond with a M-L bond results in a slightly long M-C(Cp) bond length, whereas when a M-C(Cp) bond is staggered between the M-L bonds, it is short. It is possible that such an effect also contributes to the pattern of Re-C(Cp\*) bond lengths observed here, but it must be pointed out that in cis-Cp\*Re(CO)<sub>2</sub>I<sub>2</sub> (16) none of the Re-C(Cp\*) is exactly eclipsed. Nevertheless, it is true (Fig. 30) that the longer bonds that occur for Re-C(5) and Re-C(6) are the ones most nearly eclipsed by the Re-I bonds, while the shortest (Re-C(3)) is the most staggered one. In trans-Cp\*Re(CO)<sub>2</sub>Br<sub>2</sub> (17) eclipsing does occur, but no significant statistical variation of the Re-C(Cp\*) bond lengths is observable.

The C-O bond lengths in 17 are in the expected range of 1.10-1.15(2)Å for metal carbonyl groups, but abnormally short values (0.85(3) and 1.03(3)Å) were obtained for 16. These unusually short distances are attributed to the 1-2% disordered iodine atom electron density contributions

in these regions, causing the apparent shortening. The C(1)-Re-C(2) angle in 16 is  $78.5(10)^\circ$  compared with the estimated value of  $80^\circ$  from the infrared  $\nu(\text{CO})$  intensities<sup>149</sup>. The I(2)-Re-C(1) and I(1)-Re-C(2) angles are both equal at  $75.4(7)^\circ$  however the trans angles I(1)-Re-C(1) and I(2)-Re-C(2) are not equal, being  $129.5(7)$  and  $122.4(6)^\circ$  respectively, although it would appear from the (pseudo)symmetry of the molecule (Fig. 33) that the two angles should be equal. The C(1)-Re-C(2) interbond angle in 17 is  $104.3(6)^\circ$  substantially less than that estimated  $115^\circ$ , from the  $\nu(\text{CO})$  intensities<sup>149</sup>. The Br(1)-Re-Br(2) angle is  $138.9(7)^\circ$  and the Br-Re-C angles are all very similar with the average of  $77.6(3)^\circ$ . These Br-Re-C angles are larger than the I-Re-C cis angles in 16 ( $75.3(7)^\circ$ ) and is attributed to the repulsion between the mutually cis iodine atoms contracting the I-Re-C cis angles.

#### 4.3.2 The Pentamethylcyclopentadienyl Ligand

The C-C bond lengths in the Cp\* ligands in 16 and 17 are similar and normal ranging from  $1.38(2)$ - $1.46(3)$ Å for the ring carbons and  $1.48(3)$ - $1.56(3)$ Å for the C(ring)-C(methyl) distances. The carbon atoms of the ring form well defined planes ( $\chi^2$  values = 1.71, 1.14) and there is no systematic distortion from this planarity as was observed for  $(\text{C}_5\text{Me}_5)\text{Co}(\text{CO})_2$ <sup>148</sup>. The methyl substituents are all displaced from these planes in a direction away from the Re atom with average deviations of  $0.18(2)$ Å in 16 and  $0.17(2)$ Å in 17.

A survey by Dahl<sup>148</sup> of various pentamethylcyclopentadienyl-metal complexes revealed this type of displacement to be general, with average deviation ranging from  $0.037$ Å to  $0.205$ Å. In 17 there are additional features to note, which are that the methyl carbon deviations conform to the

approximate molecular mirror plane symmetry, and the largest deviations are displayed by the methyl groups C(31), C(51) and C(71) that are closest to eclipsing the bromine positions in projection (Fig. 33). Thus, the related pair C(41) and C(61) show very small deviations (+0.04Å and +0.08Å) while those of the pair C(31) and C(71) are much larger (+0.24Å and +0.28Å) and that of C(51) is +0.19Å. In 16 the methyl carbon deviations also conform to the mirror symmetry, but C(41)-C(71) all have similar displacements of ca. 0.14-0.19Å while that of the unique methyl group C(31) is larger at 0.28Å.

## APPENDIX A

X-RAY STRUCTURE OF  $(\mu\text{-Cl})_2\text{Os}_3(\text{CO})_{10}$ A.1 Introduction

A number of X-ray crystal structures have been performed on triosmium carbonyl clusters of the general formula  $(\mu\text{-H})(\mu\text{-Y})\text{Os}_3(\text{CO})_{10}$  where  $Y = \text{H}, \text{Cl}, \text{Br}, \text{SEt}, \text{OMe}, \text{CHCH=NEt}_2$ , etc.<sup>151,152,153,154</sup> However, only in the case of the methoxide ligand<sup>154</sup> has the series  $(\mu\text{-H})_{0-2}(\mu\text{-Y})_{2-0}\text{Os}_3(\text{CO})_{10}$  been structurally characterized. In order to complete a second series of compounds, specifically that of  $(\mu\text{-H})_{0-2}(\mu\text{-Cl})_{2-0}\text{Os}_3(\text{CO})_{10}$ , we undertook to determine the X-ray structure of the missing member of this series;  $(\mu\text{-Cl})_2\text{Os}_3(\text{CO})_{10}$ . This would provide the final link in this series so valid metrical comparisons could be made.

A.2 X-ray Data Collection and Refinement

A yellow crystal, 0.25 x 0.30 x 0.22 mm, was obtained by recrystallization from hexanes from Dr R.K. Pomeroy and used for the diffraction studies. Weissenberg and precession photographs (Cu radiation) showed the crystal to be orthorhombic and systematic absences  $0kl: k = 2n+1$ ,  $h0l: l = 2n+1$ ,  $hk0: h = 2n+1$ , uniquely defined the space group as Pbca. Data were collected on a Picker FACS-I diffractometer (Mo radiation) with a scan rate of  $2^\circ \text{ min}^{-1}$  for  $3.0^\circ \leq 2\theta \leq 45^\circ$  and a symmetrical scan width of  $(1.2 + 0.692\tan\theta)^\circ$ . Peak profile analyses were performed on all reflections. Two standards were measured after every 70 reflections and showed a small overall gradual decline (4%) in average intensity; the data were scaled accordingly.

Absorption corrections were applied via an analytical correction



(transmission factors 0.039-0.070). 2838 reflections out of a total of 4619 had  $I \geq 2.3\sigma(I)$  and were regarded as observed. The positions of the osmium atoms were located by MULTAN and the chlorine, carbon and oxygen atoms were located from a series of difference Fourier synthesis and interspersed with least squares refinement. The metals and chlorine atoms were refined anisotropically and the carbonyl groups isotropically. A final difference Fourier map revealed little evidence of anisotropy around the carbonyl ligands. Computer programs are from reference 64. A summary of the crystal data and collection parameters are given in Table 61, bond distances and angles in Table 62, atom coordinates in Table 63 and anisotropic thermal parameters in Table 64.

### A.3 Results and Discussion

The crystal consists of discrete molecular units of  $(\mu\text{-Cl})_2\text{Os}_3(\text{CO})_{10}$ , with no abnormally short intermolecular contacts. The complex crystallizes in the space group Pbca with 16 molecules per unit cell giving two independent molecules in the asymmetric unit. The two molecules did not differ significantly except in the unbridged Os-Os distance, (Table 65) 3.233(1)Å versus 3.260(1)Å and in the Os-Cl-Os angle, 82.04(8)° versus 82.65(8)°. No disorder was apparent and the other bond lengths are in good agreement with each other. Only one of the molecules is presented in the ORTEP view (Fig. 34) along with the atom labelling scheme.

The molecule has approximate mirror symmetry and is based upon a triangular arrangement of osmium atoms. Os(1) is associated with four terminal carbonyl ligands and Os(2) and Os(3) are linked to three terminal carbonyl ligands and are mutually bridged by two chloride ligands.

The three Os atoms essentially define an isosceles triangle with the

FIGURE 34

ORTEP Diagram of  $(\mu\text{-Cl})_2\text{Os}_3(\text{CO})_{10}$

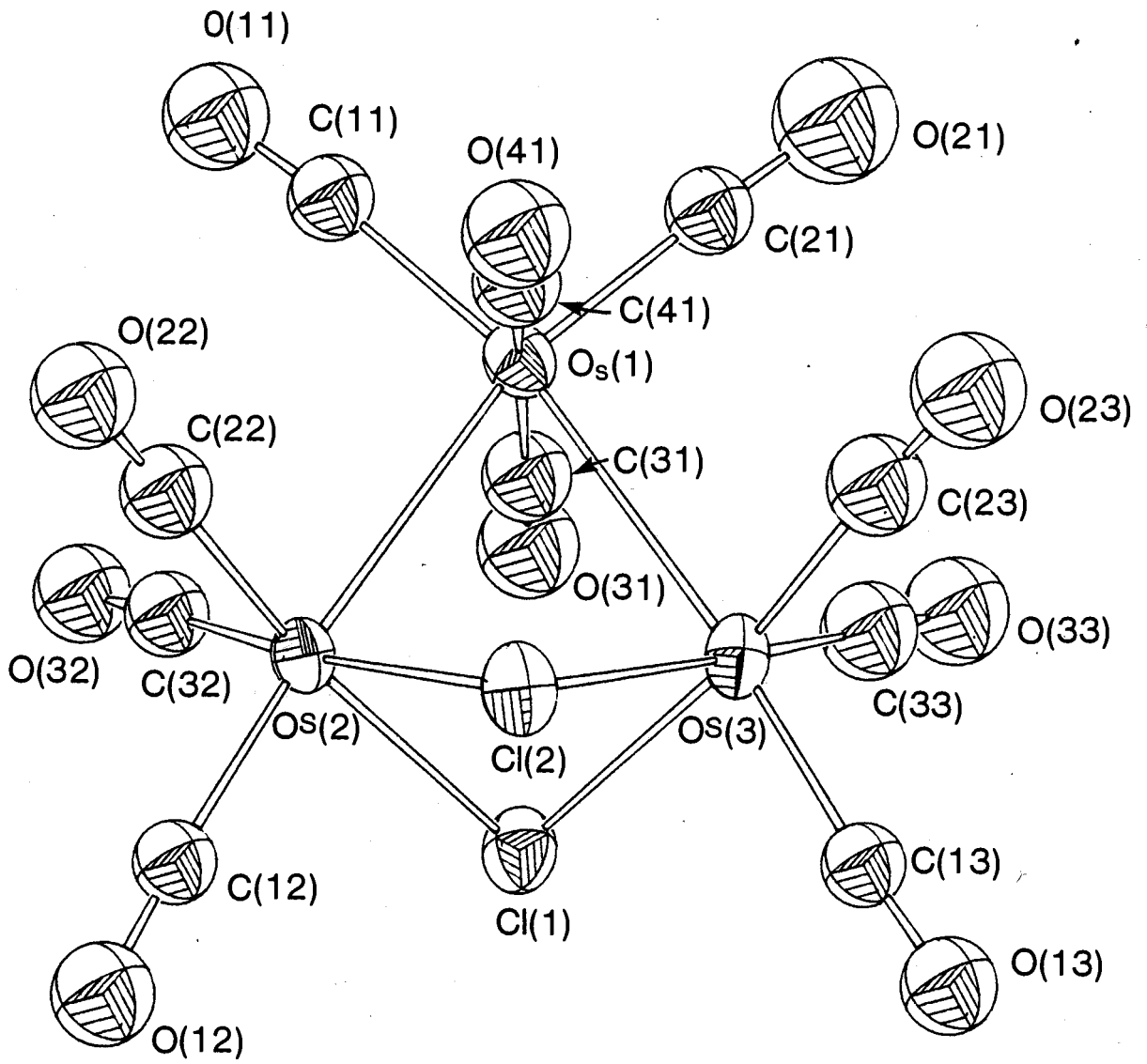


Table 61. Summary of Crystal Data and Collection Parameters for  
 $\text{Os}_3(\text{CO})_{10}(\text{Cl}_2)$  (18)

compound	$\text{C}_{10}\text{Cl}_2\text{O}_{10}\text{Os}_3$
mol. wt.	921.61
<u>a</u> , Å	25.580(8)
<u>b</u> , Å	22.832(6)
<u>c</u> , Å	12.036(3)
space group	Pbca
cell vol., Å <sup>3</sup>	7029.5
<u>Z</u>	16
$\rho$ (calcd) g cm <sup>-3</sup>	3.483
$\rho$ (obsd) g cm <sup>-3</sup>	3.25
$\mu$ (MoK $\alpha$ ) cm <sup>-1</sup>	217.0
transmission factors	0.039-0.070
total reflections, measured	4619
total reflections, $I \geq 2.3\sigma(I)$	2838
data limits (deg)	3, 45
final R	4.1
final $R_w$	4.9

Table 62. Bond Distances (Å) and Angles (deg) for Os<sub>3</sub>(CO)<sub>10</sub>Cl<sub>2</sub> (18)

## Bond Lengths

Os(1)-Os(2)	2.849(1)	Os(4)-C(44)	1.95(3)
Os(1)-Os(3)	2.855(2)	Os(5)-C(15)	1.85(2)
Os(2)-Os(3)	3.233(1)	Os(5)-C(25)	1.83(2)
Os(4)-Os(5)	2.846(1)	Os(5)-C(35)	1.91(2)
Os(4)-Os(6)	2.850(2)	Os(6)-C(16)	1.83(3)
Os(5)-Os(6)	3.260(1)	Os(6)-C(26)	1.82(2)
Os(2)-Cl(1)	2.461(6)	Os(6)-C(36)	1.82(3)
Os(2)-Cl(2)	2.462(6)	C(11)-O(11)	1.09(3)
Os(3)-Cl(1)	2.462(6)	C(21)-O(21)	1.21(3)
Os(3)-Cl(2)	2.467(6)	C(31)-O(31)	1.14(3)
Os(5)-Cl(3)	2.473(6)	C(41)-O(41)	1.12(3)
Os(6)-Cl(3)	2.479(6)	C(12)-O(12)	1.15(3)
Os(5)-Cl(4)	2.457(6)	C(22)-O(22)	1.17(3)
Os(6)-Cl(4)	2.463(6)	C(32)-O(32)	1.16(3)
Os(1)-C(11)	1.92(3)	C(13)-O(13)	1.15(3)
Os(1)-C(21)	1.85(3)	C(23)-O(23)	1.17(3)
Os(1)-C(31)	1.91(3)	C(33)-O(33)	1.18(3)
Os(1)-C(41)	1.95(3)	C(14)-O(14)	1.20(3)
Os(2)-C(12)	1.92(2)	C(24)-O(24)	1.14(3)
Os(2)-C(22)	1.87(3)	C(34)-O(34)	1.17(3)
Os(2)-C(32)	1.83(2)	C(44)-O(44)	1.14(3)
Os(3)-C(13)	1.91(2)	C(15)-O(15)	1.15(3)
Os(3)-C(23)	1.88(3)	C(25)-O(25)	1.15(3)

Table 62 (Cont'd)

Os(3)-C(33)	1.84(3)	C(35)-O(35)	1.15(3)
Os(4)-C(14)	1.83(3)	C(16)-O(16)	1.18(3)
Os(4)-C(24)	1.90(2)	C(26)-O(26)	1.17(3)
Os(4)-C(34)	1.90(2)	C(36)-O(36)	1.15(3)

## Bond Angles

Os(1)-Os(2)-Os(3)	55.56(3)	Os(2)-Os(3)-Cl(2)	48.9(1)
Os(1)-Os(3)-Os(2)	55.39(3)	Os(1)-Os(3)-Cl(1)	86.1(2)
Os(2)-Os(1)-Os(3)	69.06(4)	Os(1)-Os(3)-Cl(2)	86.2(2)
Os(4)-Os(5)-Os(6)	55.13(3)	Os(3)-Os(2)-Cl(1)	49.0(1)
Os(4)-Os(6)-Os(5)	55.04(3)	Os(3)-Os(2)-Cl(2)	49.1(1)
Os(5)-Os(4)-Os(6)	69.83(4)	Os(4)-Os(5)-Cl(3)	86.8(1)
Os(2)-Cl(1)-Os(3)	82.1(2)	Os(4)-Os(5)-Cl(4)	85.5(1)
Os(2)-Cl(2)-Os(3)	82.0(2)	Os(4)-Os(6)-Cl(3)	86.6(1)
Os(5)-Cl(3)-Os(6)	82.3(2)	Os(4)-Os(6)-Cl(4)	85.4(1)
Os(5)-Cl(4)-Os(6)	83.0(2)	Os(5)-Os(6)-Cl(3)	48.8(1)
Os(1)-Os(2)-Cl(1)	86.3(2)	Os(5)-Os(6)-Cl(4)	48.4(1)
Os(1)-Os(2)-Cl(2)	86.4(2)	Os(6)-Os(5)-Cl(3)	48.9(1)
Os(2)-Os(3)-Cl(1)	48.9(1)	Os(6)-Os(5)-Cl(4)	48.6(1)

Table 63. Atom Coordinates ( $\times 10^4$ ) for  $\text{Os}_3(\text{CO})_{10}\text{Cl}_2$ 

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>	<u>Biso</u>
Os(1)	2517.3(4)	1274.5(4)	6779.0(8)	28.7(4)*
Os(2)	1644.0(3)	782.3(4)	7913.9(8)	25.3(4)*
Os(3)	2628.3(4)	1398.4(4)	9127.9(8)	29.6(5)*
Os(4)	318.8(4)	1826.2(4)	3263.2(8)	24.3(4)*
Os(5)	-598.2(3)	1400.4(4)	2194.6(8)	26.4(4)*
Os(6)	480.0(3)	657.1(4)	2507.9(8)	25.5(4)*
Cl(1)	2284(2)	393(3)	9252(5)	34(3)*
Cl(2)	1689(2)	1642(3)	9145(5)	35(3)*
Cl(3)	-426(2)	457(3)	3136(5)	33(3)*
Cl(4)	4(2)	969(3)	830(5)	34(3)*
C(11)	2303(10)	1103(11)	5282(21)	39(5)
O(11)	2163(7)	971(9)	4461(16)	59(5)
C(21)	3138(10)	1643(11)	6421(21)	40(6)
O(21)	3563(9)	1847(10)	6214(19)	82(6)
C(31)	2858(10)	535(11)	6979(22)	42(6)
O(31)	3054(7)	91(8)	7095(15)	54(4)
C(41)	2126(10)	2009(11)	6827(21)	41(6)
O(41)	1910(7)	2436(9)	6831(16)	61(5)
C(12)	1080(9)	469(11)	8789(20)	37(5)
O(12)	768(7)	252(8)	9345(15)	56(4)
C(22)	1199(11)	1179(11)	6932(22)	45(6)
O(22)	908(8)	1443(9)	6375(16)	58(5)

Table 63 (Cont'd)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>	<u>Biso</u>
C(32)	1704(10)	164(11)	6954(20)	37(5)
O(32)	1746(7)	-241(8)	6375(15)	49(4)
C(13)	2650(9)	1435(11)	10709(20)	37(5)
O(13)	2692(7)	1474(8)	11658(15)	49(4)
C(23)	2855(11)	2144(12)	8836(24)	52(7)
O(23)	2973(8)	2636(9)	8761(18)	69(5)
C(33)	3299(11)	1119(12)	8982(23)	49(6)
O(33)	3733(8)	954(9)	8883(16)	59(5)
C(14)	74(10)	2563(12)	3565(21)	43(6)
O(14)	-91(8)	3043(9)	3763(17)	70(5)
C(24)	999(9)	1913(10)	3872(19)	33(5)
O(24)	1401(7)	1976(8)	4269(14)	51(4)
C(34)	46(9)	1488(10)	4579(18)	29(5)
O(34)	-123(6)	1250(7)	5364(13)	42(4)
C(44)	552(9)	2064(11)	1791(19)	35(5)
O(44)	687(6)	2218(7)	936(14)	44(4)
C(15)	-971(9)	1696(10)	3386(19)	33(5)
O(15)	-1195(7)	1874(7)	4142(14)	46(4)
C(25)	-636(9)	2120(10)	1522(19)	33(5)
O(25)	-662(7)	2577(9)	1121(16)	57(5)
C(35)	-1189(9)	1078(10)	1449(20)	34(5)
O(35)	-1561(7)	933(8)	986(15)	54(4)
C(16)	563(10)	-108(12)	1950(21)	44(6)



Table 63 (Cont'd)

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>	<u>Biso</u>
O(16)	617(8)	-594(9)	1644(17)	68(5)
C(26)	1109(9)	921(11)	2013(20)	37(5)
O(26)	1511(7)	1081(8)	1662(15)	53(4)
C(36)	754(10)	513(11)	3873(20)	37(5)
O(36)	946(7)	383(8)	4704(15)	52(4)

\* Anisotropic thermal parameters are found in Table 60.

Table 64. Anisotropic Thermal Parameters ( $\text{\AA}^2 \times 10^2$ ) for  
 $[(\mu\text{-Cl})_2\text{Os}_3(\text{CO})_{10}]$

<u>Atom</u>	<u><math>U_{11}</math></u>	<u><math>U_{22}</math></u>	<u><math>U_{33}</math></u>	<u><math>U_{12}</math></u>	<u><math>U_{13}</math></u>	<u><math>U_{23}</math></u>
Os(1)	3.04(5)	3.66(6)	4.12(6)	-0.07(5)	0.22(5)	0.14(5)
Os(2)	2.60(5)	3.26(5)	3.66(5)	-0.09(4)	-0.08(4)	-0.23(4)
Os(3)	3.15(6)	3.72(6)	4.30(6)	-0.07(5)	-0.77(5)	-0.44(5)
Os(4)	2.73(5)	3.11(5)	3.37(5)	-0.14(4)	-0.23(4)	-0.19(4)
Os(5)	2.57(5)	3.86(6)	3.47(6)	-0.16(4)	-0.35(4)	0.38(4)
Os(6)	3.07(5)	3.05(5)	3.53(5)	0.16(4)	-0.21(5)	0.12(4)
Cl(1)	3.7(3)	4.0(4)	5.1(4)	0.3(3)	-1.0(3)	0.2(3)
Cl(2)	3.8(4)	3.7(3)	5.5(4)	0.3(3)	-0.2(3)	-1.0(3)
Cl(3)	3.5(3)	4.4(4)	4.7(4)	-0.8(3)	-0.7(3)	1.0(3)
Cl(4)	4.3(4)	5.0(4)	3.7(3)	0.0(3)	0.0(3)	0.1(3)

angle at the unique osmium, Os(2)-Os(1)-Os(3), being  $69.04(3)^\circ$ , while the angles defined by Os(1)-Os(2)-Os(3) and Os(1)-Os(3)-Os(2) are  $55.56(3)$  and  $55.39(3)^\circ$  respectively (in the second molecule the corresponding osmium angles are  $69.83(4)$ ,  $55.13(3)$  and  $55.04(3)^\circ$ ). This compares<sup>152</sup> with the 'close to' equilateral triangle found for the three osmium atoms in  $(\mu\text{-H})(\mu\text{-Cl})\text{Os}_3(\text{CO})_{10}$  and the isosceles triangle noted<sup>151</sup> for  $(\mu\text{-H})_2\text{Os}_3(\text{CO})_{10}$ . In the former the angle at the unique osmium is  $60.32(2)^\circ$  while in the latter the same angle is  $56.93(4)^\circ$ . This is a clear demonstration of the Cl ligands pushing the two bridged osmium atoms further apart with a concomitant increase in the angle at the unique osmium.

The dibridged osmium-osmium distance in 18 is  $3.233(2)\text{\AA}$  (and  $3.260(2)\text{\AA}$  in the other molecule). As shown in Table 65, as the bridging hydride is replaced by bridging chloride ligands the corresponding Os-Os vector increases significantly. This increase in the Os-Os separation reflects the decrease in Os-Os bond order from 2 in  $(\mu\text{-H})_2\text{Os}_3(\text{CO})_{10}$  to 1 in  $(\mu\text{-H})(\mu\text{-Cl})\text{Os}_3(\text{CO})_{10}$  to 0 in  $(\mu\text{-Cl})_2\text{Os}_3(\text{CO})_{10}$ . A parallel increase, although not as large, was found for the  $(\mu\text{-H})_2$ ,  $(\mu\text{-H})(\mu\text{-OMe})$  and  $(\mu\text{-OMe})_2$  series<sup>154</sup>. However, no increase was noted for the non-bridged Os-Os distances in the methoxide series but a substantial increase is observed in this length,  $2.815(1)$  to  $2.833(1)$  to  $2.852(2)\text{\AA}$ , for the series  $(\mu\text{-H})_{2-0}(\mu\text{-Cl})_{0-2}\text{Os}_3(\text{CO})_{10}$ .

The bridging Cl atoms are bound to the triosmium core with an average bond length of  $2.463(3)\text{\AA}$  ( $2.468(3)\text{\AA}$  in molecule 2) and Os-Cl-Os angle of  $82.04(8)^\circ$ . The Os-Cl-Os angle is larger than found<sup>151</sup> with  $(\mu\text{-H})(\mu\text{-Cl})$  as bridging ligands at  $70.83(9)^\circ$  but the Os-Cl bond lengths are similar. The acute value of this angle is symptomatic of the presence of metal-metal bonding and the larger angle observed in 18 is consistent with no metal-

Table 65. Distances (Å) and Angles (deg) Within the Complexes $(\mu\text{-H})_2\text{Os}_3(\text{CO})_{10}$ ,  $(\mu\text{-H})(\mu\text{-Cl})\text{Os}_3(\text{CO})_{10}$  and $(\mu\text{-Cl})_2\text{Os}_3(\text{CO})_{10}$ 

	$(\mu\text{-H})_2$ <u>Compound<sup>a</sup></u>	$(\mu\text{-H})(\mu\text{-Cl})$ <u>Compound<sup>b</sup></u>	$(\mu\text{-Cl})_2$ <u>Compound<sup>c</sup></u>
Bridged Os-Os	2.682(1)	2.846(1)	3.233(1) 3.260(1)
Non-bridged Os-Os (av)	2.815(1)	2.833(1)	2.852(2) 2.848(3)
Os-Cl (av)	-	2.456(3)	2.463(3) 2.468(3)
Os-Cl-Os	-	70.83(9)	82.04(8) 82.65(8)

a Ref.

b Ref.

c This work; molecule 1, molecule 2.

metal bonding being present.

The carbonyl groups are in both axial and equatorial positions but because of the errors on the Os-C and C-O bond lengths, no clear distinction is possible in the present structural determination.

The present structural analysis completes the investigation of  $(\mu\text{-H})_{0-2}(\mu\text{-Cl})_{2-0}\text{Os}_3(\text{CO})_{10}$ , complexes providing the final link so that valid metrical comparisons could be made.

## APPENDIX B

COLLECTION AND TREATMENT OF THE X-RAY DIFFRACTION DATAB.1 Organization

The X-ray data of the compounds comprising this investigation were collected on two different diffractometers, a Picker FACS-I and an Enraf Nonius CAD-4. The data collected on the Picker diffractometer were solved and refined on a PDP 8/e computer and the data collected on the Enraf-Nonius diffractometer were solved and refined on a VAX 11-750 computer, both utilizing the NRC crystal structure package appropriate for each computer<sup>64,65</sup>. Since the diffractometers used, and thus the collection of data (and to a lesser extent the refinement of the structures) are not identical for every structure investigated, a short account specific to each determined structure is included in the relevant section of the thesis. Many of the specific data collection parameters are tabulated in the appropriate section.

B.2 Preliminary Investigation

Single crystals were mounted with epoxy cement either on the end of a thin glass fibre or in an appropriately sized Lindemann glass capillary with the major morphological axis, where possible, parallel to the fibre. Copper radiation was used to obtain oscillation photographs on all compounds and many (see individual sections) also had zero and upper level Weissenberg and oscillation photographs taken so as to obtain the approximate cell parameters and correct space group.

Crystals mounted on the Picker FACS-I were optically centered and strong reflections with  $\chi$  close to  $0^\circ$ , and separated in  $\phi$  by approximately  $90^\circ$ , were indexed using the photographic information. Twelve of the strongest reflections having  $2\theta > 20^\circ$  (and their Friedel equivalents) well spread over a large area of reciprocal space were used for the least-squares refinement of the cell and orientation parameters.

Crystals mounted on the Enraf-Nonius CAD-4 were optically centered and 25 reflections were found by the automatic SEARCH routine. Weak reflections were manually discarded and a wide area of reciprocal space was sampled. The reflections were centered using an automatic centering program, SETANG, and the cell and orientation parameters were then calculated.

The density of the crystals used for the X-ray work was determined by flotation of the crystal in the appropriate solvent mixture.

### B.3 Data Collection

Data on the Picker FACS-I were collected using graphite monochromated Mo-K $\alpha$  radiation ( $\lambda = 0.7093\text{\AA}$ ) at a takeoff angle of  $4^\circ$ , and a scintillation detector with pulse height analysis. A symmetrical  $\theta$ - $2\theta$  scanning mode was employed, the scan being from below  $K_{\alpha 1}$  in  $2\theta$  to above  $K_{\alpha 1}$  with a margin dependent on the mosaic spread of the particular crystal and an allowance for dispersion. Stationary-crystal stationary-counter background counts equal to 10 seconds were recorded at both limits of the scan. The appropriate  $2\theta$  limit for data collection, the scan width and measurement of standards are noted individually for each compound.

Data on the Enraf-Nonius CAD-4 were collected using graphite monochromated Mo-K $\alpha$  radiation at a takeoff angle of  $4^\circ$  (for  $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$  (11) a takeoff angle of  $3.5^\circ$  was used). An  $\omega$ - $2\theta$  scanning mode was employed

scanning symmetrically above and below the weighted mean for  $\text{MoK}\alpha$  with allowance made for dispersion corrections. The background was calculated by extending the scan by 25% on each side. A prescan was first conducted by scanning through the reflection at high speed ( $5^\circ\text{-}10^\circ/\text{min}$ ) and if the reflection was strong ( $\geq 25$  or  $30 \sigma(I)$ ) no further measurements were made. If the reflection was weak ( $\leq 1.5\sigma(I)$ ) it was rejected and when the reflections were intermediate they were remeasured at a slower scan speed according to the equation:

$$\text{NPI} = \text{NPIPRES} \times [\sigma(\text{pre})/\sigma]^2$$

where NPI = final scan speed parameter

NPIPRES = prescan speed parameter; scan speed =  $16.48/\text{NPIPRES}$  ( $^\circ/\text{min}$ )

$\sigma(\text{pre})$  = prescan acceptance parameter

$\sigma$  = the  $\sigma(I)/I$  required for final scan

A minimum speed  $\text{NPI}(\text{max})$  was determined by the maximum time set for a scan and if the calculated  $\text{NPI} > \text{NPI}(\text{max})$ , the final scan was done using  $\text{NPI}(\text{max})$ .

#### B.4 Structural Determinations

The data were corrected for Lorentz, polarization and absorption errors<sup>155,156</sup> (using either analytical corrections or empirical  $\psi$  scans). The heavy atoms were located by directed methods or by solving a three-dimensional Patterson map. Non-hydrogen atoms were located from a series of difference Fourier synthesis interspersed with least-squares refinement. Isotropic thermal parameters were used initially for several cycles and any anisotropic thermal parameters that were introduced had the form



$\exp\{-2\pi^2(U_{11}h^2a^{*2} + U_{22}k^2b^{*2} + U_{33}l^2c^{*2} + 2U_{12}hka^*b^* + 2U_{13}hla^*c^* + 2U_{23}klb^*c^*)\}$ . All hydrogen atom positions were calculated geometrically with C-H bond lengths and isotropic thermal parameters as noted for each individual structure. The weighting scheme used in the least squares cycles was of the form  $w = 1/(\sigma_{\underline{F}_O}^2 + n\underline{F}_O^2)$  and  $n$  was determined by minimizing the average  $\sum w(|\underline{F}_O| - |\underline{F}_C|)^2$  as a function of  $(\sin \theta)/\lambda$ ,  $|\underline{F}_O|$  and Miller indices. The final discrepancy indices, the residual  $R = (\sum ||\underline{F}_O| - |\underline{F}_C|| / \sum |\underline{F}_O|)$ , the weighted residual,  $R_w = [\sum w(|\underline{F}_O| - |\underline{F}_C|)^2 / \sum w|\underline{F}_O|^2]^{1/2}$  and the goodness of fit,  $GOF = [\sum w(|\underline{F}_O| - |\underline{F}_C|)^2 / (m-n)]^{1/2}$  are quoted for each crystal structure. Atomic scattering factors are taken from reference 157 and included corrections for anomalous dispersion.

APPENDIX C

LISTING OF THE OBSERVED AND  
CALCULATED STRUCTURE FACTORS

OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR  
 $\text{CpNiCoFe}(\text{CO})_5(\text{PPh}_3)(\text{C}_2\text{Ph}_2)$



COLUMNS ARE 10FO, 10FC, 10SIG

COLUMNS ARE 10FO, 10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG
5	340	367	23	8	748	795	21	11	491	505	25	10	289	256	43
7	673	693	18	9	452	432	28	14	398	390	36	8	313	294	36
8	270	288	34	10	760	756	23	12	769	927	28	11	482	439	29
9	614	654	21	14	764	763	29	1	119	1136	13	0	1142	1250	18
10	1381	1338	19	0	644	681	19	2	614	570	15	2	751	741	22
11	763	806	23	1	550	545	21	3	1280	1218	13	3	236	199	39
12	852	857	22	2	949	968	17	4	892	884	16	4	367	486	29
13	855	886	24	4	712	773	21	5	647	614	19	3	391	302	31
14	275	247	44	5	533	588	23	6	307	371	35	7	762	759	25
18	550	607	37	6	431	433	28	7	741	711	19	9	379	389	38
1	1485	1559	11	7	502	491	26	8	910	910	19	5	5	3	L
2	646	645	12	8	804	815	25	10	1121	1062	20	0	772	753	22
3	549	604	14	9	526	511	30	11	602	612	24	3	553	579	35
4	565	557	16	10	555	575	33	12	1187	1140	22	5	489	559	31
5	1803	1836	14	11	396	317	37	14	834	825	26	7	455	421	37
6	1322	1183	15	0	749	697	19	16	498	484	34	1	357	388	34
7	1326	1314	16	1	680	664	20	17	575	566	35	3	295	281	45
8	584	546	20	2	481	485	22	1	1180	1191	12	8	542	599	36
9	382	418	27	3	576	642	21	2	629	656	15	9	427	390	43
10	725	728	22	5	1025	1058	20	3	2374	2098	13	13	7	3	L
15	621	586	31	6	1015	967	21	3	623	1025	18	4	390	310	43
17	688	711	31	7	937	849	23	8	1132	1203	18	4	7	727	737
0	2534	2521	9	8	1140	1197	23	10	1115	1175	21	2	407	332	42
1	436	421	14	11	363	296	44	11	866	866	22	3	371	354	41
2	984	903	13	0	680	664	20	12	515	525	27	5	325	268	46
3	1020	1018	13	1	263	346	38	13	1465	1470	17	6	514	590	34
4	1184	1081	14	5	498	479	26	14	644	662	32	10	352	253	46
5	454	493	21	7	773	749	26	15	607	561	34	12	560	570	36
6	1150	1049	16	8	680	694	33	16	607	561	34	14	535	495	37
9	2000	1948	17	9	324	429	45	17	593	564	15	1	398	352	40
9	737	764	20	3	541	563	32	2	906	848	23	9	799	890	28
10	1332	1384	20	5	7	2	L	3	302	307	26	11	530	549	37
11	664	701	23	0	316	253	36	4	1136	1085	22	12	512	514	35
12	871	830	23	6	693	602	31	6	311	385	30	13	512	514	35
13	621	651	28	6	888	908	17	7	413	411	26	1	809	781	24
17	396	371	45	1	1631	1545	16	9	561	609	24	2	884	887	25
0	1407	1649	11	4	474	411	41	0	1355	1499	14	2	1020	1024	22
1	2087	2058	11	2	426	349	42	1	1127	1160	16	5	686	719	24
2	514	497	15	9	208	297	45	2	948	924	18	6	509	482	29
3	779	758	19	11	1557	1642	20	3	398	388	24	8	639	649	27
7	1529	1492	18	13	1065	1021	22	4	1109	1088	19	9	684	744	28
8	423	407	26	14	677	665	28	5	432	418	25	10	980	1008	25
9	1170	1072	21	15	492	490	35	6	735	732	20	12	1061	1131	26
13	488	493	33	18	448	452	43	7	935	858	21	13	464	452	39
15	751	703	30	1	1244	1118	14	8	1596	1724	20	14	439	467	42
1	354	401	21	3	193	193	47	10	1150	1182	22	15	542	600	36
2	348	358	24	4	112	399	23	15	472	377	39	1	982	936	20
4	1170	1155	16	5	1141	1082	16	1	925	966	17	2	624	642	24
5	495	446	21	6	641	690	32	3	485	474	24	3	1864	1748	19
0	1407	1649	11	7	860	793	26	4	672	715	24	8	725	786	23
1	2087	2058	11	8	700	627	28	5	672	715	24	9	725	786	23
2	514	497	15	9	1024	948	24	6	672	715	24	10	620	595	24
3	779	758	19	10	825	928	20	7	672	715	24	11	1023	1091	22
7	1529	1492	18	11	825	928	20	8	672	715	24	12	1023	1091	22

COLUMNS ARE 10FD,10FC, 10SIG

COLUMNS ARE 10FD,10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG
1	1	4	17	1	880	827	24	1	880	827	24	1	880	827	24	1	880	827	24	1	880	827	24
3	5	5	20	3	905	912	23	3	905	912	23	3	905	912	23	3	905	912	23	3	905	912	23
4	274	282	35	4	388	359	35	4	388	359	35	4	388	359	35	4	388	359	35	4	388	359	35
5	432	742	23	5	432	451	31	5	432	451	31	5	432	451	31	5	432	451	31	5	432	451	31
7	1229	1218	22	7	381	421	28	7	381	421	28	7	381	421	28	7	381	421	28	7	381	421	28
9	674	604	32	9	674	683	28	9	674	683	28	9	674	683	28	9	674	683	28	9	674	683	28
11	550	604	32	11	973	956	26	11	973	956	26	11	973	956	26	11	973	956	26	11	973	956	26
0	568	613	22	0	786	794	31	0	786	794	31	0	786	794	31	0	786	794	31	0	786	794	31
1	611	647	21	1	383	424	44	1	383	424	44	1	383	424	44	1	383	424	44	1	383	424	44
2	533	486	21	2	5	5	21	2	5	5	21	2	5	5	21	2	5	5	21	2	5	5	21
5	712	766	22	5	1026	1060	21	5	1026	1060	21	5	1026	1060	21	5	1026	1060	21	5	1026	1060	21
6	712	637	23	6	765	769	22	6	765	769	22	6	765	769	22	6	765	769	22	6	765	769	22
7	872	918	24	7	527	501	27	7	527	501	27	7	527	501	27	7	527	501	27	7	527	501	27
8	708	713	26	8	462	454	28	8	462	454	28	8	462	454	28	8	462	454	28	8	462	454	28
9	497	499	34	9	345	402	38	9	345	402	38	9	345	402	38	9	345	402	38	9	345	402	38
11	373	406	40	11	498	466	29	11	498	466	29	11	498	466	29	11	498	466	29	11	498	466	29
0	750	760	21	0	575	561	28	0	575	561	28	0	575	561	28	0	575	561	28	0	575	561	28
1	760	795	20	1	340	246	40	1	340	246	40	1	340	246	40	1	340	246	40	1	340	246	40
2	476	530	28	2	493	461	34	2	493	461	34	2	493	461	34	2	493	461	34	2	493	461	34
5	410	373	33	5	587	577	33	5	587	577	33	5	587	577	33	5	587	577	33	5	587	577	33
6	346	242	40	6	801	801	22	6	801	801	22	6	801	801	22	6	801	801	22	6	801	801	22
11	483	493	39	11	280	261	35	11	280	261	35	11	280	261	35	11	280	261	35	11	280	261	35

COLUMNS ARE 10FO,10FC, 10SIG

COLUMNS ARE 10FO,10FC, 10SIG

L	KFO	FC	SIG	L	NFO	FC	SIG	L	NFO	FC	SIG	L	KFO	FC	SIG	L	NFO	FC	SIG	
1	-6,	7,	L	10	284	197	50	12	618	609	36	3	3,	8,	L	2	603	580	29	
2	500	535	36	12	618	637	35	13	680	742	33	10	475	524	48	3	804	822	30	
3	277	271	55	1	1,	7,	L	4	8,	L	4	4,	8,	L	4	798	768	47	2,	
4	660	654	39	0	1103	1134	20	2	609	679	31	5	644	707	33	0	703	707	33	
5	-5,	7,	L	3	356	409	35	5	376	340	37	6	578	582	37	3	615	608	37	
6	493	508	31	4	309	357	41	6	502	487	33	5	514	506	43	2	624	645	35	
7	503	532	35	5	348	358	36	7	844	885	27	7	525	508	40	5	589	639	80	
8	281	133	52	6	980	994	25	8	611	600	31	0	739	857	L	6	352	376	57	
9	571	614	37	12	1042	1040	30	9	572	571	34	1	439	426	45	2,	519	489	45	
10	-4,	7,	L	10	429	453	45	2	429	453	45	2	501	479	37	3	388	330	53	
11	307	335	38	1	671	758	25	2	867	888	24	5	394	360	51	4	801	825	30	
12	346	399	38	4	331	282	35	3	598	540	28	6	330	299	67	9	530	353	42	
13	910	847	22	3	472	402	32	4	1282	1215	22	1	391	417	49	0	612	707	34	
14	321	284	36	4	1282	1215	22	5	951	940	23	1	729	809	35	2	516	566	39	
1	336	371	43	1	623	715	26	6	884	839	25	8	611	575	35	7	354	400	52	
2	476	457	34	2	623	717	28	7	874	861	26	4	441	487	40	3	645	616	33	
3	570	573	38	4	459	441	35	8	824	822	27	6	661	675	31	1	376	321	44	
4	408	421	46	8	646	647	32	9	502	452	36	7	427	416	41	1	383	386	43	
5	-3,	7,	L	10	893	786	31	10	507	543	37	10	535	527	38	5	621	555	36	
6	424	405	29	1	478	565	32	1	481	537	33	4	496	477	33	7	442	405	42	
7	964	882	22	2	666	704	28	2	426	437	32	5	750	726	30	8	310	254	54	
8	369	337	31	3	560	599	32	3	456	398	34	5	1030	964	26	5	478	420	39	
9	1504	1493	20	4	445	543	40	4	585	647	30	7	900	943	28	6	665	624	33	
10	635	624	23	5	542	497	40	5	1039	1056	25	8	537	540	34	8	659	597	34	
11	884	911	26	6	737	753	27	6	536	536	29	11	810	819	31	7	771	712	32	
12	531	486	30	7	874	861	26	7	659	648	28	13	823	870	32	7	771	712	32	
13	432	523	33	8	824	822	27	8	580	558	35	1	579	574	35	5	762	763	31	
14	710	732	29	9	502	452	36	9	580	558	35	3	691	597	27	7	383	391	45	
1	681	653	31	6	536	550	39	10	1039	1056	25	5	1064	1076	25	5	383	391	45	
2	754	713	30	6	536	550	39	0	536	536	29	11	810	819	31	1	-1,	10,	L	
3	318	325	54	0	1048	1105	26	0	481	537	33	4	496	477	33	7	442	405	42	
4	413	417	27	1	577	639	36	2	426	437	32	5	750	726	30	8	310	254	54	
5	576	416	26	2	850	792	31	3	456	398	34	5	1030	964	26	5	478	420	39	
6	1314	1274	22	1	331	296	51	4	585	647	30	7	900	943	28	6	665	624	33	
7	581	619	27	2	484	483	39	6	1039	1056	25	8	537	540	34	8	659	597	34	
8	1363	1325	22	8	565	607	40	0	536	536	29	11	810	819	31	7	771	712	32	
9	971	956	26	10	566	450	34	1	659	648	28	13	823	870	32	7	771	712	32	
10	468	520	34	8	580	558	35	3	630	690	28	13	823	870	32	7	771	712	32	
11	-1,	7,	L	1	573	600	34	8	580	558	35	3	691	597	27	7	383	391	45	
1	321	301	49	1	573	600	34	0	437	399	33	5	1064	1076	25	5	383	391	45	
2	1463	1480	19	3	631	662	32	0	437	399	33	5	1064	1076	25	5	383	391	45	
3	880	862	21	5	648	615	28	4	446	554	32	6	552	590	34	2	584	514	41	
4	872	905	23	8	388	359	42	1	446	554	32	6	552	590	34	3	584	514	41	
5	892	892	24	8	388	359	42	2	422	431	35	7	864	870	29	3	421	422	44	
6	588	615	28	1	395	390	41	3	426	369	36	8	522	567	34	4	933	944	27	
7	585	615	29	10	949	816	29	10	949	816	29	12	531	510	38	5	826	851	31	
8	897	860	21	0	486	480	37	3	486	480	37	3	323	322	43	10	479	479	43	
9	309	302	35	1	586	579	30	4	419	451	39	5	545	520	32	0	824	793	28	
10	527	530	29	5	756	816	31	7	545	471	36	7	545	471	36	0	824	793	28	
11	553	642	31	6	431	359	38	8	485	441	45	1	753	756	28	1	436	418	40	
12	983	986	23	9	881	854	28	7	485	441	45	1	753	756	28	2	727	707	31	
				11	709	687	32	9	483	470	42	42								

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COLUMNS ARE 10FO,10FC, 10SIG

L	KFO	FC	SIG	L	NFO	FC	SIG	L	NFO	FC	SIG	L	KFO	FC	SIG	L	NFO	FC	SIG
1	2,	12,	L	1	538	497	41	1	538	497	41	1	538	497	41	1	538	497	41

OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR  
 $\text{CpNiCoRu}(\text{CO})_6(\text{C}_2\text{Ph}_2)$



COLUMNS ARE 10FO, 10FC, 10SIG

COLUMNS ARE 10FO, 10FC, 10SIG

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG
1	691	704	3	1	486	488	12	2	1487	1447	6	4	1790	1854	7	-2	558	578	9	-2	558	578	9
2	631	623	5	2	505	484	12	3	1533	1533	6	1	645	638	11	-1	1060	1030	8	-1	1060	1030	8
3	2282	2315	6	3	371	383	14	6	1773	769	20	4	185	211	25	0	500	533	9	0	500	533	9
4	1253	1359	7	4	398	414	14	8	215	206	20	5	241	240	18	1	381	368	10	1	381	368	10
5	1473	1538	7	4	402	414	14	10	310	272	19	6	503	469	15	2	175	200	19	2	175	200	19
6	1664	1133	8	0	742	715	11	10	402	414	14	7	504	502	14	-10	351	340	20	-10	351	340	20
7	349	361	14	2	422	421	13	-8	432	418	15	8	776	801	20	-8	585	613	10	-8	585	613	10
8	486	467	13	3	454	421	13	-7	365	379	15	-6	446	461	15	5	407	378	13	5	407	378	13
9	1057	1094	13	4	156	192	23	-6	473	493	11	-4	329	351	14	6	294	289	16	6	294	289	16
1	2577	2504	4	5	219	183	22	-5	473	493	11	-4	329	351	14	-5	463	459	10	-5	463	459	10
2	122	150	12	6	172	151	25	-4	1363	1340	8	1	393	377	13	-3	501	490	11	-3	501	490	11
3	452	473	8	6	172	151	25	-4	1363	1340	8	1	393	377	13	-3	246	226	9	-3	246	226	9
4	1313	1315	8	1	195	210	23	-2	1985	1921	9	2	478	466	14	-2	297	287	26	-2	297	287	26
5	423	407	13	3	171	169	28	-1	720	717	7	5	218	217	23	1	328	388	17	1	328	388	17
6	503	484	14	0	412	430	8	0	412	430	8	5	218	217	23	1	451	452	6	1	451	452	6
7	2890	2787	5	0	248	218	21	1	1466	1471	7	9	592	596	13	2	923	880	6	-5	273	272	18
8	428	403	6	1	1084	1097	7	3	1084	1097	7	9	592	596	13	2	923	880	6	-5	273	272	18
9	454	454	7	2	1231	1183	8	-1	431	394	15	3	1276	1266	6	3	1276	1266	6	-4	294	275	17
10	543	497	7	3	1039	1041	8	0	307	303	17	5	823	839	8	-3	401	402	13	-3	401	402	13
1	1281	1274	7	4	1039	1041	8	0	307	303	17	5	823	839	8	-3	401	402	13	-3	401	402	13
2	678	657	11	5	254	250	15	1	181	171	24	6	344	344	12	-2	660	656	11	-2	660	656	11
3	839	858	11	5	254	250	15	1	181	171	24	6	344	344	12	-2	660	656	11	-2	660	656	11
4	839	858	11	5	254	250	15	1	181	171	24	6	344	344	12	-2	660	656	11	-2	660	656	11
5	956	925	6	3	276	281	9	2	369	383	10	2	469	478	15	-7	480	513	13	-7	480	513	13
6	632	591	8	4	309	293	9	-1	1000	966	8	3	356	390	14	-6	207	209	18	-6	207	209	18
7	299	308	14	4	309	293	9	-1	1000	966	8	3	356	390	14	-6	207	209	18	-6	207	209	18
8	424	401	13	9	229	913	4	-3	975	952	8	3	2029	1954	6	-4	555	575	9	-4	555	575	9
9	416	388	14	10	336	315	18	-2	141	157	10	4	1046	1046	8	-2	196	160	23	-2	196	160	23
1	561	531	14	10	336	315	18	-2	141	157	10	4	1046	1046	8	-2	196	160	23	-2	196	160	23
2	630	645	7	-9	250	254	20	2	1547	1472	5	8	733	741	11	1	220	211	19	1	220	211	19
3	655	670	8	-6	736	745	9	3	2057	2028	6	6	463	502	12	-9	261	265	19	-9	261	265	19
4	852	838	9	-3	299	319	12	-6	742	760	12	4	223	179	13	-6	463	502	12	-6	463	502	12
5	239	252	14	-2	240	204	5	-2	339	351	13	5	2036	1965	17	-5	180	177	20	-5	180	177	20
6	371	371	13	-1	740	704	5	-2	682	691	10	7	232	239	17	-4	686	716	9	-3	725	678	7
7	319	294	17	0	385	357	6	-1	403	353	11	9	507	516	13	-2	390	356	5	-1	390	356	5
8	219	170	24	1	508	499	6	0	263	264	13	10	1245	1224	7	-1	1245	1224	7	-1	1245	1224	7
9	1158	1126	8	5	432	430	9	2	777	774	09	9	591	591	13	0	655	665	4	0	655	665	4
1	377	376	10	6	441	436	11	-7	270	274	16	10	1472	1469	16	1	660	670	7	1	660	670	7
2	417	387	10	6	441	436	11	-7	270	274	16	10	1472	1469	16	1	660	670	7	1	660	670	7
3	497	701	10	7	397	402	13	-3	446	452	8	3	446	452	8	3	695	754	8	3	695	754	8
4	697	701	10	7	397	402	13	-3	446	452	8	3	446	452	8	3	695	754	8	3	695	754	8
5	416	440	12	-5	341	333	11	-7	328	311	17	-3	734	776	6	4	371	387	10	4	371	387	10
6	321	307	15	-3	2032	1990	6	-7	593	514	14	-1	1730	1749	5	4	371	387	10	4	371	387	10
7	203	186	17	-1	1302	1255	6	-5	325	308	15	0	2783	2835	4	9	375	381	17	-10	229	235	23
8	403	386	19	-3	493	496	12	-3	493	496	12	2	1646	1675	5	4	957	966	10	-7	475	493	12
9	1417	1346	5	-1	240	238	17	-1	240	238	17	3	639	635	6	-4	273	270	14	-3	273	270	14
10	243	265	19	1	1417	1346	5	-1	240	238	17	3	639	635	6	-4	273	270	14	-3	273	270	14





COLUMNS ARE 10FD, 10FC, 10SIG

COLUMNS ARE 10FD, 10FC, 10SIG

3	387	411	12	-6	202	222	12	7	302	330	18	5	204	212	17	2	333	368	16	13	479	474	13	16
-2	357	370	12	-4	753	771	9	-5	325	307	18	-3	1269	1244	8	-3	249	234	20	14	416	432	13	20
-1	202	178	15	-3	453	454	10	-4	410	434	14	-2	1120	1110	7	-4	258	247	21	11	529	525	11	21
0	495	509	10	-2	887	874	8	-3	481	476	13	-1	1790	1763	7	-1	242	237	21	10	619	634	10	21
2	317	337	12	-1	332	357	10	0	702	697	11	0	1311	1337	7	7	H	9	10	12	319	335	12	21
5	443	481	13	0	894	921	7	1	462	424	13	1	1266	1236	7	-2	329	343	19	20	169	168	20	20
7	271	263	19	1	259	258	11	2	502	508	12	2	673	677	8	-1	268	234	19	11	592	521	11	11
2	218	245	18	3	506	523	9	3	402	375	13	3	2173	2144	8	2	266	255	20	11	664	662	11	11
3	210	188	19	4	477	492	12	6	603	596	12	6	379	347	13	6	H	1	287	211	346	342	18	18
6	200	181	23	8	312	356	19	6	507	516	14	7	495	498	13	-6	291	287	16	16	346	342	18	18
6	H	7	8	8	H	9	9	-5	H	8	9	8	393	388	16	-5	136	135	27	10	H	5	11	16
-6	303	281	17	-9	228	205	24	-4	262	236	20	2	H	3	10	-4	524	521	10	16	408	410	16	16
-6	310	292	17	-8	524	517	14	-4	357	319	16	-6	374	374	14	-3	445	437	9	21	218	200	21	21
-4	321	216	21	-6	408	458	14	-3	394	406	15	-3	456	473	12	-2	210	197	13	11	805	805	11	11
-3	372	422	15	-5	631	632	11	-2	216	207	23	-3	633	647	9	-1	473	488	8	14	277	310	14	14
-2	393	401	14	-4	205	234	17	-5	460	471	13	-2	1437	1460	8	0	962	963	7	11	464	475	11	11
0	304	213	19	-3	1068	1095	9	5	498	400	16	-1	469	491	9	1	523	520	8	9	917	944	9	9
1	311	303	15	-2	502	523	9	-2	H	9	9	1	318	319	10	2	861	849	8	12	378	370	12	12
-4	316	336	18	0	253	247	12	-2	237	233	20	4	1095	1106	8	3	598	590	9	14	326	334	14	14
-3	243	272	21	4	168	172	19	-8	H	0	10	4	961	981	9	4	192	143	17	17	387	609	11	11
-1	287	303	16	2	316	352	15	-6	454	459	13	6	273	245	17	5	546	546	10	13	464	488	13	13
-2	199	132	24	6	516	532	13	-4	1095	1059	11	7	387	348	14	6	241	251	17	15	370	368	15	15
-1	192	191	25	4	495	516	12	-5	395	416	13	8	356	373	16	7	345	330	17	17	H	6	11	26
0	388	387	15	5	510	529	15	-3	456	445	6	-6	330	356	16	-5	535	534	11	11	181	168	11	11
3	315	326	18	-7	242	179	19	-7	2170	2098	6	-5	272	300	17	-4	301	293	12	12	707	697	12	12
-3	488	471	10	-6	323	344	16	1	662	613	7	0	662	613	7	0	662	613	7	7	273	259	19	19
-3	239	255	12	-4	715	712	10	4	405	362	8	1	140	101	20	0	420	436	9	9	476	487	13	13
3	315	326	18	-7	242	179	19	-7	2170	2098	6	-5	272	300	17	-4	301	293	12	12	707	697	12	12
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3	315	326	18	-7	242	179	19	-7	2170	2098	6	-5	272	300	17	-4	301	293	12	12	707	697	12	12
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3	315	326	18	-7	242	179	19	-7	2170	2098	6	-5	272	300	17	-4	301	293	12	12	707	697	12	12
3	315	326	18	-7	242	179	19	-7	2170	2098	6	-5	272	300	17	-4	301	293	12	12	707	697	12	12
3	315	326	18	-7	242	179	19	-7	2170	2098	6	-5	272	300	17	-4	301	293	12	12	707	697	12	12
3	315	326	18	-7	242	179	19	-7	2170	2098	6	-5	272	300	17	-4	301	293	12	12	707	697	12	12
3	315	326	18	-7	242	179	19	-7	2170	2098	6	-5	272	300	17	-4	301	293	12	12	707	697	12	12
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3	315	326	18	-7	242	179	19	-7	2170	2098														

COLUMNS ARE 10FO,10FC, 10SIG

COLUMNS ARE 10FO,10FC, 10SIG

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG
-8	364	351	16	4	405	389	13	-3	852	860	9	2	249	209	18	-1	323	195	17
-7	626	619	11	7	449	453	13	-2	277	302	13	3	310	322	17	0	640	1897	8
-6	654	679	12	4	H, 4,	12	8	3	422	427	10	4	533	558	12	1	408	390	11
-5	939	940	10	-7	214	213	21	0	720	735	8	5	325	350	17	2	502	500	11
-2	1386	1368	8	-6	344	382	16	1	383	389	11	5	H, 7,	13	9	3	339	362	12
-1	655	645	8	-3	518	493	12	3	502	488	10	-5	484	474	15	4	320	321	13
1	559	588	8	-3	362	362	13	2	560	576	10	-4	530	493	14	5	284	260	15
2	410	408	9	-2	418	393	12	4	397	419	12	-2	225	266	20	6	H, 2,	15	13
3	792	766	9	-1	327	337	13	5	395	387	13	-1	681	685	12	7	448	434	13
6	1271	1265	10	0	268	275	13	6	369	349	15	0	220	223	20	-6	567	564	11
7	666	649	12	1	140	102	22	3	H, 3,	13	15	0	635	664	12	7	162	186	22
H	535	521	13	2	434	455	13	-8	409	375	15	1	458	433	14	-2	371	356	12
-7	H, 1,	12	15	3	350	333	13	-5	389	381	14	2	352	339	15	-7	669	870	9
-6	843	870	11	-6	368	309	18	-4	346	372	12	3	195	165	22	-6	739	774	9
-4	300	290	13	-3	514	502	10	-3	514	502	10	4	271	287	18	0	390	280	13
-3	649	618	9	-4	H, 8,	13	24	-2	184	202	20	5	595	581	14	-3	512	509	10
-2	172	155	8	-1	413	444	11	-4	275	280	13	8	H, 8,	13	19	1	333	321	15
-1	346	353	8	0	397	402	11	0	391	371	15	-4	275	280	13	0	H, 3,	15	18
0	426	411	9	1	389	397	11	1	199	206	14	0	391	371	15	-7	281	253	18
1	746	744	8	2	199	208	17	1	449	445	15	1	199	206	14	-3	482	488	13
2	583	589	9	3	769	796	10	2	449	445	15	4	H, 0,	14	24	-3	258	304	17
3	619	639	9	5	593	606	11	4	H, 0,	14	24	4	H, 0,	14	24	-2	223	227	19
4	377	381	12	8	285	268	20	5	1009	995	10	-6	1009	995	10	-1	977	994	9
5	581	581	11	-8	H, 4,	13	17	-3	164	106	19	-3	164	106	19	1	414	420	12
6	328	340	14	-7	384	373	17	-2	1445	1409	8	-2	1445	1409	8	2	376	359	13
7	722	706	12	-4	409	394	15	-1	346	324	11	-1	346	324	11	3	285	318	15
H, 2,	12	12	12	-4	553	562	11	0	630	636	10	0	630	636	10	-6	482	488	13
2,	12	12	12	-1	295	292	14	0	561	571	9	0	561	571	9	4	582	576	12
2,	12	12	12	3	188	200	21	2	1458	1454	8	2	1458	1454	8	5	401	422	15
-3	470	455	12	2	320	317	13	2	1458	1454	8	2	1458	1454	8	6	239	290	20
-4	562	549	10	0	269	293	17	3	265	296	13	3	265	296	13	4	483	471	15
-2	1045	1081	8	4	982	1028	10	4	982	1028	10	5	975	943	11	6	H, 4,	15	16
-1	1491	1494	8	5	299	311	16	5	975	943	11	6	H, 4,	15	16	-3	270	233	16
0	368	352	10	7	312	262	17	7	406	383	14	7	406	383	14	-2	371	381	13
1	130	1417	8	8	388	435	17	8	371	328	18	8	371	328	18	0	664	680	10
3	721	704	9	8	H, 5,	13	23	-5	561	568	13	9	H, 1,	14	14	1	244	241	16
4	586	564	16	-5	561	568	13	-5	561	568	13	-7	703	662	12	1	574	576	12
5	695	711	12	-4	524	523	13	-4	524	523	13	-7	703	662	12	2	582	559	12
8	413	402	16	-2	290	291	17	-2	290	291	17	-5	483	432	12	2	431	435	15
-8	175	168	26	-2	292	293	15	-2	292	293	15	-4	223	231	18	5	402	395	16
-6	570	565	12	-1	336	647	11	-1	336	647	11	-3	870	880	9	7	H, 5,	15	16
-5	327	309	15	0	369	371	12	0	369	371	12	-1	982	955	6	8	H, 6,	14	21
-4	347	332	12	1	538	565	12	1	538	565	12	1	982	955	6	8	H, 6,	14	21
-3	582	547	10	2	151	119	23	2	311	319	13	2	311	319	13	-5	311	303	18
-2	1524	1547	8	3	1201	1195	12	3	1201	1195	12	3	1201	1195	12	-4	145	167	31
-1	353	371	11	4	264	264	17	4	264	264	17	4	264	264	17	-2	222	206	20
0	499	494	9	5	311	333	17	5	311	333	17	5	311	333	17	1	437	441	14
1	799	821	9	-6	516	529	13	-6	516	529	13	6	592	576	12	0	586	599	12
2	376	388	11	-5	H, 6,	13	28	7	586	591	13	6	592	576	12	-3	171	175	25
3	1203	1244	9	-4	503	511	14	-5	503	511	14	7	586	591	13	4	277	268	19
				-3	443	436	13	-4	443	436	13	8	H, 2,	14	14	2	H, 8,	14	13
				-2	559	560	12	-3	559	560	12	8	H, 2,	14	14	3	H, 1,	13	21
				-1	253	247	18	-2	253	247	18	-7	524	541	13	5	H, 6,	15	19
				0	1017	1030	10	0	1017	1030	10	-6	502	490	12	-5	598	587	11
				1	353	322	14	1	353	322	14	-5	430	429	13	-3	474	463	11
				2	442	450	11	2	442	450	11	-4	565	560	11	-4	477	478	10



COLUMNS ARE 10FD, 10FC, 10SIG

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG
0	H, 0, 22	14	14	-2	H, 5, 22	22	29	-4	663	660	14
1	464 470	12	12	1	154 177	29	29	-2	264	284	20
2	693 685	13	13	1	370 361	16	16	-1	269	284	19
3	552 534	13	13	1	H, 1, 23	17	17	0	624	612	13
5	827 806	13	13	3	1 287 259	17	17	2	364	366	16
-5	H, 1, 22	22	26	3	168 172	26	26	2	H, 2, 24		
-3	268 292	12	21	-1	H, 2, 23	21	21	-3	515	540	14
-2	755 759	18	18	-1	210 241	21	21	-1	239	214	19
0	274 254	16	16	-2	H, 3, 23	16	16	1	325	308	17
1	316 318	16	16	-2	380 392	16	16	3	258	256	22
1	540 509	13	13	-1	H, 3, 24	18	18	0	H, 3, 24		
2	283 258	18	18	0	295 312	21	21	0	260	264	19
4	504 502	13	13	1	262 269	22	22	1	H, 4, 24		
5	224 244	22	22	2	1 241 202	16	16	1	351	351	17
6	317 279	18	18	4	2 418 401	16	16	0	H, 1, 25		
-4	H, 2, 22	14	14	-3	317 308	18	18	0	216	187	22
-3	483 468	16	16	-2	H, 4, 23	16	16	-2	H, 3, 25		
-2	321 323	13	13	-1	374 377	16	16	0	482	471	15
0	446 445	13	13	0	297 300	18	18	0	268	253	18
1	192 182	24	24	1	465 463	16	16	1	283	257	18
4	223 225	19	19	1	382 358	19	19	0	H, 0, 26		
5	335 324	17	17	3	1 250 234	15	15	1	803	814	13
-3	H, 3, 22	15	15	-4	548 541	15	15	0	408	396	16
0	423 408	15	15	-3	H, 0, 24	15	15	2	356	355	17
1	177 174	26	26	-2	421 419	15	15	-1	H, 1, 26		
3	150 102	29	29	-1	713 694	13	13	2	408	391	16
5	376 364	17	17	1	278 316	20	20	1	383	378	17
-4	H, 4, 22	17	17	3	1 600 620	13	13	2	247	249	21
-1	379 376	24	24	4	0 184 136	22	22	-1	H, 2, 26		
	193 199	24	24	4	1 780 740	12	12	0	187	176	27
					H, 4, 22	14	14	0	235	263	21
					4 426 397	15	15	1	314	283	18
					H, 1, 24						

OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR  
 $\text{CpNiCoOs}(\text{CO})_6(\text{C}_2\text{Ph}_2)$





COLUMNS ARE 10FO,10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG				
1	6	0	10	1	1204	1356	8	14	8	120	598*	1	385	370	1	751	745	6	15	699	674	9	1	1162	1207	8	1	1162	1207	8	
2	330	332	10	2	215	157	22	15	207	234	26	2	1184	1201	2	983	986	6	16	134	8	2	3	1204	1261	8	2	1204	1261	8	
3	570	565	7	3	1275	1202	6	16	549	541	11	3	540	524	3	1128	1106	5	17	170	151	31	4	186	166	35	4	186	166	35	
4	1830	1879	5	4	404	404	10	17	114	172	48*	4	357	342	4	127	78	25	18	170	151	31	5	195	151	70	5	195	151	70	
5	1248	1262	6	5	569	551	8	17	6	6	6	6	357	342	6	10	351	358	10	19	76	169	72*	6	352	339	33	6	352	339	33
6	8	14	517*	6	1014	1030	6	1	707	746	10	5	1024	1043	5	11	761	765	6	19	76	169	72*	7	1287	1233	7	7	1287	1233	7
7	853	821	7	7	186	137	19	2	273	319	21	6	912	903	6	12	488	494	6	2	1162	1207	8	8	172	1530	7	8	172	1530	7
8	1009	978	7	8	741	765	19	3	330	323	18	6	694	688	6	13	1049	1046	6	3	1204	1261	8	9	824	298	19	9	824	298	19
9	359	369	16	9	596	601	8	8	374	360	18	8	674	687	6	14	189	172	6	10	166	169	72*	10	459	461	13	10	459	461	13
10	61	18	16	10	435	435	10	5	1352	1287	8	9	350	348	9	15	816	824	7	11	166	169	72*	11	535	535	16	11	535	535	16
11	587	596	11	11	912	922	17	6	8	35	757*	10	793	801	10	16	488	494	6	12	1162	1207	8	12	256	256	21	12	256	256	21
12	593	589	6	12	217	198	19	8	1408	1359	8	11	677	673	6	17	364	349	13	13	1204	1261	8	13	329	327	14	13	329	327	14
13	895	936	5	13	466	453	10	8	306	291	19	12	1001	1025	9	18	329	327	14	14	189	172	6	14	535	535	16	14	535	535	16
14	565	574	6	14	486	463	10	9	384	378	15	13	454	452	9	19	329	327	14	15	166	169	72*	15	459	461	13	15	459	461	13
15	367	386	9	15	363	383	13	9	291	272	19	14	430	413	9	20	546	535	11	16	166	169	72*	16	461	461	13	16	461	461	13
16	1154	1162	5	16	655	650	9	10	407	406	15	15	113	128	14	21	426	459	14	17	166	169	72*	17	466	468	5	17	466	468	5
17	557	564	7	17	8	176	563*	12	8	134	688*	16	577	579	8	22	465	467	13	18	166	169	72*	18	466	468	5	18	466	468	5
18	1119	1155	6	18	260	260	19	13	448	416	13	17	8	126	8	22	465	467	13	19	166	169	72*	19	466	468	5	19	466	468	5
19	665	670	6	19	457	452	12	14	310	317	18	18	1083	1090	7	1	884	945	8	20	166	169	72*	20	466	468	5	20	466	468	5
20	977	1009	6	20	116	224	47*	15	286	270	20	20	918	910	8	2	568	564	10	21	166	169	72*	21	466	468	5	21	466	468	5
21	462	438	8	21	6	4	4	1	931	929	9	21	8	29	615*	3	528	500	10	22	166	169	72*	22	466	468	5	22	466	468	5
22	180	150	20	22	224	236	24	2	41	108	139*	22	380	354	14	5	540	540	9	23	166	169	72*	23	466	468	5	23	466	468	5
23	250	243	6	23	689	672	6	3	971	1086	9	23	9	96	590*	6	127	136	28*	24	166	169	72*	24	466	468	5	24	466	468	5
24	1055	1042	16	24	118	142	38*	4	240	225	25	2	905	902	6	6	127	136	28*	25	166	169	72*	25	466	468	5	25	466	468	5
25	8	121	532*	5	893	898	7	5	616	610	12	1	1395	1394	5	8	123	123	30*	26	166	169	72*	26	466	468	5	26	466	468	5
26	1171	1186	6	6	272	247	16	6	291	253	23	2	1352	1105	5	10	148	42	24	27	166	169	72*	27	466	468	5	27	466	468	5
27	8	9	562*	7	1056	1020	6	7	385	383	18	3	1152	1105	5	10	148	42	24	28	166	169	72*	28	466	468	5	28	466	468	5
28	674	649	8	8	7	36	559*	8	258	246	26	4	1618	1593	4	11	1147	1127	6	9	166	169	72*	9	466	468	5	9	466	468	5
19	9	26	72*	9	901	891	7	9	856	824	10	5	6	12	423*	12	130	63	29*	10	166	169	72*	10	466	468	5	10	466	468	5
20	9	180	32	10	219	211	20	10	133	180	50*	6	1111	1122	5	13	469	449	10	11	166	169	72*	11	466	468	5	11	466	468	5
21	164	180	32	11	151	182	28	11	723	682	11	7	598	603	6	14	493	488	10	12	166	169	72*	12	466	468	5	12	466	468	5
1	518	578	12	12	512	510	10	12	159	55	39*	8	207	213	14	15	308	316	15	13	166	169	72*	13	466	468	5	13	466	468	5
2	1043	981	6	13	310	336	15	13	6	8	8	9	1421	1423	5	16	393	356	12	14	166	169	72*	14	466	468	5	14	466	468	5
3	529	529	7	14	545	564	10	14	260	216	25	10	215	223	14	17	985	1003	7	15	166	169	72*	15	466	468	5	15	466	468	5
4	1111	1092	5	15	856	847	8	15	306	287	21	11	465	463	8	18	319	306	16	16	166	169	72*	16	466	468	5	16	466	468	5
5	1282	1280	5	16	88	40	55*	3	131	149	48*	12	570	571	14	19	79	783	18	17	166	169	72*	17	466	468	5	17	466	468	5
6	457	471	8	17	717	709	9	4	243	224	26	13	351	349	11	20	819	783	16	18	166	169	72*	18	466	468	5	18	466	468	5
7	1167	1197	5	18	258	289	21	5	460	516	16	14	937	928	6	21	159	179	35*	19	166	169	72*	19	466	468	5	19	466	468	5
8	527	519	8	19	197	240	29	6	9	153	790*	15	479	465	6	21	159	179	35*	20	166	169	72*	20	466	468	5	20	466	468	5
9	314	326	12	20	6	5	5	7	577	571	15	16	574	565	8	1	536	575	13	21	166	169	72*	21	466	468	5	21	466	468	5
10	912	903	6	1	1294	1428	8	2	2678	2675	4	17	358	358	13	2	360	387	14	22	166	169	72*	22	466	468	5	22	466	468	5
11	369	361	11	2	297	296	19	2	1386	1417	4	18	171	207	23	3	949	947	14	23	166	169	72*	23	466	468	5	23	466	468	5
12	1201	1225	6	3	1440	1407	8	4	1386	1417	4	19	322	315	14	4	541	501	11	24	166	169	72*	24	466	468	5	24	466	468	5
13	565	566	9	4	303	254	18	6	288	295	9	20	8	126	579*	5	1616	1554	7	25	166	169	72*	25	466	468	5	25	466	468	5
14	485	496	10	5	601	580	11	8	218	212	12	21	142	192	36*	6	128	147	7	26	166	169	72*	26	466	468	5	26	466	468	5
15	248	232	17	6	489	465	12	10	499	504	7	22	642	620	10	7	1227	1189	7	27	166	169	72*	27	466	468	5	27	466	468	5
16	110	139	38*	7	356	349	15	12	479	465	8	23	9	211	699*	8	488	463	10	28	166	169	72*	28	466	468	5	28	466	468	5
17	237	217	18	8	68	15	69*	14	1398	1396	5	23	9	211	699*	8	488	463	10	29	166	169	72*	29	466	468	5	29	466	468	5
18	508	515	11	9	1003	968	7	16	1522	1522	6	1	184	226	24	9	188	164	23	30	166	169	72*	30	466	468	5	30	466	468	5
19	118	222	43*	10	534	536	10	18	422	426	11	2	981	931	7	11	839	808	7	31	166	169	72*	31	466	468	5	31	466	468	5
20	597	567	10	11	851	843	8	20	423	411	12	3	987	935																	

COLUMNS ARE 10FD,10FC, 10SIG

L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG			
6	1702	1704	4	19	401	395	11	1	491	498	10	12	175	174	32	13	1862	1889	72	6	283	249	13	15	283	249	13	15		
8	989	989	4	20	473	478	11	3	597	440	9	13	400	397	15	14	1862	1889	72	6	434	428	9	16	434	428	9	16		
10	150	165	16	21	390	404	13	3	440	400	7	14	400	440	15	15	1862	1889	72	6	137	137	10	17	137	137	10	17		
12	47	72	56*	22	181	241	28	5	228	222	19	15	720	704	10	16	904	947	31*	6	1096	1084	6	18	1096	1084	6	18		
14	346	338	9	23	182	274	31	5	715	688	8	16	180	139	33	17	129	221	31*	6	77	77	20	19	77	77	20	19		
16	917	902	6	24	288	306	19	6	527	533	9	18	393	361	15	18	129	221	31*	6	784	757	8	21	784	757	8	21		
18	1005	1003	6	24	288	306	19	6	527	533	9	18	393	361	15	18	129	221	31*	6	784	757	8	21	784	757	8	21		
20	718	729	8	1	518	539	9	8	110	102	35*	2	2	8	9	635*	20	233	282	16	8	81	81	22	22	81	81	22	22	
22	8	39	543*	2	1758	1649	6	9	1332	1333	6	3	295	343	19	22	328	394	15	23	136	192	40*	23	136	192	40*	23		
24	695	711	9	3	326	298	10	10	362	378	12	4	405	440	14	23	533	550	11	24	4	4	1	24	4	4	1	24		
1	534	494	5	5	117	136	23	12	80	107	53*	7	5	447	457	15	25	533	550	11	25	119	801*	1	25	119	801*	1	25	
2	2881	2848	3	6	544	551	6	13	525	524	9	8	122	298	56*	1	1	280	251	12	1	237	2110	2	1	237	2110	2	1	
3	38	122	53*	7	772	737	5	14	209	204	21	9	396	404	18	2	2	260	251	12	2	460	452	3	2	460	452	3	2	
4	2361	2376	3	8	999	988	5	15	826	809	8	9	9	42	757*	2	2	2337	2110	12	2	2337	2110	12	2	2337	2110	12	2	
5	249	264	3	9	910	928	5	16	360	370	14	9	241	229	27	3	3	460	452	6	3	460	452	6	3	460	452	6	3	
6	550	587	4	10	1343	1310	5	17	1126	1132	14	11	544	530	14	4	4	2307	2272	4	4	2307	2272	4	4	2307	2272	4	4	
7	863	874	4	11	640	668	6	18	152	124	34*	12	9	99	744*	5	5	1565	1573	4	5	1565	1573	4	5	1565	1573	4	5	
8	938	930	4	12	842	836	6	19	460	441	13	13	792	740	11	6	6	61	94	33*	6	61	94	33*	6	61	94	33*	6	61
9	781	783	5	13	235	227	15	20	148	124	36*	13	4	9	9	11	11	792	740	11	7	1454	1456	4	10	1454	1456	4	10	
10	1115	1119	4	14	145	84	24	21	159	93	34	1	1	129	26	44*	8	8	1656	1639	4	8	1656	1639	4	8	1656	1639	4	8
11	115	75	23	15	1011	1012	6	21	4	6	6	2	2	690	676	11	10	2	2	2	2	2	2	2	2	2	2	2	2	
12	1072	1089	5	16	189	219	22	1	439	419	12	3	213	210	30	10	10	2837	2880	4	10	547	564	4	10	547	564	4	10	
13	342	346	10	17	1242	1218	6	2	200	184	23	3	2	690	676	11	11	547	564	6	11	547	564	6	11	547	564	6	11	
14	778	780	6	18	8	44	8	3	1049	1147	8	5	9	58	717*	12	12	1256	1267	5	12	1256	1267	5	12	1256	1267	5	12	
15	498	500	8	19	547	548	8	4	148	148	32	6	281	293	23	13	13	937	954	8	13	937	954	8	13	937	954	8	13	
16	214	263	17	20	52	548	89*	5	1268	1257	7	7	167	111	37*	14	14	400	408	8	14	400	408	8	14	400	408	8	14	
17	7	107	502*	21	198	195	24	6	68	117	72*	2	3	3122	3086	3	15	3	561	582	7	15	561	582	7	15	561	582	7	15
18	310	339	13	22	486	470	12	7	412	400	13	4	2242	2249	3	16	16	1086	1078	6	16	1086	1078	6	16	1086	1078	6	16	
19	8	18	554*	23	466	490	13	8	239	216	20	6	484	489	5	17	17	159	270	26	17	159	270	26	17	159	270	26	17	
20	560	548	9	24	741*	741*	6	10	645	619	9	6	484	489	5	18	18	684	686	7	18	684	686	7	18	684	686	7	18	
21	8	24	741*	25	1567	1681	6	11	1287	1257	7	7	1595	1616	4	19	19	8	43	335*	14	19	8	43	335*	14	19	8	43	
22	932	938	7	2	475	467	11	11	1287	1257	7	7	1595	1616	4	20	20	309	308	8	20	309	308	8	20	309	308	8	20	
23	9	81	614*	3	909	901	7	12	346	344	15	12	1397	1454	4	21	21	8	156	519*	7	21	8	156	519*	7	21	8	156	
24	485	474	12	4	746	713	8	13	1112	1051	17	14	430	440	6	22	22	353	346	13	22	353	346	13	22	353	346	13	22	
1	548	524	9	5	562	554	8	14	199	115	24	16	670	695	6	23	23	264	292	20	23	264	292	20	23	264	292	20	23	
2	1392	1297	5	6	1001	986	6	15	202	177	25	18	1011	1021	7	24	24	206	295	27	24	206	295	27	24	206	295	27	24	
3	953	960	5	7	497	494	8	16	240	219	22	20	903	898	6	25	25	249	302	24	25	249	302	24	25	249	302	24	25	
4	1106	1078	4	8	431	419	6	17	164	147	33	22	354	345	13	25	25	3	3	3	25	3	3	3	25	3	3	3	25	
5	1353	1340	4	9	34	60	95*	18	179	55	30	24	154	142	31	1	1	2286	2351	5	1	2286	2351	5	1	2286	2351	5	1	
6	2348	2373	4	10	143	101	23	19	596	594	11	1	3	1	1	2	2	982	930	6	2	982	930	6	2	982	930	6	2	
7	643	651	5	11	720	696	7	1	659	622	5	1	659	622	5	3	3	561	571	6	3	561	571	6	3	561	571	6	3	
8	1886	1894	4	12	154	38	23	1	525	527	11	2	1053	1041	3	4	4	1261	1230	4	4	1261	1230	4	4	1261	1230	4	4	
9	1066	1085	5	13	1558	1585	5	2	480	485	12	3	358	350	3	5	5	267	257	10	5	267	257	10	5	267	257	10	5	
10	327	346	9	14	434	437	10	3	8	174	699*	4	2328	2302	3	6	6	2134	2131	4	6	2134	2131	4	6	2134	2131	4	6	
11	1050	1048	5	15	852	831	18	5	452	442	11	5	452	442	11	7	7	619	643	6	7	619	643	6	7	619	643	6	7	
12	537	533	7	16	248	243	18	6	307	321	6	6	2115	2155	3	8	8	1726	1721	4	8	1726	1721	4	8	1726	1721	4	8	
13	799	808	6	17	75	60	56*	6	90	35	62*	7	790	789	4	9	9	155	162	17	9	155	162	17	9	155	162	17	9	
14	869	845	6	18	340	312	14	7	898	867	6	8	1817	1859	4	10	10	35	35	3	10	35	35	3	10	35	35	3	10	
15	117	69	29*	20	214	278	26	9	1090	1037	8	10	544	546	5	11	11	962	989	5	11	962	989	5	11	962	989	5	11	
16	547	534	8	21	929	923	34	10	157	133	66*	10	11	844	840	5	12	12	1506	1508	5	12	1506	1508	5	12	1506	1508	5	12
17	219	261	18	22	9	49	646*	11	456	432	13	12	1094	1155	5	14	14	1115	1111	5	14	1115	1111	5	14	1115	1111	5	14	

COLUMNS ARE 10FU,10FC, 10SIG

L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	
15	283	249	13	15	283	249	13	15	283	249	13	15	283	249	13	15	283	249	13	15	283	249	13	15	283	249	13	15
16	434	428	9	16	434	428	9	16																				



COLUMNS ARE 10FO,10FC, 10SIG

COLUMNS ARE 10FO,10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	
1	1202	1136	4	2	223	2361	5	6	410	385	11	22	814	788	7	
2	557	545	5	3	664	646	7	7	14	102	302*	23	261	141	570*	
3	1827	1720	3	4	2887	2793	5	8	305	273	14	24	1093	1076	7	
4	192	221	9	5	744	744	6	9	1589	1569	6	25	488	487	13	
5	601	600	4	6	681	686	6	10	488	487	13	26	454	456	12	
6	1136	1163	3	7	564	573	6	11	1717	1661	6	0	0	4	L	
7	268	261	3	8	441	431	7	12	497	498	10	0	3556	3809	4	
8	1913	1952	3	9	1413	1412	5	13	815	789	8	1	123	146	19	
9	1250	1249	4	10	244	252	12	14	599	603	9	2	1107	1133	3	
10	1570	1568	4	11	1078	1092	5	15	24	224	197*	3	998	987	3	
11	1286	1324	4	12	158	186	19	16	8	35	546*	4	642	616	4	
12	534	538	6	13	690	681	6	18	178	144	28	5	729	738	4	
13	763	745	5	14	347	347	10	19	479	477	12	6	818	840	4	
14	1271	1260	5	15	363	360	10	20	245	211	22	7	1405	1384	3	
15	94	205	34*	16	595	620	8	21	390	374	15	8	1875	1887	3	
16	1375	1366	5	17	349	346	12	22	108	150	57*	9	707	697	4	
17	739	787	7	18	664	660	8	23	358	342	18	10	2141	2112	4	
18	881	907	6	19	549	572	9	24	282	252	22	11	755	747	5	
19	447	499	11	20	8	140	8	26	0	0	L	12	864	854	5	
20	8	94	538*	21	802	811	8	27	2998	2993	1	13	1266	1272	4	
21	227	221	18	22	8	160	595*	28	4	377	356	14	13	1003	1021	5
22	463	464	10	23	439	473	13	29	3729	3655	2	15	1210	1185	5	
23	8	154	620*	24	481	476	13	30	1931	1920	3	16	867	873	6	
24	688	665	9	25	8	140	546*	31	10	412	435	17	17	728	720	7
25	9	9	791*	26	1063	1098	6	32	0	0	L	18	764	722	6	
26	564	544	11	27	571	584	8	33	84	850	8	19	410	425	10	
1	1663	1605	5	28	237	240	16	34	16	186	43*	20	936	912	6	
2	824	818	5	29	287	286	13	35	11	216	203	13	20	555	530	9
3	117	103	19	30	2190	2136	5	36	8	822	844	6	21	383	378	11
4	247	263	4	31	699	705	7	37	20	1460	1516	6	22	366	343	11
5	1564	1547	4	32	2101	2097	5	38	22	686	698	7	23	104	93	38*
6	376	384	4	33	135	121	25	39	24	301	300	15	24	8	48	534*
7	1280	1321	4	34	1016	1008	6	40	1191	1150	7	25	164	199	30	
8	118	135	19	35	160	179	21	41	0	1	L	26	541	528	11	
9	699	696	5	36	99	99	517*	42	339	301	4	1	1408	1389	6	
10	861	856	5	37	1009	1014	6	43	3752	3715	2	2	1130	1099	6	
11	462	471	7	38	534	534	9	44	2	68*	46*	3	1332	1349	6	
12	1325	1343	4	39	1082	1063	6	45	3	743	717	2	4	1587	1683	4
13	756	790	6	40	645	631	8	46	4	3301	3239	2	5	1772	1784	4
14	920	930	5	41	626	607	8	47	5	993	928	3	6	1951	1960	6
15	943	963	5	42	130	107	32*	48	7	571	608	4	7	860	906	6
16	7	44	470*	43	154	169	29	49	8	551	552	4	8	784	807	4
17	730	736	6	44	390	357	13	50	9	1839	1791	5	9	848	823	4
18	707	688	7	45	91	142	55*	51	10	584	600	6	10	669	636	5
19	497	509	9	46	22	467	450	52	11	707	734	10	11	1610	1593	5
20	966	956	6	47	418	409	14	53	12	1126	1123	4	12	354	342	11
21	8	106	529*	48	1	764	798	54	13	6	37	386*	13	958	930	5
22	476	461	11	49	7	494	456	55	14	844	885	5	14	789	788	6
23	447	469	12	50	2	455	486	56	15	772	774	5	15	592	574	6
24	9	204	633*	51	3	1768	1832	57	16	1133	1125	6	16	799	795	12
25	575	577	11	52	4	605	580	58	17	876	872	6	17	681	702	7
-1	-1	4	L	53	5	1513	1495	59	18	826	840	6	18	1182	1194	5
				54	6	454	436	60	19	107	83	31*	19	787	736	8
				55	7	454	436	61	20	234	248	16	20	780	737	8
				56	8	130	463*	62	21	130	130	463*	21	105	102	36*
				57	9	130	463*	63	22	105	102	36*	22	105	102	36*

COLUMNS ARE 10F0,10FC, 10S16

COLUMNS ARE 10F0,10FC, 10S16

L	23	0	9	5	L	KFO	448	L	8	SIG	14	FC	461	L	KFO	123	FC	141	SIG	15	L	KFO	749	FC	745	L	7	SIG	4	L	KFO	58	FC	61	SIG	7	FC	1191	SIG	7																								
		0	6	27			263		10		19		245			683		699		4		8		244		244																																						
		1	114	47			340		11		16		342			542		553		5		9		562		586																																						
		2	258	260			236		12		23		231			807		814		5		10		349		350																																						
		3	411	416			696		13		10		687			1469		1496		5		11		1186		1190																																						
		4	495	486			175		14		33		161			912		916		5		12		348		336																																						
		5	1697	1802			781		15		10		770			405		447		8		13		1301		1304																																						
		6	228	224			0		16		682*		117			379		385		10		14		964		941																																						
		7	1574	1609			9						9			708		710		7		15		309		313																																						
		8	308	295			580		1		11		558			645		649		7		16		8		96																																						
		9	853	833			249		2		13		476			554		580		8		17		8		8																																						
		10	488	513			834		3		22		222			278		288		16		18		890		903																																						
		11	858	848			192		4		9		834			138		70		14		19		890		903																																						
		12	212	264			731		5		30		221			664		641		23		20		356		352																																						
		13	1687	1643			428		6		10		684			8		51		5		21		1067		1035																																						
		14	66	212			358		7		15		316			9		37		7		22		162		206																																						
		15	1337	1293			314		8		18		288			767		781		8		23		575		551																																						
		16	432	436			9		9		20		79			1		202		9		24		199		196																																						
		17	8	114			246		10		20		266			392		365		19		25		58		88																																						
		18	193	187			487		11		26		524			967		938		7		26		58		88																																						
		19	607	574			0		12		15		524			2		593		4		27		321		320																																						
		20	9	23			603		13		12		589			621		2203		4		28		43		73																																						
		21	634	601			167		14		36		103			602		610		3		29		656		691																																						
		22	0	7			496		15		45*		443			1586		1572		4		30		56		33																																						
		23	469	453			333		16		18		35			1385		1375		5		31		1031		1046																																						
		24	653	640			341		17		19		261			100		99		6		32		1031		1046																																						
		25	933	909			259		18		20		261			1771		1773		7		33		199		196																																						
		26	913	859			1		19		20		79			1770		1773		8		34		199		196																																						
		27	515	499			1651		20		21		1434			2388		2360		9		35		199		196																																						
		28	343	325			4676		21		22		4449			833		845		10		36		199		196																																						
		29	400	429			4058		22		23		3884			2388		2360		11		37		199		196																																						
		30	172	153			728		23		24		702			566		587		12		38		199		196																																						
		31	1300	1301			2083		24		25		2073			846		833		13		39		199		196																																						
		32	175	154			2997		25		26		2982			1338		1325		14		40		199		196																																						
		33	396	399			1625		26		27		1642			425		445		15		41		199		196																																						
		34	180	287			281		27		28		302			794		793		16		42		199		196																																						
		35	597	605			60		28		29		30			170		209		17		43		199		196																																						
		36	371	343			361		29		30		346			534		540		18		44*		199		196																																						
		37	1019	1000			853		30		31		863			587		577		19		45		199		196																																						
		38	197	203			158		31		32		42			423		457		20		46		199		196																																						
		39	561	537			1		32		33		1			259		330		21		47		199		196																																						
		40	0	8			3013		33		34		3424			147		164		22		48		199		196																																						
		41	188	163			487		34		35		466			587		577		23		49		199		196																																						
		42	745																																																													

COLUMNS ARE 10FD\*10FC; 10SIG

COLUMNS ARE 10FD\*10FC; 10SIG

16	1368	1370	5	18	70	92	51*	18	8	117	530*	0	803	798	8	12	1310	1342	5	
18	20	92	156*	19	726	708	7	19	811	811	7	0	439	432	12	13	1346	1348	5	
20	134	97	25	20	580	615	8	20	116	98	37*	1	909	973	8	4	72	52	40*	
21	1018	1003	8	21	846	832	7	22	84	136	60*	3	809	829	8	6	1488	1493	5	
22	772	746	10	23	8	60	551*	23	428	427	13	4	914	970	8	17	7	48	523*	
24	582	586	10	24	379	381	13	24	2	167	597*	5	563	589	11	18	1422	1412	6	
25	130	145	38*	25	130	145	38*	0	449	450	10	6	728	753	9	14	202	210	9	
0	1186	1245	4	0	299	288	11	1	420	425	10	8	671	678	9	20	619	636	9	
1	466	447	4	1	47	66	59*	2	650	674	6	9	255	255	19	6	325	323	14	
2	2886	2827	3	2	236	232	11	3	1375	1493	6	10	328	337	15	20	802	785	6	
3	180	202	3	3	1276	1269	4	4	699	726	7	11	253	282	19	8	15	503*	609*	
4	205	216	7	4	873	865	4	5	1864	1897	5	12	8	29	573*	24	760	769	9	
5	559	553	4	4	1952	1921	4	6	205	170	16	13	601	614	10	3	1014	992	6	
6	381	365	5	5	955	942	4	7	1237	1274	5	14	8	119	549*	0	1086	1182	6	
7	1631	1623	3	6	555	571	5	8	437	436	9	15	757	755	9	1	712	710	5	
8	191	227	11	8	539	536	6	9	65	70	49*	16	163	35	31	2	1680	1667	6	
9	2332	2320	4	9	695	672	5	10	288	286	12	17	879	850	8	3	642	628	8	
10	1908	1891	4	10	2093	2068	4	11	540	539	8	18	169	243	34	4	2257	2180	15	
11	425	420	22	11	573	572	6	12	144	84	23	2	2	8	L	5	365	361	6	
12	1124	1106	6	12	2052	2046	4	13	841	846	7	0	566	574	11	6	1210	1196	4	
13	425	420	22	13	745	746	6	14	7	42	517*	1	1298	1240	7	7	947	952	4	
14	714	709	6	14	174	168	17	15	784	786	7	2	441	444	13	8	481	459	5	
15	1124	1106	6	15	766	777	6	16	217	181	19	3	610	598	11	9	561	566	5	
16	1329	1347	5	16	1058	1058	6	17	826	798	7	4	895	922	9	10	6	121	129	20
17	7	293	606*	17	830	845	7	18	8	121	525*	5	8	706	720	11	12	1485	1441	4
18	1329	1347	5	18	1238	1216	6	19	182	155	24	6	706	720	11	12	1485	1441	4	
19	204	695	17	19	434	456	10	20	100	179	47*	7	733	768	10	13	6	117	438*	5
20	271	350	17	20	470	462	10	21	407	393	13	8	209	219	23	14	1642	1635	5	
21	541	570	9	21	296	302	15	22	230	250	23	9	509	534	12	15	597	634	7	
22	166	220	27	22	8	115	555*	10	239	228	22	16	874	849	6	16	712	690	7	
23	295	291	16	23	766	760	9	23	341	346	17	17	499	489	8	17	411	397	10	
24	291	297	18	24	812	792	9	24	1519	1528	6	12	427	426	14	18	143	137	23	
25	291	297	18	25	812	792	9	25	85	75	50*	13	154	190	36*	19	36	278	116*	5
26	124	124	26	26	755	738	7	0	662	685	9	14	370	360	16	20	840	854	10	
0	377	391	4	0	2836	2864	5	1	2	76	631*	21	288	275	21	21	8	271	719*	7
1	684	671	4	2	492	500	8	3	114	130	35*	0	501	496	13	22	1270	1293	6	
2	832	791	7	3	1512	1531	5	4	153	102	25	1	8	24	644*	23	22	174	770*	8
3	244	218	4	4	312	322	9	5	776	795	8	2	686	679	10	24	471	468	11	
4	627	618	4	5	338	345	9	6	72	87	53*	3	831	822	10	25	9	52	594*	39*
5	1983	2013	3	6	1164	1147	5	7	1069	1057	7	4	193	209	32	0	297	361	11	
6	414	411	6	8	1977	1936	6	8	810	786	10	3	810	786	10	1	462	468	7	
7	2103	2104	4	9	549	534	5	9	124	341	14	6	124	45	50*	2	229	223	9	
8	1239	1217	4	10	108	163	27*	10	629	607	9	7	293	273	22	3	510	510	5	
9	1174	1172	4	11	508	507	7	11	369	379	13	8	475	459	15	4	444	439	5	
10	1295	1277	4	12	582	576	7	13	185	207	55	6	23	182	280*	5	832	864	4	
11	545	575	7	13	791	780	8	14	791	780	8	11	99	144	64*	6	68	126	28*	26*
12	2580	2559	4	14	954	931	6	15	8	55	595*	0	2	10	L	7	819	803	4	
13	359	405	10	15	290	300	13	16	952	939	8	8	9	136	674*	8	582	610	5	
14	1449	1454	5	16	463	454	9	17	247	259	22	1	426	407	16	10	1916	1905	10	
15	690	692	7	17	147	141	23	18	652	601	11	2	29	220	217*	11	626	644	6	

COLUMNS ARE 10FU,10FC, 10SIG

13	17	18	17	96	128	48*	L	KFO	FC	SIG
14	614	614	7	9	46	590*	4	1807	1687	4
15	21	67	163*	19	562	580	8	941	881	7
16	113	131	32*	20	9	101	10	1391	1345	4
17	516	518	9	3,	7,	L	12	784	772	5
18	362	351	11	0	300	313	14	307	311	10
19	531	537	9	1	918	929	21	332	387	0
20	192	172	21	2	686	676	18	36	94	8
21	520	535	11	3	351	362	20	948	938	16
22	78	61	64*	5	182	243	22	1025	1038	7
23	767	757	9	6	344	320	24	597	597	10
0	459	432	10	7	1060	1110	0	1042	1020	4
1	1571	1614	6	8	421	429	2	601	614	7
2	509	537	10	9	986	997	3	500	538	5
3	718	748	7	10	614	583	3	88	57	21*
4	472	497	9	11	429	413	4	2165	2174	4
5	780	812	7	12	300	282	5	5	83	377*
6	64	83	53*	13	8	67	6	1490	1467	4
7	1817	1798	5	14	97	38	8	542	506	5
8	409	418	10	15	532	505	9	1053	1038	4
9	1589	1593	18	16	179	81	9	135	137	17
10	193	161	15	17	736	703	10	322	314	8
11	393	393	18	0	809	827	11	822	802	5
12	200	211	18	1	368	342	12	547	529	6
13	343	336	12	2	589	584	13	772	748	6
14	8	24	50*	3	1068	1038	14	855	851	6
15	660	643	8	4	119	164	15	117	149	28*
16	8	109	510*	5	922	968	16	857	891	6
17	770	738	8	6	593	616	17	337	388	13
18	329	328	14	7	318	333	18	900	927	7
19	500	498	11	8	613	639	19	571	604	9
20	60	41	78*	9	483	508	20	187	235	23
21	217	197	23	10	346	345	21	165	197	26
22	106	276	53*	11	666	668	22	381	360	13
0	110	161	40*	12	9	125	23	9	80	789*
1	788	784	6	13	398	399	24	1116	1110	8
2	310	296	16	14	118	112	0	1826	1950	6
3	1107	1163	663*	0	296	297	1	131	127	19
4	1281	1341	51*	2	843	832	2	1183	1176	4
5	81	25	51*	3	275	273	3	225	204	10
6	808	804	7	4	644	620	5	192	202	11
7	194	183	22	5	373	377	6	616	594	5
8	170	272	25	6	424	401	7	1146	1151	5
9	138	133	30*	7	810	847	8	384	388	7
10	1131	1129	7	8	127	47	9	1102	1119	5
11	131	172	34*	9	425	423	10	589	599	6
12	1157	1132	7	10	269	282	11	296	283	10
13	8	55	561*	4	0,	L	12	1180	1161	5
14	643	627	9	0	2439	2225	13	747	737	6
15	197	248	24	2	3402	3200	14	938	956	6
16	197	248	24	2	3402	3200	15	958	966	6

COLUMNS ARE 10FU,10FC, 10SIG

17	19	158	18	494	530	507*	L	KFO	FC	SIG
18	20	342	19	533	530	28	16	158	167	30
19	175	175	20	682	698	15	17	860	845	7
20	433	441	21	174	141	29	18	656	651	8
21	787	786	22	433	441	10	19	8	86	539*
22	262	268	23	787	786	8	20	1154	1161	7
23	986	946	24	262	268	8	21	332	387	16
0	986	946	0	262	268	8	22	837	852	8
1	350	328	1	986	946	8	23	320	343	15
2	933	916	2	1039	1088	8	24	368	430	16
3	282	260	3	1363	1391	6	25	153	134	33*
4	136	136	4	797	796	7	26	4	3,	L
5	160	160	5	1068	1115	6	27	320	308	13
6	180	210	6	1115	1115	6	28	775	778	7
7	4	8,	7	16	180	210	29	210	206	30
8	485	490	8	0	485	490	30	601	614	4
9	364	354	9	1	364	354	31	1522	1498	5
10	737	741	10	2	737	741	32	1371	1348	5
11	543	556	11	3	543	556	33	1138	1173	5
12	591	587	12	4	591	587	34	426	409	8
13	70	30	13	5	70	30	35	1594	1608	5
14	320	283	14	6	320	283	36	234	241	13
15	467	477	15	7	467	477	37	910	891	17
16	334	351	16	8	334	351	38	793	777	7
17	194	232	17	9	194	232	39	554	575	9
18	698	698	18	10	698	698	40	107	210	41*
19	481	473	19	11	481	473	41	388	401	13
20	419	411	20	12	419	411	42	354	366	15
21	719	697	21	13	719	697	43	280	297	19
22	156	156	22	14	156	156	44	1091	1130	7
23	1767	1593	23	15	1767	1593	45	1501	1607	7
24	1603	1500	24	16	1603	1500	46	984	1028	7
25	3689	3493	25	17	3689	3493	47	453	461	9
26	2615	2331	26	18	2615	2331	48	391	376	9
27	184	141	27	19	184	141	49	1537	1518	14
28	2027	1937	28	20	2027	1937	50	234	213	14
29	772	767	29	21	772	767	51	232	2403	5
30	149	128	30	22	149	128	52	494	500	8
31	168	175	31	23	168	175	53	705	707	11
32	92	503*	32	24	92	503*	54	328	335	11
33	341	331	33	25	341	331	55	574	567	11
34	2189	2189	34	26	2189	2189	56	436	431	9
35	990	993	35	27	990	993	57	1035	1028	6
36	2230	2255	36	28	2230	2255	58	85	168	45*
37	2230	2255	37	29	2230	2255	59	739	718	7
38	2230	2255	38	30	2230	2255	60	8	66	500*
39	2230	2255	39	31	2230	2255	61	436	415	11



COLUMNS ARE 10FU\*10FC, 10SIG

COLUMNS ARE 10FD\*10FC, 10SIG

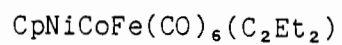
L	KFD	FC	SIG	L	NFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG
3	5	1	11	7	732	740	6	14	366	386	14	14	837	833	7	2	121	195	43*
4	136	146	18	9	814	809	6	15	1069	1070	17	15	604	620	9	3	268	252	18
5	178	161	14	10	894	945	234*	16	359	347	15	16	1067	1055	7	4	607	632	10
6	1328	1318	4	11	755	777	6	17	307	282	17	17	8	41	542*	5	1322	1325	7
7	6	46	430*	12	609	605	7	19	280	289	20	19	413	407	13	6	310	307	15
8	2485	2415	4	13	913	908	6	0	5	6	34*	2	734	736	9	8	223	222	20
9	518	500	7	14	114	140	34*	4	154	157	34*	4	6	3	9	9	815	856	8
10	1248	1222	5	15	927	967	7	1	559	561	12	10	575	619	10	10	662	673	34*
11	32	14	90*	16	349	336	13	2	76	75	70*	10	934	936	6	11	131	143	16
12	419	416	9	17	307	324	14	3	7	62	705*	12	762	752	7	12	289	292	16
13	377	353	10	18	665	659	13	4	1525	1536	6	14	1525	1536	6	13	143	174	31
14	519	514	8	19	219	216	21	5	165	101	28	16	705	712	8	14	175	116	26
15	825	819	7	20	510	519	11	6	138	100	33*	18	100	145	41*	4	468	463	12
16	355	356	11	21	222	247	11	7	281	278	17	20	92	118	50*	5	240	233	22
17	271	259	15	22	295	277	18	8	100	55	47*	22	139	124	36*	6	706	692	10
18	513	529	9	5	5	4	1	9	805	782	8	22	6	1	6	7	6	6	6
19	389	387	12	0	171	192	28	9	8	25	562*	0	370	351	8	8	6	60	710*
20	594	605	9	1	1326	1385	7	11	1169	1201	7	1	207	178	13	9	1196	1254	9
21	430	461	12	2	353	361	13	12	143	33	34*	2	939	958	5	10	702	696	8
22	457	470	12	3	1000	998	7	13	702	699	9	3	647	619	6	11	178	156	22
23	191	281	29	4	1084	1086	6	14	459	465	13	4	1880	1908	5	12	932	990	8
0	102	79	31*	5	221	219	16	15	334	302	16	5	506	493	7	13	232	221	18
1	179	169	16	6	639	676	6	16	147	230	39*	6	1585	1573	5	14	571	578	36*
2	1415	1364	5	7	1058	1056	8	17	1016	984	8	7	215	187	14	15	417	429	11
3	586	546	6	8	7	39	490*	8	5	7	7	8	270	235	12	16	83	77	51*
4	1890	1858	4	9	1353	1330	6	9	486	478	13	9	63	94	50*	17	438	441	30
5	463	465	7	10	378	381	11	10	527	526	12	10	1661	1635	5	18	431	422	10
6	1231	1243	5	11	959	971	6	11	482	497	13	11	365	354	11	19	525	527	13
7	115	77	23	12	735	731	7	12	542	566	13	12	1588	1558	5	20	578	587	11
8	332	347	9	13	181	149	22	13	330	340	19	13	246	246	16	21	6	4	7
9	473	461	7	14	129	103	31*	14	366	401	16	14	681	650	7	0	511	568	13
10	349	353	10	15	345	371	13	15	425	436	14	15	302	289	14	1	963	1050	9
11	889	907	6	16	8	83	504*	16	488	515	13	16	356	356	12	2	569	588	10
12	934	939	6	17	790	777	8	17	583	590	11	17	541	521	9	3	1670	1696	17
13	542	543	8	18	202	201	23	18	434	433	14	18	648	631	10	4	114	81	36*
14	984	976	6	19	629	624	9	19	362	337	16	19	349	350	14	5	949	979	7
15	100	108	38*	20	89	122	57*	20	71	66	77*	20	624	632	10	6	518	532	15
16	646	678	8	21	344	369	16	21	9	144	682*	21	9	139	593*	7	101	83	16
17	926	956	7	5	5	5	9	12	875	860	9	6	2	2	7	8	856	869	17
18	120	185	35*	13	486	502	11	13	163	176	37*	0	710	689	7	9	772	787	8
19	604	621	9	1	565	608	11	14	5	8	37*	1	1161	1129	5	10	499	502	10
20	396	398	13	2	1026	1104	8	2	213	210	27	2	53	70	54*	11	952	944	7
21	147	181	33*	3	1062	1095	8	3	432	448	15	3	726	707	6	12	392	384	12
22	484	484	12	4	451	450	11	4	218	184	27	4	138	133	20	13	571	565	9
0	1295	1403	7	5	743	779	8	5	258	262	24	5	239	235	13	14	449	451	11
1	619	605	8	6	485	489	10	6	763	787	11	6	728	747	6	15	306	338	16
2	703	694	7	7	1021	1007	9	7	163	20	40*	7	7	88	459*	16	394	379	13
3	861	882	6	8	373	569	9	8	727	779	12	8	1328	1324	5	17	97	121	51*
4	105	96	27*	9	312	314	14	9	229	231	28	9	140	187	24	18	169	178	30
5	751	758	6	10	276	282	15	10	169	68	37	10	1142	1130	6	19	300	278	18
6	916	928	6	11	310	321	15	11	528	535	14	11	399	404	10	20	299	299	20
7	1133	1130	6	12	330	336	14	12	398	381	17	12	203	181	19	21	471	509	14
8	1133	1130	6	13	1133	1130	7	13	5	9	7	13	734	717	7	22	471	509	14



COLUMNS ARE 10FD,10FC, 10SIG

L	KFO 10,	FC 1,	SIG L	L	KFO 10,	FC 2,	SIG L	L	KFO 10,	FC 3,	SIG L
3	128	113	40*	0	640	676	10	7	75	62	70*
4	840	815	8	1	468	478	12	0	28	71	215*
5	343	318	16	2	683	675	9	1	415	367	15
6	996	984	8	3	502	504	12	2	466	451	13
7	258	258	21	4	98	76	52*	3	80	105	67*
8	433	439	13	5	298	265	18	4	677	698	10
9	56	19	93*	6	450	440	13	5	25	77	211*

OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR



COLUMNS ARE 10FO,10FC, 10SIG

1	1	9	1	185	183	114	19	19	6	1	236	212	9		
2	1	10	45	227*	21	367	357	4	173*	7	564	570	5		
3	3	215	211	11	207	12	4	173*	7	170	9	335	254*		
4	4	10	90	238*	319	8	11	485	463	7	1070	1055	4		
5	4	5	78	83	5	-7,	4,	7	4,	7	5	55	16		
6	10	10	80	221*	364	21	21	21	78	7	8	82	194*		
7	-9,	2,	2	216	375	114	17	109	114	8	801	823	5		
8	252	255	10	10	336	234	9	339	336	9	196	201	10		
9	82	88	26*	19	54	216*	7	293	294	8	477	481	6		
10	4	10	21	200*	101	33	9	246	254	9	92	61	21*		
1	253	262	10	200*	494	476	7	376	373	7	288	275	8		
2	-8,	0,	L	137	128	16	13	149	150	13	211	198	11		
3	72	30	25*	-7,	0,	L	11	205	197	11	-6,	2,	L		
4	304	308	8	485	488	6	10	103	95	18	469	477	5		
5	125	130	16	357	369	5	12	199	205	12	1014	1035	5		
6	346	341	8	253	253	8	3	163	165	10	202	202	10		
7	228	225	11	698	705	5	1	570	587	6	302	314	6		
8	-8,	1,	L	1015	1000	6	2	324	321	8	56	66	26*		
9	250	247	9	249	235	10	3	40	23	45*	6	106	134	16	
10	312	321	8	-7,	1,	L	4	350	350	7	181	178	10		
1	9	49	215*	555	551	6	5	111	117	18	8	238	241	8	
2	372	368	7	372	368	7	7	9	9	210*	8	9	204*	8	
3	704	692	5	704	692	5	7	344	346	8	10	604	595	6	
4	145	146	12	145	146	12	8	218	218	10	11	325	323	8	
5	99	43	17	99	43	17	9	279	269	9	9	12	367	365	8
6	83	17	23*	209	205	9	10	177	177	13	13	-6,	3,	L	
7	404	405	7	333	340	7	1	-7,	6,	L	1	183	186	9	
8	115	142	19	162	159	13	2	216	227	10	1	2	368	370	6
9	585	566	7	9	51	51	3	194	201	11	2	3	274	272	7
10	-8,	2,	L	63	62	32*	3	83	61	22*	3	4	108	27	17
1	94	69	20	732	709	6	4	92	118	23*	4	5	420	412	7
2	277	283	9	176	156	13	5	343	347	8	5	6	145	134	14
3	445	452	7	-7,	2,	L	6	438	455	7	6	6	82	49	19*
4	31	54	65*	149	148	70	7	238	242	10	8	9	392	384	7
5	13	20	154*	124	149	175	8	185	175	12	9	10	44	222*	8
6	140	132	14	8	70	214*	13	10	44	222*	10	11	96	89	20
7	333	329	8	601	590	6	1	386	401	7	11	12	147	157	15
8	119	110	17	124	149	175	2	9	79	216*	12	13	298	292	9
9	90	72	23*	63	62	32*	3	56	61	34*	13	1	-6,	4,	L
10	378	368	8	474	481	6	4	373	384	8	4	5	521	524	5
1	468	472	7	281	270	8	5	333	83	62*	6	6	490	508	6
2	9	42	228*	641	629	6	6	109	115	19	1	2	163	152	13
3	131	133	15	71	38	26*	7	304	301	9	3	3	-6,	9,	L
4	9	42	211*	215	209	11	2	-6,	0,	L	4	4	212	216	8
5	358	355	9	-7,	3,	L	5	1706	1743	4	5	5	253	254	8
6	332	328	8	766	800	5	2	361	379	6	6	6	310	321	7
7	91	57	20*	699	722	5	6	558	568	5	7	8	130	126	14
8	566	556	7	218	230	9	8	92	42	17	9	9	251	255	8
9	278	271	9	341	347	7	10	638	649	6	6	10	181	180	7
10	-8,	4,	L	247	249	8	12	418	414	7	11	11	285	288	9
1	287	247	10	426	427	7	6	374	388	7	7	7	250	247	7
2	426	427	7	453	462	7	1	-6,	1,	L	12	12	431	433	7
3	453	462	7	90	58	20*	1	7	41	201*	12	12	431	433	7

COLUMNS ARE 10FO,10FC, 10SIG

1	9	9	9	227	212	9	227	212	9	227	212	9	227	212	9	227	212	9	
2	11	11	8	335	344	7	335	344	7	335	344	7	335	344	7	335	344	7	
3	13	13	13	485	463	7	485	463	7	485	463	7	485	463	7	485	463	7	
4	1	1	1	21	78	92*	6	114	113	14	39	39	1561	1601	4	1561	1601	4	
5	2	2	2	109	114	17	8	801	823	5	7	7	1503	1519	4	1503	1519	4	
6	3	3	3	339	336	9	9	196	201	10	8	8	473	485	6	473	485	6	
7	5	5	5	293	294	8	8	477	481	6	9	9	168	172	8	168	172	8	
8	6	6	6	246	254	9	9	92	61	21*	10	10	250	256	7	250	256	7	
9	7	7	7	376	373	7	7	376	373	7	7	7	79	79	18*	79	79	18*	
10	8	8	8	149	150	13	13	211	198	11	12	12	252	246	10	252	246	10	
1	9	9	9	205	197	11	11	-6,	2,	L	10	10	347	354	7	347	354	7	
2	10	10	10	103	95	18	18	469	477	5	1	1	204	204	10	204	204	10	
3	11	11	11	199	205	12	12	1014	1035	5	2	2	215	211	9	215	211	9	
4	-7,	5,	L	-7,	5,	L	3	163	165	10	3	3	202	202	10	202	202	10	
5	1	1	1	570	587	6	6	302	314	6	4	4	287	280	8	287	280	8	
6	2	2	2	324	321	8	8	56	66	26*	5	5	359	361	7	359	361	7	
7	3	3	3	40	23	45*	6	106	134	16	6	6	416	418	7	416	418	7	
8	4	4	4	350	350	7	7	181	178	10	7	7	68	64	25*	68	64	25*	
9	5	5	5	111	117	18	18	8	238	241	8	8	230	236	10	230	236	10	
10	6	6	6	9	9	210*	9	9	9	204*	8	8	764	757	4	764	757	4	
1	7	7	7	344	346	8	8	10	604	595	6	6	69	56	27*	69	56	27*	
2	8	8	8	218	218	10	10	11	325	323	8	8	10	42	219*	10	42	219*	
3	9	9	9	279	269	9	9	12	367	365	8	8	-6,	7,	L	-6,	7,	L	
4	10	10	10	177	177	13	13	-6,	3,	L	1	1	171	180	12	171	180	12	
5	-7,	6,	L	216	227	10	1	183	186	9	2	2	144	155	13	144	155	13	
6	2	2	2	194	201	11	2	368	370	6	3	3	256	260	9	256	260	9	
7	3	3	3	83	61	22*	3	274	272	7	4	4	108	27	17	108	27	17	
8	4	4	4	92	118	23*	4	269	283	7	5	5	420	412	7	420	412	7	
9	5	5	5	343	347	8	5	81	78	18*	6	6	145	134	14	145	134	14	
10	6	6	6	438	455	7	6	82	49	19*	8	8	451	443	7	451	443	7	
1	7	7	7	238	242	10	6	263	252	17	9	9	118	109	17	118	109	17	
2	8	8	8	185	175	12	8	392	395	7	1	1	600	592	4	600	592	4	
3	9	9	9	10	44	222*	10	392	384	7	2	2	819	831	4	819	831	4	
4	1	1	1	386	401	7	11	96	89	20	3	3	376	373	7	376	373	7	
5	2	2	2	9	79	216*	12	147	157	15	3	3	196	188	11	196	188	11	
6	3	3	3	56	61	34*	13	298	292	9	4	4	236	216	10	236	216	10	
7	4	4	4	373	384	8	4	-6,	4,	L	5	5	387	385	8	387	385	8	
8	5	5	5	333	83	62*	6	521	524	5	6	6	102	83	19	102	83	19	
9	6	6	6	109	115	19	2	490	508	6	7	7	9	140	230*	9	9	140	230*
10	7	7	7	304	301	9	3	-6,	9,	L	8	8	163	152	13	163	152	13	
1	8	8	8	-6,	0,	L	4	212	216	8	1	1	244	243	10	244	243	10	
2	9	9	9	1706	1743	4	5	253	254	8	2	2	66	47	32*	66	47	32*	
3	10	10	10	361	379	6	6	310	321	7	3	3	335	333	8	335	333	8	
4	11	11	11	558	568	5	7	130	126	14	4	4	47	28	43*	47	28	43*	
5	12	12	12	92	42	17	8	251	255	8	5	5	-5,	0,	L	-5,	0,	L	
6	13	13	13	638	649	6	6	9	9	217*	1	1	790	743	4	790	743	4	
7	14	14																	

COLUMNS ARE 10FO,10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	
1	149	160	10	6	349	352	7	13	179	176	12	7	1428	1464	4	1	300	295	5	
2	185	193	10	7	349	68	224*	14	531	508	7	8	548	561	5	2	718	708	4	
3	438	453	5	8	137	139	214*	15	128	129	18	9	316	312	7	3	250	254	6	
4	19	31	75*	9	9	21	212*	1	368	388	5	10	15	111	27*	4	963	974	4	
5	882	891	12	1	9	21	9	2	368	388	5	10	15	507	7	5	1316	1362	4	
6	143	152	12	2	250	241	9	3	93	121	12	11	3	405	411	6	405	411	5	
7	197	200	9	3	31	48	55*	4	820	838	4	12	1	810	801	7	913	940	4	
8	8	5	213*	4	217	218	10	5	4	1038	1056	4	13	7	697	703	5	703	5	
9	182	165	11	5	332	340	8	6	568	579	4	4	3	447	421	8	124	129	13	
10	342	348	7	6	115	1130	4	7	1	486	496	5	4	4	139	127	10	127	129	13
11	415	398	15	8	600	600	5	8	2	144	139	11	5	383	387	10	59	71	30*	
12	137	138	15	9	318	322	6	9	3	358	360	6	6	321	317	5	173	178	12	
13	508	477	17	10	384	380	6	10	4	362	353	6	7	164	161	8	372	353	8	
1	451	462	6	11	105	92	21	11	5	101	102	17	8	482	504	5	106	45	19	
2	449	446	6	12	369	364	7	12	6	718	711	5	9	450	462	5	3	5	L	
3	152	158	11	13	83	83	8	13	7	58	55	35*	10	19	119	119	66	28	17*	
4	87	63	18*	14	275	266	6	14	8	290	292	8	11	161	160	11	669	671	4	
5	99	76	6	1	706	709	4	1	4	211	208	10	12	166	163	12	3	1143	1165	4
6	367	370	6	2	779	776	4	2	5	162	158	13	13	691	662	6	4	201	209	7
7	484	484	6	3	334	338	5	3	6	144	156	14	14	86	86	5	5	129	138	11
8	270	262	8	4	384	393	6	4	7	48	8	8	15	417	379	8	6	792	787	5
9	86	53	22*	5	30	44	4	5	8	132	127	15	15	437	439	5	7	593	606	5
10	269	243	9	6	491	493	5	6	9	279	279	7	1	968	955	3	8	69	78	22*
11	139	150	16	7	491	493	5	7	10	358	394	7	2	627	635	3	9	28	70	59*
12	274	271	7	8	706	709	4	8	11	209	209	9	3	682	674	3	10	433	423	6
1	274	271	7	9	779	776	4	9	12	141	135	16	4	911	917	3	11	361	364	4
2	224	232	9	10	334	338	5	10	13	141	135	16	5	908	926	4	12	138	131	15
3	451	477	6	11	465	459	6	11	14	130	122	16	6	346	351	5	13	159	166	14
4	451	477	6	12	568	568	6	12	15	296	279	9	7	1004	1037	4	14	51	39	41*
5	474	474	6	13	85	83	8	13	16	4	8	8	8	475	479	5	1	479	480	5
6	297	302	8	14	297	296	9	14	17	177	171	10	10	156	169	11	2	877	857	4
7	297	302	8	1	204	206	5	1	1	452	440	6	11	177	171	10	3	383	396	6
8	451	477	6	2	391	404	8	2	2	519	508	6	12	87	75	22*	4	370	358	6
9	451	477	6	3	493	498	5	3	3	187	195	11	13	87	100	23*	5	935	953	4
10	274	271	7	4	598	601	5	4	4	288	291	8	14	412	378	8	6	160	163	10
11	274	271	7	5	64	82	23*	5	5	167	138	12	15	4	3	3	7	537	548	5
12	274	271	7	6	64	82	23*	6	6	9	42	225*	1	331	322	4	8	291	283	7
13	274	271	7	7	117	110	4	7	7	123	122	16	2	331	322	4	9	81	86	22*
14	274	271	7	8	489	494	5	8	8	61	101	32*	3	963	948	3	10	146	145	13
15	274	271	7	9	340	353	7	9	9	172	172	13	4	477	482	4	11	349	358	8
1	482	503	4	10	314	320	7	10	1	209	218	11	5	328	334	5	12	93	84	22*
2	261	261	8	11	211	213	10	11	2	75	123	28*	6	807	843	4	13	256	244	10
3	64	24	18*	12	211	213	10	12	3	150	155	14	7	225	235	5	1	128	125	12
4	957	954	4	13	9	38	235*	13	4	267	274	9	8	497	492	5	2	28	28	186*
5	570	599	4	14	150	169	15	14	5	4	4	9	9	359	349	6	3	583	594	5
6	160	165	8	15	4	6	19	15	6	145	153	15	10	477	494	6	4	96	109	16
7	120	120	197*	1	615	609	5	1	5	195	198	12	11	8	10	215*	5	8	10	207*
8	7	52	194*	2	176	180	9	2	6	3	40	249*	12	8	10	215*	6	332	333	7
9	718	729	5	3	234	239	7	3	7	3	0	L	13	376	362	7	7	287	292	7
10	230	245	8	4	995	1021	5	4	8	1863	1803	3	14	38	42	55*	8	335	336	7
11	278	278	8	5	465	465	8	5	9	80	78	14	15	291	269	10	9	173	123	28*
12	284	288	8	6	227	223	8	6	10	836	814	3	15	3	4	L	10	201	204	11

COLUMNS ARE 10FO,10FC, 10SIG









COLUMNS ARE 10FO,10FD, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG
5	4	63	222*	0	63	100	32*	3	8	58	233*	0	5	5	7
6	8	845	5	4	350	353	6	4	526	500	7	0	369	18	244*
7	107	138	20	5	122	137	10	2	286	275	10	1	304	35	244*
8	400	406	7	6	182	185	12	3	150	141	17	2	242	15	44*
9	75	141	30*	8	315	309	9	9	196	180	13	3	78	364	9
10	9	30	250*	9	326	323	9	0	199	209	13	5	391	7	7
11	88	61	255*	9	5	4	5	6	372	384	8	6	454	4	7
	4	9	9	0	88	85	18	0	648	635	5	7	74	39	30*
	7	9	7	1	288	291	17	2	268	272	8	0	650	627	6
	1	151	10	2	100	115	17	4	225	212	10	1	415	403	7
	3	331	14	3	56	47	30*	8	232	235	11	3	462	446	7
	3	117	16	4	264	267	8	6	6	1	1	3	111	103	20
	4	54	30*	5	285	284	8	0	689	682	5	4	115	135	21
	5	77	20*	6	233	224	10	1	326	324	6	5	77	107	31*
	6	194	11	7	161	176	16	2	79	79	22*	6	10	51	269*
	7	246	10	8	210	213	13	3	8	7	227*	6	6	102	116
	8	84	25*	9	0	123	128	4	176	190	12	0	9	34	233*
	9	9	250*	0	565	562	5	6	348	340	8	1	105	113	22
	10	164	15	1	362	357	7	7	6	111	115	2	194	202	13
		5	6	2	495	482	6	8	326	323	9	3	171	167	14
		6	4	3	8	85	245*	6	6	2	6	4	84	82	27*
		7	5	4	298	292	9	0	557	547	5	0	398	403	8
		8	6	5	223	231	11	1	284	283	7	1	345	331	9
		9	7	6	57	55	40*	2	353	363	7	2	242	225	11
		10	8	7	132	143	19	3	487	484	6	3	143	120	17
			9	8	5	6	16	4	221	227	10	7	7	0	1
			10	0	253	241	8	5	111	79	18	1	212	218	10
				1	511	521	6	6	191	191	12	3	263	274	9
				2	411	406	7	7	460	458	7	5	142	130	16
				3	94	8	21*	8	284	299	10	7	7	1	1
				4	298	290	9	0	6	3	1	0	160	157	12
				5	116	156	20	1	131	108	13	1	395	405	7
				6	137	156	18	2	544	520	6	2	338	353	8
				7	506	496	7	3	222	216	9	3	221	203	10
				8	230	242	12	4	217	219	10	4	108	104	20
				9	5	7	249*	5	98	60	19	5	9	69	234*
				10	711	690	5	5	574	590	7	6	245	248	11
					447	422	7	6	395	388	8	0	624	643	6
					323	310	8	7	9	91	275*	1	252	252	7
					85	166	29*	8	285	297	10	2	158	168	14
					452	448	7	0	858	851	5	3	316	322	8
					119	73	20	1	170	162	12	3	580	570	7
					259	271	11	2	504	519	6	5	9	18	237*
					97	106	26*	3	396	405	7	6	24	35	95*
					5	8	132*	4	205	211	11	0	106	157	20
					17	60	268*	5	9	49	248*	7	3	1	6
					9	43	273*	6	74	32	29*	1	632	616	6
					449	439	8	7	383	398	9	0	243	228	11
					81	67	30*	8	0	0	0	0	87	5	1
					128	112	12	4	153	170	10	5	243	228	11
					163	166	10	5	0	0	0	0	243	228	11

COLUMNS ARE 10FO,10FD, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG
5	4	63	222*	0	63	100	32*	3	8	58	233*	0	5	5	7
6	8	845	5	4	350	353	6	4	526	500	7	0	369	18	244*
7	107	138	20	5	122	137	10	2	286	275	10	1	304	35	244*
8	400	406	7	6	182	185	12	3	150	141	17	2	242	15	44*
9	75	141	30*	8	315	309	9	9	196	180	13	3	78	364	9
10	9	30	250*	9	326	323	9	0	199	209	13	5	391	7	7
11	88	61	255*	9	5	4	5	6	372	384	8	6	454	4	7
	4	9	9	0	88	85	18	0	648	635	5	7	74	39	30*
	7	9	7	1	288	291	17	2	268	272	8	0	650	627	6
	1	151	10	2	100	115	17	4	225	212	10	1	415	403	7
	3	331	14	3	56	47	30*	8	232	235	11	3	462	446	7
	3	117	16	4	264	267	8	6	6	1	1	3	111	103	20
	4	54	30*	5	285	284	8	0	689	682	5	4	115	135	21
	5	77	20*	6	233	224	10	1	326	324	6	5	77	107	31*
	6	194	11	7	161	176	16	2	79	79	22*	6	10	51	269*
	7	246	10	8	210	213	13	3	8	7	227*	6	6	102	116
	8	84	25*	9	0	123	128	4	176	190	12	0	9	34	233*
	9	9	250*	0	565	562	5	6	348	340	8	1	105	113	22
	10	164	15	1	362	357	7	7	6	111	115	2	194	202	13
		5	6	2	495	482	6	8	326	323	9	3	171	167	14
		6	4	3	8	85	245*	6	6	2	6	4	84	82	27*
		7	5	4	298	292	9	0	557	547	5	0	398	403	8
		8	6	5	223	231	11	1	284	283	7	1	345	331	9
		9	7	6	57	55	40*	2	353	363	7	2	242	225	11
		10	8	7	132	143	19	3	487	484	6	3	143	120	17
			9	8	5	6	16	4	221	227	10	7	7	0	1
			10	0	253	241	8	5	111	79	18	1	212	218	10
				1	511	521	6	6	191	191	12	3	263	274	9
				2	411	406	7	7	460	458	7	5	142	130	16
				3	94	8	21*	8	284	299	10	7	7	1	1
				4	298	290	9	0	6	3	1	0	160	157	12
				5	116	156	20	1	131	108	13	1	395	405	7
				6	137	156	18	2	544	520	6	2	338	353	8
				7	506	496	7	3	222	216	9	3	221	203	10
				8	230	242	12	4	217	219	10	4	108	104	20
				9	5	7	249*	5	98	60	19	5	9	69	234*
				10	711	690	5	5	574	590	7	6	245	248	11
					447	422	7	6	395	388	8	0	624	643	6
					323	310	8	7	9	91	275*	1	252	252	7
					85	166	29*	8	285	297	10	2	158	168	14
					452	448	7	0	858	851	5	3	316	322	8
					119	73	20	1	170	162	12	3	580	570	7
					259	271	11	2	504	519	6	5	9	18	237*
					97	106	26*	3	396	405	7	6	24	35	95*
					5	8	132*	4	205	211	11	0	106	157	20
					17	60	268*	5	9	49	248*	7	3	1	6
					9	43	273*	6	74	32	29*	1	632	616	6
					449	439	8	7	383	398	9	0	243	228	11
					81	67	30*	8	0	0	0	0	87	5	1
					128	112	12	4	153	170	10	5	243	228	11
					163	166	10	5	0	0	0	0	243	228	11

OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR  
 $\text{CoMn(CO)}_6(\text{MeC}_2\text{Ph})_2$

COLUMNS ARE 10FO,10FC, 10SIG

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L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG
2	-17	0	140*	2	760	746	2	4	1720	1732	2	0	688	686	2	0	985	982	3
4	7	60	138*	4	858	847	2	8	649	655	3	2	1978	1947	2	2	49	17	10*
2	-14	0	L	6	492	466	3	10	905	934	3	4	404	377	2	4	290	296	3
4	218	219	5	8	492	485	3	12	564	561	4	6	506	492	2	6	500	518	3
6	145	150	7	10	91	99	9	14	412	401	4	8	642	629	3	8	390	395	5
2	-15	0	L	14	397	410	4	16	103	94	9	10	701	701	3	10	123	118	8
4	50	107	18*	2	152	143	7	2	-1	0	L	12	524	511	3	12	77	64	12
6	6	29	121*	2	524	516	2	6	1125	1100	1	14	315	312	4	11	236	224	5
8	155	147	7	4	57	500	3	8	986	975	2	16	167	174	7	2	78	91	11
2	-14	0	L	6	512	500	3	8	601	602	3	0	1234	1231	2	0	344	366	3
4	132	135	7	8	363	356	3	10	431	414	3	2	645	612	2	6	361	358	3
6	151	155	6	10	485	487	3	12	158	171	6	4	709	725	2	8	213	201	4
8	245	257	5	14	75	78	11	14	81	22	9	6	904	933	2	10	76	31	11
10	100	111	10	16	255	263	5	16	8	29	109*	8	385	389	3	12	205	213	6
2	-13	0	L	2	104	105	4	2	2402	2577	1	12	292	303	4	0	236	218	4
4	6	48	131*	4	431	416	3	4	427	427	2	14	312	306	4	2	130	126	5
6	214	235	5	6	244	243	3	6	925	912	2	4	6	0	L	2	43	12	17*
8	568	556	3	8	250	250	3	8	800	761	3	0	1088	1073	2	6	123	122	7
10	403	392	4	10	441	454	3	10	41	59	15*	2	1561	1508	2	8	85	85	10
2	703	710	3	12	394	388	4	12	140	137	6	4	746	724	2	10	307	304	5
4	460	450	3	14	78	82	12	14	184	194	5	6	199	215	3	0	13	0	L
6	334	336	4	16	399	382	5	16	7	49	147*	8	980	982	3	0	389	396	4
8	64	108	14*	2	-5	0	L	2	1	0	L	10	1137	1145	3	2	6	46	121*
10	6	36	134*	4	693	675	2	4	1165	1147	1	12	537	519	4	4	348	355	5
12	315	320	5	6	1030	1012	2	6	540	584	2	14	232	242	6	6	172	171	6
2	-11	0	L	8	907	899	2	8	342	328	3	0	815	778	2	10	106	116	10
4	605	624	3	10	589	601	3	10	80	72	8	2	1073	1050	2	0	14	0	L
6	585	570	3	12	83	87	10	12	112	109	7	4	1019	1058	2	0	41	28	18*
8	902	885	3	14	116	105	8	14	242	247	5	6	856	856	3	2	145	138	6
10	548	512	4	16	378	361	5	16	314	316	5	8	517	529	3	4	295	306	4
12	6	33	126*	2	-4	0	L	2	2	0	L	10	199	199	5	6	354	349	4
2	890	909	3	4	1060	1019	2	0	1615	1547	1	12	265	270	5	8	305	296	5
4	953	963	3	6	426	414	2	6	928	906	2	14	173	175	6	0	15	0	L
6	665	668	3	8	245	243	3	8	28	22	13*	0	1368	1358	2	2	291	300	4
8	32	41	20*	10	509	514	3	10	34	46	17*	2	445	426	3	4	378	388	4
10	458	448	4	12	894	899	3	12	332	327	4	4	17	19	28*	6	268	287	5
12	93	124	10	14	729	719	3	14	232	232	5	8	513	534	3	6	161	0	L
2	-9	0	L	16	208	203	6	16	120	104	9	10	870	860	3	0	201	202	5
4	84	77	6	2	-3	0	L	2	3	0	L	12	209	218	3	4	92	146	12
6	449	454	3	4	226	219	2	4	433	414	1	14	25	12	36*	4	270	265	5
8	215	219	4	6	842	819	2	6	48	53	7	14	9	0	L	0	17	0	L
10	364	361	4	8	912	949	2	8	966	957	2	0	411	407	3	2	260	262	5
12	6	31	130*	10	699	712	3	10	124	118	4	2	976	994	3	2	421	441	5
2	74	59	3	12	758	744	3	12	366	336	3	4	1034	1058	3	1	17	1	L
4	605	570	3	14	5	18	145*	8	366	336	3	6	615	610	3	2	7	28	133*
6	585	570	3	16	271	273	4	10	385	378	3	6	247	248	4	3	81	88	11
8	902	885	3	18	219	202	6	12	92	91	9	8	247	248	4	3	68	77	14*
10	548	512	4	2	-2	0	L	14	548	560	4	10	159	166	6	4	100	113	10
12	6	33	126*	4	1622	1613	1	16	571	560	4	12	74	33	10	4	-16	1	L

COLUMNS ARE 10FO,10FC, 10SIG

L	NFO	FC	SIG	L	NFO	FC	SIG	L	NFO	FC	SIG	L	NFO	FC	SIG	L	NFO	FC	SIG	L	NFO	FC	SIG	
1	393	393	3	10	77	91	10	16	64	53	14*	2	1041	1015	1	0	893	864	1	2	213	196	3	
2	418	431	3	11	158	169	6	1	-4	1	14*	3	340	330	2	1	583	559	2	3	226	217	3	
3	165	156	5	12	127	133	9	1	89	95	4	3	131	125	3	2	414	402	2	4	472	456	2	
4	524	525	3	13	272	271	126*	2	275	734	2	4	346	339	5	1	426	416	1	5	319	331	4	
5	274	270	3	14	272	271	126*	3	207	199	2	5	1192	1210	2	2	1326	1316	2	6	759	771	5	
6	173	164	5	15	73	78	13*	4	1546	1529	2	6	121	119	2	3	442	432	3	7	451	457	6	
7	390	381	4	1	-7	1	L	5	533	541	2	8	651	653	4	4	239	235	4	8	199	218	7	
8	34	23	21*	2	378	382	3	6	551	569	3	9	419	406	6	5	893	864	5	9	567	564	8	
9	380	378	4	3	506	502	2	7	186	213	3	10	113	121	3	6	1680	1713	6	10	86	92	9	
10	6	40	144*	4	673	655	2	8	44	64	4	11	355	369	14*	7	58	57	14*	11	358	345	10	
11	6	24	11	5	650	649	2	9	51	59	11*	12	146	146	6	8	116	116	12	12	121	118	11	
12	75	24	11	6	461	466	3	10	447	457	3	13	209	196	6	9	1326	1316	13	13	143	141	12	
13	97	96	10	7	5	14	111*	11	589	584	3	14	224	218	5	10	442	432	14	14	180	171	13	
1	748	762	3	8	682	700	3	12	378	382	4	15	213	198	5	11	221	241	15	15	204	193	14	
2	33	63	17*	9	395	383	3	13	143	143	7	16	255	248	5	12	221	241	16	16	199	218	15	
3	524	530	3	10	411	412	3	14	146	143	7	1	1680	1713	5	13	893	864	17	17	576	545	16	
4	160	159	4	11	363	354	4	15	7	90	148*	2	3	583	574	1	14	583	559	18	1246	1218	17	
5	455	458	3	12	153	163	6	1	-3	1	L	3	413	401	2	15	414	402	19	19	419	414	18	
6	159	154	5	13	70	6	137*	2	820	605	2	3	1118	1094	2	16	414	402	20	20	821	787	19	
7	70	48	8	14	6	51	137*	3	449	449	2	4	359	349	2	17	1326	1316	21	21	609	619	20	
8	200	194	4	15	7	13	135*	4	595	598	2	5	194	170	3	18	442	432	22	22	444	464	21	
9	437	445	4	1	75	80	5	5	351	352	2	6	813	813	2	19	221	241	23	23	495	505	22	
10	237	237	5	2	463	454	5	6	308	303	2	7	547	546	3	20	13	27	24	24	517	517	23	
11	405	417	4	3	85	76	6	7	817	833	2	8	705	703	3	21	255	248	25	25	748	752	24	
12	260	265	5	4	337	338	2	8	242	232	3	9	648	636	3	22	461	451	26	26	444	464	25	
13	88	86	10	5	93	119	5	9	651	665	3	10	118	122	6	23	461	451	27	27	495	505	26	
14	184	182	6	6	329	327	3	10	348	354	3	11	262	261	4	24	221	241	28	28	112	117	27	
1	356	361	3	7	395	391	3	11	5	12	116*	12	130	108	6	25	221	241	29	29	748	752	28	
2	868	877	3	8	242	253	3	12	410	401	4	13	130	108	6	26	143	154	30	30	5	47	137*	
3	62	41	8	9	41	68	14*	13	292	292	4	14	261	248	5	27	156	148	31	31	416	422	29	
4	612	607	3	10	187	178	8*	14	469	463	4	15	87	62	10	28	156	148	32	32	82	44	30	
5	712	698	3	11	14	78	48*	15	235	237	5	16	284	288	5	29	32	64	33	33	19	137*	31	
6	800	801	3	12	67	67	12*	16	233	231	6	1	1764	1748	1	30	1642	1593	34	34	125	137	32	
7	556	548	3	13	67	37	12*	17	-2	1	L	2	1795	768	1	31	1411	1408	35	35	281	299	33	
8	250	257	4	14	95	113	9	18	1272	1260	1	3	1222	1173	1	32	144	146	36	36	472	456	34	
9	125	126	6	15	280	274	5	19	842	829	1	4	613	585	2	33	764	741	37	37	760	764	35	
10	154	147	6	1	1015	995	2	20	830	808	2	5	1034	1047	2	34	164	159	38	38	389	367	36	
11	348	336	9	2	623	608	2	21	867	867	2	6	185	171	3	35	600	592	39	39	562	586	37	
12	99	114	6	3	864	862	2	22	1328	1338	2	7	190	176	3	36	230	242	40	40	235	237	38	
13	249	235	5	4	17	26	20*	23	437	427	3	8	213	231	3	37	484	479	41	41	87	99	39	
14	131	120	7	5	371	371	2	24	958	969	3	9	253	264	3	38	51	88	42	42	189	197	40	
1	761	756	2	6	566	563	2	25	650	659	3	10	142	153	3	39	146	152	43	43	201	198	41	
2	175	153	3	7	579	581	3	26	288	296	3	11	433	433	3	40	105	94	44	44	344	342	42	
3	432	430	3	8	292	298	3	27	484	484	3	12	32	28	5	41	504	496	45	45	111	111	43	
4	164	159	4	9	773	768	3	28	169	162	5	13	32	28	5	42	6	6	46	46	163	154	44	
5	56	43	9	10	409	422	3	29	169	162	5	14	32	28	5	43	226	226	47	47	272	272	45	
6	143	147	5	11	339	336	4	30	251	242	5	15	6	28	6	44	226	226	48	48	833	809	46	
7	397	411	5	12	341	339	4	31	58	56	15*	16	6	28	6	45	186	186	49	49	500	506	47	
8	156	153	5	13	224	233	5	32	84	75	10	17	52	74	19*	17	5	5	5	51	514	507	48	
9	183	189	5	14	255	262	5	33	96	89	10	18	2	1	L	18	743	714	50	50	321	335	49	
				15	212	207	6	34	-1	1	L	19	2	295	289	2	1	365	353	51	51	184	170	50
									271	273	1	20	823	786	1		296	293	52	52	303	305	51	

COLUMNS ARE 10FO,10FC, 10SIG

COLUMNS ARE 10FO,10FC, 10SIG

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L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG		
8	228	222	4	2	-17	2, L	18	140*	10	-12	2, L	7	278	289	4	14	42	41	18*		
7	307	317	4	1	93	99	12	12	228	240	5	8	675	696	3	15	167	176	6		
8	327	342	5	1	211	207	6	15*	11	57	52	9	209	209	5	1	-4	2, L	1		
9	720	725	3	2	133	109	8	127*	12	193	184	6	10	627	624	1	195	186	2		
10	234	248	5	3	-16	2, L	4	127*	1	-11	2, L	4	11	241	250	2	3	783	724	2	
11	200	212	4	4	97	110	10	10	2	218	219	9	12	161	155	5	3	585	566	2	
12	53	28	16*	1	2	59	16	10	3	278	274	4	13	113	120	6	4	1091	1088	2	
13	202	198	6	2	161	159	7	7	4	143	147	5	14	172	127	8	5	78	81	6	
14	60	75	15*	3	92	63	10	9	5	398	411	3	1	-7	2, L	6	4	466	512	3	
9	150	146	7	4	116	118	9	9	6	71	114	11	1	322	312	3	7	455	453	3	
10	236	242	5	5	77	49	12	12	2	30	33	25*	2	861	904	2	8	601	598	3	
11	274	279	5	6	15	2, L	12	146*	7	70	39	39	3	332	321	3	9	542	556	3	
13	13	1, L	5	8	-15	2, L	14	12	8	55	16	13*	5	4715	737	2	10	74	68	9	
0	206	203	5	1	6	11	143*	5	9	174	175	5	5	36	70	15*	11	128	134	6	
2	413	402	4	2	491	480	4	4	10	85	118	10	6	738	745	11	12	70	75	11	
3	154	169	6	3	237	226	5	4	11	41	112	24*	7	293	292	3	13	126	145	7	
4	103	118	6	4	399	400	4	4	12	296	301	5	8	5	5	119*	14	131	122	7	
5	648	627	3	5	4	399	400	4	13	141	127	7	9	129	127	6	15	77	76	12	
6	81	79	8	6	5	6	34	138*	5	-10	2, L	4	10	311	311	4	16	58	63	16*	
7	114	103	6	7	293	290	5	5	1	207	208	4	11	31	19	23*	1	84	86	4	
8	278	297	4	8	181	195	6	6	2	289	292	3	12	423	421	4	2	580	561	2	
9	160	167	6	9	138	149	8	8	3	176	174	4	13	71	134	14*	3	219	213	2	
10	303	293	4	10	-14	2, L	10	14*	4	157	157	5	14	320	316	5	4	214	209	3	
11	218	223	5	1	139	140	6	6	5	325	322	3	15	99	113	10	5	264	277	3	
12	159	170	7	2	272	105	12*	12*	6	271	269	4	1	-6	2, L	6	2	242	241	3	
13	115	122	9	3	292	303	4	4	7	134	128	6	1	190	185	3	7	4	3	104*	
0	57	71	10*	4	32	58	27*	27*	8	252	249	4	2	175	180	3	8	98	105	6	
1	131	128	5	5	39	41	20*	20*	9	72	58	11	3	496	476	2	9	508	524	3	
2	174	176	4	6	67	52	14*	14*	10	399	388	4	4	192	212	3	10	52	17	11*	
3	256	256	3	7	162	151	79*	79*	10	399	388	4	5	4	20	110*	11	303	311	4	
4	447	457	3	8	158	170	7	7	11	83	57	10	6	989	1009	3	12	347	349	4	
5	425	420	3	9	160	159	7	7	12	305	308	4	7	43	57	13*	13	199	193	5	
6	260	262	4	10	186	172	6	6	13	61	39	15*	8	912	906	4	14	330	329	5	
7	461	478	3	11	-13	2, L	11	14*	13	-9	2, L	3	9	217	218	4	15	90	44	9	
8	147	150	6	12	37	31	22*	22*	1	307	312	3	10	589	610	3	16	311	304	5	
9	81	74	8	13	517	513	4	4	2	676	686	3	11	422	411	2	1	1077	1061	2	
10	198	184	5	14	199	195	5	5	3	344	332	4	12	296	303	4	3	1131	1115	2	
11	190	196	6	15	320	319	4	4	4	772	758	3	13	185	203	5	4	404	393	2	
12	203	205	6	16	150	142	6	6	5	65	33	8	14	99	88	6	5	465	476	7	
13	91	77	11	17	131	138	7	7	6	255	253	4	15	186	188	8	6	60	58	8	
11	11	1, L	11	18	79	76	11	11	7	99	118	7	16	186	188	8	7	730	724	2	
0	750	751	3	19	-12	2, L	12	12*	8	52	43	12*	17	-5	2, L	9	8	404	393	2	
1	557	558	3	20	263	258	4	4	8	136	140	6	18	422	411	2	10	730	705	2	
2	521	505	3	21	270	278	4	4	9	345	349	4	19	1500	1508	2	11	465	476	7	
3	245	258	4	22	154	153	5	5	10	345	349	4	20	4	555	567	3	12	60	58	8
4	48	118	16*	23	154	153	5	5	11	6	27	126*	21	4	230	208	3	13	730	724	2
5	6	18	126*	24	263	258	4	4	12	501	497	4	22	4	46	106*	4	14	541	535	3
6	85	77	9	25	270	278	4	4	13	102	104	6	23	5	567	567	3	15	466	470	3
7	251	243	4	26	171	162	8	8	14	433	419	4	24	2	1500	1508	2	16	350	349	4
8	119	115	7	27	154	153	5	5	15	-8	2, L	4	25	7	57	33	8	17	59	25	10*
9	266	268	5	28	84	71	9	9	1	301	291	3	26	53	16	9*	18	117	114	4	
10	126	107	7	29	335	335	4	4	2	327	330	3	27	179	180	9*	19	269	279	5	
11	17	55	48*	30	335	335	4	4	3	327	330	3	28	68	78	9	20	169	158	6	
12	59	26	15*	31	131	135	7	7	4	305	293	3	29	82	69	9	21	75	60	13*	
12	12	1, L	5	32	131	135	7	7	5	386	378	3	30	113	122	7	22	381	369	5	
12	12	1, L	5	33	6	319	4	4	6	675	677	3	31	16	101	52*	23	-1	2, L	2	

COLUMNS ARE 10FU,10FC, 10SIG















OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR  
 $\text{CpRe}(\text{CO})_2(\text{C}_2\text{Ph}_2)$

COLUMNS ARE 10FD,10FC, 10SIG

1	-11,	0,	L	1	473	505	33	13	L	13	411	-8,	0,	L	2	944	983	18	L	2	1249	1282	SIG
5	820	821	22	4	678	664	24	3	14	18	302	255	46	3	734	690	20	3	3	1747	1728	12	
1	-11,	1,	L	2	515	466	30	4	16	27	609	559	28	6	1022	1041	16	4	4	562	558	11	
2	840	864	23	5	564	544	28	5	24	27	537	528	1, L	1	299	264	44	5	5	746	755	16	
1	413	354	38	10	573	564	28	6	1	2	1340	1374	14	9	1108	1110	16	6	6	1264	1261	12	
3	309	243	46	5	507	506	28	7	2	2	861	857	16	9	369	347	37	7	7	1662	1664	11	
4	536	490	31	1	570	582	28	7	32	32	438	422	32	10	491	500	29	8	8	773	752	16	
5	635	680	28	1	523	534	31	9	33	33	653	657	20	11	475	488	30	9	9	408	414	28	
6	373	356	43	6	839	840	22	10	30	30	443	430	28	14	833	816	20	10	10	336	333	34	
4	-11,	2,	L	7	429	377	36	11	47	47	312	320	41	2	781	740	20	11	11	278	278	45	
4	483	532	35	10	437	430	37	13	45	45	605	609	23	3	-8,	6,	L	14	14	433	410	30	
6	535	526	31	14	-10,	6,	L	14	39	38	389	368	34	3	744	738	21	15	15	689	680	22	
7	333	111	43	2	447	410	35	13	-9,	45	289	139	5,	5	1008	1023	17	16	16	501	487	29	
8	730	693	24	3	617	566	27	1	13	13	724	687	21	6	678	713	23	17	17	297	288	47	
1	-11,	3,	L	4	548	552	31	2	3	4	908	921	19	17	438	410	35	18	18	297	288	47	
2	456	482	34	5	419	370	38	3	2	4	295	226	47	18	645	557	26	19	19	743	699	23	
4	708	732	25	8	470	473	35	5	4	4	310	376	44	2	333	279	40	20	20	547	526	31	
5	345	361	48	3	-10,	7,	L	5	2	2	633	647	20	11	534	535	28	1	1	961	954	14	
1	386	383	42	6	386	383	42	6	5	5	758	744	21	5	1477	1469	12	2	2	1438	1431	12	
2	456	482	34	7	419	370	38	7	6	6	561	565	27	6	353	342	35	3	3	660	665	18	
3	344	284	48	4	328	372	50	8	7	7	564	532	27	7	399	409	31	4	4	541	523	20	
5	482	476	35	5	-9,	0,	L	9	9	9	673	667	24	9	1608	1578	13	10	10	1465	1474	13	
1	530	500	26	11	530	500	26	11	11	11	429	447	31	11	429	447	31	12	12	1069	1060	15	
2	1458	1506	16	13	1383	1453	15	13	13	13	507	483	28	13	1203	1248	16	13	13	764	722	20	
2	925	908	19	1	924	922	17	1	17	17	587	619	6,	17	415	387	37	14	14	-7,	2,	L	
10	581	580	29	3	555	533	28	2	2	2	528	529	29	13	507	483	28	15	15	514	506	22	
12	581	580	29	4	284	292	46	4	4	4	759	785	19	13	1534	1523	13	16	16	1465	1474	13	
-10,	1,	L	15	7	527	531	30	7	3	3	274	141	41	5	913	926	17	17	17	1069	1060	15	
1	534	526	29	8	-9,	1,	L	8	8	8	321	328	45	5	429	447	31	18	18	764	722	20	
2	916	907	20	10	286	209	44	10	4	4	491	482	32	6	522	514	24	19	19	-7,	3,	L	
3	387	473	39	11	290	352	47	11	11	11	504	443	31	8	647	606	21	20	20	1701	1713	11	
5	532	531	28	7	608	634	24	7	9	9	384	327	33	9	384	327	33	21	21	514	500	21	
6	1088	1056	18	8	801	781	19	1	11	11	909	884	20	11	455	428	29	22	22	370	338	20	
7	764	745	22	9	833	824	19	3	12	12	439	416	35	12	302	346	46	23	23	573	583	29	
9	321	111	44	11	663	656	23	5	13	13	849	847	21	13	954	951	18	24	24	1546	1583	12	
10	643	617	26	12	632	628	25	7	36	36	426	432	36	14	491	502	31	25	25	813	815	16	
11	329	358	47	13	664	660	24	8	36	36	553	493	28	15	506	485	30	26	26	473	458	24	
13	299	67	48	13	-9,	2,	L	10	44	44	345	126	44	17	636	626	26	27	27	601	578	20	
-10,	2,	L	2	2	588	590	24	2	17	17	-9,	8,	L	17	636	626	26	28	28	764	783	19	
1	898	942	20	3	328	413	42	1	46	46	328	317	46	1	353	324	36	29	29	378	423	34	
3	726	737	23	4	1551	1530	14	4	47	47	322	336	47	3	594	613	23	30	30	566	550	24	
4	540	513	27	7	427	420	31	5	36	36	430	455	36	4	916	924	17	31	31	509	534	29	
11	928	911	20	8	696	705	21	6	32	32	490	462	32	5	433	467	30	32	32	429	412	35	
-10,	3,	L	12	12	413	406	35	9	30	30	563	589	30	6	1093	1075	15	33	33	-7,	4,	L	
3	777	819	22	12	-9,	3,	L	9	30	30	563	589	30	6	1093	1075	15	34	34	601	578	20	
5	517	506	29	1	771	754	20	2	22	22	887	885	22	7	286	238	44	35	35	764	783	19	
6	498	419	30	2	536	535	26	2	8	8	681	670	21	8	681	670	21	36	36	378	423	34	
7	902	882	20	3	471	454	30	2	9	9	656	669	22	8	656	669	22	37	37	897	906	15	
8	305	358	50	4	285	248	46	4	9	9	911	904	16	9	1031	1034	17	38	38	1231	1216	13	
9	370	379	41	5	286	328	47	4	16	16	1483	1463	13	10	1311	1314	17	39	39	550	540	22	
11	319	363	48	6	575	576	25	8	13	13	1108	1092	14	11	1877	1906	12	40	40	884	879	16	
12	371	354	41	8	1052	1044	17	10	14	14	1283	1279	14	15	434	402	34	41	41	636	602	20	
-10,	4,	L	12	12	1062	1056	18	12	32	32	408	386	32	1	-8,	5,	L	42	42	281	281	44	
											295	311	46	1	295	311	46	43	43	1354	1360	14	

COLUMNS ARE 10FD,10FC, 10SIG



COLUMNS ARE 10FO,10FC, 10SIG

Table with 10 columns: L, KFO, FC, SIG, L, KFO, FC, SIG, L, KFO, FC, SIG. The table contains numerical data for 10 different sets of columns. Each set includes values for L, KFO, FC, and SIG, often with subscripts like 'L' or '5'. The values range from 1 to 1000.

COLUMNS ARE 10FO,10FC, 10SIG

Table with 10 columns: L, KFO, FC, SIG, L, KFO, FC, SIG, L, KFO, FC, SIG. The table contains numerical data for 10 different sets of columns. Each set includes values for L, KFO, FC, and SIG, often with subscripts like 'L' or '5'. The values range from 1 to 1000.



COLUMNS ARE 10FO,10FC, 10SIG

COLUMNS ARE 10FD,10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG				
18	768	759	19	16	610	580	22	6	631	631	21	15	1102	1110	13	13	692	717	15	13	755	784	18
19	314	332	43	17	519	519	27	8	382	341	32	18	624	603	22	14	484	492	23	16	885	886	18
20	522	535	26	18	631	650	24	9	298	254	41	19	326	349	41	15	978	1026	14	18	308	257	43
22	727	715	23	21	302	235	50	10	1311	1322	14	22	412	360	35	18	1046	1026	16	20	851	855	21
1	2073	2060	8	22	338	384	51	14	357	350	40					20	301	343	46				
2	531	514	13	1	2052	2074	9	1	-3,	11,	7,	1	2481	2411	5	21	306	275	47	3	1159	1194	11
3	2145	2123	7	3	1677	1710	10	2	866	855	18	2	1264	1254	7	22	754	726	23	5	504	529	19
4	1259	1247	8	4	729	740	14	3	1016	1025	16	3	1262	1190	7	1	-2,	5,	20	6	686	671	15
5	2189	2182	8	5	1009	1024	11	4	279	150	45	5	3882	3899	6	2	1522	1485	8	7	270	270	15
6	1603	1584	8	6	301	253	28	5	604	634	23	6	1002	971	7	3	2157	2140	7	8	753	760	15
7	404	412	19	7	529	521	19	6	413	402	31	7	1708	1694	7	4	819	816	10	9	921	938	14
8	563	572	15	8	1478	1528	11	11	592	584	25	8	761	747	10	5	723	748	12	10	664	684	18
10	1385	1427	10	10	592	566	13	14	497	502	33	9	2124	2160	7	6	499	475	15	11	555	568	21
11	1042	1098	12	11	1169	1149	13	15	902	911	21	10	411	395	18	7	2116	2129	8	12	1164	1186	14
13	1418	1388	11	12	774	774	17	17	-3,	12,	12,	11	769	771	12	9	1410	1437	10	13	633	692	22
14	694	699	18	13	334	357	36	3	519	512	27	12	606	621	15	10	1617	1644	10	15	519	532	26
15	436	474	27	15	1362	1376	14	5	618	622	24	15	645	657	18	11	241	265	38	16	681	692	22
16	578	587	23	19	741	729	23	7	1111	1131	17	16	403	434	29	12	294	294	32	17	328	300	42
17	1420	1444	14	20	371	335	42	9	714	689	22	17	1025	1028	15	14	1798	1788	11	18	334	332	43
18	304	21	41		-3,	8,	11	11	877	904	21	19	1378	1356	14	15	764	798	18				
18	355	304	38	1	274	375	35	23	13,	13,	24	23	821	788	22	18	548	534	23	1	493	492	21
1	324	303	22	2	336	348	28	1	706	675	24	23	821	788	22	18	548	534	23	2	566	552	19
2	2019	2023	8	4	777	765	14	2	986	1001	18	30	2333	2322	6	19	726	708	21	3	1905	1925	10
3	612	612	13	5	1235	1261	11	3	524	563	30	1	2333	2322	6	21	613	644	27	4	276	395	38
4	476	472	16	6	1152	1142	12	5	564	539	27	2	1994	1965	6	2	-2,	6,	9	5	943	939	14
5	1810	1791	18	7	1097	1123	13	6	463	398	32	3	193	130	29	1	1636	1614	9	7	1788	1764	11
6	767	772	12	8	484	494	23	10	345	353	45	4	1868	1855	7	2	1643	1608	8	8	425	394	22
7	782	766	18	9	673	691	18	1	-3,	14,	14,	5	896	882	7	3	508	470	15	9	524	515	27
8	479	479	18	10	939	971	15	6	329	331	49	7	993	994	9	5	1898	1895	8	10	421	470	30
9	2228	2254	9	13	297	140	39	6	885	875	22	8	1522	1525	8	6	1416	1389	9	11	628	637	21
10	811	799	14	13	140	140	39	6	-2,	0,	0,	9	1351	1348	8	7	1522	1523	10	13	350	318	39
11	893	924	13	14	413	410	31	2	4818	4841	5	10	1008	988	10	8	1289	1263	10	16	329	211	43
12	440	442	24	16	317	322	42	4	2025	2004	6	11	1104	1118	10	9	386	372	24	17	768	802	22
13	913	896	14	19	596	610	28	6	3380	3324	6	13	1749	1751	10	10	446	474	22				
14	1136	1153	14	20	464	514	37	8	672	656	10	14	938	954	14	11	1452	1443	11	1	327	276	33
16	331	349	38	1	535	531	22	14	647	652	16	16	628	649	20	13	804	811	16	2	398	430	28
17	515	476	27	2	2268	2284	10	18	2033	1986	12	17	690	690	20	16	817	838	18	3	898	930	15
18	603	603	23	3	562	563	20	20	481	535	30	20	651	673	23	17	666	666	22	7	550	539	22
21	764	749	22	6	931	952	15	22	882	863	19	22	300	300	50	19	486	482	31	8	383	436	32
1	924	955	12	7	779	795	16	24	340	353	47	23	463	436	35	22	391	366	43	9	517	524	25
2	401	392	20	9	416	460	29	1	-2,	1,	1,	25	-2,	4,	4,	15	-2,	7,	12	11	1124	1104	15
3	1334	1323	9	10	580	580	22	1	3007	2962	5	1	434	419	15	1	1785	1758	9	12	593	609	24
4	1198	1204	10	12	493	505	27	2	2961	2920	5	2	1238	1238	8	2	695	685	13	13	943	945	18
6	1458	1440	9	13	360	388	37	3	1833	1807	6	3	1346	1332	7	4	2469	2497	8	15	599	605	26
7	1700	1689	9	14	1182	1177	16	4	1464	1418	6	4	1574	1569	7	5	323	328	26	17	493	506	33
8	419	380	21	15	333	315	42	5	297	298	19	5	1716	1659	7	6	1339	1353	10	18	331	44	46
9	771	764	14	16	854	869	30	10	1557	1544	8	6	3588	3611	7	8	2104	2140	9	2	665	690	21
10	898	914	14	18	546	550	31	9	2439	2462	8	7	287	293	24	9	883	892	14	3	794	803	18
11	1121	1119	13	11	1425	1399	8	11	1425	1399	8	10	1839	1811	7	8	1020	1015	13	4	1308	1324	14
12	1265	1284	13	13	1224	1223	20	13	1224	1223	11	10	1439	1451	9	11	1020	1015	13	5	379	405	34
				14	2021	2009	43	14	2021	2009	10	11	945	943	12	12	343	317	32	6	297	262	41



COLUMNS ARE 10FD,10FC, 10SIG

L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG
0	442	421	24	1	3665	3770	4	0	2390	2337	6	0	1993	1974	8
8	894	894	15	3	2593	2537	5	1	233	170	29	17	553	531	17
10	1059	1076	14	5	4544	4561	5	2	2599	2518	7	2	719	715	14
11	694	734	19	7	1337	1326	8	3	1341	1314	8	3	853	844	13
12	805	812	17	9	1567	1558	8	4	592	576	13	4	945	928	11
13	701	694	20	11	362	362	22	5	1011	1006	9	5	1287	1254	10
17	353	363	40	15	1663	1725	11	7	309	296	22	7	311	329	29
18	533	564	29	17	941	964	15	8	557	553	14	8	645	634	16
19	550	562	30	19	1841	1843	13	9	1748	1727	9	9	1443	1427	11
20	564	559	30	21	299	215	47	10	2918	2942	8	10	633	606	17
22	744	685	24	23	817	792	22	11	821	849	13	11	747	782	16
0	1555	1511	10	0	444	422	9	14	1984	2014	11	12	1218	1211	13
1	906	891	15	0	3211	3259	4	15	931	974	16	13	1003	1016	15
2	1184	1156	9	2	1577	1551	5	24	585	596	23	14	699	702	20
4	1910	1871	8	3	238	253	22	0	386	470	48	15	727	759	20
5	1335	1323	9	3	719	719	9	0	1143	1100	8	18	848	846	21
6	1485	1487	9	3	927	923	9	1	3389	3404	7	19	497	547	33
7	1282	1264	10	6	1048	1036	11	2	1928	1894	8	18	361	371	30
8	1745	1723	13	7	894	912	10	4	324	346	22	1	349	368	26
10	1102	1116	13	8	703	728	11	3	1303	1265	9	2	642	655	17
11	1169	1164	13	8	894	912	10	4	1422	1414	8	3	2942	2972	13
14	816	805	17	9	1902	1956	8	5	2514	2523	7	2	642	655	17
16	753	763	20	10	2035	2052	9	6	1422	1414	8	3	2942	2972	13
17	828	793	20	11	1676	1680	10	8	885	895	11	6	370	336	25
19	509	493	30	12	1456	1481	11	9	577	554	16	7	1635	1662	10
20	686	640	24	13	1344	1341	12	10	796	796	13	8	275	305	36
0	247	268	32	14	520	648	26	11	779	793	14	9	826	850	15
2	2843	2850	8	15	523	628	36	12	1538	1504	11	10	1206	1203	12
3	812	801	12	16	1747	1737	5	13	929	949	14	11	271	142	38
4	303	292	26	17	184	195	30	16	635	682	21	12	290	280	39
5	1492	1490	9	18	453	4715	5	17	725	700	19	13	309	354	41
6	779	781	13	19	1038	1023	8	17	641	663	23	13	944	925	17
7	979	984	12	20	1059	1086	18	15	1059	1086	18	15	339	503	43
9	1447	1469	11	21	1267	1270	8	23	475	527	38	17	1058	1063	17
10	1151	1139	12	22	3820	3802	6	0	1267	1273	8	19	416	428	37
12	1068	1095	14	23	224	221	29	1	2571	2563	8	21	802	801	24
13	461	406	26	24	361	356	19	2	1058	1070	10	0	1	87	87
14	1088	1083	15	25	1779	1771	8	3	460	456	17	1	673	695	17
16	881	920	18	26	1068	1085	10	4	2115	2104	8	3	462	475	23
17	324	347	42	27	1146	1085	18	6	595	591	14	3	462	475	23
18	463	458	31	28	1267	1318	32	7	1113	1110	11	5	419	441	24
21	598	642	29	29	1057	1054	12	8	1270	1291	10	6	848	882	15
0	675	681	16	30	1344	1361	12	9	1552	1536	10	7	1302	1320	12
1	758	751	13	31	495	495	34	0	0	0	0	0	0	0	0
2	988	1035	12	32	1779	1771	8	3	460	456	17	1	673	695	17
3	1021	1019	11	33	1068	1085	10	4	2115	2104	8	3	462	475	23
5	516	516	19	34	1146	1085	18	6	595	591	14	3	462	475	23
6	1846	1838	10	35	1267	1318	32	7	1113	1110	11	5	419	441	24
7	1333	1348	11	36	1057	1054	12	8	1270	1291	10	6	848	882	15

COLUMNS ARE 10FD,10FC, 10SIG

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L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	
0	402	431	37	10	422	427	21	10	448	442	18	18	503	471	32	9	741	718	23	
1	432	413	39	11	2688	2720	19	19	2738	2731	9	9	1112	1127	18	4	2011	2013	8	
2	535	465	33	13	887	883	14	21	350	351	24	12	350	346	44	5	608	593	15	
3	738	684	25	14	716	728	19	5	2386	2371	9	14	403	451	43	6	486	482	17	
4	1044	1035	19	15	454	447	26	6	1666	1620	10	10	2, 12, L	2, 12, L	16	7	942	985	12	
5	399	368	19	14	714	715	23	6	1666	1620	10	10	2268	2207	16	8	1434	1448	10	
7	662	657	26	21	764	767	23	8	343	313	26	2	1983	1994	11	9	387	402	24	
8	744	720	24	23	531	583	34	8	388	390	24	4	1734	1764	12	4	1284	1262	11	
0	828	839	22	0	556	539	12	11	2244	2235	10	3	516	536	24	6	800	824	10	
5	592	567	31	1	363	358	19	13	1173	1193	13	5	1257	1221	13	10	311	255	23	
8	769	745	25	2	723	707	12	15	285	372	47	8	581	603	22	2	906	886	23	
0	1358	1330	6	3	2344	2304	8	16	1595	1551	13	7	359	303	32	0	384	409	40	
2	321	328	20	5	1246	1236	6	11	858	880	18	5	858	880	18	5	690	660	27	
4	1286	1254	8	6	941	933	11	20	712	704	25	12	826	815	19	6	642	632	29	
6	2367	2534	7	7	491	487	16	21	625	654	28	14	338	421	43	9	704	708	27	
8	2815	2821	8	8	1581	1550	9	15	2, 6, L	2, 6, L	11	15	1051	1050	17	10	634	557	29	
10	2387	2379	8	9	809	845	12	0	994	1002	11	11	445	501	34	11	410	367	45	
12	2462	2450	9	10	1180	1205	11	2	1668	1683	10	17	667	649	26	1	1054	1026	21	
14	321	342	31	11	776	778	14	3	1488	1461	11	18	863	891	22	5	630	637	33	
20	1136	1158	17	12	347	348	28	4	1368	1389	11	1	389	383	29	1	3720	3734	6	
22	310	53	45	13	1140	1141	13	5	1524	1554	11	2	458	481	28	3	909	905	11	
24	1008	1007	21	14	708	723	18	6	270	258	10	10	1628	1644	13	5	1247	1240	9	
0	2141	2067	5	15	1491	1567	13	7	1253	1285	12	5	394	336	33	11	914	891	13	
1	761	733	9	16	848	854	18	8	378	386	27	8	414	436	32	13	2657	2710	10	
2	3000	2958	6	17	953	963	17	9	760	732	16	8	1950	1956	12	17	1112	1159	16	
3	3046	3023	6	19	788	776	21	10	843	834	16	10	407	389	35	3	1, L	1, L	17	
4	2097	2070	7	20	376	439	42	11	695	681	19	10	582	604	26	2	772	763	11	
6	1084	1104	9	22	503	527	34	13	1256	1208	14	13	868	913	20	3	1557	1547	8	
7	890	909	10	2	785	776	10	14	1161	1115	15	14	314	319	47	4	1522	1543	9	
8	440	415	17	0	785	776	10	17	583	588	26	14	2, 10, L	2, 10, L	19	5	3271	3216	8	
9	383	396	20	1	350	355	20	18	700	671	24	0	713	703	19	5	1266	1265	10	
10	220	192	36	4	1384	1340	9	19	646	667	27	1	2334	2353	12	6	1266	1265	10	
11	950	983	12	5	1390	1374	9	20	319	403	53	2	653	677	24	7	719	709	13	
12	1385	1391	11	6	1698	1695	9	2	987	975	12	2	637	664	24	8	1729	1732	9	
14	250	272	41	7	1094	1105	11	0	2, 7, L	2, 7, L	12	4	1762	1782	14	9	2307	2299	9	
15	1281	1318	13	8	2737	2704	9	1	400	427	24	5	340	420	43	10	905	956	13	
16	1738	1740	12	9	600	598	16	2	311	339	32	6	387	370	37	11	637	644	17	
17	722	763	21	10	474	486	21	3	2737	2733	10	7	387	370	37	12	314	329	17	
19	604	594	26	11	1070	1105	13	4	784	755	16	10	296	321	48	14	314	329	17	
20	839	815	20	12	1724	1742	11	5	250	270	39	13	1015	994	19	16	501	477	26	
21	421	473	39	13	518	534	23	6	1699	1718	11	14	520	502	31	17	860	852	19	
0	1021	1015	8	14	316	359	37	7	780	784	16	15	436	472	38	18	519	490	28	
1	496	481	14	15	1029	999	15	8	1458	1473	12	17	806	828	25	20	349	418	46	
2	2888	2831	7	16	312	105	40	10	1989	2039	11	2	579	567	46	20	991	960	20	
3	348	284	22	17	561	567	26	11	546	578	24	0	322	45	24	22	652	666	29	
4	1027	1041	10	18	468	483	32	12	457	475	28	4	757	792	25	23	324	320	53	
5	1053	1035	10	20	787	765	22	13	420	398	32	5	681	667	25	23	324	320	53	
7	3440	3457	17	21	444	496	38	14	756	748	20	6	1129	1085	18	0	4198	4176	6	
8	664	656	13	0	1946	1916	8	15	839	834	33	8	651	640	25	2	831	820	11	

COLUMNS ARE 10FD,10FC, 10SIG

COLUMNS ARE 10FO,10FC, 10SIG

COLUMNS ARE 10FO,10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	
20	529	574	31	4	993	987	16	9	507	507	38	10	335	35	4	320	288	51	47	0	389	39	9	4	389	39	31	
0	640	635	14	5	696	689	20	3	14	635	32	2	617	635	3	433	415	19	10	1070	1077	15	3	1645	1675	13		
1	1410	1405	10	6	1207	1189	14	2	617	635	32	4	1481	1460	3	4	674	677	15	4	451	428	31	4	1085	1094	17	
2	1679	1716	10	7	1763	1770	13	3	336	197	53	5	438	435	5	6	603	574	20	5	1595	1624	11	5	1085	1094	17	
3	1831	1809	14	8	509	511	27	6	793	754	28	6	735	729	7	2059	2040	10	12	958	948	19	7	502	446	32		
4	1146	1151	12	10	932	951	18	0	1050	1031	9	8	678	686	8	359	371	29	17	921	917	23	9	502	446	32		
5	2547	2560	10	11	650	669	24	2	4378	4387	7	9	1162	1173	9	698	661	17	13	562	531	29	10	562	531	29		
6	904	925	14	13	512	517	30	2	4378	4387	7	11	1202	1242	10	526	551	23	13	13	551	540	30	13	551	540	30	
7	1829	1799	11	14	480	492	33	6	3973	3990	9	10	2137	2098	11	603	574	20	15	15	358	287	47	15	358	287	47	
8	1650	1654	12	16	513	505	25	11	513	505	21	11	806	798	17	1464	1445	13	17	17	921	917	23	17	921	917	23	
9	811	834	17	17	812	835	25	14	1563	1575	13	12	869	867	18	1581	1609	13	18	18	4	10	4	10	4	10	4	10
10	673	676	22	19	3	9	9	16	445	444	32	13	842	833	19	668	673	21	19	19	562	535	4	19	562	535	4	
11	858	861	20	0	516	517	22	18	1586	1581	15	15	16	537	587	29	350	328	37	21	21	667	686	23	21	667	686	23
12	995	1007	19	2	1845	1858	12	20	895	869	21	16	302	276	47	11	763	744	20	2	2	667	686	23	2	667	686	23
13	358	400	46	3	497	509	27	22	551	486	32	18	485	437	37	12	1220	1232	16	7	7	300	286	45	7	300	286	45
14	744	730	26	4	280	258	46	22	4	1	4	22	485	437	37	13	908	877	19	13	13	300	286	45	13	300	286	45
15	1709	1671	9	5	643	627	23	0	1135	1140	9	0	1048	1056	10	1048	1056	10	13	13	562	535	4	13	562	535	4	
16	651	638	15	8	678	679	24	9	1270	1280	9	1	631	647	14	631	647	14	15	15	667	686	23	15	667	686	23	
17	1583	1615	10	9	351	413	44	2	526	553	15	2	2720	2728	9	1088	1070	13	16	16	667	686	23	16	667	686	23	
18	1820	1770	11	10	847	876	21	3	677	681	13	3	1106	1125	12	8	495	482	33	17	17	300	286	45	17	300	286	45
19	357	392	30	11	489	421	30	5	739	726	13	5	1448	1443	11	1088	1070	13	18	18	300	286	45	18	300	286	45	
20	1159	1167	13	12	1265	1254	16	6	2078	2074	9	6	2245	2244	11	1450	1450	12	19	19	300	286	45	19	300	286	45	
21	1712	1689	11	13	1098	1085	18	7	1550	1528	11	7	1015	1018	15	1015	1018	15	20	20	300	286	45	20	300	286	45	
22	1038	1002	14	14	811	808	23	8	2477	2454	10	8	1487	1526	12	1487	1526	12	21	21	300	286	45	21	300	286	45	
23	991	1015	15	15	2	1194	1191	16	10	1487	1526	12	11	1015	1018	15	1015	1018	15	22	22	300	286	45	22	300	286	45
24	1016	990	16	16	862	867	20	13	610	576	21	13	472	461	28	13	839	827	18	23	23	300	286	45	23	300	286	45
25	971	925	16	17	968	939	19	14	1397	1353	13	13	839	827	18	14	1239	1215	15	24	24	300	286	45	24	300	286	45
26	1710	1711	21	18	1801	1839	14	15	682	639	21	14	1239	1215	15	15	614	593	25	25	25	300	286	45	25	300	286	45
27	981	995	18	18	709	708	24	20	346	324	47	15	614	593	25	16	616	639	26	26	26	300	286	45	26	300	286	45
28	630	624	26	19	987	992	19	22	515	455	35	16	616	639	26	18	928	939	20	27	27	300	286	45	27	300	286	45
29	659	665	26	20	771	803	23	23	514	526	37	18	928	939	20	20	731	696	25	28	28	300	286	45	28	300	286	45
30	315	206	52	21	3	11	11	25	4	2	4	21	343	393	53	20	731	696	25	29	29	300	286	45	29	300	286	45
31	464	439	20	22	627	619	25	0	769	785	11	21	343	393	53	21	343	393	53	30	30	300	286	45	30	300	286	45
32	1833	1826	10	23	829	838	21	1	2953	2988	18	22	747	743	14	22	747	743	14	31	31	300	286	45	31	300	286	45
33	734	696	16	24	979	969	20	2	1023	1055	11	23	959	949	12	0	959	949	12	32	32	300	286	45	32	300	286	45
34	1656	1657	11	25	355	396	47	3	2948	2932	8	24	959	949	12	1	959	949	12	33	33	300	286	45	33	300	286	45
35	1499	1493	12	26	471	454	38	4	491	468	20	25	1888	1925	10	2	939	914	12	34	34	300	286	45	34	300	286	45
36	869	861	16	27	893	857	23	5	2186	2160	10	26	1888	1925	10	3	1888	1925	10	35	35	300	286	45	35	300	286	45
37	392	355	41	28	3	12	12	38	246	348	40	27	1954	1935	11	4	1954	1935	11	36	36	300	286	45	36	300	286	45
38	1025	974	15	29	392	405	38	9	534	539	21	28	1954	1935	11	5	581	541	24	37	37	300	286	45	37	300	286	45
39	657	618	21	30	398	369	39	10	388	398	27	29	1245	1252	13	6	391	410	34	38	38	300	286	45	38	300	286	45
40	1856	1888	12	31	1202	1188	19	11	285	200	38	30	1245	1252	13	7	1245	1252	13	39	39	300	286	45	39	300	286	45
41	304	245	44	32	642	621	29	12	620	653	21	31	1245	1252	13	8	303	302	37	40	40	300	286	45	40	300	286	45
42	1429	1442	16	33	1343	1306	18	13	780	768	18	32	1245	1252	13	9	819	816	17	41	41	300	286	45	41	300	286	45
43	513	558	32	34	582	573	32	14	1507	1505	14	33	1245	1252	13	10	1811	1755	12	42	42	300	286	45	42	300	286	45
44	988	999	13	35	3	13	13	15	555	582	28	34	1245	1252	13	11	819	816	17	43	43	300	286	45	43	300	286	45
45	1099	1116	14	36	552	576	31	16	1453	1401	16	35	1245	1252	13	12	621	631	25	44	44	300	286	45	44	300	286	45
46	407	372	27	37	691	708	27	17	1717	1718	12	36	1245	1252	13	13	621	631	25	45	45	300	286	45	45	300	286	45
47	1849	1847	8	38	370	436	49	18	717	718	12	37	1245	1252	13	14	1204	1218	16	46	46	300	286	45	46	300	286	45
48	1849	1847	8	39	1849	1847	8	19	1849	1847	8	38	1245	1252	13	15	1043	1054	15	47	47	300	286	45	47	300	286	45
49	1849	1847	8	40	1849	1847	8	20	1849	1847	8	39	1245	1252	13	16	1043	1054	15	48	48	300	286	45	48	300	286	45

COLUMNS ARE 10FO,10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG
0	1539	1547	9	3	1432	1471	11	13	547	520	31	17	858	823	23
2	1444	1432	10	4	1177	1197	12	16	572	586	32	18	331	326	53
3	3001	3022	9	7	2133	2134	11	0	535	479	39	19	505	471	37
4	2316	2275	9	8	912	899	16	0	2973	3034	9	0	1782	1806	11
5	626	611	17	9	970	951	16	2	757	767	15	1	387	410	27
7	324	335	31	10	682	718	22	3	440	402	30	4	1185	1192	13
9	488	464	23	11	870	871	18	8	914	908	17	3	1460	1504	12
10	580	545	21	12	498	487	28	5	656	668	23	10	1720	1700	13
11	978	954	16	13	645	656	24	7	446	468	33	12	1178	1176	16
12	1195	1248	15	14	820	854	30	8	490	506	31	16	743	745	24
13	1144	1163	16	18	589	584	30	9	588	572	27	18	524	527	34
15	1019	1025	18	19	623	617	30	10	647	633	26	18	296	296	34
16	1098	1072	18	21	691	655	29	13	1128	1141	19	11	296	309	48
17	759	745	23	3	591	574	23	14	684	683	27	0	962	974	12
18	356	436	46	3	1445	1481	12	15	660	707	29	2	1020	1020	12
19	524	530	33	4	921	895	15	17	484	467	38	2	952	977	13
20	351	288	48	5	497	507	24	5	9	9	9	3	792	793	15
0	446	449	19	6	1778	1748	12	0	288	328	44	4	2000	2032	11
1	861	864	12	7	547	541	25	2	538	591	28	5	1919	1946	11
2	708	674	14	8	875	892	18	3	532	519	28	6	1485	1530	12
3	691	720	15	9	459	466	30	4	613	612	26	8	1192	1165	14
4	2084	2046	10	10	1365	1363	15	5	544	517	29	9	1018	969	16
5	1000	995	13	11	708	708	22	6	1396	1361	14	10	689	684	21
6	1433	1487	12	13	597	592	27	8	338	612	27	10	365	339	40
7	883	899	15	14	326	378	50	9	357	300	45	17	801	782	24
8	2784	2854	10	17	706	675	26	11	1096	1268	18	18	1105	1106	25
10	607	850	21	18	558	491	32	12	340	368	48	19	388	357	45
11	795	801	19	19	392	461	46	16	364	58	46	0	716	722	15
12	1326	1398	15	0	1227	1262	12	0	995	1006	18	2	2406	2455	10
15	489	497	32	1	1795	1765	11	1	788	809	21	6	746	745	15
17	556	578	30	2	1485	1484	12	2	1255	1269	17	9	1799	1820	13
20	1132	1112	20	5	925	926	16	4	664	595	25	10	448	471	31
0	2521	2498	9	6	1008	981	16	5	463	510	36	13	1764	1725	14
1	262	262	31	7	690	658	21	6	564	554	31	15	355	420	46
2	2308	2200	9	8	1013	1010	16	10	584	619	32	17	688	686	26
3	1030	1065	12	11	709	671	24	11	458	501	40	19	359	350	47
4	1449	1436	11	12	718	733	24	12	403	360	44	0	915	892	13
5	1505	1469	11	13	436	458	37	14	1021	978	23	0	1868	1878	10
6	1518	1496	11	14	837	833	22	4	343	250	49	2	296	310	34
7	278	252	38	15	640	635	28	4	1091	1067	20	3	459	461	23
9	1043	1040	15	17	366	410	49	5	503	527	37	6	1248	1266	13
11	1013	1012	16	18	404	373	43	7	547	463	33	5	2526	2561	11
12	1321	1299	15	0	957	981	15	9	1102	1067	21	6	673	667	20
15	430	400	36	2	543	573	23	10	491	422	39	7	400	411	31
16	1670	1627	14	3	660	664	21	11	455	431	43	8	356	338	36
17	822	817	22	4	660	664	21	11	455	431	43	8	356	338	36
20	604	618	32	7	1539	1489	13	1	1263	1283	18	10	738	719	21
0	1015	986	12	9	347	350	40	5	661	687	30	12	599	620	27
1	783	781	14	11	518	531	32	5	13	13	13	15	447	422	36
2	1089	1089	12	12	877	886	21	4	374	343	50	16	645	669	28

COLUMNS ARE 10FO,10FC, 10SIG

COLUMNS ARE 10FO,10FC, 10SIG



COLUMNS ARE 10FD, 10FC, 10SIG

L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG
10	458	412	40	7	961	990	22	4	742	719	27
12	534	538	36	10	349	330	52	3	964	978	22
	10,	1,	L	11	613	645	34	5	689	727	27
5	460	511	36	10,	4,	L	19	7	821	799	25
6	936	938	22	2	966	991	19	0	582	569	30
7	756	767	25	4	347	422	48	3	576	621	32
8	519	472	34	5	592	557	31	0	11,	2,	L
10	885	889	24	8	369	366	48	0	322	351	51
11	739	735	28	9	566	532	33	1	331	323	48
	10,	2,	L	10,	5,	L	24	4	1086	1046	20
1	988	1016	19	1	720	712	24	6	350	192	49
3	1055	1120	19	3	466	469	35	0	886	934	22
5	475	462	34	5	551	546	33	1	364	343	45
6	346	286	47	6	668	701	29	2	421	460	42
7	430	455	40	6	10,	6,	L	5	498	464	36
8	503	437	35	1	374	415	45	0	11,	4,	L
11	519	527	37	2	533	503	32	2	357	361	46
	10,	3,	L	3	630	629	29	0	713	710	28
0	783	766	22	6	347	347	52	4	358	426	51
1	308	298	49	7	389	402	48				
3	578	601	29	10,	7,	L					
4	522	510	32	0	922	896	21				



OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR  
 $\text{CpRe}(\text{CO})_2(\text{N}_2\text{Ph}_2)$



COLUMNS ARE 10FD, 10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG				
1	722	779	4	4	506	503	9	9	213	203	9	5	5	1010	984	7			
2	1091	1132	4	4	261	293	13	5	6	277	259	8	4	521	525	7			
4	434	451	5	6	329	316	12	7	476	753	5	7	673	665	7				
5	307	351	8	7	436	434	11	8	839	666	7	8	511	502	9				
6	544	525	7	8	351	572	8	9	240	238	11	10	1	7	L	6			
7	699	717	7	9	3	9	11	9	325	331	13	9	825	837	6				
8	218	204	14	9	383	404	11	10	333	348	13	10	712	688	7				
9	507	520	9	10	487	474	9	11	579	571	8	1	523	512	8				
1	1156	1122	4	1	2	5	L	1	1413	1384	3	3	609	614	8				
2	262	335	7	3	925	926	5	3	571	591	4	4	441	424	L				
3	1014	998	5	4	1056	1036	5	4	705	732	5	5	1	1183	1174	4			
4	299	282	7	5	498	504	7	5	393	393	6	6	2	558	551	5			
5	685	672	6	6	531	531	7	6	794	786	6	7	4	907	894	5			
6	579	587	7	7	564	565	8	8	533	533	8	8	3	245	240	9			
7	304	203	14	8	490	499	10	9	329	364	10	9	6	762	753	6			
8	615	617	8	9	2	6	L	10	267	268	15	10	7	240	248	12			
1	452	459	6	2	925	929	6	2	1341	1370	4	2	8	468	468	12			
2	906	906	5	3	190	168	12	3	752	757	4	3	9	351	319	9			
4	958	959	6	4	4	760	754	4	4	627	618	4	4	0	1166	1182	5		
5	408	416	8	5	595	598	7	5	942	941	5	5	1	1068	1073	5			
6	393	379	8	6	284	300	12	6	226	236	11	6	2	845	842	6			
7	585	600	8	7	495	498	9	7	328	336	11	7	3	464	449	7			
9	491	493	10	8	2	7	L	8	757	768	15	8	4	464	449	7			
1	810	815	6	9	706	715	7	9	149	148	22	9	5	877	877	6			
2	186	192	11	10	3	4	8	10	341	317	12	10	6	615	591	8			
3	332	950	6	1	4	386	381	9	1	1181	1171	1	7	245	277	17			
4	281	261	9	2	5	438	440	9	1	305	298	2	8	0	5	L			
5	621	609	7	3	6	522	506	9	2	1049	1038	3	9	1	1047	1065	5		
6	447	435	8	4	8	420	420	11	3	426	421	4	1	2	470	482	7		
7	293	258	11	5	2	645	644	8	4	451	468	5	2	3	574	590	7		
8	498	513	9	6	3	168	146	17	5	801	788	6	3	4	774	772	6		
1	393	401	8	7	4	366	378	10	6	445	429	7	4	5	189	191	14		
2	712	704	6	8	5	485	481	9	7	237	238	8	6	6	646	630	7		
3	133	105	17	9	6	248	254	15	8	1	256	244	7	7	161	160	18		
4	794	798	7	10	7	477	468	11	9	1279	1267	15	8	8	416	434	10		
5	291	296	11	1	8	2	9	L	3	589	561	6	9	0	6	L			
6	339	347	10	2	1	634	650	8	4	805	791	6	10	0	964	975	6		
7	458	467	9	3	3	489	488	9	5	656	686	6	11	2	800	795	6		
1	694	682	7	4	4	263	262	13	6	262	253	11	12	3	578	582	7		
2	252	270	11	5	5	312	305	13	7	549	555	8	13	4	446	446	9		
3	626	653	18	6	6	456	452	10	8	144	139	22	14	5	622	623	8		
4	153	162	18	7	7	2	10	L	9	323	317	13	15	6	459	455	10		
5	482	496	9	8	8	501	510	10	10	1	5	L	16	7	251	263	17		
6	401	396	10	9	4	412	404	11	1	816	844	5	17	8	0	7	L		
7	193	185	17	10	5	321	319	13	2	289	306	6	18	1	944	930	7		
8	418	390	11	1	6	461	477	11	3	1036	1014	7	19	2	385	370	9		
9	315	309	11	2	7	3	433	429	4	4	590	614	8	20	3	554	546	8	
1	1	1	1	3	8	4	333	329	5	5	355	340	9	21	4	672	689	8	
2	1	1	1	4	9	1	196	216	6	6	651	655	10	22	5	485	490	10	
3	1	1	1	5	10	2	124	114	7	7	393	400	11	23	6	366	348	12	
4	1	1	1	6	11	3	1588	1585	8	8	239	223	12	24	7	0	8	L	
5	1	1	1	7	12	4	960	958	9	9	1	6	L	25	8	0	759	776	7
6	1	1	1	8	13	5	2	3	2	861	851	6	26	9	1	154	150	18	
7	1	1	1	9	14	6	1	4	3	1	1	6	27	10	2	0	0	0	0



COLUMNS ARE 10FD, 10FC, 10SIG

COLUMNS ARE 10FD, 10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG	
6	441	454	9	3	272	254	9	4	272	254	9	3	272	254	9	4	272	254	9	
4	508	496	8	4	508	496	8	2	452	436	8	2	452	436	8	1	578	572	7	
0	1103	1083	5	5	237	230	14	3	347	332	10	3	347	332	10	2	289	294	12	
1	398	381	7	6	443	450	10	4	391	396	10	4	391	396	10	0	535	526	8	
2	795	787	6	5	5	1	1	6	6	1	1	6	6	1	1	1	267	292	12	
4	522	517	7	0	713	717	6	0	578	578	7	0	578	578	7	1	321	324	11	
5	359	349	10	2	674	686	6	1	163	151	15	15	163	151	15	3	349	355	11	
6	191	175	15	3	413	424	8	2	450	447	8	8	450	447	8	2	424	428	9	
7	361	368	11	4	311	317	11	3	424	428	9	9	424	428	9	0	176	175	17	
0	477	475	6	5	463	463	9	4	190	191	16	16	190	191	16	1	406	398	9	
1	781	789	6	5	5	2	1	5	397	402	11	11	397	402	11	0	287	292	12	
2	155	174	14	0	118	130	16	16	6	2	2	16	6	2	2	1	471	484	8	
3	556	568	7	1	747	753	6	6	471	484	8	8	471	484	8	1	471	484	8	
4	328	324	10	2	346	354	8	6	377	379	9	9	377	379	9	2	377	379	9	
5	331	330	10	3	523	541	7	7	325	326	11	11	325	326	11	3	426	419	9	
6	406	394	10	4	544	543	8	4	397	392	10	10	397	392	10	4	238	233	13	
0	705	695	6	5	237	228	13	3	667	673	7	7	667	673	7	0	667	673	7	
1	249	259	10	6	420	425	10	10	426	435	9	9	426	435	9	2	402	435	9	
2	648	633	7	0	695	713	6	6	426	419	9	9	426	419	9	3	426	419	9	
3	157	164	16	2	661	671	7	7	238	233	13	13	238	233	13	4	238	233	13	
4	418	407	9	3	339	337	9	9	384	373	11	11	384	373	11	5	384	373	11	
5	423	429	9	4	384	389	9	9	6	4	4	9	6	4	4	0	474	472	8	
6	200	201	16	5	469	458	9	9	474	472	8	8	474	472	8	1	474	472	8	
0	372	367	8	6	5	4	1	7	354	317	10	10	354	317	10	2	354	317	10	
1	710	715	8	1	718	714	7	7	331	334	11	11	331	334	11	3	331	334	11	
3	556	544	8	2	194	184	14	14	413	413	10	10	413	413	10	4	413	413	10	
4	303	298	12	3	422	427	9	9	6	5	5	9	6	5	5	0	509	493	8	
5	350	348	11	4	409	398	9	9	223	229	14	14	223	229	14	2	223	229	14	
6	407	388	11	5	205	199	16	16	385	373	10	10	385	373	10	2	385	373	10	
0	683	681	7	0	632	641	7	7	377	366	10	10	377	366	10	3	377	366	10	
2	564	561	8	2	511	495	8	8	177	175	19	19	177	175	19	4	177	175	19	
3	230	212	14	3	325	326	11	11	412	414	10	10	412	414	10	1	412	414	10	
4	402	404	10	4	348	328	10	10	275	270	13	13	275	270	13	2	275	270	13	
5	352	350	12	5	454	434	10	10	315	276	12	12	315	276	12	3	315	276	12	
0	355	258	13	1	613	600	8	8	6	7	7	8	6	7	7	0	478	464	9	
1	484	503	9	2	213	187	14	14	0	478	464	9	0	478	464	9	0	478	464	9
3	413	419	10	3	397	391	10	10	2	376	350	11	2	376	350	11	2	376	350	11
4	241	267	16	4	389	370	10	10	6	8	8	11	6	8	8	1	370	364	11	
4	4	9	9	0	493	514	9	9	7	0	7	9	7	0	7	0	163	199	18	
0	430	435	10	2	435	445	10	10	0	163	199	18	0	163	199	18	0	163	199	18
1	139	106	21	3	264	235	13	13	1	461	453	9	1	461	453	9	1	461	453	9
2	445	443	11	4	297	273	13	13	2	354	342	9	2	354	342	9	2	354	342	9
0	230	214	16	0	175	170	18	18	3	170	160	18	3	170	160	18	3	170	160	18
1	399	416	11	1	453	454	10	10	0	532	533	8	0	532	533	8	0	532	533	8
0	265	262	8	3	332	320	12	12	1	236	255	13	1	236	255	13	1	236	255	13
1	741	755	6	0	461	452	10	10	2	315	318	11	2	315	318	11	2	315	318	11
0	265	262	8	0	461	452	10	10	3	329	335	11	3	329	335	11	3	329	335	11
1	741	755	6	2	385	378	12	12	7	2	2	12	7	2	2	0	461	452	10	

OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR  
cis-Cp\*Re(CO)<sub>2</sub>I<sub>2</sub>







COLUMNS ARE 10FO, 10FC, 10SIG

COLUMNS ARE 10FU, 10FC, 10SIG

5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118	119	120	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136	137	138	139	140	141	142	143	144	145	146	147	148	149	150	151	152	153	154	155	156	157	158	159	160	161	162	163	164	165	166	167	168	169	170	171	172	173	174	175	176	177	178	179	180	181	182	183	184	185	186	187	188	189	190	191	192	193	194	195	196	197	198	199	200	201	202	203	204	205	206	207	208	209	210	211	212	213	214	215	216	217	218	219	220	221	222	223	224	225	226	227	228	229	230	231	232	233	234	235	236	237	238	239	240	241	242	243	244	245	246	247	248	249	250	251	252	253	254	255	256	257	258	259	260	261	262	263	264	265	266	267	268	269	270	271	272	273	274	275	276	277	278	279	280	281	282	283	284	285	286	287	288	289	290	291	292	293	294	295	296	297	298	299	300	301	302	303	304	305	306	307	308	309	310	311	312	313	314	315	316	317	318	319	320	321	322	323	324	325	326	327	328	329	330	331	332	333	334	335	336	337	338	339	340	341	342	343	344	345	346	347	348	349	350	351	352	353	354	355	356	357	358	359	360	361	362	363	364	365	366	367	368	369	370	371	372	373	374	375	376	377	378	379	380	381	382	383	384	385	386	387	388	389	390	391	392	393	394	395	396	397	398	399	400	401	402	403	404	405	406	407	408	409	410	411	412	413	414	415	416	417	418	419	420	421	422	423	424	425	426	427	428	429	430	431	432	433	434	435	436	437	438	439	440	441	442	443	444	445	446	447	448	449	450	451	452	453	454	455	456	457	458	459	460	461	462	463	464	465	466	467	468	469	470	471	472	473	474	475	476	477	478	479	480	481	482	483	484	485	486	487	488	489	490	491	492	493	494	495	496	497	498	499	500	501	502	503	504	505	506	507	508	509	510	511	512	513	514	515	516	517	518	519	520	521	522	523	524	525	526	527	528	529	530	531	532	533	534	535	536	537	538	539	540	541	542	543	544	545	546	547	548	549	550	551	552	553	554	555	556	557	558	559	560	561	562	563	564	565	566	567	568	569	570	571	572	573	574	575	576	577	578	579	580	581	582	583	584	585	586	587	588	589	590	591	592	593	594	595	596	597	598	599	600	601	602	603	604	605	606	607	608	609	610	611	612	613	614	615	616	617	618	619	620	621	622	623	624	625	626	627	628	629	630	631	632	633	634	635	636	637	638	639	640	641	642	643	644	645	646	647	648	649	650	651	652	653	654	655	656	657	658	659	660	661	662	663	664	665	666	667	668	669	670	671	672	673	674	675	676	677	678	679	680	681	682	683	684	685	686	687	688	689	690	691	692	693	694	695	696	697	698	699	700	701	702	703	704	705	706	707	708	709	710	711	712	713	714	715	716	717	718	719	720	721	722	723	724	725	726	727	728	729	730	731	732	733	734	735	736	737	738	739	740	741	742	743	744	745	746	747	748	749	750	751	752	753	754	755	756	757	758	759	760	761	762	763	764	765	766	767	768	769	770	771	772	773	774	775	776	777	778	779	780	781	782	783	784	785	786	787	788	789	790	791	792	793	794	795	796	797	798	799	800	801	802	803	804	805	806	807	808	809	810	811	812	813	814	815	816	817	818	819	820	821	822	823	824	825	826	827	828	829	830	831	832	833	834	835	836	837	838	839	840	841	842	843	844	845	846	847	848	849	850	851	852	853	854	855	856	857	858	859	860	861	862	863	864	865	866	867	868	869	870	871	872	873	874	875	876	877	878	879	880	881	882	883	884	885	886	887	888	889	890	891	892	893	894	895	896	897	898	899	900	901	902	903	904	905	906	907	908	909	910	911	912	913	914	915	916	917	918	919	920	921	922	923	924	925	926	927	928	929	930	931	932	933	934	935	936	937	938	939	940	941	942	943	944	945	946	947	948	949	950	951	952	953	954	955	956	957	958	959	960	961	962	963	964	965	966	967	968	969	970	971	972	973	974	975	976	977	978	979	980	981	982	983	984	985	986	987	988	989	990	991	992	993	994	995	996	997	998	999	1000
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COLUMNS ARE 10FO, 10FC, 10SIG

COLUMNS ARE 10FO, 10FC, 10SIG

1	0	492	499	15	16	924	933	12	12	459	409	12	12	10	10	1210	1237	9	9	23	368	333	27	27	SIG	
2	8	435	493	17	18	250	306	28	28	1438	1421	8	8	11	11	309	277	17	17	24	734	778	18	18	FC	
3	9	668	683	14	23	713	765	17	10	912	897	9	9	15	13	455	410	14	14	25	252	156	38	38	K	
4	10	789	732	13	24	350	602	35	12	174	65	28	19	19	551	560	20	10	0	1603	1600	9	9	9	KFO	
5	11	219	145	31	27	455	455	24	13	1536	1549	10	10	1	1	421	459	19	19	0	690	718	11	11	K	
6	12	506	437	18	1	555	524	7	15	237	227	27	15	2	225	165	32	13	3	1137	1139	10	10	10	SIG	
7	13	863	817	14	0	2388	2371	5	17	1613	1532	11	11	3	314	371	25	13	4	974	971	10	10	10	FC	
8	14	697	694	16	1	3388	279	33	18	242	279	33	18	5	670	634	14	26	6	505	454	14	14	14	K	
9	16	531	510	19	2	311	264	12	20	338	303	25	25	6	260	307	30	22	8	788	792	12	12	12	FC	
10	17	550	527	18	3	1193	1210	6	21	1014	1012	14	14	7	217	218	35	23	9	560	545	14	14	14	K	
11	18	456	471	21	5	3361	3470	6	23	645	643	17	17	8	377	364	22	24	10	1057	1031	11	11	11	FC	
12	20	311	344	30	6	2067	2046	8	1	2067	2046	8	8	13	801	831	14	2	12	1279	1300	11	11	11	K	
13	21	449	438	23	9	1776	1737	7	3	1620	1685	8	8	14	350	265	24	0	13	821	827	13	13	13	FC	
14	1	614	627	15	10	883	871	9	4	315	294	16	16	14	850	785	15	7	13	821	827	13	13	13	K	
15	2	955	980	13	11	2465	2517	8	5	1734	1622	8	8	1	443	473	21	1	14	486	440	16	16	16	FC	
16	3	599	548	15	12	853	840	10	6	1033	1056	9	9	2	357	352	25	2	15	593	534	16	16	16	K	
17	5	354	355	20	13	735	727	11	7	1414	1412	9	9	1	443	473	21	3	16	899	876	14	14	14	FC	
18	6	654	682	15	14	256	296	24	8	666	647	11	11	2	357	352	25	4	17	2309	2321	20	20	20	K	
19	7	684	711	15	16	596	585	14	9	898	873	11	11	3	597	579	16	5	18	492	475	28	28	28	FC	
20	10	641	604	15	17	1550	1527	10	11	1644	1644	10	10	4	476	442	19	6	19	18	492	475	28	28	28	K
21	11	814	819	14	22	1155	1164	13	12	697	668	13	13	5	472	430	20	7	20	340	371	25	25	25	FC	
22	12	343	288	23	22	611	568	16	13	605	636	14	14	6	568	612	19	8	22	723	711	17	17	17	K	
23	14	532	516	19	23	297	280	28	14	370	330	19	19	0	3128	3003	5	10	1	733	759	13	13	13	FC	
24	15	694	672	17	25	260	356	36	15	1290	1306	11	11	1	1414	1341	6	12	2	1518	1534	10	10	10	K	
25	16	239	108	32	26	249	258	37	16	230	64	29	16	2	890	990	6	12	3	446	486	18	18	18	FC	
26	0	549	564	18	1	268	171	15	17	370	387	22	15	3	188	212	21	16	4	1520	1474	10	10	10	K	
27	1	669	689	16	2	1404	1442	6	23	536	529	19	19	5	2534	2541	6	16	5	1397	1363	11	11	11	FC	
28	3	607	640	17	3	2651	2555	6	24	251	236	35	35	6	1097	1180	7	17	6	413	418	20	20	20	K	
29	4	444	405	19	4	789	860	8	24	251	236	35	35	7	1884	1866	7	21	7	217	168	34	34	34	FC	
30	5	737	727	15	6	243	266	17	0	1096	1104	10	10	8	1477	1334	8	22	8	282	290	29	29	29	K	
31	7	603	516	17	7	2039	2003	12	0	1096	1104	10	10	9	1380	1329	8	24	9	242	235	33	33	33	FC	
32	8	379	341	23	8	418	403	12	3	1733	1712	9	9	10	317	329	9	26	10	474	545	22	22	22	K	
33	9	633	701	18	9	1411	1389	8	4	618	601	13	13	11	1323	1233	9	2	11	497	490	21	21	21	FC	
34	11	401	372	24	11	876	845	9	5	949	965	11	11	12	294	273	19	1	12	274	208	17	17	17	K	
35	12	340	341	27	12	304	228	18	7	622	665	13	13	13	906	887	11	2	13	1519	1523	8	8	8	FC	
36	15	1684	1635	4	15	1385	1395	10	8	456	406	16	16	14	1174	1176	10	3	14	0	799	872	13	13	13	K
37	16	3055	3033	4	16	670	683	14	9	1152	1139	11	11	15	554	569	15	4	15	1108	1124	12	12	12	FC	
38	2	2584	2582	5	17	307	283	22	11	473	459	16	16	16	1413	1389	11	5	16	228	212	35	35	35	K	
39	3	1182	1322	6	19	1755	1753	11	12	395	409	19	19	18	1343	1373	12	6	17	1237	1165	12	12	12	FC	
40	5	1594	1556	6	20	656	617	16	13	1242	1222	12	12	20	840	821	13	7	18	746	780	15	15	15	K	
41	6	866	901	8	21	265	273	31	15	584	570	17	17	21	1058	1073	13	8	19	887	896	16	16	16	FC	
42	7	2398	2376	6	23	764	799	19	17	611	611	16	16	24	266	306	35	9	20	482	484	20	20	20	K	
43	8	440	453	13	27	482	469	21	21	466	437	20	20	1	598	585	1	10	1	1028	1032	11	11	11	FC	
44	9	1859	1876	7	1	636	683	14	2	2369	2272	20	20	2	2369	2272	6	12	2	555	621	18	18	18	K	
45	10	1264	1372	8	0	677	731	8	3	1	636	683	14	3	2624	2647	6	13	3	1267	1301	25	25	25	FC	
46	11	1449	1445	8	1	1199	1161	7	4	249	197	18	18	4	249	197	18	14	4	771	793	15	15	15	K	
47	12	444	461	11	2	664	680	9	5	596	587	13	13	5	1601	1690	7	15	5	601	604	17	17	17	FC	
48	13	444	461	14	2	1139	1157	8	6	613	601	14	14	7	1946	1993	7	16	6	509	544	10	10	10	K	
49	14	1873	1813	9	5	1921	1883	7	7	804	804	13	13	8	158	118	26	19	7	693	737	9	9	9	FC	
50	15	1308	1289	10	6	400	385	13	10	391	399	21	21	9	2665	2649	8	20	8	1050	1055	8	8	8	K	

COLUMNS ARE 10FO, 10FC, 10SIG												COLUMNS ARE 10FU, 10FC, 10SIG											
K	KFO	FC	SIG	K	KFO	FC	SIG	K	KFO	FC	SIG	K	KFO	FC	SIG	K	KFO	FC	SIG				
5	271	316	18	12	806	865	12	8	939	908	13	0	1043	1099	2	4	5	1296	1265	12	12		
6	1586	1579	8	14	836	668	14	9	639	646	16	11	1405	1400	10	1	845	859	17	17	17		
7	411	458	14	15	1480	1434	11	11	263	290	31	20	541	541	15	2	1073	1123	19	19	19		
8	739	692	10	18	362	337	22	12	467	533	20	19	312	1229	10	5	231	147	8	8	8		
9	1567	1597	9	20	590	600	16	13	544	571	19	17	1220	1260	11	6	1398	1386	13	13	13		
10	1134	1111	10	36	3	3	36	17	774	744	17	3	456	455	17	7	352	356	10	10	10		
11	850	844	11	30	660	663	17	1	851	869	17	7	851	869	12	10	1002	1032	26	26	26		
12	977	985	11	19	427	507	22	8	281	223	24	8	1266	1307	24	11	1163	1151	12	12	12		
13	776	770	12	9	743	767	11	3	743	767	34	9	820	775	14	14	1163	1151	13	13	13		
14	477	523	17	9	233	168	34	4	233	168	34	11	558	496	16	15	488	492	14	14	14		
15	1515	1468	11	20	727	639	15	5	727	639	15	12	1013	1022	13	16	815	809	17	17	17		
16	261	262	33	2	663	681	16	6	663	681	16	13	347	341	26	17	295	321	34	34	34		
17	985	961	13	3	594	623	18	15	594	623	18	15	239	211	35	18	302	238	27	27	27		
18	264	175	31	4	497	471	19	16	497	471	19	16	780	807	15	19	297	362	29	29	29		
19	862	870	14	5	382	414	24	10	382	414	24	18	427	417	22	20	359	422	22	22	22		
20	394	404	25	6	438	504	23	11	438	504	23	20	342	301	28	21	242	262	39	39	39		
21	682	632	17	7	416	369	12	0	416	369	14	14	731	757	17	21	5	5	14	14	14		
22	536	543	20	8	1319	1308	14	3	1319	1308	14	9	4	4	4	1	449	470	17	17	17		
23	3294	3290	7	10	1620	1582	9	1	1620	1582	9	9	749	724	13	2	482	478	15	15	15		
24	182	154	23	11	442	417	14	5	442	417	14	14	236	211	30	3	484	492	17	17	17		
25	692	669	10	12	1125	1068	12	6	1125	1068	12	24	1313	1300	11	4	756	774	13	13	13		
26	2392	2350	8	13	235	256	32	7	242	298	24	24	261	235	28	5	348	399	23	23	23		
27	534	513	11	14	686	675	15	8	204	170	29	29	682	691	14	6	612	598	15	15	15		
28	1070	1139	9	16	802	758	14	9	1506	1437	10	10	627	644	14	7	240	36	18	18	18		
29	1533	1562	9	17	396	419	22	10	391	429	17	17	1346	1284	11	8	1121	1084	12	12	12		
30	705	722	11	18	397	425	22	11	240	289	32	32	937	921	13	9	781	822	15	15	15		
31	787	758	12	19	505	506	20	13	1748	1649	11	11	643	701	16	13	365	385	25	25	25		
32	1439	1475	10	2	887	912	12	15	463	441	19	19	253	217	32	14	563	563	18	18	18		
33	1154	1136	11	3	738	700	15	16	472	500	19	19	325	237	22	15	264	270	34	34	34		
34	607	629	15	4	231	167	28	17	703	717	15	15	606	599	17	16	313	345	29	29	29		
35	389	351	21	5	675	637	14	18	363	316	23	23	269	260	34	17	255	360	41	41	41		
36	1098	1111	12	6	566	572	15	20	347	353	24	24	786	813	17	18	800	806	17	17	17		
37	357	325	23	7	872	857	12	21	596	595	18	18	4	4	4	5	5	5	10	10	10		
38	235	264	32	8	675	637	14	23	359	318	27	27	282	282	26	1	376	436	21	21	21		
39	546	546	18	9	354	412	21	4	629	635	15	15	1141	1109	12	2	624	626	15	15	15		
40	509	509	21	10	785	805	13	1	1420	1396	9	9	468	499	19	3	689	700	14	14	14		
41	384	312	26	11	403	385	20	2	1315	1260	9	9	788	770	14	4	705	659	15	15	15		
42	347	312	26	12	569	547	16	3	651	685	12	12	546	558	19	5	598	551	16	16	16		
43	481	461	24	14	669	673	15	4	511	542	14	14	296	255	31	6	1357	1328	12	12	12		
44	1114	1148	9	15	587	552	17	5	427	417	15	15	242	164	37	7	734	728	15	15	15		
45	2822	2831	8	16	381	374	23	7	964	997	11	11	4	4	4	8	255	251	33	33	33		
46	573	573	12	19	842	805	16	11	1729	1752	22	22	455	444	19	9	710	691	16	16	16		
47	670	667	11	20	515	513	22	13	316	317	25	25	347	357	23	10	869	814	14	14	14		
48	954	973	10	0	939	982	13	14	764	748	15	15	737	788	15	13	414	391	22	22	22		
49	739	737	11	1	385	425	20	15	1249	1231	13	13	419	402	20	14	370	359	24	24	24		
50	416	436	13	4	1007	1002	13	16	335	383	26	26	757	780	15	15	237	192	35	35	35		
51	516	491	14	5	828	873	14	18	361	298	17	17	255	221	31	16	669	695	18	18	18		
52	1184	1211	10	6	210	184	33	19	707	733	23	23	4	4	4	0	1168	1101	12	12	12		
53	938	930	12	7	598	561	16	23	841	805	17	17	578	597	19	11	218	236	33	33	33		

K	KFO	FC	SIG	K	KFO	FC	SIG	K	KFO	FC	SIG	K	KFO	FC	SIG	K	KFO	FC	SIG				
5	656	640	17	5	656	640	17	5	656	640	17	5	656	640	17	5	656	640	17	5	656	640	17
6	457	388	780	6	457	388	780	6	457	388	780	6	457	388	780	6	457	388	780	6	457	388	780
7	786	780	16	7	786	780	16	7	786	780	16	7	786	780	16	7	786	780	16	7	786	780	16
8	248	134	36	8	248	134	36	8	248	134	36	8	248	134	36	8	248	134	36	8	248	134	36
9	429	411	24	9	429	411	24	9	429	411	24	9	429	411	24	9	429	411	24	9	429	411	24
10	877	654	16	10	877	654	16	10	877	654	16	10	877	654	16	10	877	654	16	10	877	654	16
11	253	260	34	11	253	260	34	11	253	260	34	11	253	260	34	11	253	260	34	11	253	260	34
12	246	235	35	12	246	235	35	12	246	235	35	12	246	235	35	12	246	235	35	12	246	235	35
13	494	488	19	13	494	488	19	13	494	488	19	13	494	488	19	13	494	488	19	13	494	488	19
14	245	306	36	14	245	306	36	14	245	306	36	14	245	306	36	14	245	306	36	14	245	306	36
15	303	276	28	15	303	276	28	15	303	276	28	15	303	276	28	15	303	276	28	15	303	276	28
16	374	366	25	16	374	366	25	16	374	366	25	16	374	366	25	16	374	366	25	16	374	366	25
17	802	792	15	17	802	792	15	17	802	792	15	17	802	792	15	17	802	792	15	17	802	792	15
18	735	765	16	18	735	765	16	18	735	765	16	18	735	765	16	18	735	765	16	18	735	765	16
19	429	411	24	19	429	411	24	19	429	411	24	19	429	411	24	19	429	411	24	19	429	411	24
20	533	555	19	20	533	555	19	20	533	555	19	20	533	555	19	20	533	555	19	20	533	555	19
21	673	691	18	21	673	691	18	21	673	691	18	21	673	691	18	21	673	691	18	21	673	691	18
22	7	7	7	22	7	7	7	22	7	7	7	22	7	7	7	22	7	7	7	22	7	7	7
23	605	605	17	23	605	605	17	23	605	605	17	23	605	605	17								

OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR  
trans-Cp\*Re(CO)<sub>2</sub>Br<sub>2</sub>



COLUMNS ARE 10FD,10FC, 10SIG

L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	
6	451	657	6	1	348	548	11	4	247	225	20	13	307	347	15	9	770	803	8	
8	203	183	16	2	663	548	10	5	471	457	12	14	368	376	14	11	353	370	15	
12	719	735	8	3	1406	1402	5	1	436	454	14	7	427	432	13	1	-3	9	9	
14	978	935	8	4	1103	1094	5	2	361	397	15	14	-3	4	1	1	1606	1525	6	
1	117	26	18	5	1035	1088	6	3	999	1037	5	1	-3	4	1	2	372	314	11	
2	1768	1775	4	6	575	591	7	4	185	218	28	5	1151	1162	5	5	468	398	11	
3	432	452	7	7	558	601	8	5	-3	0	1	6	1423	1488	5	7	175	171	23	
4	2400	2386	4	8	309	327	12	2	2213	2211	3	8	6	1310	1337	8	186	166	22	
5	331	318	9	9	399	415	11	4	1507	1457	4	7	1248	1298	5	9	545	524	10	
6	1038	999	5	10	347	384	13	6	1565	1593	5	8	1226	1242	5	11	588	552	11	
8	495	508	7	11	469	464	10	8	2067	2150	5	9	683	696	7	10	507	482	9	
10	903	920	8	12	340	323	14	10	1024	994	6	10	507	482	9	1	163	159	24	
11	153	156	25	13	293	289	18	12	212	184	19	14	188	141	25	3	500	538	10	
12	724	698	8	14	4	6	12	14	428	392	12	1	-3	5	1	4	276	263	15	
1	974	1004	5	1	1767	1764	5	1	-3	1	1	2	1863	1933	4	5	591	604	9	
2	2402	2418	4	2	799	814	6	1	874	919	4	2	896	950	5	6	283	309	16	
3	658	662	6	3	1043	1028	6	2	2956	2947	4	3	1275	1325	5	7	780	769	8	
4	770	723	5	4	274	270	12	3	425	469	6	4	503	502	6	8	174	221	27	
5	147	128	18	5	180	154	19	4	1169	1119	4	5	849	838	5	9	627	618	10	
6	542	592	7	6	770	723	14	6	535	538	6	6	468	490	7	6	762	807	8	
7	147	183	20	7	11	341	391	7	230	219	14	8	234	229	14	1	544	574	10	
10	158	105	24	8	12	251	283	8	576	574	7	9	926	908	6	2	544	574	10	
12	816	809	8	9	3	1024	1009	6	1443	1378	6	10	648	691	6	3	589	605	9	
13	859	809	9	10	4	344	346	6	11	233	212	11	718	768	8	4	326	341	14	
1	140	149	19	11	5	1172	1159	6	12	1017	1006	17	12	482	483	11	5	256	300	18
2	949	939	5	12	6	204	240	18	14	486	478	11	13	390	420	13	8	204	221	24
3	983	1026	5	13	7	615	638	8	1	291	305	8	1	145	6	1	4	354	348	15
4	2057	2040	5	14	8	494	537	10	2	818	876	4	2	527	508	6	5	407	429	14
5	1106	1126	5	15	10	182	226	24	3	560	601	5	3	1222	1198	5	6	417	440	14
6	1086	1104	5	16	11	555	526	10	4	1949	1957	4	4	389	388	8	8	1706	1787	4
7	363	383	9	17	1	1640	1576	6	5	1922	997	5	5	1359	1391	5	1	319	313	17
8	247	292	14	18	2	182	213	7	6	1934	1928	4	6	445	475	8	2	469	462	13
9	203	212	17	19	3	962	988	7	7	679	739	6	6	1277	1300	6	3	277	287	20
10	631	625	8	20	4	694	720	8	8	1609	1620	5	8	601	631	6	4	1670	1589	3
11	317	382	14	21	5	157	44	24	408	434	9	9	859	823	7	2	2	0	1	6
12	592	564	9	22	6	874	806	6	10	874	806	6	10	303	340	14	4	2477	2350	4
13	232	216	20	23	7	164	150	24	13	164	150	24	13	177	90	27	6	1538	1593	4
1	1411	1397	5	24	8	182	105	22	14	182	105	22	14	7	7	1	8	845	870	5
2	1679	1665	5	25	9	307	290	16	1	1857	1849	5	10	1421	1393	5	10	1421	1393	5
3	985	995	3	26	10	3	3	1	287	291	11	12	1073	1028	6	12	1073	1028	6	
4	624	633	6	27	11	1328	1388	4	3	1423	1427	5	14	947	855	7	14	947	855	7
5	221	176	13	28	12	1624	1459	4	4	159	116	18	18	-2	1	1	1	163	170	10
6	139	156	21	29	13	1006	1088	4	5	636	611	7	1	163	170	10	10	163	170	10
7	186	178	17	30	14	1048	1056	4	7	164	153	21	4	470	432	5	5	902	930	5
10	220	234	18	31	15	523	546	7	9	703	707	8	5	257	288	8	6	1312	1328	5
11	275	335	16	32	16	488	484	6	11	821	806	8	6	2685	2737	4	7	1867	1883	5
12	537	547	10	33	17	351	337	9	12	351	337	9	6	523	582	6	8	1172	1162	5
13	448	498	12	34	18	561	563	7	13	140	140	23	8	649	644	4	9	1170	1170	4
1	141	149	19	35	19	191	195	27	10	1122	1084	6	3	818	837	9	9	289	281	14
2	949	939	5	36	20	552	637	9	11	552	637	9	5	1139	1105	6	10	289	281	14
3	983	1026	5	37	21	897	888	7	12	897	888	7	7	940	948	7	12	172	128	22

COLUMNS ARE 10FD,10FC, 10SIG

L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	L	KFD	FC	SIG	
6	451	657	6	1	348	548	11	4	247	225	20	13	307	347	15	9	770	803	8	
8	203	183	16	2	663	548	10	5	471	457	12	14	368	376	14	11	353	370	15	
12	719	735	8	3	1406	1402	5	1	436	454	14	7	427	432	13	1	-3	9	9	
14	978	935	8	4	1103	1094	5	2	361	397	15	14	-3	4	1	1	1606	1525	6	
1	117	26	18	5	1035	1088	6	3	999	1037	5	1	-3	4	1	2	372	314	11	
2	1768	1775	4	6	575	591	7	4	185	218	28	5	1151	1162	5	5	468	398	11	
3	432	452	7	7	558	601	8	5	-3	0	1	6	1423	1488	5	7	175	171	23	
4	2400	2386	4	8	309	327	12	2	2213	2211	3	8	6	1310	1337	8	186	166	22	
5	331	318	9	9	399	415	11	4	1507	1457	4	7	1248	1298	5	9	545	524	10	
6	1038	999	5	10	347	384	13	6	1565	1593	5	8	1226	1242	5	11	588	552	11	
8	495	508	7	11	469	464	10	8	2067	2150	5	9	683	696	7	10	507	482	9	
10	903	920	8	12	340	323	14	10	1024	994	6	10	507	482	9	1	163	159	24	
11	153	156	25	13	293	289	18	12	212	184	19	14	188	141	25	3	500	538	10	
12	724	698	8	14	4	6	12	14	428	392	12	1	-3	5	1	4	276	263	15	
1	974	1004	5	1	1767	1764	5	1	-3	1	1	2	1863	1933	4	5	591	604	9	
2	2402	2418	4	2	799	814	6	1	874	919	4	2	896	950	5	6	283	309	16	
3	658	662	6	3	1043	1028	6	2	2956	2947	4	3	1275	1325	5	7	780	769	8	
4	770	723	5	4	274	270	12	3	425	469	6	4	503	502	6	8	174	221	27	
5	147	128	18	5	180	154	19	4	1169	1119	4	5	849	838	5	9	627	618	10	
6	542	592	7	6	770	723	14	6	535	538	6	6	468	490	7	6	762	807	8	
7	147	183	20	7	11	341	391	14	230	219	14	8	234	229	14	1	544	574	10	
10	158	105	24	8	12	251	283	8	576	574	7	9	926	908	6	2	544	574	10	
12	816	809	8	9	3	1024	1009	6	1443	1378	6	10	648	691	6	3	589	605	9	
13	859	809	9	10	4	344	346	6	11	233	212	17	11	718	768	8	4	326	341	14
1	140	149	19	11	5	1172	1159	6	12	1017	1006	17	12	482	483	11	5	256	300	18
2	949	939	5	12	6	204	240	18	14	486	478	11	13	390	420	13	8	204	221	24
3	983	1026	5	13	7	615	638	8	1	291	305	8	1	145	6	1	4	354		





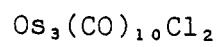




COLUMNS ARE 10FO, 10FC, 10SIG

L	KFO	FC	SIG	L	KFO	FC	SIG	L	KFO	FC	SIG
0	248	226	16	4	365	341	13	1	749	712	9
1	635	624	8	6	195	116	24	0	814	782	8
2	563	562	9	0	950	923	7	0	602	557	11
3	572	611	9	1	407	428	12	2	322	282	18
4	548	591	9	2	1296	1325	7	4	302	308	17
5	521	541	10	3	376	396	13	0	488	496	12
6	303	321	16	4	707	737	9	4	934	935	9
7	428	427	12	5	252	279	19	2	621	638	10
0	551	522	10	7	177	201	29	0	272	275	19
1	761	756	8	0	1183	1148	7	1	581	561	11
2	398	348	11	1	419	449	12	2	261	209	20
3	206	230	21	2	192	239	23	3	218	256	26
7	257	293	20	4	198	204	24	4	239	230	19
1	586	603	9	7	7	4	9	0	418	396	14
2	277	267	16	0	735	741	9	0	432	430	13
3	703	712	9	1	785	775	8	2	415	445	14
4	304	262	15	2	849	855	8	3	765	780	10
5	521	523	11	3	640	635	9	4	518	511	11
6	6	8	11	4	432	474	12	0	432	430	13
1	643	641	9	5	394	427	13	1	333	291	17
3	190	171	25	6	307	349	17	2	216	220	24
6	540	515	11	0	770	736	8	0	299	286	18
3	622	577	10	1	643	650	9	2	218	225	24
6	10	1	12	0	360	334	15	3	197	188	29
1	477	459	12	0	890	863	8	0	572	559	11
7	0	0	7	1	407	381	13	1	594	552	11
0	1245	1165	7	2	727	706	9	2	0	0	0
2	1507	1457	6	3	193	199	25	0	0	0	0
4	782	775	8	4	7	7	10	0	0	0	0
6	563	599	10	0	307	270	16	0	0	0	0
7	1	1	6	1	608	602	10	0	0	0	0
0	1538	1532	6	1	0	0	6	0	0	0	0

OBSERVED AND CALCULATED  
STRUCTURE FACTOR LISTINGS FOR





COLUMNS ARE 10FO,10FC, 10S1G

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG
6	4989	4699	42	4	5520	5589	70	5	1455	1418	93	11	1342	1099	143
8	6611	6200	44	6	5057	4880	67	8	3399	3250	73	12	2108	2062	109
9	2021	1916	50	8	3135	3091	70	6	3626	3526	70	14	4399	4181	92
10	5989	5847	46	10	2926	2934	55	9	1784	1846	89	15	H, 22,	1	
11	1429	1467	60	14	3064	2958	66	10	839	823	124	0	4602	4827	78
13	1297	1220	75	17	1102	965	112	12	2259	2178	99	3	1585	1493	101
14	9454	8894	54	18	2153	2158	78	13	3400	3521	76	5	1467	1521	109
15	1536	1439	77	20	1896	1978	88	14	1809	1675	103	7	2030	1895	103
18	6023	5942	62	H, 10,	1		63	16	3611	3450	78	9	1217	1102	155
20	2364	2287	75	2	3487	3534	63	18	4036	3889	81	10	3569	3308	94
21	1641	1695	95	4	2449	2405	63	18	1774	1672	112	12	1716	1956	144
24	3878	3791	87	7	1656	1471	76	H, 17,	1			2	1631	1549	108
25	1847	1606	134	8	4517	4304	63	1	1589	1625	92	2	4250	4410	67
26	1102	850	172	8	3102	2891	65	2	4250	4410	67	2	1832	1585	102
H, 7,	1			10	2544	2400	65	3	1540	1624	97	3	1677	1445	112
1	863	888	71	11	970	980	111	4	4445	4273	65	8	1895	1643	116
2	2674	3041	46	12	807	816	132	6	5231	5154	71	0	H, 24,	1	
3	1060	1185	70	13	1925	1775	81	8	4996	4820	74	0	1586	1616	122
4	14026	14569	43	14	4179	4038	70	9	2003	1801	89	2	1994	1848	97
5	2788	2782	49	17	1398	1467	107	12	1291	1360	128	3	1331	1142	120
6	7045	6845	44	20	2373	2422	93	13	1456	1411	126	0	H, 0,	2	
8	1418	1321	57	21	1596	1811	117	14	1619	1374	105	0	16186	18953	29
9	2688	2911	51	22	1916	2059	108	15	2059	1974	97	1	5060	4889	29
11	961	903	78	H, 14,	1		70	17	1310	1311	136	2	7890	8051	29
12	2717	2856	55	0	2094	2223	81	18	1835	1664	109	3	3662	3629	36
13	4831	4165	56	1	1724	1630	72	1	1701	1613	84	5	4968	1755	43
14	10491	9815	56	2	5424	5686	58	2	5465	5793	67	7	2286	2216	42
15	3332	3140	64	3	3309	3315	64	3	880	1088	143	8	6323	6109	40
17	2818	2676	67	4	4810	4579	64	5	2821	2723	80	9	2990	2876	45
18	1728	1653	81	6	6593	6372	63	8	2740	2563	84	10	13613	13315	43
19	1372	1542	99	8	1402	1410	67	8	2463	2372	91	12	3352	3291	49
20	1058	986	112	10	1620	1555	85	9	5998	5958	78	14	5990	5721	52
21	1451	1662	105	11	3187	3156	72	12	966	801	165	15	2903	2728	57
22	2097	2282	98	13	1500	1413	99	14	2768	2769	95	16	1287	1157	81
24	3928	4005	85	16	2880	2989	80	16	H, 19,	1		17	4503	4259	61
26	2216	2037	102	18	3349	3229	85	2	1826	1923	90	18	3737	3587	63
H, 8,	1			22	1376	1213	136	4	2318	2405	83	19	3713	3517	65
0	2583	2761	47	H, 15,	1		113	6	1561	1518	113	20	3752	3937	69
1	3453	3530	52	3	979	796	89	6	1073	989	145	21	2782	2779	79
2	3372	3374	47	5	1536	1507	88	9	4650	4630	84	22	2616	2617	86
3	5078	5007	50	6	2618	2416	77	10	1875	1832	108	26	2729	2943	100
4	2301	2282	50	7	3301	2402	76	12	H, 20,	1		1	2793	2749	31
5	2350	2143	52	9	3430	3383	69	13	1365	1390	109	2	1671	1700	34
6	3297	3145	48	13	1854	1959	88	14	3169	3131	81	3	1642	1763	42
7	3009	2980	50	16	2399	2370	94	17	2416	2473	85	4	2141	2179	40
8	5034	4755	49	17	1351	1407	128	19	3545	3542	92	5	2086	2107	46
9	1828	1713	56	19	1609	1360	109	20	1651	1558	111	6	1688	1568	47
10	2543	2431	54	20	3545	3542	92	23	1460	1498	132	8	4795	4619	42
11	1281	1110	89	21	1267	1228	143	8	H, 16,	1		9	1876	1832	52
16	2605	2543	73	0	2926	2741	69	10	1301	986	137	11	3965	3880	45
18	1886	1935	107	1	3391	2372	72	11	H, 21,	1		10	1067	1069	65
22	H, 9,	1		2	4539	4445	69	4	3890	3760	78	6	4289	4498	93
1	1896	1963	58	3	2384	2458	74	6	2011	2053	103	11			

COLUMNS ARE 10FO,10FC, 10S1G

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG
6	4989	4699	42	4	5520	5589	70	5	1455	1418	93	11	1342	1099	143
8	6611	6200	44	6	5057	4880	67	8	3399	3250	73	12	2108	2062	109
9	2021	1916	50	8	3135	3091	70	6	3626	3526	70	14	4399	4181	92
10	5989	5847	46	10	2926	2934	55	9	1784	1846	89	15	H, 22,	1	
11	1429	1467	60	14	3064	2958	66	10	839	823	124	0	4602	4827	78
13	1297	1220	75	17	1102	965	112	12	2259	2178	99	3	1585	1493	101
14	9454	8894	54	18	2153	2158	78	13	3400	3521	76	5	1467	1521	109
15	1536	1439	77	20	1896	1978	88	14	1809	1675	103	7	2030	1895	103
18	6023	5942	62	H, 10,	1		63	16	3611	3450	78	9	1217	1102	155
20	2364	2287	75	2	3487	3534	63	18	4036	3889	81	10	3569	3308	94
21	1641	1695	95	4	2449	2405	63	18	1774	1672	112	12	1716	1956	144
24	3878	3791	87	7	1656	1471	76	H, 17,	1			2	1631	1549	108
25	1847	1606	134	8	4517	4304	63	1	1589	1625	92	2	4250	4410	67
26	1102	850	172	8	3102	2891	65	2	4250	4410	67	2	1832	1585	102
H, 7,	1			10	2544	2400	65	3	1540	1624	97	3	1677	1445	112
1	863	888	71	11	970	980	111	4	4445	4273	65	8	1895	1643	116
2	2674	3041	46	12	807	816	132	6	5231	5154	71	0	H, 24,	1	
3	1060	1185	70	13	1925	1775	81	8	4996	4820	74	0	1586	1616	122
4	14026	14569	43	14	4179	4038	70	9	2003	1801	89	2	1994	1848	97
5	2788	2782	49	17	1398	1467	107	12	1291	1360	128	3	1331	1142	120
6	7045	6845	44	20	2373	2422	93	13	1456	1411	126	0	H, 0,	2	
8	1418	1321	57	21	1596	1811	117	14	1619	1374	105	0	16186	18953	29
9	2688	2911	51	22	1916	2059	108	15	2059	1974	97	1	5060	4889	29
11	961	903	78	H, 14,	1		70	17	1310	1311	136	2	7890	8051	29
12	2717	2856	55	0	2094	2223	81	18	1835	1664	109	3	3662	3629	36
13	4831	4165	56	1	1724	1630	72	1	1701	1613	84	5	4968	1755	43
14	10491	9815	56	2	5424	5686	58	2	5465	5793	67	7	2286	2216	42
15	3332	3140	64	3	3309	3315	64	3	880	1088	143	8	6323	6109	40
17	2818	2676	67	4	4810	4579	64	5	2821	2723	80	9	2990	2876	45
18	1728	1653	81	6	6593	6372	63	8	2740	2563	84	10	13613	13315	43
19	1372	1542	99	8	1402	1410	67	8	2463	2372	91	12	3352	3291	49
20	1058	986	112	10	1620	1555	85	9	5998	5958	78	14	5990	5721	52
21	1451	1662	105	11	3187	3156	72	12	966	801	165	15	2903	2728	57
22	2097	2282	98	13	1500	1413	99	14	2768	2769	95	16	1287	1157	81
24	3928	4005	85	16	2880	2989	80	16	H, 19,	1		17	4503	4259	61
26	2216	2037	102	18	3349	3229	85	2	1826	1923	90	18	3737	3587	63
H, 8,	1			22	1376	1213	136	4	2318	2405	83	19	3713	3517	65
0	2583	2761	47	H, 15,	1		113	6	1561	1518	113	20	3752	3937	69
1	3453	3530	52	3	979	796	89	6	1073	989	145	21	2782	2779	79
2															

COLUMNS ARE 10FO,10FC, 10SIG

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG
1	1394	1474	48	19	1794	1735	87	19	1710	1884	105	7	1818	1842	78	9	1603	1394	105
2	5728	6237	40	20	1584	1539	97	20	2521	2663	89	7	1818	1842	78	9	1603	1394	105
3	1624	1654	53	0	5361	5468	46	21	1024	1085	154	8	6313	6083	64	11	1609	1764	110
4	3997	3984	46	1	2663	2852	51	10	4515	4465	69	10	4515	4465	69	12	1897	1834	100
5	3929	3374	47	2	4750	4907	49	11	4448	4316	73	11	4448	4316	73	18	1858	1823	115
6	5822	5630	47	4	1827	1833	58	2	3686	3807	57	14	1609	1377	92	16	1688	1610	119
7	6559	6488	48	5	5217	5187	52	3	1670	1733	66	16	1688	1610	119	17	1343	1451	155
8	2310	2399	58	6	2414	2347	59	6	3104	3115	59	18	5636	5600	80	2	2615	2668	83
9	4958	4413	51	7	2698	2765	56	7	2113	2154	70	19	1627	1817	119	4	4354	4390	70
10	5943	4664	53	8	4388	4208	55	8	3287	3282	62	19	1627	1817	119	6	3568	3536	76
11	5097	4937	56	9	4994	4798	55	8	3287	3282	62	20	2549	2462	97	8	2653	2596	83
12	1685	1491	70	10	3442	3395	60	9	1918	2009	72	21	1831	1770	123	8	2653	2596	83
13	3798	3653	62	11	3442	3395	60	11	1370	1430	93	21	1831	1770	123	12	1833	1732	113
14	4682	4573	63	12	1662	1761	78	12	1662	1761	78	14	3000	3100	94	12	1833	1732	113
15	1512	1353	83	13	5775	5506	60	12	1662	1761	78	14	3000	3100	94	14	3000	3100	94
16	4808	4847	68	14	3527	3378	65	13	4139	4024	71	16	1688	1610	119	16	1688	1610	119
17	1912	1753	100	15	6353	6137	64	16	1866	2025	97	16	1866	2025	97	17	1343	1451	155
18	2836	2934	92	16	1996	1920	81	18	1866	2025	97	18	1866	2025	97	17	1343	1451	155
19	4092	4161	41	17	3056	3051	78	19	1722	1775	100	19	1722	1775	100	4	5242	5233	61
20	831	838	44	18	6353	6137	64	19	1722	1775	100	6	2626	2689	69	4	5242	5233	61
21	2346	2359	47	19	1996	1920	81	20	1722	1775	100	6	2626	2689	69	6	2626	2689	69
22	5728	6237	40	20	2688	2614	93	21	1471	1583	130	7	1951	1962	76	7	1951	1962	76
23	1912	1753	100	21	451	4585	83	23	1359	1129	130	11	3458	3438	79	7	1951	1962	76
24	4092	4161	41	22	2385	2467	100	23	1359	1129	130	11	3458	3438	79	5	1712	1798	96
0	492	4161	41	23	451	4585	83	0	2089	2078	61	14	3388	3264	81	5	1712	1798	96
1	831	838	44	24	4719	4976	52	1	2826	2777	63	14	3388	3264	81	6	809	886	174
2	3146	3338	44	1	4719	4976	52	1	2826	2777	63	15	1776	1738	107	6	809	886	174
3	2739	2839	47	2	8063	8532	51	4	2169	2355	59	16	1872	1760	96	7	4078	3992	81
4	7149	7205	47	5	2044	1980	59	6	4633	4572	59	16	1872	1760	96	10	2762	2638	89
5	3025	3132	50	6	6163	6214	54	7	1844	1862	73	19	2439	2442	103	10	2762	2638	89
6	1139	1019	65	7	1121	1111	79	8	1625	1563	81	19	2439	2442	103	11	1706	1743	119
7	3589	3465	51	8	7110	7046	55	8	1625	1563	81	0	6912	6818	66	12	1262	1432	120
8	3658	3589	51	9	1968	2045	63	9	3246	3346	70	2	2011	2084	76	3	2380	2384	85
9	1785	1738	61	10	3448	3344	57	10	2239	2255	74	3	3653	3546	74	3	2380	2384	85
10	2615	2599	56	11	3219	3197	62	11	2239	2255	74	3	3653	3546	74	6	8429	9014	44
11	1441	1452	91	12	1996	1919	76	12	1887	2014	98	4	3966	3936	64	6	8429	9014	44
12	1753	1617	80	13	2753	2656	62	13	1887	2014	98	4	3966	3936	64	7	1471	1551	60
13	3774	3853	67	14	4335	4203	64	14	1887	2014	98	6	1626	1591	88	8	1471	1551	60
14	1441	1452	91	15	4335	4203	64	15	1887	2014	98	6	1626	1591	88	9	3705	3696	52
15	1614	1667	96	16	3824	3619	68	16	1887	2014	98	6	1626	1591	88	9	3705	3696	52
16	1614	1667	96	17	2096	2058	79	17	1887	2014	98	8	1626	1591	88	10	7312	7408	50
17	3080	3272	90	18	1696	1736	100	18	1887	2014	98	8	1626	1591	88	12	4520	4359	54
18	3774	3853	67	19	1696	1736	100	19	1887	2014	98	10	4314	4216	72	12	4520	4359	54
19	1614	1667	96	20	2096	2058	79	20	1887	2014	98	10	4314	4216	72	13	1393	1301	78
20	1614	1667	96	21	2096	2058	79	21	1887	2014	98	10	4314	4216	72	14	1545	1180	72
21	1614	1667	96	22	2096	2058	79	22	1887	2014	98	12	1589	1496	115	14	1545	1180	72
22	1614	1667	96	23	2096	2058	79	23	1887	2014	98	12	1589	1496	115	15	1900	1560	108
23	1614	1667	96	24	2096	2058	79	24	1887	2014	98	12	1589	1496	115	16	1900	1560	108
24	1614	1667	96	25	2096	2058	79	25	1887	2014	98	12	1589	1496	115	17	1545	1613	110
25	1614	1667	96	26	2096	2058	79	26	1887	2014	98	12	1589	1496	115	18	1545	1613	110
26	1614	1667	96	27	2096	2058	79	27	1887	2014	98	12	1589	1496	115	19	1545	1613	110
27	1614	1667	96	28	2096	2058	79	28	1887	2014	98	12	1589	1496	115	20	1545	1613	110
28	1614	1667	96	29	2096	2058	79	29	1887	2014	98	12	1589	1496	115	21	1545	1613	110
29	1614	1667	96	30	2096	2058	79	30	1887	2014	98	12	1589	1496	115	22	1545	1613	110
30	1614	1667	96	31	2096	2058	79	31	1887	2014	98	12	1589	1496	115	23	1545	1613	110
31	1614	1667	96	32	2096	2058	79	32	1887	2014	98	12	1589	1496	115	24	1545	1613	110
32	1614	1667	96	33	2096	2058	79	33	1887	2014	98	12	1589	1496	115	25	1545	1613	110
33	1614	1667	96	34	2096	2058	79	34	1887	2014	98	12	1589	1496	115	26	1545	1613	110
34	1614	1667	96	35	2096	2058	79	35	1887	2014	98	12	1589	1496	115	27	1545	1613	110
35	1614	1667	96	36	2096	2058	79	36	1887	2014	98	12	1589	1496	115	28	1545	1613	110
36	1614	1667	96	37	2096	2058	79	37	1887	2014	98	12	1589	1496	115	29	1545	1613	110
37	1614	1667	96	38	2096	2058	79	38	1887	2014	98	12	1589	1496	115	30	1545	1613	110
38	1614	1667	96	39	2096	2058	79	39	1887	2014	98	12	1589	1496	115	31	1545	1613	110
39	1614	1667	96	40	2096	2058	79	40	1887	2014	98	12	1589	1496	115	32	1545	1613	110
40	1614	1667	96	41	2096	2058	79	41	1887	2014	98	12	1589	1496	115	33	1545	1613	110
41	1614	1667	96	42	2096	2058	79	42	1887	2014	98	12	1589	1496	115	34	1545	1613	110
42	1614	1667	96	43	2096	2058	79	43	1887	2014	98	12	1589	1496	115	35	1545	1613	110
43	1614	1667	96	44	2096	2058	79	44	1887	2014	98	12	1589	1496	115	36	1545	1613	110
44	1614	1667	96	45	2096	2058	79	45	1887	2014	98	12	1589	1496	115	37	1545	1613	110
45	1614	1667	96	46	2096	2058	79	46	1887	2014	98	12	1589	1496	115	38	1545	1613	110
46	1614	1667	96	47	2096	2058	79	47	1887	2014	98	12	1589	1496	115	39	1545	1613	110
47	1614	1667	96	48	2096	2058	79	48	1887	2014	98	12	1589	1496	115	40	1545	1613	110
48	1614	1667	96	49	2096	2058	79	49	1887	2014	98	12	1589	1496	115	41	1545	1613	110
49	1614	1667	96	50	2096	2058	79	50	1887	2014	98	1							







COLUMNS ARE 10FD,10FC, 10SIG

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	
1	954	1110	114	3	1673	1686	64	22	2683	2650	95	2	1103	1047	82	21	2439	2375	109	
11	4008	4163	63	4	4092	4178	51	1	2106	2112	59	3	3697	3733	85	1	4036	4121	67	
13	2308	2387	76	5	1385	1501	68	2	2411	2419	56	5	1153	1292	85	4	2161	2373	80	
14	1560	1432	100	6	916	988	82	7	5162	5155	56	7	2830	2939	63	5	2132	2167	77	
15	2276	2249	78	7	5546	5797	56	8	1622	1842	75	6	1672	1842	75	6	2890	3015	70	
16	4142	4219	72	8	1702	1719	67	8	5847	6045	56	9	1221	1189	82	10	2464	2518	72	
17	1645	1786	93	9	3656	3654	56	6	3945	4035	57	11	8042	8275	62	10	3037	2986	70	
18	1650	1712	101	10	3295	3266	57	7	6423	6511	57	11	2584	2642	71	11	1668	1515	84	
19	1252	1264	128	11	3127	3111	60	8	2077	2041	68	13	2118	2292	78	15	3272	3302	76	
20	2045	2245	113	12	4839	4877	59	9	2515	2519	68	13	2118	2292	78	15	3272	3302	76	
21	1441	1734	151	13	3127	3011	67	10	1388	1336	76	14	1847	1984	85	20	2184	2197	101	
22	3617	3508	47	14	3927	3992	68	11	5024	5212	62	15	2380	2384	78	14	15	19	5	101
23	6673	6822	47	15	1728	1772	90	12	2810	2821	66	18	2602	2560	86	0	2674	2660	70	
24	2364	2128	47	16	2492	2711	88	14	5608	5641	69	19	1945	1703	96	1	3698	4019	71	
3	3501	3485	51	17	956	939	140	15	1686	1617	88	20	1043	1108	145	2	1238	1316	95	
4	1290	1311	64	18	1039	998	140	16	1782	1718	89	21	3419	3661	93	3	2305	2237	81	
5	13198	14270	51	19	1698	1698	120	17	3340	3370	77	22	3283	3413	93	3	2305	2237	81	
6	4879	4825	52	20	1591	1698	120	18	1528	1535	105	22	3283	3413	93	5	2150	2198	78	
7	4940	5053	54	21	2999	3208	94	19	2670	2624	92	1	3250	3379	64	6	1410	1606	98	
8	2230	2295	61	22	5	5	5	21	4022	4238	87	2	1251	1309	87	7	1220	1209	90	
9	5875	5677	55	23	1871	1874	59	22	1559	1557	133	3	4350	4856	66	9	2801	2897	72	
10	3370	3244	56	24	3912	3856	54	23	3067	3157	104	5	1432	1434	79	10	2677	2774	76	
11	2926	2822	61	25	2453	2499	59	24	2159	2337	118	6	1386	1533	85	13	2040	2303	88	
12	1187	1113	112	26	3969	3964	56	25	802	804	93	7	3530	3708	63	14	1348	1364	108	
13	5401	5271	68	27	2214	2214	63	0	802	804	93	8	3073	2991	71	16	1430	1401	114	
14	5596	5660	72	28	3852	3852	59	1	6640	6717	55	10	2979	2920	75	20	1681	1818	110	
15	4278	4198	78	29	3852	3852	59	4	2242	2448	61	11	4107	4140	67	7	15	5	5	94
16	1761	1840	99	30	3592	3827	58	5	1566	1535	60	12	2452	2501	74	1	3179	3345	74	
17	1554	1641	101	31	1824	1858	68	6	1331	1321	79	13	1153	843	99	3	1061	1170	139	
18	4356	4644	76	32	1503	1412	75	8	1597	1739	73	14	2343	2326	76	4	1124	1146	113	
19	3383	3435	47	33	5117	4876	65	9	2108	2093	66	17	1871	1992	93	6	2079	2239	87	
20	1163	1143	69	34	1028	1004	114	11	4934	4934	60	18	1253	1229	127	7	1745	1938	87	
21	4023	4826	49	35	1239	1283	112	13	1974	1920	76	19	2398	2349	94	8	945	1260	131	
22	2405	2526	55	36	1855	2044	97	14	1101	1031	101	20	2271	2347	97	9	1254	1314	106	
23	4713	4767	54	37	1959	1854	105	16	1232	1214	114	21	2134	2034	108	10	1563	1590	96	
24	5023	5347	56	38	990	923	182	17	1444	1521	105	22	1694	1657	131	11	2664	1722	96	
1	5509	5455	56	39	2567	2644	53	19	1915	2032	124	0	5950	6203	59	13	2664	2673	86	
2	3967	3983	74	40	6435	6435	51	21	1594	1420	124	1	2064	2220	72	15	1850	1674	105	
3	3548	3515	59	41	3548	3515	59	2	6350	6574	61	2	6350	6574	61	16	1909	2006	96	
4	3548	3515	59	42	2230	2230	59	3	1267	1327	78	3	1867	1812	79	17	3056	3232	87	
5	5562	5928	55	43	2230	2230	59	4	2236	2253	65	5	4062	4006	67	0	2394	2383	79	
6	1789	1706	66	44	1465	1698	91	5	1465	1698	91	5	4062	4006	67	2	3062	3160	78	
7	4059	3969	58	45	1826	1913	79	6	1062	980	95	6	1062	980	95	3	1351	1126	107	
8	3291	3540	59	46	1826	1913	79	7	4078	4078	65	7	4078	4078	65	5	5966	6013	69	
9	1129	1203	85	47	2189	2254	80	8	4180	4267	63	8	4180	4267	63	7	3203	3338	77	
10	7394	7446	67	48	1676	1614	88	10	4076	3960	65	8	1676	1614	88	8	1478	1539	107	
11	1947	1978	67	49	2036	2042	80	11	3747	3845	67	10	4076	3960	65	9	7003	7192	55	
12	1099	1030	96	50	1509	1461	113	12	3880	3945	67	12	3880	3945	67	6	2248	2208	59	
13	4423	4518	69	51	1809	1809	96	13	1938	1999	93	13	3880	3945	67	7	3203	3338	77	
14	2479	2649	102	52	1852	1619	124	14	1792	1826	91	14	1792	1826	91	15	2483	2453	63	
15	1317	1169	110	53	1852	1619	124	15	1852	2012	85	15	1852	2012	85	16	2479	2447	91	
16	2018	2105	87	54	1874	1666	65	16	1874	1666	65	16	1874	1666	65	17	2479	2447	91	
17	1617	1738	108	55	1874	1666	65	17	1665	1782	99	17	1665	1782	99	18	2479	2447	91	
18	1617	1738	108	56	1874	1666	65	18	1665	1782	99	18	1665	1782	99	19	2479	2447	91	
19	1617	1738	108	57	1874	1666	65	19	1665	1782	99	19	1665	1782	99	20	2479	2447	91	
20	1617	1738	108	58	1874	1666	65	20	1665	1782	99	20	1665	1782	99	21	2479	2447	91	
21	1617	1738	108	59	1874	1666	65	21	1665	1782	99	21	1665	1782	99	22	2479	2447	91	
22	1617	1738	108	60	1874	1666	65	22	1665	1782	99	22	1665	1782	99	23	2479	2447	91	
23	1617	1738	108	61	1874	1666	65	23	1665	1782	99	23	1665	1782	99	24	2479	2447	91	
24	1617	1738	108	62	1874	1666	65	24	1665	1782	99	24	1665	1782	99	25	2479	2447	91	
25	1617	1738	108	63	1874	1666	65	25	1665	1782	99	25	1665	1782	99	26	2479	2447	91	
26	1617	1738	108	64	1874	1666	65	26	1665	1782	99	26	1665	1782	99	27	2479	2447	91	
27	1617	1738	108	65	1874	1666	65	27	1665	1782	99	27	1665	1782	99	28	2479	2447	91	
28	1617	1738	108	66	1874	1666	65	28	1665	1782	99	28	1665	1782	99	29	2479	2447	91	
29	1617	1738	108	67	1874	1666	65	29	1665	1782	99	29	1665	1782	99	30	2479	2447	91	
30	1617	1738	108	68	1874	1666	65	30	1665	1782	99	30	1665	1782	99	31	2479	2447	91	
31	1617	1738	108	69	1874	1666	65	31	1665	1782	99	31	1665	1782	99	32	2479	2447	91	
32	1617	1738	108	70	1874	1666	65	32	1665	1782	99	32	1665	1782	99	33	2479	2447	91	
33	1617	1738	108	71	1874	1666	65	33	1665	1782	99	33	1665	1782	99	34	2479	2447	91	
34	1617	1738	108	72	1874	1666	65	34	1665	1782	99	34	1665	1782	99	35	2479	2447	91	
35	1617	1738	108	73	1874	1666	65	35	1665	1782	99	35	1665	1782	99	36	2479	2447	91	
36	1617	1738	108	74	1874	1666	65	36	1665	1782	99	36	1665	1782	99	37	2479	2447	91	
37	1617	1738	108	75	1874	1666	65	37	1665	1782	99	37	1665	1782	99	38	2479	2447	91	
38	1617	1738	108	76	1874	1666	65	38	1665	1782	99	38	1665							



COLUMNS ARE 10FO,10FC, 10SIG

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG
1	1918	1967	67	2	681	126	114	0	1761	1717	74	13	1055	824	126	1	1254	1326	142
2	2250	2211	63	3	4703	4770	62	1	1610	1734	81	16	1843	1898	101	2	2630	2086	100
3	1682	1642	77	4	2733	2818	65	2	1082	1003	98	0	2015	2002	83	3	2271	2232	104
4	2592	2551	63	5	1298	1220	89	3	3092	3124	74	4	2031	2045	118	4	0	0	0
5	3562	3504	85	6	976	989	109	0	3009	2981	69	0	1930	1919	93	0	5639	5904	64
6	1934	1898	66	7	3547	3563	67	5	2619	2611	76	2	1663	1607	92	1	7857	8387	46
7	3738	3683	76	8	2790	2902	73	6	897	859	126	3	1700	1617	101	2	4519	4849	62
8	1030	1009	85	9	1490	1519	98	7	4926	4949	71	5	2746	2865	82	3	2727	2580	70
9	1317	1249	115	10	1689	1760	82	8	3128	3242	70	6	1686	1723	101	4	3991	4331	63
10	2572	2545	86	11	2960	2874	72	11	1953	2100	79	7	1700	1723	101	5	2472	2444	70
11	1297	1424	125	12	1656	1611	90	10	1660	1974	91	8	1824	1245	112	6	1404	1383	90
12	1183	1409	149	13	3418	3488	83	14	2397	2493	78	11	1124	1245	112	8	4461	4735	66
13	1442	1481	125	14	1765	1860	103	17	1804	1922	97	12	1196	1215	114	10	2538	2543	81
14	1860	1666	89	15	2781	2873	104	18	4337	4400	79	14	1210	1251	120	10	9451	9240	70
15	1938	1987	65	16	1635	1634	78	19	2072	2128	96	16	2324	2222	88	11	9451	9240	70
16	1918	1967	67	17	2610	2697	63	0	2023	1094	107	17	989	1033	158	11	3860	4181	74
17	1682	1642	77	18	4392	4362	66	2	1023	1094	107	0	2388	2330	87	14	3450	3505	76
18	2592	2551	63	19	3313	3376	74	3	1860	1856	83	1	1244	1332	117	15	2803	2877	76
19	1030	1009	85	20	3078	3231	72	4	2913	2727	76	2	1655	1675	99	18	3015	3050	89
20	1317	1249	115	0	1735	1735	71	5	3088	3094	83	4	858	804	157	19	2437	2530	89
21	1442	1481	125	1	3598	3988	60	6	3825	3746	74	6	2112	2222	93	21	3133	3199	93
22	1860	1666	89	2	1574	1547	75	7	2984	2860	73	9	3295	3271	80	1	3600	3726	70
23	1938	1987	65	3	2191	2118	75	8	2080	2138	80	16	2495	2400	91	2	1266	1541	92
24	1918	1967	67	4	2069	2142	76	9	1540	1528	97	0	2388	2330	87	3	3945	3941	65
25	1682	1642	77	5	4623	4613	64	10	3443	3496	73	0	1874	1915	93	5	871	798	124
26	2592	2551	63	6	3283	3167	69	11	998	914	127	1	1874	1915	93	7	3723	3542	68
27	1030	1009	85	7	1735	1735	71	14	1279	1024	100	2	2514	2341	84	8	2243	2461	75
28	1317	1249	115	8	3078	3231	72	15	4088	4150	73	3	2731	2988	85	11	3500	3491	79
29	1442	1481	125	9	1920	1874	84	16	1280	1362	120	5	2344	2509	92	12	2074	1839	84
30	1860	1666	89	10	2290	2256	76	0	4739	4598	66	6	1135	1048	119	12	1142	1298	120
31	1938	1987	65	11	1003	1001	106	1	1794	1762	88	7	1238	1175	130	16	1142	1298	120
32	1918	1967	67	12	3328	3276	76	14	1794	1762	88	8	2005	1855	88	17	1802	1684	94
33	1682	1642	77	13	1989	2022	99	15	5308	5371	67	9	1379	1268	117	18	767	633	163
34	2592	2551	63	14	1263	1333	135	16	3904	3914	75	10	1564	1632	114	19	1588	1688	113
35	1030	1009	85	15	1827	1986	90	0	1179	1188	113	13	3224	3378	85	21	1541	1528	136
36	1317	1249	115	16	1860	1666	89	1	2084	2054	85	14	1225	1359	127	0	1942	1913	74
37	1442	1481	125	17	3189	3080	62	2	2084	2054	85	1	1586	1555	104	1	742	654	154
38	1860	1666	89	18	4722	4622	75	3	1802	1738	93	3	1586	1555	104	2	3034	2986	63
39	1938	1987	65	19	1666	1623	81	4	3104	3335	72	4	1151	1374	139	4	1524	1653	82
40	1918	1967	67	20	3320	3368	68	5	3325	3374	79	9	2051	2179	99	7	2639	2428	71
41	1682	1642	77	21	1495	1646	85	12	3423	3467	79	10	1919	1987	102	10	1621	1512	89
42	2592	2551	63	22	2014	2132	78	13	2911	2888	80	0	2561	2607	90	10	2010	2019	86
43	1030	1009	85	23	1230	1116	107	14	2370	2369	82	2	2250	2237	98	12	2786	2785	79
44	1317	1249	115	24	1349	1356	110	16	1058	1187	141	6	2433	2525	87	13	1505	1434	107
45	1442	1481	125	25	1121	1337	129	1	2085	2144	83	10	2339	2595	94	14	2148	2128	85
46	1860	1666	89	26	1973	2065	109	2	2305	2367	83	3	1863	1865	103	14	2148	2128	85
47	1938	1987	65	27	1915	2067	111	3	3205	3267	83	4	1693	1865	103	0	2485	2421	74
48	1918	1967	67	28	1153	1272	110	4	1486	1527	94	1	1863	1865	103	3	2485	2421	74
49	1682	1642	77	29	1317	1249	115	6	2737	2856	86	1	1693	1865	103	4	4638	4673	63
50	2592	2551	63	30	4870	4870	119	7	1084	1256	120	6	1713	1723	111	5	1125	1253	97
51	1030	1009	85	31	2241	2189	77	8	2075	2209	92	8	1430	1447	118	6	4691	4625	66
52	1317	1249	115	32	1183	1209	149	9	3687	3723	79	9	2835	2804	91	7	2813	2849	74
53	1442	1481	125	33	1635	1666	89	10	3687	3723	79	9	2835	2804	91	8	2033	2109	81
54	1860	1666	89	34	1860	1959	95	11	920	1021	135	0	3076	2982	95	9	2033	2109	81
55	1938	1987	65	35	1930	1021	95	11	920	1021	135	0	3076	2982	95	10	2226	2144	81

H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG	H	KFO	FC	SIG
1	1918	1967	67	2	681	126	114	0	1761	1717	74	13	1055	824	126
2	2250	2211	63	3	4703	4770	62	1	1610	1734	81	16	1843	1898	101
3	1682	1642	77	4	2733	2818	65	2	1082	1003	98	0	2015	2002	83
4	2592	2551	63	5	1298	1220	89	3	3092	3124	74	4	2031	2045	118
5	3562	3504	85	6	976	989	109	0	3009	2981	69	0	1930	1919	93
6	1934	1898	66	7	3547	3563	67	5	2619	2611	76	2	1663	1607	92
7	3738	3683	76	8	2790	2902	73	6	897	859	126	3	1700	1617	101
8	1030	1009	85	9	1490	1519	98	7	4926	4949	71	5	2746	2865	82
9	1317	1249	115	10	1689	1760	82	8	3128	3242	70	6	1686	1723	101
10	2572	2545	86	11	2960	2874	72	11	1953	2100	79	7	1700	1723	101
11	1297	1424	125	12	1656	1611	90	10	1660	1974	91	8	1824	1245	112
12	1183	1409	149	13	3418	3488	83	14	2397	2493					

## COLUMNS ARE 10FO,10FC, 10SIG

1	H	KFO	FC	SIG	H	KFO	FC	SIG	H	NFO	FC	SIG	H	NFO	FC	SIG	H	NFO	FC	SIG	H	NFO	FC	SIG	H	NFO	FC	SIG	H	NFO	FC	SIG						
18	1242	1250	127	88	6	2412	2334	88	2	2022	2033	94	12	1504	1559	120	12	1504	1559	120	11	1112	1019	144	11	976	955	158	11	976	955	158						
	H,	7,	8	99	7	1813	1910	99	4	1665	1573	104	14	1693	1691	97	15	3540	3538	81	15	3540	3538	81	H,	14,	9	152	0	970	904	152						
1	1267	1179	96	99	9	1723	1644	99	7	1720	1825	109	16	2177	2278	89	16	3949	3986	82	17	1398	1283	110	3	1666	1542	102	3	1666	1542	102						
6	1259	1571	114	105	10	1432	1314	105	8	2358	2393	91	8	2358	2393	91	0	4470	4563	69	0	4470	4563	69	H,	8,	9	98	5	2335	2134	98						
8	859	1060	141	119	11	1399	1451	119	11	1375	1487	141	11	1375	1487	141	1	858	868	144	1	858	868	144	5	1931	1852	113	5	1931	1852	113						
9	1038	1104	117	85	12	1354	1556	127	12	1354	1556	127	12	1354	1556	127	2	2716	3709	73	2	2716	3709	73	6	2479	2411	102	6	2479	2411	102						
10	1727	1900	89	65	3	1677	1740	112	3	1677	1740	112	5	3440	3223	74	3	1701	1699	92	3	1701	1699	92	7	1741	1607	110	7	1741	1607	110						
13	1494	1574	115	88	7	2154	2212	100	7	2154	2212	100	6	1617	1440	94	6	1617	1440	94	8	1455	1481	125	8	1455	1481	125	8	1455	1481	125						
14	1508	1656	95	88	8	1427	1349	122	8	1427	1349	122	8	3340	3199	73	6	1986	1854	91	6	1986	1854	91	9	1443	1605	132	9	1443	1605	132						
15	1778	1869	93	85	8	1071	1141	123	8	1071	1141	123	8	3340	3199	73	6	1986	1854	91	6	1986	1854	91	H,	15,	9	125	1	1212	1266	125						
17	1990	2047	94	85	4	2985	2975	85	4	2985	2975	85	9	2467	2521	87	8	2336	2489	94	9	2467	2521	87	9	2467	2521	87	9	2467	2521	87	9	2467	2521	87		
20	1451	1368	121	84	5	4977	5034	79	5	4977	5034	79	0	3129	3098	88	9	1448	1255	114	9	1448	1255	114	11	2111	2251	149	11	2111	2251	149						
0	2457	2450	73	93	6	3725	3819	84	6	3725	3819	84	1	3058	3004	92	11	3269	3407	92	11	3269	3407	92	H,	9,	9	106	2	2047	2053	106						
1	2317	2376	80	101	7	1917	1967	100	7	1917	1967	100	14	1501	1535	105	1	1935	1924	91	1	1935	1924	91	6	1107	960	150	6	1107	960	150						
2	2177	2154	72	101	8	1917	1967	100	8	1917	1967	100	4	1479	1439	128	18	1552	1580	109	2	1893	1996	98	8	2338	2336	99	8	2338	2336	99						
4	2736	2741	73	92	9	1784	1708	101	9	1784	1708	101	1	H,	1,	9	3405	3418	78	18	1552	1580	109	6	1810	1791	98	6	1810	1791	98	6	1810	1791	98			
5	7324	7135	68	112	10	2101	2170	92	10	2101	2170	92	1	H,	1,	9	3405	3418	78	18	1552	1580	109	6	1810	1791	98	6	1810	1791	98	6	1810	1791	98			
6	4913	4771	69	100	15	1626	1544	112	15	1626	1544	112	2	2908	2930	72	2	2908	2930	72	19	2116	2224	110	8	1496	1341	121	8	1496	1341	121	8	1496	1341	121		
7	1616	1573	95	100	16	2409	2373	92	16	2409	2373	92	2	2908	2930	72	2	2908	2930	72	19	2116	2224	110	8	1496	1341	121	8	1496	1341	121	8	1496	1341	121		
8	1553	1422	95	86	17	1914	1998	100	17	1914	1998	100	5	1034	816	119	5	1034	816	119	19	2116	2224	110	8	1496	1341	121	8	1496	1341	121	8	1496	1341	121		
9	3112	3063	83	86	2	2121	2226	86	2	2121	2226	86	6	1741	1674	81	6	1741	1674	81	2	1968	2054	79	15	1335	1372	117	2	1968	2054	79	15	1335	1372	117		
10	3276	3121	73	128	3	1116	1173	128	3	1116	1173	128	8	3954	3870	72	8	3954	3870	72	2	1968	2054	79	15	1335	1372	117	2	1968	2054	79	15	1335	1372	117		
12	1010	1146	147	87	4	2168	2178	90	4	2168	2178	90	9	3458	3410	79	9	3458	3410	79	8	3607	3537	74	3	3840	3797	77	3	3840	3797	77	3	3840	3797	77		
14	1318	1315	109	87	6	3144	3236	87	6	3144	3236	87	10	2187	2150	90	10	2187	2150	90	9	3607	3537	74	3	3840	3797	77	3	3840	3797	77	3	3840	3797	77		
15	2102	2086	88	86	7	2912	2837	86	7	2912	2837	86	14	1664	1572	98	14	1664	1572	98	9	3607	3537	74	3	3840	3797	77	3	3840	3797	77	3	3840	3797	77		
16	4093	4178	76	100	8	2107	1978	100	8	2107	1978	100	16	1948	1832	97	13	2132	2206	94	8	4244	4200	81	0	2789	2906	81	0	2789	2906	81	0	2789	2906	81		
17	1955	1973	101	98	10	1464	1568	125	10	1464	1568	125	18	2414	2487	98	18	2414	2487	98	8	4244	4200	81	0	2789	2906	81	0	2789	2906	81	0	2789	2906	81		
1	2301	2457	81	100	11	1842	1831	98	11	1842	1831	98	19	1654	1484	109	19	1654	1484	109	18	1660	1787	112	12	1454	1145	110	2	4396	4273	74	2	4396	4273	74		
3	2874	2876	75	89	13	2007	2214	100	13	2007	2214	100	0	H,	2,	9	5696	5808	67	0	H,	2,	9	5696	5808	67	3	855	587	162	3	855	587	162	3	855	587	162
7	2458	2523	82	104	16	2403	2260	89	16	2403	2260	89	1	3213	3196	76	1	3213	3196	76	0	1117	1218	111	4	2863	2937	88	12	3774	3987	89	12	3774	3987	89		
8	1422	1513	102	80	0	1902	1956	104	0	1902	1956	104	2	4878	4970	69	2	4878	4970	69	1	1279	1051	104	5	2408	2003	94	13	1368	1363	137	13	1368	1363	137		
9	3072	2904	79	80	1	3906	3860	80	1	3906	3860	80	3	3939	3835	71	3	3939	3835	71	2	3559	3546	71	7	1892	2052	121	15	2505	2428	90	15	2505	2428	90		
10	1525	1476	106	80	2	4264	4440	80	2	4264	4440	80	4	2260	2426	80	4	2260	2426	80	3	4924	4783	71	8	1288	1411	153	18	16	17	10	18	16	17	10	18	
11	1422	1622	97	123	2	4556	4523	123	2	4556	4523	123	6	2613	2669	72	6	2613	2669	72	4	2573	2538	78	9	902	849	181	1	3441	3320	90	1	3441	3320	90		
13	3660	3743	97	85	5	3325	3388	85	5	3325	3388	85	8	1214	1095	106	8	1214	1095	106	6	2457	2431	81	14	2313	2426	87	3	2819	2763	83	3	2819	2763	83		
14	2590	2764	81	98	7	1778	1660	107	7	1778	1660	107	7	4473	4260	78	7	4473	4260	78	6	2457	2431	81	14	2313	2426	87	3	2819	2763	83	3	2819	2763	83		
17	4605	4648	76	98	8	2087	1966	98	8	2087	1966	98	10	5008	4893	76	10	5008	4893	76	8	3927	4055	77	15	4008	3962	79	6	2246	2187	91	6	2246	2187	91		
19	1505	1315	133	95	9	2436	2342	95	9	2436	2342	95	12	2885	2846	89	12	2885	2846	89	8	3927	4055	77	15	4008	3962	79	6	2246	2187	91	6	2246	2187	91		
0	1570	1613	88	115	10	1366	1012	115	10	1366	1012	115	13	5441	5383	79	13	5441	5383	79	10	1504	1292	111	0	5958	5123	81	7	1271	1271	133	7	1271	1271	133		
1	2958	2834	76	88	11	2389	2349	88	11	2389	2349	88	12	2085	2026	95	12	2085	2026	95	3	2757	3082	87	9	2203	2035	92	9	2203	2035	92	9	2203	2035	92		
5	4979	4955	74	165	12	3899	3910	78	12	3899	3910	78	14	796	853	163	14	796	853	163	13	2219	2395	90	4	2939	1784	108	10	1783	1784	108	10	1783	1784	108		
8	1846	1921	96	124	14	1107	802	124	14	1107	802	124	19	2168	2012	101	19	2168	2012	101	17	1149	1179	136	5	3043	3155	138	11	2009	1921	138	11	2009	1921	138		
9	1646	1692	99	131	1	1107	802	124	1	1107	802																											

COLUMNS ARE 10FD, 10FC, 10SIG

COLUMNS ARE 10FD, 10FC, 10SIG

H	NFO	FC	SIG	H	NFO	FC	SIG	H	NFO	FC	SIG	H	NFO	FC	SIG	H	NFO	FC	SIG
13	1903	1626	114	12	2224	2172	94	11	1240	1292	154	13	3075	2875	97	1	1676	1631	119
16	880	850	151	15	1424	1283	122	13	3075	2875	97	13	3075	2875	97	5	3380	3294	99
3	1020	993	115	1	1135	1067	126	1	2743	2861	92	0	3638	3299	90	0	3638	3299	90
4	3436	3696	77	2	1368	1360	113	3	1582	1456	101	4	5562	5108	88	4	5562	5108	88
6	3437	3274	78	3	829	914	171	4	976	984	154	5	1854	1774	130	6	1725	1595	121
7	2048	1996	96	4	1804	1737	100	6	1399	1288	115	6	3254	2849	96	8	1153	1265	171
10	2070	1816	100	7	1190	1067	141	7	1963	1933	112	8	2774	2665	106	1	1240	1024	128
11	1116	770	145	9	2618	2639	97	11	2965	2694	93	10	2394	2354	112	2	2534	2480	100
12	1748	1722	111	13	1370	1398	133	11	2965	2694	93	10	2394	2354	112	4	2909	2668	96
14	2816	2716	86	13	1616	1669	112	0	2870	2812	88	1	1753	1576	126	5	1661	1602	127
15	1950	2036	96	14	2477	2323	90	8	2257	2138	100	4	3378	3258	98	8	3088	2980	96
0	2525	2672	86	2	1639	1580	104	10	2691	2452	102	10	2691	2452	102	6	2470	2366	103
1	1707	1799	99	3	2844	2794	83	11	2849	2571	104	11	2849	2571	104	0	3257	3352	84
2	1532	1615	109	5	2080	2083	101	11	2849	2571	104	11	2849	2571	104	3	1690	1789	102
5	2113	1960	88	7	2000	2052	107	2	1983	1894	102	2	1983	1894	102	4	1212	1232	130
6	1250	803	111	8	929	863	167	3	1914	1818	97	3	1914	1818	97	7	1918	1838	108
10	1774	1811	108	13	1141	1296	148	7	923	875	168	7	923	875	168	8	1379	1171	119
11	2993	2909	93					12	1561	1685	132	12	1561	1685	132	10	3000	2827	98
12	1692	1610	108	4	2988	2996	89												
15	1714	1723	113	6	2051	2013	108	0	3257	3352	84	0	3257	3352	84	13	1254	1340	138
1	3451	3513	79	4	2021	1851	108	3	1690	1789	102	3	1690	1789	102	1	917	1017	159
4	2523	2494	84	5	2978	2832	91	4	1212	1232	130	4	1212	1232	130	2	1975	1927	98
9	3389	3352	88	6	2635	2458	98	7	1918	1838	108	7	1918	1838	108	7	2546	2464	98
14	2995	2851	85	9	1345	1285	144	8	1379	1171	119	8	1379	1171	119	8	2829	2733	95
								10	3000	2827	98	10	3000	2827	98	11	2009	1732	111
1	1794	1866	98	10	1774	1811	108	13	1254	1340	138	13	1254	1340	138	11	2009	1732	111
4	2218	2208	91	2	1397	1463	124	13	1254	1340	138	13	1254	1340	138	0	2789	2670	91
6	3035	3047	80	6	2076	2044	114	0	3257	3352	84	0	3257	3352	84	1	1902	1754	103
7	2327	2261	95	9	1342	1365	149	1	917	1017	159	1	917	1017	159	5	1162	1105	141
9	2469	2332	97	2	2487	2321	98	2	1975	1927	98	2	1975	1927	98	10	1711	1556	124
10	1509	1480	120	3	2354	2467	100	7	2546	2464	98	7	2546	2464	98	0	2789	2670	91
11	1122	816	142	5	1940	1872	112	8	2829	2733	95	8	2829	2733	95	1	1902	1754	103
12	2807	2590	93	7	3245	3214	97	11	2009	1732	111	11	2009	1732	111	5	1162	1105	141
13	2724	2591	91					0	2789	2670	91	0	2789	2670	91	10	1711	1556	124
16	1551	1454	111	1	1710	1413	112	1	1902	1754	103	1	1902	1754	103	0	2789	2670	91
								10	1711	1556	124	10	1711	1556	124	1	917	1017	159
1	2542	2527	87	5	2075	1882	116	10	1711	1556	124	10	1711	1556	124	0	2789	2670	91
3	1373	1446	108																
5	1894	1837	112	1	1830	1621	106												
11	1310	1119	130	2	2953	2910	88	2	1197	1378	137	2	1197	1378	137	2	1864	1695	111
15	2032	2160	96	4	1471	1299	117	4	1864	1695	111	4	1864	1695	111	6	1863	1861	111
								6	2542	2447	91	6	2542	2447	91	8	2397	2166	102
0	2672	2556	84	6	2542	2447	91	8	2397	2166	102	8	2397	2166	102	0	5342	5137	85
1	1390	1263	110	8	3833	3561	87	8	3833	3561	87	8	3833	3561	87	0	5342	5137	85
2	1687	1582	95	9	1372	1262	130	9	1372	1262	130	9	1372	1262	130	3	2249	2087	98
3	1112	1109	118	11	1880	1649	110	11	1880	1649	110	11	1880	1649	110	4	1215	1109	142
4	2718	2686	88																
5	4032	4173	82	0	2127	2375	100	0	2127	2375	100	0	2127	2375	100	7	2441	2368	105
6	3169	3033	85	2	1480	1441	115	2	1480	1441	115	2	1480	1441	115	3	1479	1361	135
9	2625	2527	101	5	1181	878	126	5	1181	878	126	5	1181	878	126	4	1505	1407	133
								6	3169	3033	85	6	3169	3033	85	7	2101	1967	113
								8	1228	1092	144	8	1228	1092	144				

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